



Thermodynamics of the Corn-Ethanol Biofuel Cycle

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THIS WEB VERSION
is being periodically updated

New: Appendix D on fuel cells, consistent use of fuel HHVs,
corrected theoretical yield of ethanol from starch, equivalent CO₂
emissions and CExC adjusted to Patzek's, not Shapouri et al.'s inputs,
added eroded soil humus oxidation
Increased ethanol yield to ~2.5 gal/*wet* bushel, 91.5% of theoretical yield
Appendix E on free energy consumed to produce machinery

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Contents

1	Introduction	1
1.1	Corn Highlights	2
1.2	Energy Inputs to Corn Production	3
1.3	Layout	4
I	Mass & Energy Balance	5
1	Introduction	5
2	Mass Balance of Corn	5
3	Major Energy Inputs to Corn Production	7
3.1	Field Chemicals	7
3.1.1	Specific Energy Requirements for Nitrogen Fertilizer	8
3.1.2	Other Energy Inputs to Fertilizer Production	10
3.1.3	Specific Energy Requirements for Phosphorus Fertilizers	10
3.1.4	Specific Energy Requirements for Potassium Fertilizers	11
3.1.5	Specific Energy Requirements for Calcinated Lime	11
3.1.6	Specific Energy Requirements for Herbicides and Insecticides	12
3.2	Specific Energy Requirements for Fossil Fuels	12
3.3	Use of Electricity	14
3.4	Averages Can Be Misleading.	15
3.5	Energy in Human Labor	15
3.6	Energy in Corn Seeds	16
3.7	Energy in Irrigation	16
3.8	Energy in Transportation	17
3.8.1	Personal Commute	18
3.9	Machinery & Infrastructure	18
3.10	Fossil Energy Inputs into Corn Production	19
3.11	Solar Energy Input into Corn Production	19
3.12	Soil Humus and Micro-Element Depletion by Corn Production	20
4	Major Energy Inputs to Ethanol Production	22
4.1	Corn Mass Balance Revisited	22
4.2	Transport in Ethanol Refineries	24
4.3	Fossil Energy Inputs to Ethanol	24
4.4	Energy Credits	25
4.5	Overall Energy Balance of the Corn-Ethanol Process	25
II	Sustainability & Renewability	28
1	Introduction	28
2	Disclaimer	28

3 Preliminaries	29
4 Laws of Thermodynamics	29
5 Thermodynamics and Economics	31
6 Economic Activity	32
7 Agriculture	34
8 Industrial Production	35
9 Waste	36
10 Sustainability	36
10.1 The Earth is an Open System to Heat Flow	38
10.2 Conclusions	39
III Sustainability of Corn-Ethanol Cycle	41
1 Introduction	41
2 Available Free Energy	41
2.1 Introduction to Exergy	41
2.2 Change of B_{th} between Two States	42
2.3 An Industrial Flow Process	42
2.4 Cumulative Exergy Consumption (CExC)	44
3 The Ideal and Real Corn-Ethanol Cycle	44
4 System Boundary	46
5 The Carbon Cycle	46
5.1 Net CO ₂ Emissions	47
5.2 Conclusions	49
6 Water Cycle	50
7 Exergy Analysis of the Ideal Corn Ethanol Cycle	52
7.1 Chemistry of the CO ₂ -Glucose-EtOH Cycle	52
7.1.1 The Maximum Cycle Output per Unit Mass of Corn	53
8 Exergy Analysis of the Modified Ideal Corn-Ethanol Cycle	53
9 Resource Consumption and Waste Generation in the Industrial Corn-Ethanol Cycle	54
9.1 Cleanup of BOD in Ethanol Plant Wastewater	55
9.2 Cleanup of Contaminated Field Runoff Water	56
9.3 Restoration Work	57

10 Conclusions	58
IV Other Problems with Corn-Ethanol	60
1 Introduction	60
2 First-Law View of Corn-Ethanol Production in 2004	60
3 Second-Law View of Corn-Ethanol Production in 2004	61
4 Public Health Problems	63
V Summary & Conclusions	64
A Examples of Entropy Production and Disposal	69
B Availability and Irreversibility in Thermal Systems	72
C Is Economic Sustainability Possible?	75
D Efficiency of a Fuel Cell System	77
E Cumulative Exergy Consumption in Steel Production	79
E.1 Steel component manufacturing	79
VI Tables	87

List of Tables

1	Corn kernel composition	87
2	Application rates of field chemicals	87
3	Specific energies and application rates of nitrogen fertilizer	87
4	Energy consumption in superphosphate production	88
5	Specific energy and application rate of phosphorus fertilizers	88
6	Energy consumption in potassium fertilizer production	88
7	Specific energy and application rates of potassium fertilizer	88
8	Specific energy and application rates of calcinated lime	89
9	Specific energy and application rates of herbicides	89
10	Specific energy and application rates of insecticides	89
11	Average high and low heating values of fuels	90
12	Calorific values and specific volumes of gasoline	90
13	Calorific values and specific volumes of diesel fuel	91
14	Calorific values and specific volumes of LPG fuel	91
15	Calorific values and specific volumes of methane	91
16	Electricity use in corn farming	92
17	Energy used in transportation related to corn farming	93
18	Corn kernel composition	93

19	Specific CO ₂ emissions	93
20	Chemical exergies of compounds in the ideal corn-ethanol cycle	94
21	Product exergies after each step of the ideal corn-ethanol cycle	94
22	Product exergies after each step of the ideal corn-ethanol-hydrogen cycle	94
23	CExC of major non-renewable resources used in the industrial corn-ethanol cycle . .	95
24	First Law summary of the U.S. corn-ethanol production in 2004	95
25	Second Law summary of the U.S. corn-ethanol production in 2004	96
26	Primary energy consumption in steel production from ore	96
27	Primary energy consumption in steel production from scrap	96
28	Estimates of primary energy embedded in steel and its constituents	97
29	Estimates of CExC in steelmaking	97

List of Figures

1	Starch molecule	6
2	Mean ethanol yield	7
3	History of energy efficiency of ammonia production	8
4	Seven largest ammonia plants in the U.S.	9
5	Various estimates of the unit energy consumption to produce ammonium nitrate . .	10
6	The overall fertilizer application rates	12
7	The overall application rates of herbicides and insecticides	13
8	The overall fossil fuel volumes used in corn farming	14
9	By-state and average use of methane in corn farming	14
10	By-state and average use of electricity in corn farming	15
11	Energy use in labor	16
12	Specific energy use in transport related to corn farming	17
13	Major fossil energy inputs into corn farming	18
14	Solar energy dominates all other energy inputs to corn farming	20
15	soil nutrient losses with corn grain and stover removal	21
16	Mass balance of dry corn and efficiency of ethanol production	23
17	Average wet and dry corn yields	23
18	Average fossil energy inputs to ethanol production in a wet milling plant	24
19	The overall energy balance of ethanol production	26
20	Fossil energy gain/loss in corn ethanol production	27
21	Net energy yield in corn production	27
22	The boundary, process, and environment	29
23	The Second Law efficiency of copper production	32
24	The 2001 per capita energy consumption in the world	33
25	The 1990 per capita water consumption in the world	33
26	The 1999 per capita carbon emissions in the world	34
27	August 14, 2003, power blackout in New York	35
28	A linear process in industry	36
29	A linear process in industrial agriculture	37
30	Energy and mass flow in an ecosystem	38
31	Thermodynamic cycles	38
32	Exergy balance in an isothermal, ideal flow machine	42
33	Exergy balance in ideal and real nonisothermal industrial process	43

34	The ideal corn-ethanol cycle	44
35	Average U.S. corn yield from 1866 to 1939	45
36	The industrial corn-ethanol cycle	46
37	The carbon cycle	47
38	Specific CO ₂ emissions in the industrial corn-ethanol cycle	48
39	Total CO ₂ emissions in the industrial corn-ethanol cycle	49
40	The water cycle	50
41	Exergy flow the ideal CO ₂ -Glucose-EtOH cycle	52
42	Exergy flow the ideal CO ₂ -Glucose-EtOH-H ₂ cycle	54
43	Some of the useful work from the industrial corn-ethanol cycle is diverted to “undo” mining of the environment by this cycle	55
44	CExC by the industrial corn-ethanol cycle.	56
45	The minimum cumulative exergy consumption by the industrial corn-ethanol cycle and its maximum useful work	57
46	Societal costs of ethanol production in the U.S.	61
47	The cumulative one-hour exceedances of maximum legal ozone level in Southern California	62
48	Electricity obtained from solar cells and ethanol powered fuel cells	76

Abstract

In this paper I define sustainability, sustainable cyclic processes, and quantify the degree of non-renewability of a major biofuel: ethanol produced from industrially-grown corn.

First, I demonstrate that more fossil energy is used to produce ethanol from corn than the ethanol's calorific value. Analysis of the carbon cycle shows that all leftovers from ethanol production must be returned back to the fields to limit the irreversible mining of soil humus. Thus, production of ethanol from whole plants is unsustainable. In 2004, ethanol production from corn will generate 11 million tonnes of incremental CO₂, over and above the amount of CO₂ generated by burning gasoline with 115% of the calorific value of this ethanol.

Second, I calculate the cumulative exergy (available free energy) consumed in corn farming and ethanol production, and estimate the minimum amount of work necessary to restore the key non-renewable resources consumed by the industrial corn-ethanol cycle. This amount of work is compared with the maximum useful work obtained from the industrial corn-ethanol cycle. It appears that if the corn ethanol exergy is used to power a car engine, the minimum restoration work is about 7 times the maximum useful work from the cycle. This ratio drops down to 2.4, if an ideal (but nonexistent) fuel cell is used to process the ethanol.

Third, I estimate the U.S. taxpayer subsidies of the industrial corn-ethanol cycle at \$3.3 billion in 2004. The parallel subsidies by the environment are estimated at \$1.9 billion in 2004. The latter estimate will increase manifold when the restoration costs of aquifers, streams and rivers, and the Gulf of Mexico are also included.

Finally, I estimate that (per year and unit area) the inefficient solar cells produce ~100 times more electricity than corn ethanol. We need to rely more on sunlight, the only source of renewable energy on the earth.

KEY WORDS: biofuel, ethanol, fossil fuels, corn, sustainability, thermodynamics, energy, entropy, exergy, solar

Nullis in verba (Take nobody's word)

The motto of the Royal Society of London, 1662

1 Introduction

In the Preface to *What is Life?* – one of the great science classics of all times – ERWIN SCHRÖDINGER (1967) observed: “A scientist is supposed to have a complete and thorough knowledge, at first hand, of *some* subjects and, therefore, is usually expected not to write on any topic of which he is not a master. This is regarded as a matter of *noblesse oblige*.”

The principle of non-interference with the far-away fields of science often precludes the scientists from seeking to explain the *universal* aspects of science, which are of paramount importance to the society at large. For example, the sophisticated technological models of biofuel production, e.g., HEMELINCK (2004), cannot be formulated alone, without welding them first to a detailed analysis of the possibilities of depleting the environment in the long-term and destroying the valuable ecosystems¹. This example is not merely of academic interest. A *United Nations Bioenergy Primer* (Kartha and Larson, 2000) states: “In the most biomass-intensive scenario, [modernized] biomass energy contributes...by 2050...about one half of total energy demand in developing countries. . . . The IPCC's² biomass intensive future energy supply scenario includes 385 million hectares of biomass energy plantations globally in 2050 (equivalent to about one quarter of current planted agricultural area), with three quarters of this area established in developing countries.” The magic

¹For a definition, see Footnote 33.

²Intergovernmental Panel on Climate Change.

word “sustainable” appears 130 times in this *Primer*, without ever being defined³. What will happen if the developing countries entrust their fragile ecosystems and societies to a fundamentally flawed, unsustainable energy supply scheme? What if the distributed generation of solar power is a significantly better alternative to biofuels?

So here I renounce the noblesse and embark on a synthesis of facts and theories related to the production of a common biofuel, ethanol from corn, albeit with second-hand knowledge of some of these facts – and at a risk of making a fool of myself. I hope that some or most of this paper will be read by the concerned farmers, engineers, environmentalists, and policymakers. In particular I wish to reach the fellow scientists, who – for most part – remain blissfully unaware of the astronomical *real* problems with supplying energy to over 6 billion people, but who often vigorously analyze the peripheral issues (which in addition are tackled in isolation and out of context).

Most traditional biofuels, such as ethanol from corn, wheat, or sugar beets, and biodiesel from oil seeds, are produced from classic agricultural food crops that require high-quality agricultural land for growth. A significant portion of the sunlight these crops capture is diverted to produce seeds and store sugar, and their growing seasons are short. The net energy yield of corn⁴, ~100-130 GJ/ha-crop (Part I of this paper), is significantly lower than those⁵ of perennial crops and grasses (200-300 GJ/ha-crop), and sugarcane (~400 GJ/ha-crop) (Rogner, 2000). Also, the environmental costs of annual crops are very high: they cause more soil erosion (up to 100-fold), require 7-10 times more pesticides, and more fertilizers than perennial grasses or wood (Berndes et al., 2003). Finally, industrial *manufacturing* of hybrid seeds is very energy-intensive.

In this paper, I will describe in some detail the unfavorable thermodynamics of the industrial production of ethanol from one particular food crop, corn. I will use the Second Law of thermodynamics to track what is happening to us (or, is it U.S.?) as mere years pass, and the precious resources the sun and the earth have been making and storing for millions of years are being squandered in front of our eyes.

1.1 Corn Highlights

The U.S. is the single largest corn producer in the world. Large overproduction of subsidized cheap corn forces corn producers and processors to invent new ingenious uses for their product⁶. In terms of their large negative impact on the society and the environment, two corn products – ethanol and high-fructose syrup – stand out (Pollan, 2002; Elliott et al., 2002). About 13% of the U.S. corn production is now diverted to produce ethanol. Hence, in this paper I will *de facto* argue that the U.S. corn production should be reduced by at least 13% with significant benefits to the taxpayers and the planet. A telegraphic description of the U.S. corn farming and processing is as follows:

- Corn is the single largest U.S. crop (a record 300 million tonnes of moist corn grain in 2004).
- Corn is harvested from ~30 million hectares, roughly the area of Poland or Arizona, and a bit less than 1/4 of all harvested cropland in the U.S.
- The recent average yield⁷ of moist corn grain has been ~8600 kg/ha (and a record 10100 kg/ha in 2004).

³The endlessly repeated harvest of whole plants that grow on the same soil would have to be sustainable.

⁴The energy of dry corn grain minus the fossil energy inputs per hectare and per crop.

⁵The reported net energy yields of perennial grasses, sugarcane, etc., seem somewhat high to me.

⁶“Ethanol production makes huge amounts of the nation’s corn disappear – some 1.4 billion bushels went into ethanol production in 2004 – and that affects overall corn supply and helps shore up corn prices nationwide.” National Corn Growers Association, <http://www.ncga.com/ethanol/main/economics.htm>, accessed July 2, 2005.

⁷Source: USDA NASS database: www.usda.gov/nass/

- 42% of world's 708 million tonnes of moist corn grain⁸ in 2004 was produced in U.S.
- All of the U.S. corn fields are fertilized.
- Corn requires more fertilizer than any other major crop; 40% of all nitrogen fertilizer goes to corn (Frink et al., 1999).
- Corn erodes soil much faster than it can rejuvenate by natural processes.
- Corn needs ~100 cm water, 15% of corn is irrigated.
- Between 1995 and 2003, USDA distributed \$37.4 billions, or ~\$2 – \$7 billions per year, in corn crop subsidies. Recipients of payments made through most cooperatives, and the amounts, have not been made public⁹.
- From 1995 to 2003, the top 10 percent of corn subsidy recipients were paid 68 percent of all corn subsidies. The mean payments were \$465 172 each for the top first percent, and \$176 415 each for the top tenth percent of recipients. The bottom 80% of farmers received mean payments of \$4763 each.
- Over 12 billion liters of corn ethanol was produced in the U.S. in 2004.
- U.S. goal: Produce 20 billion liters of ethanol from corn annually.
- Ethanol producers receive ~\$3 billion annually from the federal government and state governments, and extract ~\$2 billion from the environment.

1.2 Energy Inputs to Corn Production

Fossil energy is essential to industrial agriculture. The following are the major energy inputs to industrial corn farming:

- Nitrogen fertilizers (all fossil energy)
- Phosphate, potash, and lime (mostly fossil energy)
- Herbicides and insecticides (all fossil energy)
- Fossil fuels: diesel, gasoline, liquified petroleum gas (LPG), and natural gas (NG)
- Electricity (almost all fossil energy)
- Transportation (all fossil energy)
- Corn seeds and irrigation (mostly fossil energy)
- Infrastructure (mostly fossil energy)
- Labor (mostly fossil energy)

Corn produced at a large expense of fossil energy is then transformed, with even more fossil energy, into pure ethanol.

⁸Source: 2004 world production of corn: www.cbot.com/cbot/pub/static/files/c_wprod.gif.

⁹Source: Farm Subsidy Database, <http://www.ewg.org/farm>, accessed July 2, 2005

1.3 Layout

This paper is divided into five parts, each of which can be read more-less independently. In Part I, I discuss the mass balance of corn processing, and the energy and mass balances of corn farming and ethanol production. Any First Law analysis of the corn-ethanol production process is *fundamentally incomplete*, and gives rise to confusion and arguments, which become moot once a more complete Second Law analysis is performed. Therefore, in Part II, I overview the fundamentals of thermodynamics, define the linear processes and cycles, irreversibility and sustainability, as well as *exergy* (the free energy available relative to the environmental conditions). In Part III, I apply the concepts developed in Part II to the industrial corn-ethanol cycle and answer the following questions:

1. Is ethanol production from corn a sustainable process?
2. If it is not sustainable, how unsustainable is it?
3. Can process changes result in making ethanol production from corn sustainable?

In particular, in Part III, I discuss the Carbon Cycle, the Water Cycle, the Ideal and Industrial Corn-Ethanol Cycles, and calculate the minimum work required to restore the nonrenewable resources consumed to produce corn ethanol. In Part IV, I estimate the various subsidies lavished on the transnational agribusiness corporations by the U.S. federal and local governments, and the huge subsidy extracted by these corporations from the U.S. environment: the rural population, soil, groundwater, rivers, the Gulf of Mexico, air, plants, and wildlife. Part V lists all major conclusions from this work.

Through my analysis, I hope to put to rest the sweeping statements made by some scientists, such as the following (Deluga et al., 2004):

Fast and efficient fuel reforming is one of the critical steps in producing H₂ for fuel cells and the “hydrogen economy,” and ethanol is now the most available and economically renewable fuel...

...Recent studies indicate that the energy in the fuel-ethanol is at least 1.34 times the energy used in its production.

Part I

Mass & Energy Balance

1 Introduction

Here, I revisit the classical story of ethanol from corn: the cumulative mass and energy balance of corn farming and subsequent ethanol production. There have been several well-known predecessors who told their versions of this story before. Professor DAVID PIMENTEL of Cornell is the world-famous agricultural expert from Cornell University, author of the *CRC Handbook of Energy Utilization in Agriculture* (Pimentel, 1990), the book *Food, Energy, and Society* (Pimentel, 1996), and dozens of publications on the subjects of ethics and energy efficiency of agriculture, e.g., (Pimentel et al., 1988; Pimentel and Dazhong, 1990; Pimentel, 1991; Pimentel et al., 1994; Pimentel, 2001). Here I will use only his most recent analysis of production of corn-ethanol (Pimentel, 2003). Doctor MICHAEL WANG, CHRISTOPHER SARICKS, and MAY WU are the authors of the 1997 Argonne National Laboratory Report, *Fuel-Cycle Fossil Energy Use and Greenhouse Gas Emissions of Fuel Ethanol Produced from the U.S. Midwestern Corn* (Wang et al., 1997), which told quite a different version of the corn-ethanol story. Doctor HOSEIN SHAPOURI, JAMES DUFFIELD, and MICHAEL WANG co-wrote the most recent 2002 USDA Report: *The Energy Balance of Corn Ethanol: An Update* (Shapouri et al., 2002a), which was less rosy than the Argonne Report, but considerably more optimistic than Professor PIMENTEL's analysis. I joined the corn-ethanol story in late 2002, after reading the thought-provoking book, *Food, Energy, and Society* by PIMENTEL. Subsequently, in Spring 2003, I decided to teach a Freshman Seminar at Cal on the subject of corn ethanol, and published with the students our own version of the story, *Ethanol from Corn: Clean Renewable Fuel for the Future, or Drain on Our Resources and Pockets?* (?), highly critical of the Argonne Report, and also critical of the 2002 USDA Report.

The new twist in my current story is the consistent split of all energy flows into the *specific energies* in MJ/kg, and the *mass fluxes* in kg/ha-crop¹⁰. This split will make my explanations clearer, and the remaining errors easier to spot and correct. However nicely told, the classical ethanol-from-corn story is fundamentally incomplete and should be treated as Part I of the *Ethanol Biofuel Trilogy*. Parts II and III of this Trilogy must be told in the somewhat exotic language of the Second Law of thermodynamics, which describes the passage of Time and limits the possible directions of natural and industrial processes.

2 Mass Balance of Corn

Mass will be usually expressed in kilograms per hectare. One hectare is 10000 m² or 2.47 acres.

It is safe to assume that the average corn yield in the U.S. has increased 5-fold over the last 70 years (NASS, 2004b). For example, in Indiana (Nielsen, 2002), the average corn yield was ~ 30 ¹¹ bushels per acre in 1930 and 156 bushels per acre in 2001. The steadily improving yield resulted mostly from the increased fertilizer use¹² and better corn genetics.

In 2001-2003, corn in the U.S. was harvested from $\sim 28 \times 10^6$ hectares (NASS, 2004a). The average corn yield varied from 130 bushels per acre in 2002, to 138.2 bushels/acre in 2001, and

¹⁰The notation kg/ha-crop stresses that the time unit in the average fluxes is the duration of a single crop, e.g., 120 days in the case of corn plants.

¹¹The lowest average yield of corn in the U.S., ~ 18.2 bushels/acre, was recorded in 1901 and 1934, USDA, <http://usda.mannlib.cornell.edu/data-sets/crops/96120/trackrec2003.txt>.

¹²Between 1960 and 2000, the use of nitrogen fertilizer increased 5-fold (USDA-NASS database).

142.2 bushels/acre in 2003. (NASS, 2004a).

Definition 1 One equivalent bushel is defined as 56 pounds (25.4 kg) of corn grain that contains 15 percent of moisture by weight (Bender and Hill, 1997). □

Therefore, the mean corn yield over the last 3 years was just below 8600 kg/ha of *moist* corn grain, or $8600 \times 0.85 = 7300$ kg/ha of *dry* corn grain.

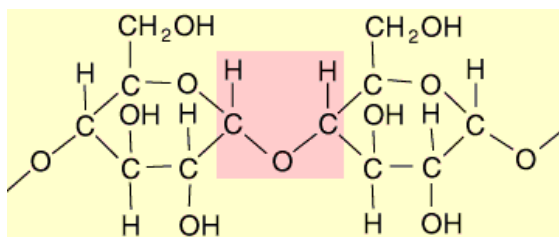


Figure 1: A typical starch molecule is constructed from α -glucosidic bonds (purple background), each of which links two dehydrated glucose molecules. These molecules form either unbranched or somewhat branched polymer chains with up to 360 or 1000 glucose units, respectively (Avers, 1976). In hydrolysis, the glucosidic bonds are broken, and each glucose unit gains one water molecule.

Since ethanol (EtOH) is made from the hydrolyzed starch, see **Figure 1**, with the theoretical efficiency of 0.51, 1 kg of dry corn grain may yield $(0.66 \times 180/162) \times 0.51 = 0.374$ kg of water-free EtOH with zero losses, see **Table 1**. Therefore, from 1 hectare, one may theoretically produce 2730 kg EtOH, given the dry corn yield above. Of course there are losses in the corn-to-ethanol conversion process, and the practical efficiency will be lower.

Remark 1 In the U.S., the customary unit of reporting efficiency of corn conversion to ethanol is gallons EtOH per bushel, e.g., (Shapouri et al., 2002a). With the standard¹³ EtOH density of 0.787 kg/L of anhydrous EtOH at 25°C, the theoretical efficiency of 0.374 kg EtOH/kg dry grain (0.475 L EtOH/kg dry grain) yields 3.19 gallons EtOH/dry bushel = 2.71 gallons EtOH/equivalent bushel of corn with 15% moisture¹⁴. □

The 2000 – 2004 mean yield has been calculated as follows:

$$\frac{\text{Industry reported gallons of denatured EtOH} - \text{Imports from Brazil}}{\text{Industry reported bushels of corn for ethanol}} \times 0.95 \quad (1)$$

The factor of 0.95 accounts for 5% by volume of gasoline denaturant in the ethanol shipped from biorefineries.

Remark 2 The 2000-2004 mean ethanol yield in the U.S. is $2.477 \approx 2.5$ gallons of pure EtOH (and up to 2% of longer-chain alcohols) per nominal wet bushel with 15% moisture. This yield has been calculated directly from the industry data. For details see (Patzek, 2006). □

USDA has arrived at the volume-averaged estimate of 2.682 gallons of anhydrous ethanol per wet bushel (Shapouri et al., 2002b). This estimate is too high (Patzek, 2006), Section 3.

¹³(Lide, 1994), p. 15-46.

¹⁴I would like to thank Dr. DREW RONNEBERG of Technology and Management Services, Inc. for pointing out my error in not accounting for the gain of one water molecule per glucose unit in starch hydrolysis, and referring me to the report by Dr. MICHAEL S. GRABOSKI (Graboski, 2002).

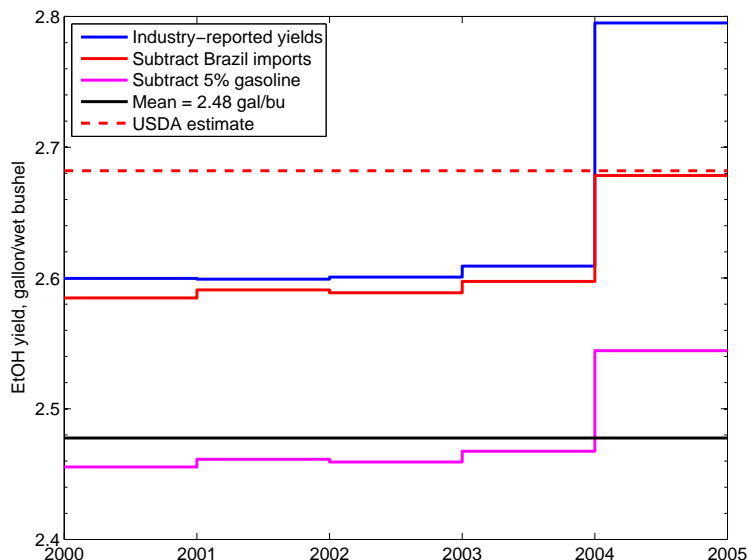


Figure 2: The mean of industry-reported ethanol yields. Source: (Patzek, 2006), Section 3.3.

3 Major Energy Inputs to Corn Production

Most energy inputs will be expressed in MJ/kg of active ingredient in the input. For example, ammonia contains 82% nitrogen (active component); therefore, the specific energy input in MJ to obtain one kg of ammonia will be divided by 0.82.

3.1 Field Chemicals

- **Nitrogen** is a component of many important structural, genetic and metabolic compounds in plant cells. It is a major component of chlorophyll, amino acids, cell energy carriers (ATP/ADP), and genetic material (DNA/RNA).
- **Phosphorus** is one of the primary structural components of cell membranes. It is involved in the photosynthesis (ADP/ATP), synthesis of proteins and vitamins, and it occurs in important enzymes.
- **Potassium** activates enzymes that produce proteins and sugars. It maintains water content and, hence, the turgor (rigidity) of plant cells.
- **Calcinated lime** is used to increase the pH of soil acidified by nitrogen fertilizer. The ideal pH for corn is 6.6.
- **Herbicides**, such as Atrazine, Acetochlor, S-Metolachlor, Dicamba, Nicosulfuron, etc. are used to protect corn from weeds.
- **Insecticides**, such as Chlorpyrifos, Terbufos, Carbofuran, Tefluthrin, etc. are used to protect corn from insects.

The average application rates of major field chemicals used in corn farming are reported in **Table 2**.

3.1.1 Specific Energy Requirements for Nitrogen Fertilizer

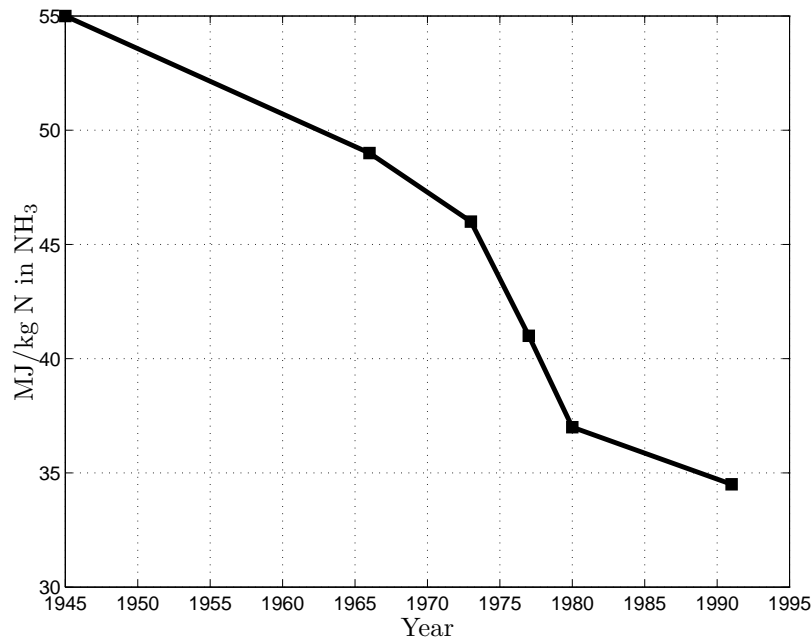


Figure 3: History of energy efficiency of ammonia production in MJ/kg N. Source: G. KONGSHAUG (1998).

Nitrogen fertilizers are derived from ammonia, nitric acid, and carbon dioxide. Practically all ammonia is produced from natural gas and nitrogen in the HABER-BOSCH process (Worrell et al., 1994; Kongshaug, 1998; Worrell et al., 2000). The energy efficiency of the HABER-BOSCH process has been improved by 1/3 over the last 60 years, see **Figure 3**. Therefore, the age of the ammonia-producing plant does matter. ERNST WORRELL et al. (2000) have compiled the ages and outputs of the 44 largest U.S. ammonia plants, see **Figure 4**. Most of these plants were built in the 1960's, and some were later modernized and expanded. The fact is that the major U.S. plants were built 40 years ago, and some were revamped 20-30 years ago. Another example comes from Europe: In 1995, ammonia synthesis in modern European plants consumed approx. 36.93 MJ/kg N, while older plants needed about 43.08 MJ/kg N (Biermann et al., 1999).

Remark 3 For nitrogen fertilizer production, I will use the average efficiency of 30-year old plants. I will also assume that all nitrogen fertilizer applied to the U.S. corn fields is represented by ammonium nitrate. □

KONGSHAUG (1998) analyzed energy efficiency of ammonia production and divided ammonia plants into three classes: “Modern,” “Average European plants,” and “30-years old plants.” Using his terminology, the major nitrogen fertilizers are produced with the following specific energy inputs *per unit mass of nitrogen* in them.

Ammonia, NH₃, has 82% of nitrogen by mass. Following KONGSHAUG (1998), I will assume the following *net* energy consumption to produce ammonia:

30 Years Old	47	MJ/kg N	
Average Europe	39	MJ/kg N	(2)
Modern	34.5	MJ/kg N	

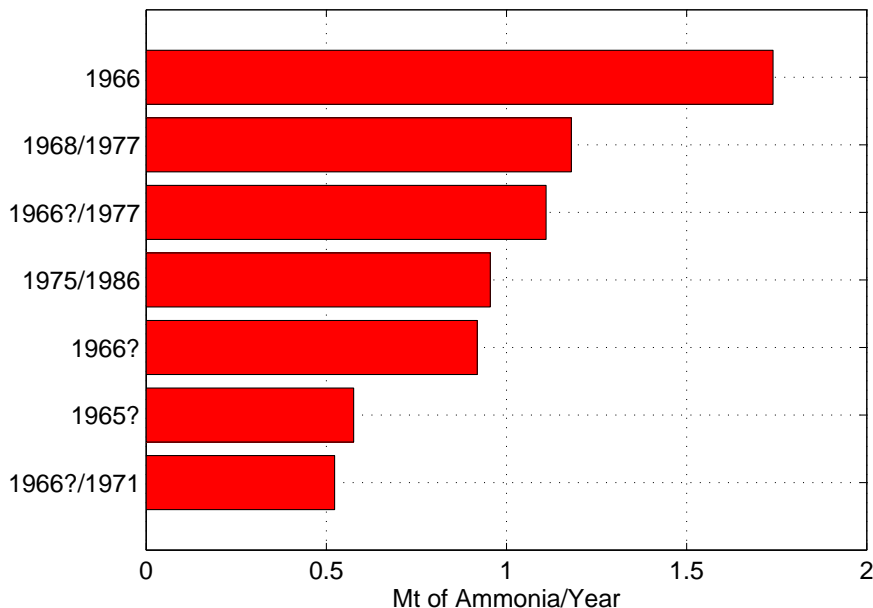


Figure 4: Together, these seven largest plants produce 40% of the U.S. ammonia. The first dates refer to plant opening. Some of the plants were later expanded and revamped, as indicated by the second dates. Source: ERNST WORRELL et al., (2000).

The November 1981 process description for Haldor Topsøe plants published in *Hydrocarbon Processing* (p. 129) gives 35.6 MJ/kg N as the total energy requirement for ammonia production.

According to Prof. VACLAV SMIL (1985), total energy expended on ammonia production in the U.S. is 55 MJ/kg N for the largest new plants, and up to 65 MJ/kg N for small pre-1969 units, the weighted average for all plants being 58 MJ/kg N (see p. 163-165). The European and Asian plants are more energy-efficient.

Urea, $\text{CO}(\text{NH}_2)_2$, has 45% of nitrogen by mass, and is obtained from ammonia and carbon dioxide: $2\text{NH}_3 + \text{CO}_2 \rightarrow \text{CO}(\text{NH}_2)_2 + \text{H}_2\text{O}$. The net energy consumption (Kongshaug, 1998) is:

$$\begin{array}{rcl}
 \text{30 Years Old Ammonia} & + & 10 & = & 57 \text{ MJ/kg N} \\
 \text{Average Europe Ammonia} & + & 9 & = & 48 \text{ MJ/kg N} \\
 \text{Modern Ammonia} & + & 7.2 & = & 42 \text{ MJ/kg N}
 \end{array} \quad (3)$$

SMIL's (1985) estimate (p. 164) of the U.S. urea production costs is 70 MJ/kg N.

Ammonium Nitrate, NH_4NO_3 , has 35% nitrogen by mass, and is produced from nitric acid and ammonia: $\text{HNO}_3 + \text{NH}_3 \rightarrow \text{NH}_4\text{NO}_3$. Nitric acid is obtained by burning ammonia over catalyst to produce NO_x . One of the by-products of ammonium nitrate production is nitrous oxide N_2O , a potent greenhouse gas. With 97% conversion of ammonia to AN, the energy consumption is

$$\begin{array}{rcl}
 \text{30 Years Old Ammonia} & + & 4 & = & 51 \text{ MJ/kg N} \\
 \text{Average Europe Ammonia} & + & 2 & = & 41 \text{ MJ/kg N} \\
 \text{Modern Ammonia} & + & 0.43 & = & 35 \text{ MJ/kg N}
 \end{array} \quad (4)$$

SMIL's (1985) estimate (p. 164) of the U.S. ammonium nitrate production costs is 72-90 MJ/kg N.

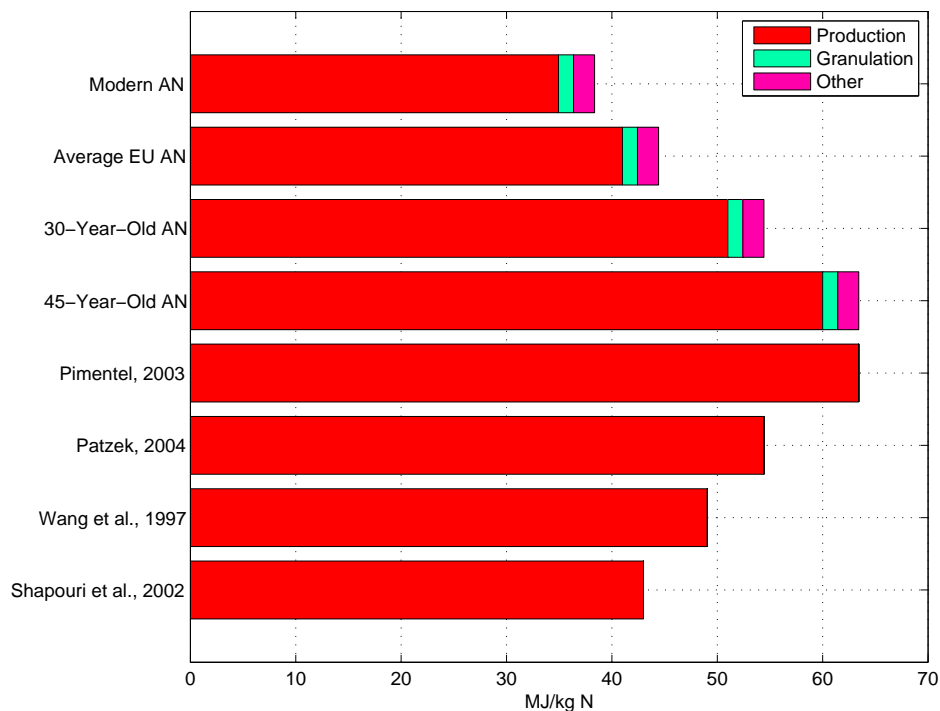


Figure 5: The various estimates of the unit energy consumption to produce ammonium nitrate (AN). The estimates by SHAPOURI et al. are listed in (Shapouri et al., 2002a), by WANG et al. in (Wang et al., 1997), and PIMENTEL's estimate (Pimentel, 2003) of 78.1 MJ/kg N was adjusted down to the average 45-year-old AN plant in Europe. The latter estimate is the middle of the well-researched estimates by SMIL (1985) for the U.S. plants.

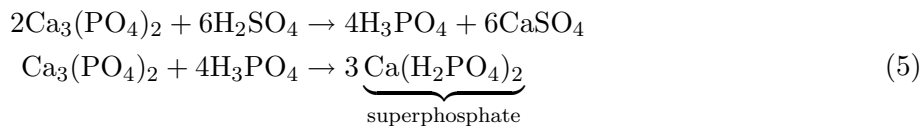
3.1.2 Other Energy Inputs to Fertilizer Production

Primary particulation of fertilizers is carried out in prilling and granulation processes. The granulation units can also be used for a second pass of product building blocks (for mixed fertilizers), in addition to compaction and bulk blending units. The energy requirement for primary particulation is $\sim 0.25\text{--}0.5$ MJ/kg product, and for secondary granulation and compaction 0.7-1.1 MJ/kg product (Kongshaug, 1998). Here I have used 0.5 MJ/kg of AN. The energy costs of natural gas recovery, compression, purification and transportation, and fertilizer packaging, if any, are estimated cautiously at 2 MJ/kg N. The various estimates of net energy inputs into nitrogen fertilizer production are shown in **Figure 5** and listed in **Table 3**. In all the tables below, the estimates by SHAPOURI et al. come from their 2002 report (Shapouri et al., 2002a), by WANG et al. from (Wang et al., 1997), and by BERTHIAUME et al. from (Berthiaume et al., 2001). Since BERTHIAUME et al. have reported only the specific exergy consumption, I have used my energy consumption estimates to represent theirs. I have also corrected and/or amended the various estimates as noted below.

3.1.3 Specific Energy Requirements for Phosphorus Fertilizers

Phosphate and phosphoric acid are produced from the igneous *fluorapatite* $\text{Ca}_{10}(\text{PO}_4)_6(\text{F}, \text{OH})_2$, and the sedimentary *francolite* $\text{Ca}_{10}(\text{PO}_4)_6\text{x}(\text{CO}_3)\text{x}(\text{F}, \text{OH})_{2+\text{x}}$. For example, superphosphate may

be produced as follows:



Calcium sulphate (gypsum) produced in reaction (5)₁ may be precipitated as a *dihydrate*, using milder conditions, 26-32% P₂O₅ at 70-80°C, or a *hemihydrate*, using 40-52% P₂O₅ at 90-110°C. Even though it demands more energy, the hemihydrate reaction is preferred because it optimizes use of sulfuric acid. The energy costs of gypsum disposal are omitted here.

Unlike the European Fertilizer Manufacturers Association (EFMA), which maintains an informative web site, the U.S. does not have clear statistics on the types of phosphorus (and other) fertilizers used. From the USGS and U.S. Census Bureau documents, e.g., (Jasinski, 2002) and references therein, it may be deduced that calcium phosphates (triple superphosphate and single superphosphate) dominate in the U.S., followed by Di-Ammonium Phosphate (DAP).

Just as with nitrogen fertilizers, phosphate fertilizers have different contents of active ingredient, here P₂O₅. For example, triple superphosphate is 46% P₂O₅, single superphosphate 18-20%, and di-ammonium phosphate 46%.

In this analysis, I will use the single superphosphate as the reference phosphorus fertilizer. The typical energy consumption for the 30 years old technology is given by KONGSHAUG (1998), see **Table 4**. In contrast to the energy input for nitrogen fertilizer, the data for phosphorus vary widely between 5.1 MJ/kg P₂O₅ and 26.4 MJ/kg P₂O₅ (Pimentel and Dazhong, 1990). EFMA uses 15.8 MJ/kg P₂O₅ (Biermann et al., 1999).

Using the data in Table 4, the typical U.S. energy consumption is 0.3+6.5 = 6.8 MJ/kg P₂O₅ as single superphosphate. The specific energy consumption for phosphorus and its application rates are listed in **Table 5**.

3.1.4 Specific Energy Requirements for Potassium Fertilizers

The energy requirements for potassium vary from 4.0 MJ/kg K₂O (Mudahar and Hignett, 1982) to 13.7 MJ/kg K₂O (Patyk and Reinhardt, 1997). EFMA uses 9.3 MJ/kg K₂O (Biermann et al., 1999). The typical energy requirements to produce different potassium fertilizers are listed in **Table 6**. Here I will use the muriate of potash (MOP), or KCl, as the typical potassium fertilizer. Production of KCl is mainly carried out by shaft mining and beneficiation. The most energy efficient potash processes, based on a high quality sylvanite salt, require only 1.5 MJ/kg MOP (2.5 MJ/kg K₂O). The energy consumption, however, varies a lot, and estimates up to 6 MJ/kg MOP are reported for mining of more mixed salts (Kongshaug, 1998).

Here I will treat KCl as an admixture to a superphosphate fertilizer, and give it an identical energy consumption of 6.8 MJ/kg K₂O. The specific energy inputs and application rates for the potassium fertilizers are listed in **Table 7**.

3.1.5 Specific Energy Requirements for Calcinated Lime

Lime must be added to de-acidify soil after heavy use of nitrogen fertilizers. Calcinated lime, CaO, is obtained from limestone and dolomites. The calcination process is energy-intensive, and generates one mole of CO₂ per one mole of CaO. EFMA uses 2.97 MJ/kg CaO (Biermann et al., 1999), but here I will follow the Australian Academy of Sciences and assume that only 1.75 MJ/kg CaO is used to produce calcinated lime.

Lime application rate is not commonly reported by the USDA. The suggested application rate is 1.8 times the application rate of nitrogen (Tisdale et al., 1985), but there are reports of several times higher application rates, e.g., (Pimentel, 2003).

The total application rates of nitrogen, phosphorus and potash fertilizers, as well as that of calcinated lime are shown in **Figure 6**.

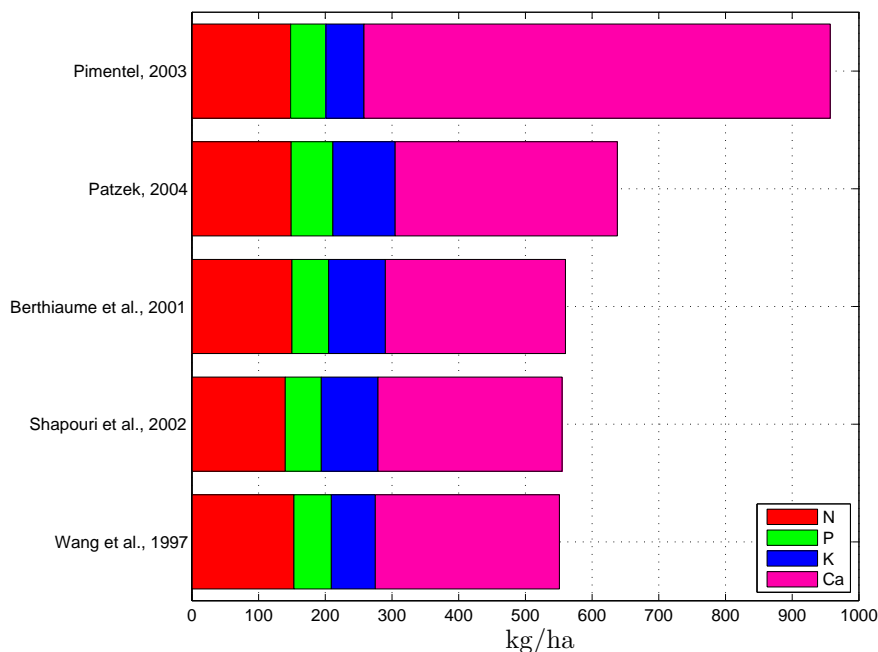


Figure 6: The total fertilizer application rates listed in Tables 3, 5, 7, and 8.

3.1.6 Specific Energy Requirements for Herbicides and Insecticides

There are many active ingredients in commercial herbicides and insecticides, but all have very high specific energy requirements for their production. Here I have used the data collected in Table 5 in SHAPOURI et al.'s report (2002a).

The specific energy inputs and application rates of herbicides and insecticides are listed in **Tables 9** and **10**. The overall application rates of herbicides and insecticides are also shown in **Figure 7**.

3.2 Specific Energy Requirements for Fossil Fuels

A unit mass of a fossil fuel gives out a specific amount of heat (its calorific value) when burned. The *Low Heating Value* (LHV), or *Net Calorific Value* (NCV), of a fossil fuel assumes that combustion products contain the water of combustion as vapor. The heat contained in this water is not recovered. Outside of power stations and fuel cells, water remains as vapor after combustion. The *High Heating Value* (HHV), or *Gross Calorific Value* (GCV), assumes that combustion water is entirely condensed. The heat contained in this water is recovered. PIMENTEL (2003), SHAPOURI et al. (2002a) seem to use HHV for all fuels. WANG et al. (1997) give only the total amount of energy for each fossil fuel; therefore, their specific values are back-calculated for each fuel and agree with my estimates.

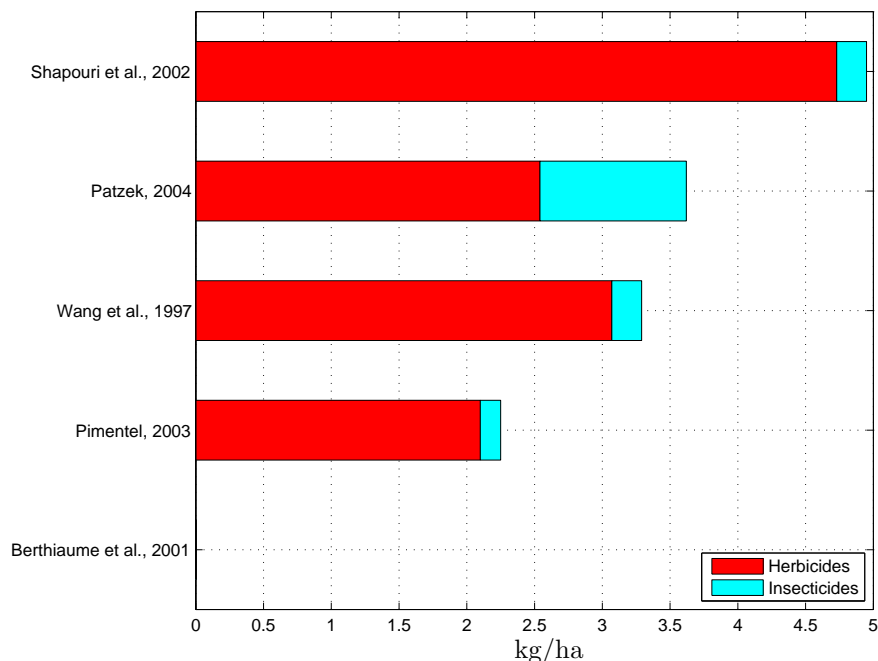


Figure 7: The total herbicide and insecticide application rates listed in Tables 9 and 10.

In prior work, I used the low heating values of fuels in combustion engines to perform the First Law energy balance calculations (Patzek, 2004; Patzek et al., 2005). More recently, however, I was swayed by BOSSEL (2003b) to use the high heating values. The reason is simple: Regardless of a machine (an internal combustion engine or a fuel cell) we use to burn a fuel, the fuel’s full energetic potential *could* be realized *if* we improved this machine. In other words, in determining sustainability, we check what possibly *could* be done, not what actually *is* done.

Remark 4 (Use of High Heating Values) From now on, the high heating values of all fuels will be used to determine whether a technological process is sustainable according to the First *and* Second Law of thermodynamics. □

After deciding on a high heating value for each fossil fuel, one needs to find the standard values for “average” gasoline, diesel fuel, LPG, and natural gas. Finding consensus was more difficult than I expected. The International Energy Agency’s (IEA) standards up to the year 2000 are published in (IEA, 2000). IEA changed its standards for gasoline equivalent and diesel fuel equivalent in 2003. Finally, I decided to use the high heating values measured and compiled by Professor DIETRAM CASTORPH et al. at the Technical University of Munich (Castorph et al., 1999), see **Table 11**. For reference, the standard densities of liquid fuels used in this work are: gasoline, 0.74 kg/L; diesel fuel, 0.84 kg/L; LPG, 0.58 kg/L; natural gas, 0.84 kg/sm³, and ethanol 0.79 kg/L.

The calorific values and average volumes of fossil fuels used in corn farming are listed in **Tables 12 – 15**. The cumulative volumes of all fossil fuels are shown in **Figure 8**.

Notice that not all sources account for all five fossil fuels, especially for LPG and natural gas. Both LPG and natural gas are used for corn drying and as fuel to power water pumps in irrigation. Their uses vary greatly from one state to another, and from one season to another, see **Figure 9**.

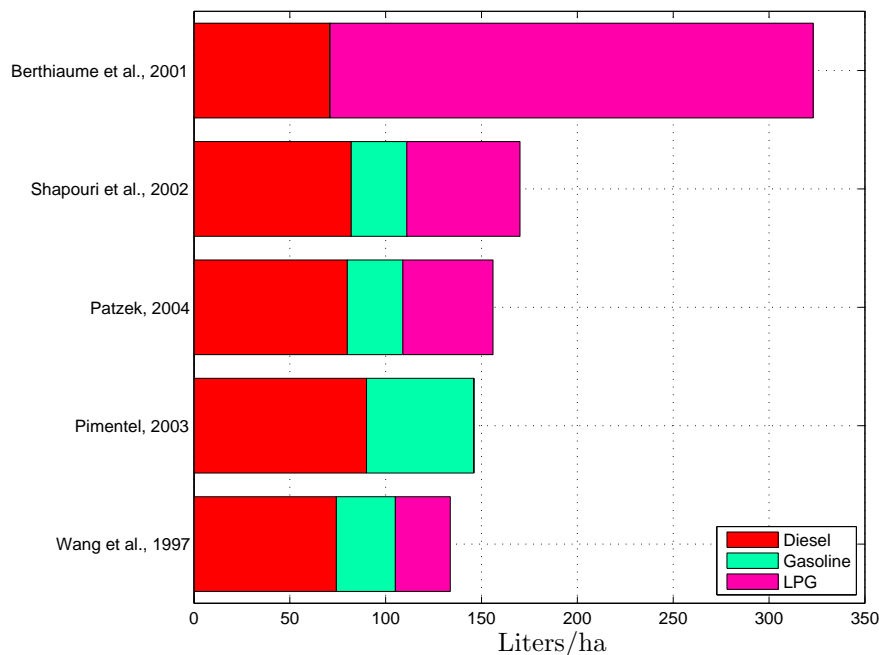


Figure 8: The total fossil fuel volumes listed in Tables 12 - 15.

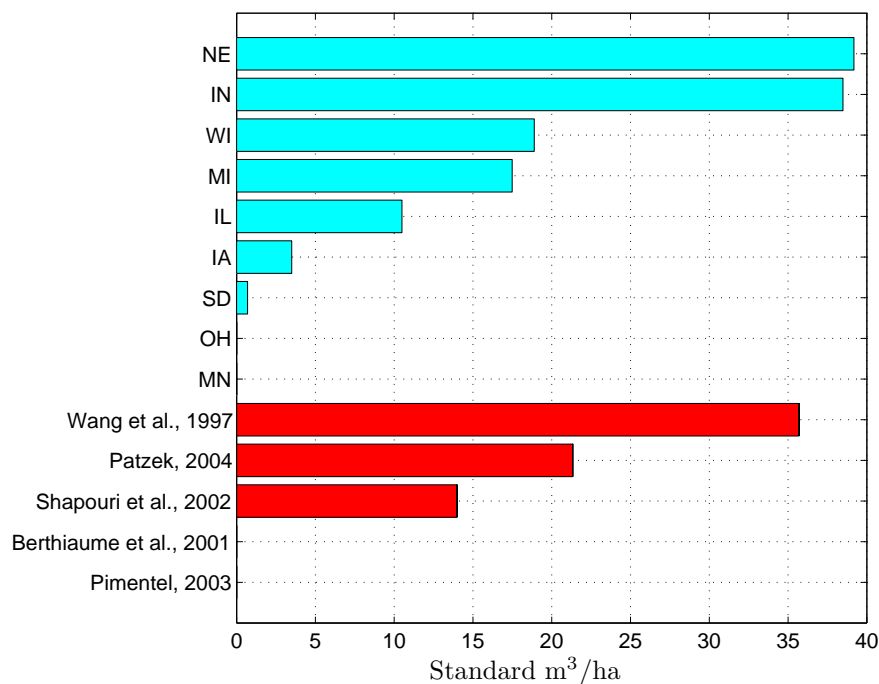


Figure 9: By-state and average use of methane in corn farming. The 1996 methane volume data from the largest corn-producing states are from SHAPOURI et al. (2002a). Note the large variability of methane use depending on wet/dry weather.

3.3 Use of Electricity

The average electric energy spent on farming 1 hectare of corn is listed in **Table 16**. SHAPOURI et al. (2002a) have attempted to include the efficiency of fossil energy conversion into electricity in

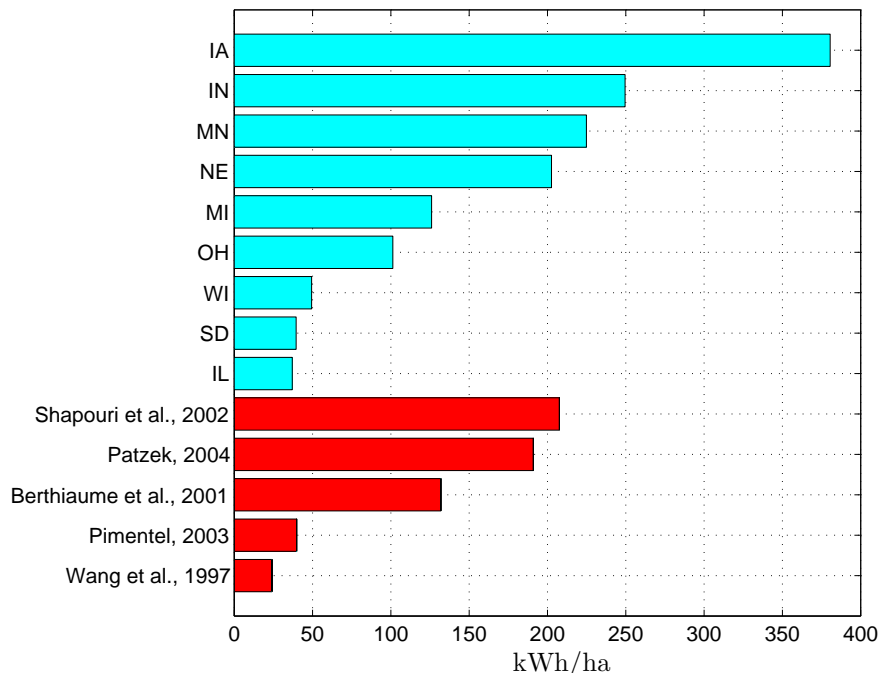


Figure 10: By-state and average use of electricity in corn farming. The 1996 electricity data from the largest corn-producing states are from SHAPOURI et al. (2002a). Note the large variability of electricity use depending on wet/dry weather.

their estimates. I have not done it here in the energy balance, but will consider all the steps in free energy conversion in the exergy balance in Part III. The average and the state data are compared in **Figure 10**. As with natural gas, variability in electricity use is very large.

3.4 Averages Can Be Misleading...

- Depending on the weather, *state-to-state* (regional) and *year-to-year* (temporal) variations in the use of electricity, liquid petroleum gas, and methane can be extreme.
- Therefore, a *single calculation* of energy efficiency of corn production for all states and all years is insufficient for the generalizations readily made from such a calculation. In other words, in the same corn growing season, the energy balance of corn ethanol production may be highly unfavorable in one state, and only somewhat unfavorable in another one.

3.5 Energy in Human Labor

Average labor time is 6.2 hrs/ha per growth season (Pimentel, 2003). I will assume that a physical worker is on a 4000 kcal/day diet. The specific energy to sustain a worker is then

$$4 \times 10^6 \text{ cal/day} \times 4.186 \frac{\text{J}}{\text{cal}} \times \frac{1}{24} \frac{1}{\text{hr}} = 0.7 \frac{\text{MJ}}{\text{hr}}, \quad (6)$$

and it is negligible.

In my estimate, I have used SHAPOURI et al.'s (2002a) custom labor estimate plus the worker sustenance energy, see **Figure 11**.

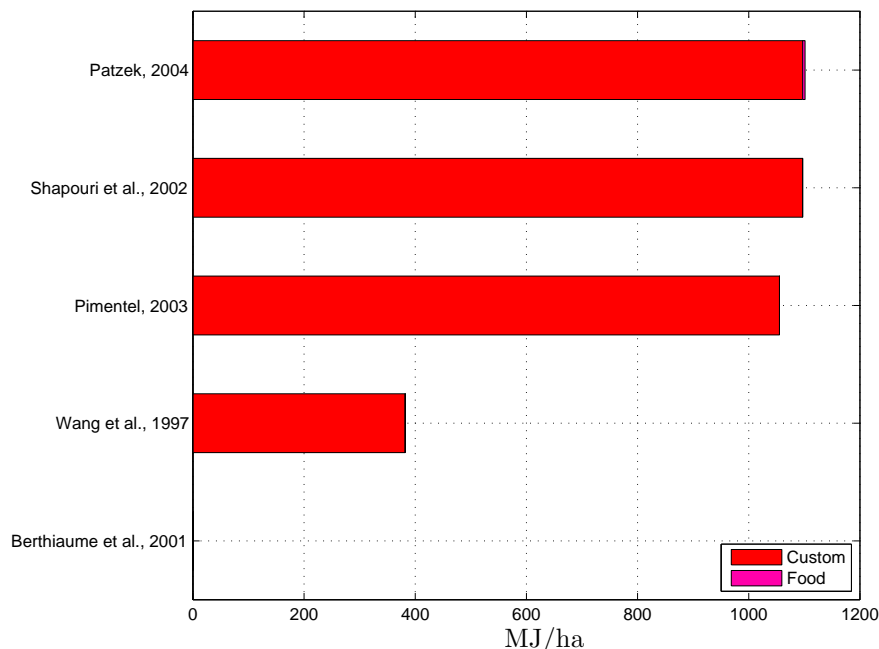


Figure 11: Energy use in labor.

3.6 Energy in Corn Seeds

The production of *hybrid* corn seeds is very energy intensive (Pimentel, 2004a). Two corn genotypes are needed. The pollen from one genotype is used, and the pollen from the other genotype must be prevented from pollinating the genotype from which the corn seed will be harvested. Because of all these manipulations, about 7 times *more* energy is required to produce hybrid seeds compared with the energy in the same mass of corn grain.

The following data have been used to estimate the specific energy requirements for corn seeds:

- The seed rate is 57,000 (Corn Belt) –74,000 (West) seeds per ha (White and Johnson, 2003), p. 255.
- The mean weight of 1000 corn seeds is 0.332 kg (White and Johnson, 2003), p. 202.
- The caloric value of corn grain with 13-15% moisture is 15-16 MJ/kg.
- The production of hybrid corn requires about 104 MJ/kg (Heichel, 2004).

Both SHAPOURI et al. (2002a), and WANG et al. (1997) ignored the hybrid seed energy, and assigned just the diesel fuel energy for planting the seeds. I have amended their calculations with the seed energy, based on their reported seed rates, and moved their estimates of diesel fuel use in corn planting to the fuel category.

3.7 Energy in Irrigation

Corn gets water mostly from rainfall, but some irrigation is generally required. In a dry season, the use of irrigation systems may increase substantially, as seen from the electricity use by state reported in Figure 10. I have made the following assumptions to estimate the energy cost of corn irrigation:

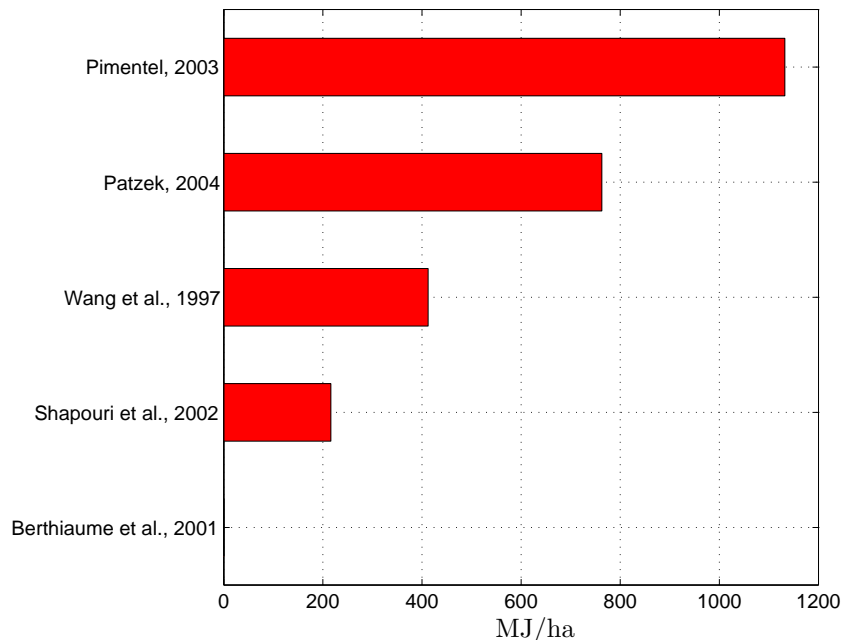


Figure 12: Specific energy use in transport related to corn farming. Note that PIMENTEL's estimate may contain an additional single commute to and from the field.

- Only 15% of crop is irrigated, USDA-NASS, 1997, (Pimentel, 2003).
- On average 8.1 cm of water is used per acre, USDA-NASS, 1997, (Pimentel, 2003).
- Water is pumped on average from the depth of 100 m.
- Pump efficiency, etc., is 0.75.

Then the specific energy requirement for irrigation is

$$\frac{1 \text{ m}}{100 \text{ cm}} \times 10,000 \frac{\text{m}^2}{\text{ha}} \times 1000 \frac{\text{kg}}{\text{m}^3} \times 100 \text{ m} \times 9.81 \frac{\text{m}}{\text{s}^2} / 0.75 = 131 \text{ MJ/cm-ha} \quad (7)$$

I have lowered PIMENTEL's 2003 estimate of irrigation energy to my estimate. Both SHAPOURI et al., and WANG et al. have buried the irrigation energy in their use of electricity and methane, so they account for the irrigation indirectly, for the particular mixture of states and weather they considered. Since I use SHAPOURI et al.'s estimates for the fossil fuels, I have not included the irrigation energy in my calculations.

3.8 Energy in Transportation

The specific energy use in transportation related to corn farming has been estimated by WANG et al. (Wang et al., 1997), see **Table 17**. The total energy use is about 720 kJ/kg of field chemicals (fertilizers, lime, fuel, etc.) transported into the field (~ 400 MJ/ha-crop), plus personal commutes, see **Figure 12**. This estimate is sensitive to the number of commutes to and from the field by personnel using motor vehicles.

3.8.1 Personal Commute

At 6.2 hr/ha/crop of labor, 20 l/100km gasoline use, and a 30 km round trip, the energy cost of commuting is

$$\frac{6.2 \text{ hr/ha/crop}}{9 \text{ hr work day}} \times 60 \text{ km} \times 20 \frac{\text{liter}}{100\text{km}} \times \frac{1}{100} \times 0.74 \frac{\text{kg}}{\text{liter gasoline}} \times 47 \frac{\text{MJ}}{\text{kg gasoline}} \quad (8)$$

$$= 288 \text{ MJ/ha-crop}$$

Therefore a single commute *nearly doubles* the overall transportation energy costs. This issue should be investigated further.

3.9 Machinery & Infrastructure

Industrial agriculture requires heavy machinery (trucks, tractors, ploughs, cranes, railroad cars, airplanes, locomotives, barges, ships, etc.), which must be replaced periodically. Industrial agriculture also requires an extensive infrastructure with a large environmental footprint (spare parts, machine shops, machine manufacturing factories, access roads, railroad tracks, ports, silos, pumps, driers, electricity generators, air-conditioners, etc.). Industrial corn is the single largest crop in the U.S., and its share of this infrastructure should be highest. The energy inputs as hardware have been estimated in **Appendix E** at 68-168 MJ/kg of operational machinery, close to the 110 MJ/kg estimated by PIMENTEL (2003). The mass of hardware assigned by PIMENTEL to corn farming is 55 kg/ha. Note that this estimate includes only a *tiny* part of the huge infrastructure listed above.

Both SHAPOURI et al. (2002a) and WANG et al. (1997) omit this input altogether, but I have not amended their calculations.

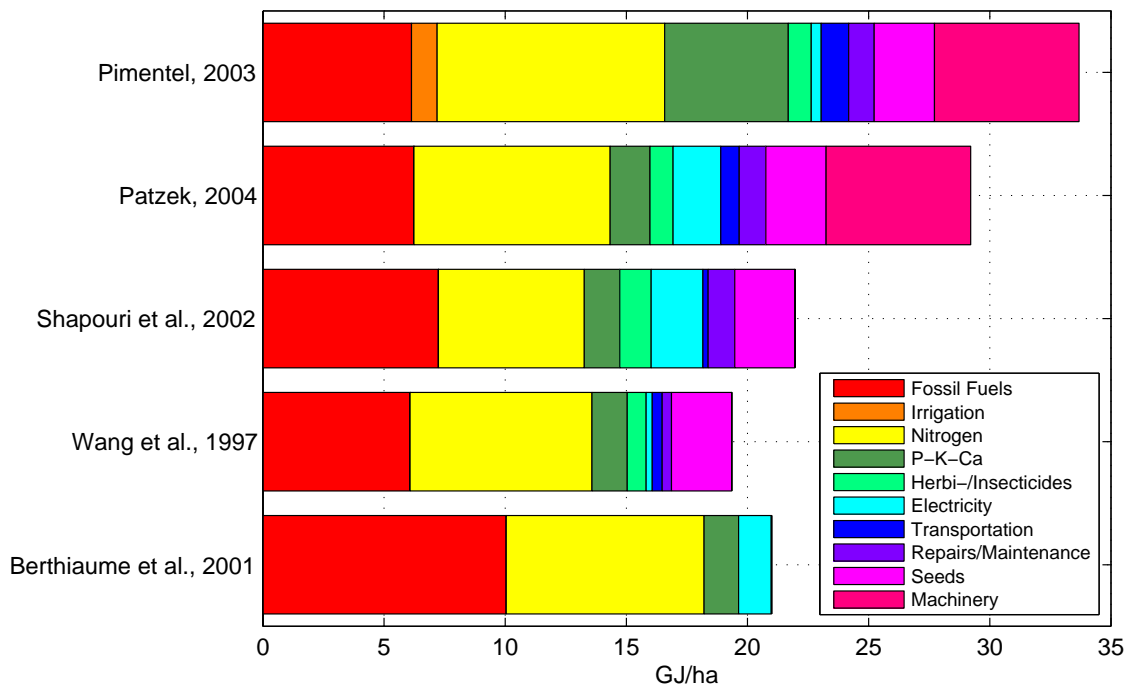


Figure 13: Major fossil energy inputs into corn farming.

3.10 Fossil Energy Inputs into Corn Production

The specific fossil energy requirements in industrial corn farming are shown in **Figure 13**. A few comments are in order.

- The lowered PIMENTEL's 2003 estimate of fossil fuel energy plus irrigation is identical with that of SHAPOURI et al.
- The lowered PIMENTEL's 2003 estimate of nitrogen fertilizer energy is higher than the uncorrected one by SHAPOURI et al., which is too low. My estimate is in the middle.
- PIMENTEL's 2003 lime application rate is twice those of everyone else's. It reflects the 1997 USDA average.
- PIMENTEL's 2003 transportation energy is higher than everyone else's. It may reflect 1-2 more commutes/ha/crop.
- SHAPOURI and WANG et al. have underestimated the fossil energy in seeds and left out the machinery and infrastructure.
- I have added my seed energy estimate to their inputs, but left the machinery out.
- BERTHIAUME et al. (2001) have not included several of the energy inputs in corn farming, so their estimate is presented here only for comparison.
- The estimates of fossil energy inputs range from 19 GJ/ha (WANG) to 33 GJ/ha (PIMENTEL). My estimate is 28 GJ/ha.
- WANG et al.'s estimates are consistently too low. SHAPOURI et al.'s and my estimates are almost identical. The only significant difference is the inclusion of machinery into my estimates.
- The fossil energy use in corn farming is large and equivalent to 0.4 (WANG et al.) to 0.7 (PIMENTEL) metric tonne of gasoline per hectare and per crop.
- The average energy use in corn farming does not tell the whole story because of the very large variability of energy use by state, depending on the local weather conditions.
- All estimates, including mine, have had errors and/or omissions at one stage or another. I hope that by bringing the approaches of all authors into a common framework, most of the deficiencies of the prior analyses have been eliminated.

3.11 Solar Energy Input into Corn Production

The amount of solar energy that irradiates 1 average hectare of corn field in the U.S. during the growth season is gigantic, and it dwarfs all fossil energy inputs *and* the calorific value of the 8600 kg of corn grain harvested from this hectare, see **Figure 14**. In fact, during the 120-day growth season, roughly only $\sim 0.7\%$ of the solar energy is converted by corn plants into biomass (Biermann et al., 1999). In contrast, solar cells can collect sunlight all year long, see **Appendix C**. On an annual basis, the solar efficiency of corn plants drops by a factor of 3, i.e., only $\sim 0.2\%$ of the solar energy is captured by an average corn crop¹⁵. In summary, the solar energy does *not* limit corn

¹⁵Two tenth of one percent is 20 parts in 10000 parts of mean insolation. Roughly half of these 20 parts becomes corn grain.

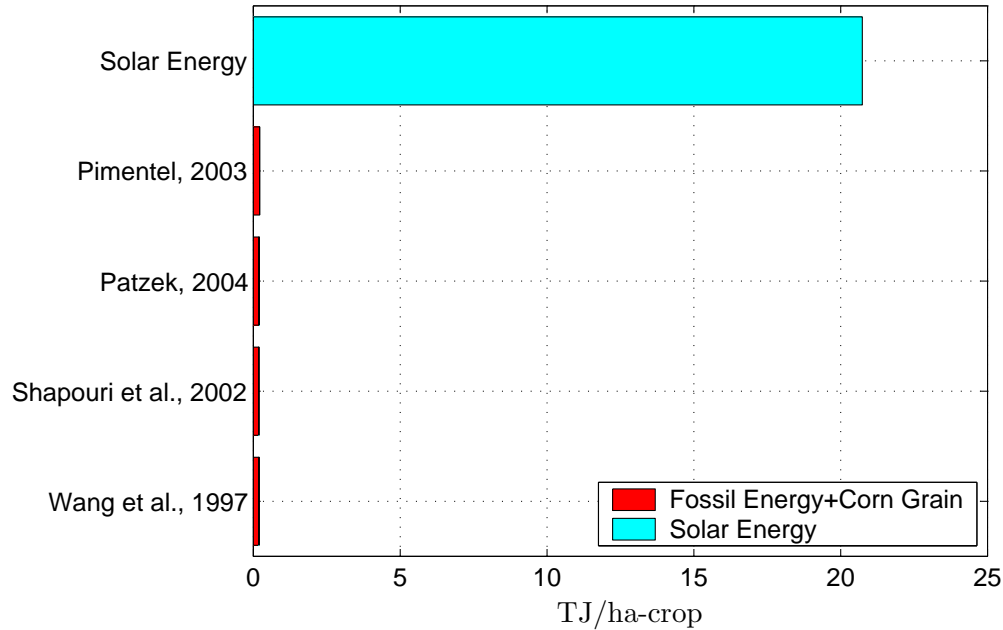


Figure 14: Comparison of the cumulative solar energy input and the fossil+grain energy inputs per hectare and per crop (120 days).

production – soil, water, and the dissolved nutrients do. I will discuss this crucially important point later. Here it suffices to say that if the sun were the limiting factor, the Sahara desert would be the best place to grow corn (Georgescu-Roegen, 1971).

3.12 Soil Humus and Micro-Element Depletion by Corn Production

Because good soil quality sustains farming, one needs to account for the extraction by corn plants of soil humus and metals.

Definition 2 Humus is the soil component which develops over time by the decomposition of organic matter. □

Soil humus is among the principal carbon pools on continents. One hectare of corn produces 8600 kg of corn grain, cf. Section 2, and 8600 kg of stalk, leaves and roots (Pimentel, 2004a). By scaling MICHAUD's (1995) results to the average U.S. corn yield, I have estimated that about 2600 kg/ha-crop of soil humus is extracted by the corn monoculture grown on tilled soil. Similarly, I have calculated that decomposing **all** 8600 kg of plant parts returns about 2100 kg/ha of humus (4:1 ratio). Therefore, with the full decomposition of plant leftovers, corn depletes ~500 kg/ha-crop of humus.

The net humus depletion of 500 kg/ha-crop could be replenished by returning to the field 2000 kg/ha of corn grain leftovers from ethanol production. Otherwise, industrial corn farming is also *soil mining*, in addition to being crude oil, natural gas, and coal mining. I will come back to this point when I discuss the energy credits for ethanol production assumed by WANG et al. and SHAPOURI et al.

Based on the information in TROEH & THOMPSON (1993), Prof. PIMENTEL (2004b) has calculated that only 1630 kg of humus would be added to the soil from the 8600 kg/ha of corn residues.

On the other hand, he points out that with appropriate field practices¹⁶ and *continuous use* of synthetic fertilizers, it is possible to keep soil humus at a constant value. The 20-year study of conventional corn and soybean production system in Rodale, Pennsylvania, following Penn. State recommendations, confirms that the soil humus has remained relatively constant at about 1.7% of soil carbon over most of the 20 years. At Rodale, the application rate of synthetic nitrogen fertilizer has been 140 kg/ha; phosphorus and potassium have been applied too¹⁷.

STANLEY A. BARBER (1977) estimated that harvesting about 9000 kg/ha of corn grain removes ~150 kg/ha of N and ~30 kg/ha of P, see **Figure 15**. Currently farmers are applying ~150 kg/ha of N and ~50kg/ha of P. The additional ~9500 kg/ha of corn stover contains ~80 kg/ha of N and ~10 kg/ha of P. Back in the 1930's - 1940's, with corn yields below 2000 kg/ha, only 10–20 kg/ha of N was harvested with the grain. However, nitrogen was also being lost by leaching, denitrification, and erosion. With soil erosion rates of 20 to 30 t/ha/yr, from 50 to 90 kg/ha of nitrogen were being lost by erosion alone. Since then, soil erosion rates in the U.S. corn fields have declined to 10 to 15 t/ha/yr. Soil erosion is deadly when it comes to nutrient and water loss (Larson, 1979; Lindstrom et al., 1979).

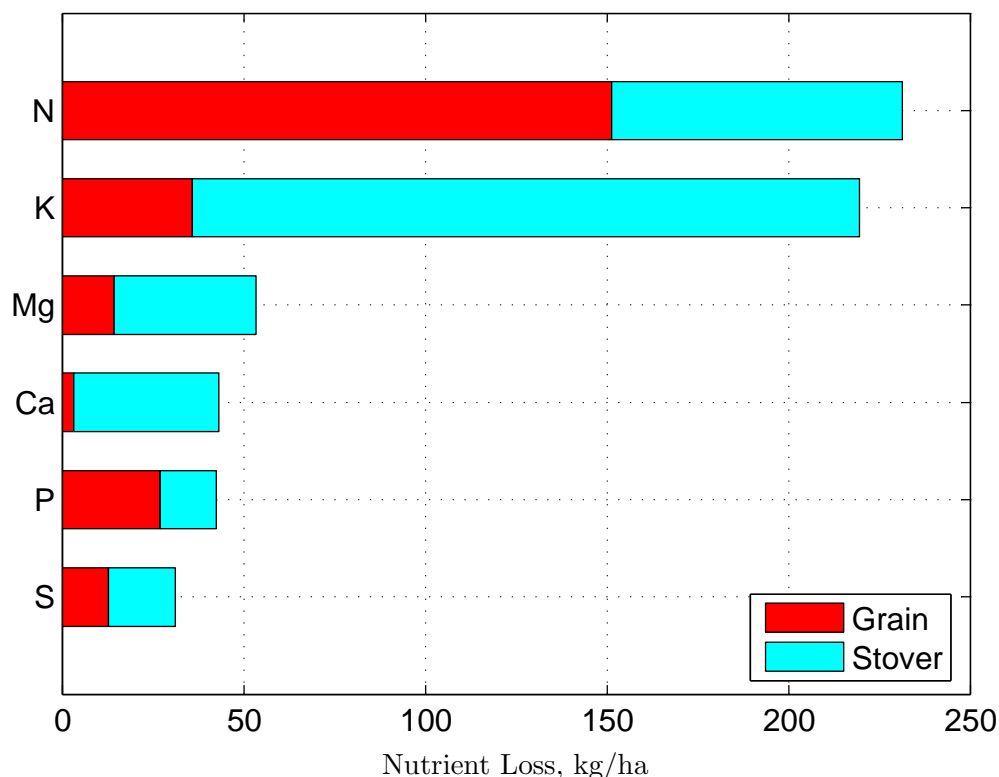


Figure 15: Soil nutrient losses with corn grain and stover removal. From BARBER (1977), scaled to 9 t/ha of grain and 9.75 t/ha of stover.

All corn plant parts, other than the starch in corn grain, should be decomposed and recycled to recover their N, P, K, C, Ca, Mg, Zn, B, Cu, Mn, etc., and diminish the degree of unsustainability of corn farming. However, most of the nitrogen and phosphorus and some other nutrients are

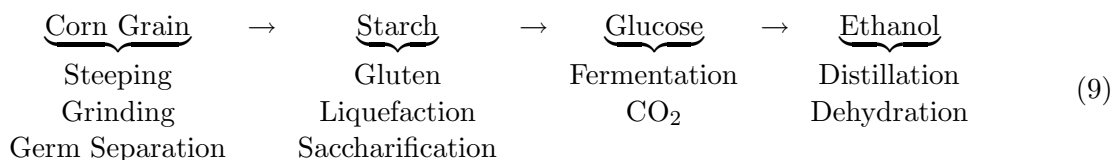
¹⁶Good field practices might include: (1) Soybean/corn crop rotation to naturally bind nitrogen and limit synthetic fertilizer use; (2) Corn stover decomposition in the field to conserve soil nutrients and limit erosion; and (3) Moderate soil tilling, or no tilling, to diminish loss of soil carbon and erosion.

¹⁷In Part II, I will show that the use of synthetic fertilizers makes farming *unsustainable*.

translocated from vegetative plant parts to the developing grain later in the season. According to Figure 15, a corn crop harvested with no recycling removes more than 1.5 times as much nitrogen, 1.6 as much phosphorus, 4 times as much potassium, 13 times as much calcium and 6 times as much magnesium as when this crop is harvested for grain. Other estimates are even more unfavorable (Wheaton et al., 1993). Whole plant harvest also removes most of the soil metals essential to the well-being of corn plants. The need to recycle plant parts and limit soil erosion largely negates the now fashionable attempts to produce ethanol from whole corn plants by harvesting everything from the corn field, see e.g., (NREL, 2002; Sheehan et al., 2004). Every ecosystem on the earth is highly optimized to recycle almost all mass it generates; otherwise life would not persist.

4 Major Energy Inputs to Ethanol Production

Conversion of corn grain into 100% ethanol (EtOH) is a fossil energy-intensive process, which also generates significant gas emissions, as well as liquid and solid waste. Here I will consider only wet-milling of corn to convert it into glucose, which is subsequently fermented to industrial beer, and distilled to 96% ethanol. The final water removal is achieved in molecular sieves that exclude water, or by distillation with benzene, see Eq. (9). Fermentation is a slightly exothermic catalytic burning of aqueous glucose, in which 49% of its mass is converted to carbon dioxide gas. The main liquid reaction product, ethanol, retains most of the free energy of the glucose. Dry milling is energetically similar, and need not be considered.



4.1 Corn Mass Balance Revisited

In Section 2, I calculated the theoretical efficiency of corn conversion into ethanol, in which every step is 100% efficient. Here, in agreement with the USDA estimate (Shapouri et al., 2002b), I will assume that the conversion of corn grain into 100% ethanol incurs 8.5% losses¹⁸ by mass, see **Figure 16**.

Two important conclusions can be drawn from Figure 16:

1. The average yield of anhydrous ethanol from corn is now 0.435 L EtOH/kg dry corn grain, or 2.914 gallons of EtOH per 56 lbs of dry corn grain (“dry bushel”), or 2.477 gallons of EtOH per nominal wet bushel with 15% moisture.
2. The reported field corn yields must be multiplied by 0.85 to convert the harvested corn to water-free or “dry” corn, see **Figure 17**.
3. In the literature, the USDA estimate of 2.682 gallons EtOH/bushel has been multiplied by the moist corn grain yields; this is *incorrect* and leads to an overestimation of the corn-ethanol yield by 8% ($\sim 1/4$ of the positive fossil energy balance claimed by USDA).

¹⁸This 8.5% overall loss lumps the losses in broken corn kernels and foreign matter (nominally 3% by mass for No. 2 yellow corn), in starch separation and hydrolysis, fermentation, distillation/rectification, and ethanol transportation and distribution. The fermentation process has several byproducts: n-propyl, isobutyl, amyl, iso-amyl, 1,2,3-propanetriol (glycerol) and higher alcohols; acetic aldehyde and acid; etc., see (White and Johnson, 2003), page 710; also see **Appendix D** for more details. The fermentation selectivity to ethanol can be less than 90%.

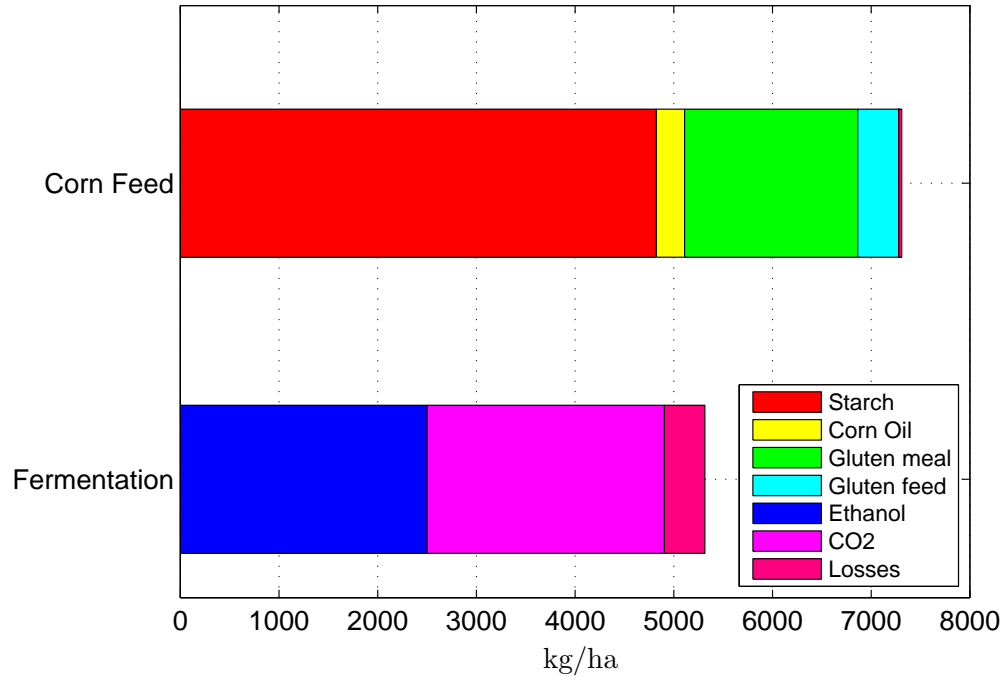


Figure 16: The result of practical corn conversion into ethanol with 8.5% losses is 0.435 L EtOH/kg dry corn grain = 2.48 gal EtOH/**wet** bushel with 15% moisture. Note that the dry starch is swollen by a factor of 180/162 caused by hydrolysis to glucose.

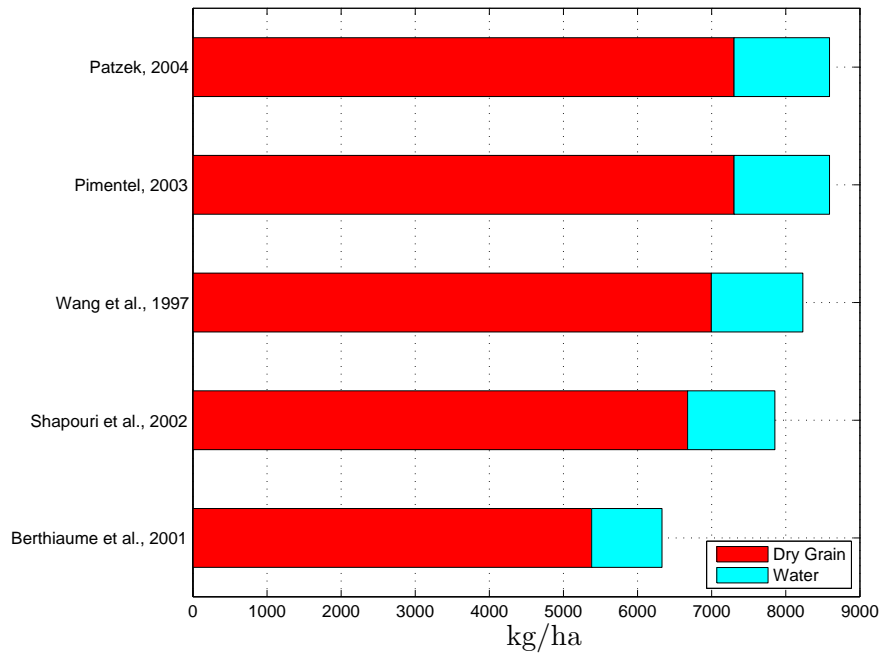


Figure 17: Average wet and dry corn yields.

4.2 Transport in Ethanol Refineries

Transport of materials and people in-and-out of an ethanol plant requires energy, and there is some disagreement between PIMENTEL and PATZEK on one hand, and SHAPOURI et al. and WANG et al. on the other. Here it suffices to state that

- Corn grain (8600 kg/ha) and fuel (e.g., ~1200 kg/ha of coal) must be transported *in*.
- Ethanol (~2200 kg/ha), gluten meal and feed (~2600 kg/ha) must be transported *out*
- Workers must travel *in-and-out*.
- Both SHAPOURI et al. and WANG et al. seem to underestimate these transport costs by a factor of 3-4.

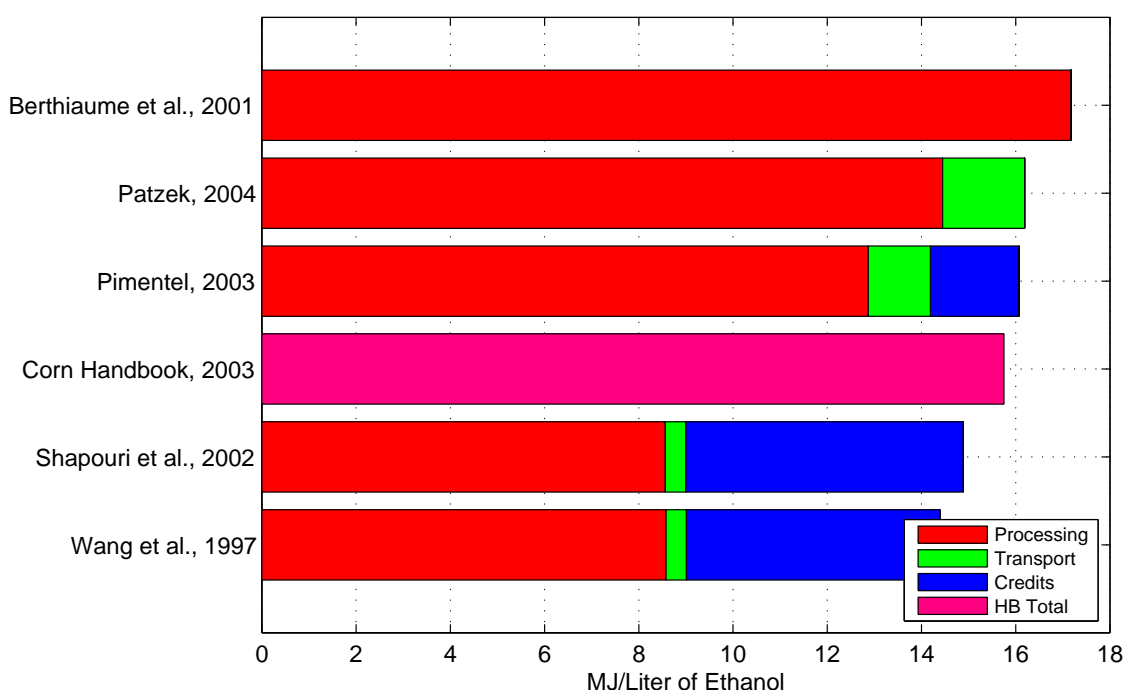


Figure 18: The average fossil energy inputs to ethanol production in a wet milling plant. The length of each bar is the total energy outlay to produce 1 liter of EtOH, and the blue parts denote the size of energy credits assumed by the different authors. The modern dry mill plants use 11.36 MJ/L as steam and 3.12 MJ/L as electricity, 14.5 MJ/L total, not counting transportation costs.

4.3 Fossil Energy Inputs to Ethanol

Because transportation is but a small fraction of the total energy outlay in ethanol production, there is little disagreement in the various estimates of the total energy used to produce ethanol from corn, which are all close to 15 MJ/L EtOH, see **Table 18**. This is easily seen when the total lengths of the bars in **Figure 18** are compared¹⁹.

¹⁹For comparison, a recent feasibility study for a new ethanol plant (International, 2001) projects 13.08 MJ/L EtOH in methane, and 1.675 MJ/L EtOH in electricity, for the total of 14.8 MJ/L EtOH, excluding transport and

4.4 Energy Credits

Major disagreements surface when it comes to energy credits used by the different authors to offset the high energy cost of ethanol production. The idea is to somehow use the market or energy value of gluten meal and gluten feed, see Table 1, both by-products of wet milled corn, to offset the fossil energy used to produce ethanol. Gluten meal is a more valuable by-product, but it is 4 times less abundant than gluten feed. The key assumptions made by the different authors are as follows:

- WANG et al. (1997) assume an energy credit of 30% of all energy inputs into ethanol production.
- SHAPOURI et al. (2002a) assume an energy credit of 5.9 MJ/L of ethanol, justifying it as the replacement value of soybean meal with corn gluten feed and meal.
- PIMENTEL (2003; 2004a), assumes an energy credit of 1.9 MJ/L of ethanol, justifying it as the replacement value of distiller's dried grain in *dry* milling, which is similar in composition to soybean meal (it contains 30% protein, 8% fat). Pimentel also argues that gluten feed does not replace soybean meal; in addition soybean plants fix nitrogen, and corn does not, so the replacement is really difficult to justify.
- I give ethanol **zero energy credit**, and want the ethanol refineries to bear the transportation and disposal costs of gluten feed and meal, as well as all other solid and liquid waste from ethanol production. Some of the environmental restoration costs will be included in the discussion of the carbon and water cycles in ethanol production from corn. In Section 3.12, I have already argued that all of the ethanol processing leftovers should be returned to the field to replenish soil humus and micro-elements.

4.5 Overall Energy Balance of the Corn-Ethanol Process

With the energy credits or without, ethanol production from corn is a fossil energy losing proposition, even if the energy costs of environmental damage are neglected, see **Figures 19** and **20**. In addition, the net energy²⁰ gained from corn production is *small*, see **Figure 21**; several times lower than those of perennial grasses and sugarcane (Rogner, 2000). As I have demonstrated above, corn grain is not the sun's gift to the producers of corn ethanol, but it is one of the fossil energy inputs. The calorific value of corn grain is therefore shown in Figure 19 to demonstrate that a large fraction of the energy inputs into ethanol production is dissipated on fermentation, distillation, and farming. The horizontal lines in the middle of each bar in Figure 19 denote the corrected ethanol yield, cf. Section 4.1, given the corn yields shown in Figure 17.

Remark 5 It appears that the high net energy yields of corn production – up to 200 GJ/ha-crop – reported in the literature, e.g., (Rogner, 2000), are unrealistic, and result from an incomplete thermodynamic analysis of industrial corn production. □

Non-scientists should stop looking for shortcuts around the hard work of learning the science

— CHRISTOPHER ESSEX and ROSS MCKITRICK
*Taken by Storm, The Troubled Science,
 Policy and Politics of Global Warming* (BPR Publishers, 2002)

commute costs. This study seems to contain some mistakes. For example, the efficiency of grain conversion into ethanol is overestimated, and the amount of feed water is significantly underestimated.

²⁰The energy of dry corn grain minus the fossil energy inputs per hectare and per crop.

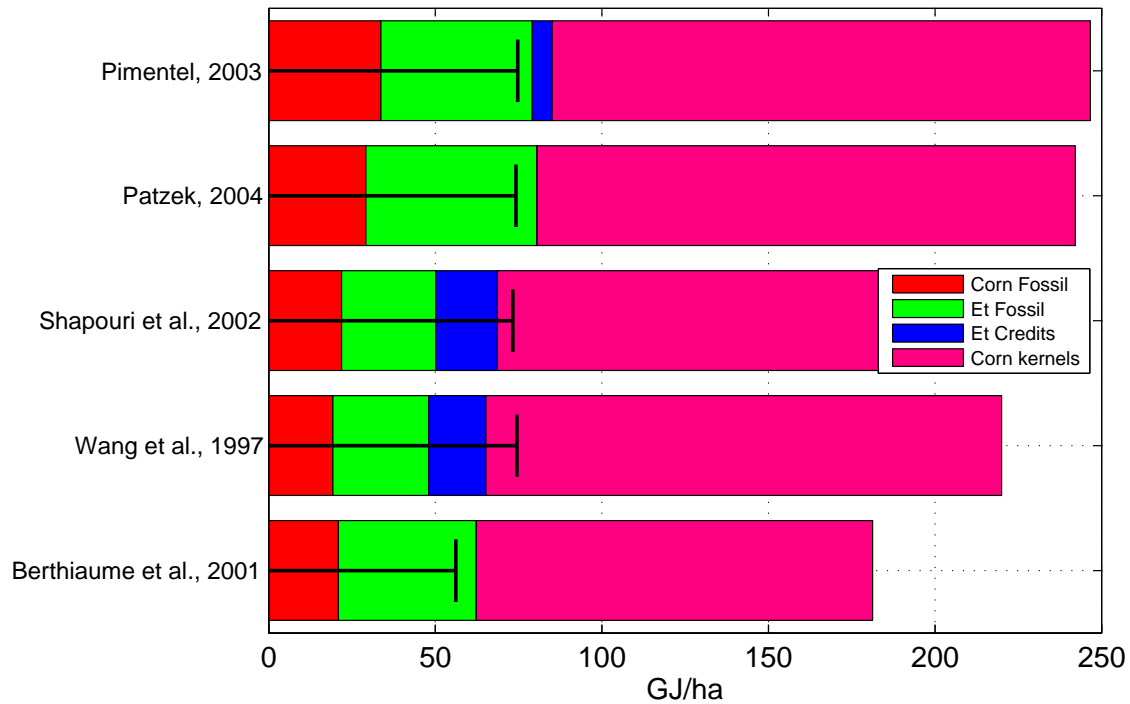


Figure 19: The overall energy balance of ethanol production. The two or three leftmost parts of each bar represent the specific fossil energy used in corn farming and ethanol production. The fossil energy inputs into ethanol production are the sum of the green part and the blue energy credit part for some authors. The rightmost part is the calorific value of corn grain harvested from 1 hectare. The total lengths of the horizontal bars represent all energy inputs into ethanol production. The horizontal lines with the vertical anchors represent the calorific value of ethanol obtained from one hectare of corn. Note that the total energy inputs into ethanol production are equivalent to ~ 4 – 5 metric tonnes of gasoline per hectare. The ethanol's calorific value is equal to 1 – 1.3 metric tonnes of gasoline.

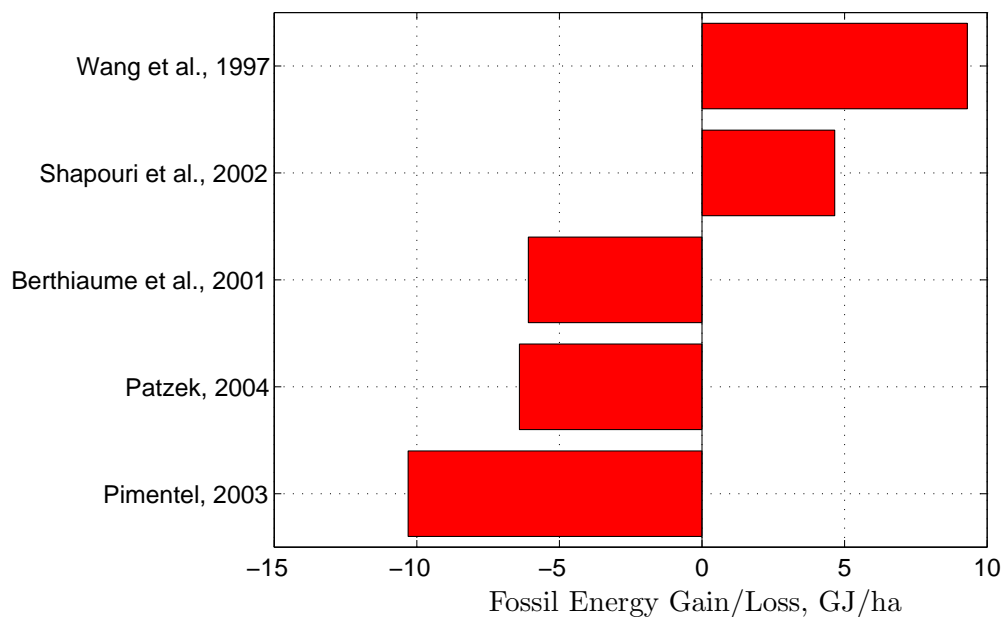


Figure 20: Fossil energy gain/loss in corn ethanol production. Note that the dubious energy credits described in Section 4.4 do not eliminate the use of fossil fuels in the first place, but present alternative useful outcomes of this use.

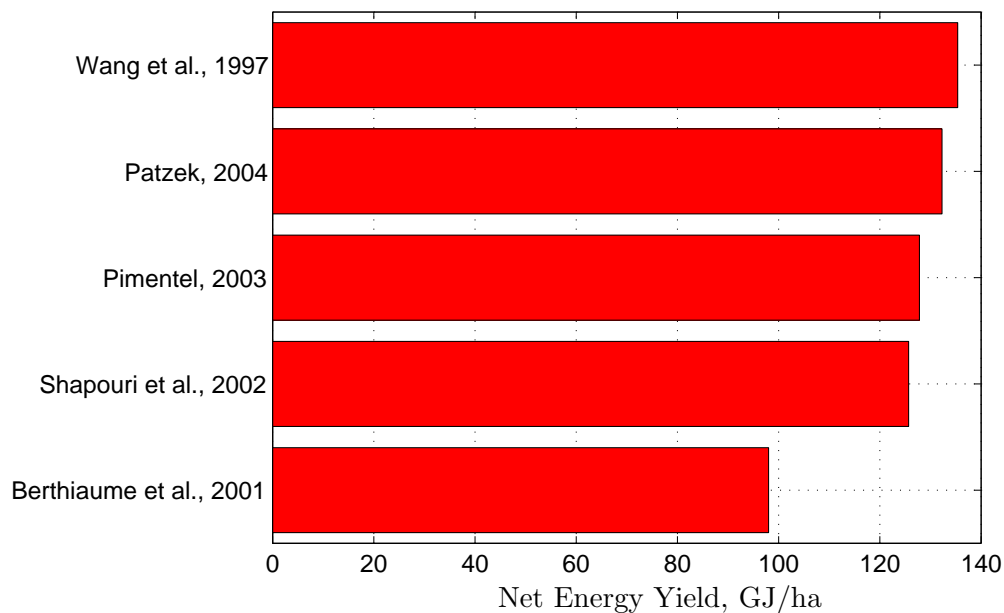


Figure 21: The net energy yield in industrial corn grain production is relatively small, 100 – 135 GJ/ha-crop. The HHV of dry corn grain is 18.8 MJ/kg, based on the mean of the values reported by SCHNEIDER & SPRAQUE (1955), p. 496, 2033 kcal/lb; and MILLER (1958), p. 639, 2059 kcal/lb. 1 thermochemical kcal = 4.184 kJ.

Part II

Sustainability & Renewability

1 Introduction

The following type of reasoning (Sheehan et al., 2004) is not uncommon in environmental literature:

(Page 118: ...Sustainability is fundamentally an ethical issue, the technological context... is not adequate to fully assess the sustainability of ethanol or any other fuel choices. ...The stakeholders²¹ established a list of indicators that they felt should be used to measure the relative sustainability of switching from gasoline to [corn] stover-derived ethanol to fuel our cars²².

More broadly, an informal check of Amazon.com, performed on August 16, 2003, revealed 4454 book titles containing the word “sustain★.” In particular, there were 573 book titles with “sustain★” and “★culture”²³. The phrase *sustainable development* is firmly rooted in our consciousness. Therefore, one must ask the following question: Is sustainable *anything* possible in nature? In the economy? Also, how *sustained* are the processes deemed by some as “sustainable”?

Human nature, being what it is, destines us to choose a “truly great but brief, not a long and dull, career²⁴” on the earth. After our eventual demise, the earth will be home to other less ambitious and impatient species. The name of the game, therefore, is to make the human presence on the earth as happy as possible, albeit not too short²⁵. These two tasks require careful thought and delicate balance of human actions. No country has demonstrated an adequate implementation of either. In fact the opposite may be true. As the entropy on the earth increases, the actions of governments and societies resemble more and more episodic spasms, with ever less forethought and deliberation. The current hot button issues: the Hydrogen Economy, Ethanol from Corn, and the War on Terrorism are good examples.

2 Disclaimer

The next eight sections of this paper are punctuated with verbatim quotations from, and my digestion of the most important book I have read in decades: *The Entropy and the Economic Process* by NICHOLAS GEORGESCU-ROEGEN (1906-1994), who was a great twentieth century thinker, economist, mathematician, historian, and philosopher. The impact of this book on my thinking has been profound.

²¹ “A group of farmers, environmentalists, automakers, grain processors, and government researchers.”

²² Therefore, any fuel or technology can be declared as “sustainable,” whenever there exists a group of people who feel good about it, and say that it is!

²³ My favorite: *Permaculture: Principles and Pathways Beyond Sustainability* by David Holmgren, published by Holmgren Design Services (July 2003).

²⁴ GEORGESCU-ROEGEN (1971), page 304.

²⁵ Preaching alone will not do. People will never choose less fulfilling life styles without coercion. This is why communism, or any other totalitarian “ism,” can never work; they strive to convert the thinking individuals into slaves or working animals. Says PERCY WILLIAMS BRIDGMAN, (1955), p.114, italics mine: “The *individual* is the unit in terms of which all our social concepts ultimately find their meanings.”

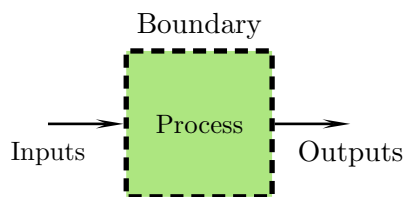


Figure 22: The *boundary* separates the process from the environment at any time (it is the *interface*), and it defines the *duration* of the process. We *may not* describe a process by what happens inside or outside of it, but only by what crosses its boundary.

3 Preliminaries

In science we divide actuality into two slices: one representing the *partial process* determined by our interest, and the second, its *environment*, see **Figure 22**. These two parts are separated by an *analytic*²⁶ *boundary*. The boundary has two attributes. The first separates the process from the environment at any time (we can call this attribute the *interface*, or the *frontier*), and the second defines the *duration* of the process²⁷. Often the terms *process* and *environment* are used interchangeably with *system* and *surroundings*. We may not describe a process by what happens inside or outside of it, but only by what crosses its boundary. Anything of interest crossing the boundary from the environment into the process is an *input*, and anything crossing the boundary in the opposite direction is an *output*. Solar energy is a typical example of only an input for any terrestrial process. The various materials abbreviated as “waste” are examples of only outputs.

4 Laws of Thermodynamics

The three empirical laws of *Classical Thermodynamics*²⁸ were originated by JOULE, CLAUDIUS, THOMSON, PLANCK and NERNST, and are often formulated as follows:

First Law or the Energy Conservation Law (Joule, Clausius, Thomson)

- Energy can neither be created or destroyed;
- The energy of the universe remains constant; or
- You can't win.

Second Law or the Entropy Law (Clausius)

- Without the compensating changes elsewhere, heat can flow only from a hotter to a colder body; or
- With passing chronological time, the entropy of the universe tends towards a maximum; or
- You can't break even.

²⁶The word *analytic* means well-defined mathematically in space and/or time.

²⁷The process is not defined outside its time interval.

²⁸Started in 1824 with a memoir, *Réflexions sur la puissance motrice du feu et sur les machines propres à développer cette puissance*, on the efficiency of steam engine by a French engineer, SADI CARNOT (1943).

Third Law (Nernst and Planck)

- The entropy of any condensed substance, i.e., liquid or solid, has at zero absolute temperature the value zero; or
- Zero absolute temperature cannot be reached; or
- You have to stay in the game.

The story of *Classical Thermodynamics* is rather simple if we ignore the fine print. Energy comes in two qualities: (1) *free* or *available*, and (2) *bound* or *dissipated*. Free energy can be transformed into mechanical work²⁹. Like heat, free energy dissipates itself, without any loss, into bound energy. The material universe, therefore, changes spontaneously in such a way that free energy is degraded. The final outcome is a state where all energy is dissipated, the *Heat Death* as it was called in the earliest thermodynamics³⁰.

For some technical reasons, which need not bother us now, *entropy* was defined by the formula:

$$\Delta S = \delta Q/T \quad (10)$$

where ΔS is the entropy *increment*, δQ is the *quantity* of heat transferred from a hotter to a colder body, and T is the absolute temperature at which the transfer is made. The entropy increments are always determined in the direction of *Time*³¹, from the *earlier* to the *later* moment in Time. So, if chronological Time τ_2 is later than another Time τ_1 , then the entropy of the universe (or another closed system) is

$$S(\tau_2) > S(\tau_1) \quad (11)$$

The *Entropy Law* is the simplest (and thus far the only) law known to science, by which the existence of true *happening* in nature is recognized. The Entropy Law defines the boundaries of what *cannot* happen, but does not prescribe entirely what *can*. This loophole resulted in life. All life³² feeds on an environment of low-to-moderate entropy (highly organized ecosystems³³, such as shallow sea water, a river or lake, a jungle, etc.), and is extinguished when the entropy of the environment becomes too high (the degraded environment of sterile desert or chemically polluted water).

²⁹Initially *free heat* was defined as the heat which can be exchanged between a hotter body and a colder one, and which can move a turbine in power station. If the hotter body, i.e., a steam boiler, remains unchanged, and the colder body, i.e., the environment, warms up, the amount of free heat goes down and the power station produces less electricity, as happened in Europe during the record-breaking summer heat wave of 2003.

³⁰BRIDGMAN disagrees, (1955), p. 265.

³¹Time is a notion of extraordinary complexity. Here I use the historical or chronological *Time*, as opposed to the mechanical clock *time*.

³²Not only biological life, but also social life. Consider the following comment made by JOSEPH SAMAHA in the daily *Al Safir* of Beirut about the desperately dissipative, high-entropy political situation in the Middle East: "Let us expel every mediator. Let us banish every international organization. Let the situation collapse. Let electricity and water be cut off. Let the pumping of oil stop. Let theft prevail. Let the universities and schools close down. Let businesses go bankrupt. Let civilian life break down." Source NYT, August 21, 2003, page A13.

³³*eco* – From late Latin *oeco*- household, from Greek *oik*-, *oiko*-, from *oikos* house habitat or environment (Webster, 1993). *Ecosystems* are the earth's households, and *ecology* is the study of these households.

5 Thermodynamics and Economics

Classical thermodynamics and *economics*³⁴ are thoroughly bonded. In fact, thermodynamics has been mostly a physical theory of economic value from its inception by CARNOT. Just as physical life, our whole economic life *feeds on low entropy*, e.g., on grain, lumber, steel, copper, aluminum, cloth, computers, etc., all of which are highly organized structures. A slab of electrolytic aluminum is worth much more than the dispersed molecules of aluminum oxide, diffused so much as to be of no use to us.

In 1860 or so, WILLIAM STANLEY JEVONS described the phenomenon of resource degradation and dilution as follows: “The expression “exhaustion of our coal mines,” states the subject in the briefest form, but is sure to convey erroneous notions to those who do not reflect upon the long series of changes in our industrial condition which must result from the gradual deepening of our coal mines and the increased price of fuel. Many persons perhaps entertain a vague notion that some day our coal seams will be found emptied to the bottom, and swept clean like a coal-cellar. Our fires and furnaces, they think, will then be suddenly extinguished, and cold and darkness will be left to reign over a depopulated country. It is almost needless to say, however, that our mines are literally inexhaustible. We *cannot get* to the bottom of them; and though we may some day have to pay dear for fuel, it will never be *positively wanting*³⁵.” Today we may substitute “crude oil” or “natural gas” for “coal”, and JEVONS’ statement will be as true³⁶.

In particular, thermodynamics explains to us (while economics usually does not³⁷) why *land* has economic *value*, as opposed to *price*. Even though land cannot be consumed, it derives its economic value from two physical facts: (1) land allows humans to catch the most vital form of low entropy – sunlight – and (2) the earth stopped making land in large quantities. Other physical and chemical characteristics of fertile land³⁸ are *scarce* in a different sense: (1) the amount of low entropy in the soil is finite and it decreases³⁹ *continuously* and *irrevocably*, and (2) a given amount of low entropy can be used *only once*.

A different example of irrevocably lost low entropy is given by mining copper ore and transforming it into pure metallic copper. A sheet of manufactured copper metal is made of: (1) copper ore, (2) other raw and processing materials, (3) electricity from fossil fuels, and (4) mechanical work. All these factors ultimately resolve into an orderly structure of primary materials (e.g., highly concentrated copper atoms in the ore, fuel, electricity), i.e., to *environmental low entropy* and nothing else. The free energy used in production to deliver mechanical work, or to heat the ore, is *irrevocably lost*.

We cannot *bootleg* any entropy⁴⁰ by means of an ingenious process or device. Just like a

³⁴A social science concerned chiefly with description and analysis of the production, distribution, and consumption of goods and services; from Greek *oikonomica* – skilled in the management of a household (Webster, 1993).

³⁵(Jevons, 1866), *Preface*, pp. vi-vii, my italics. JEVONS could not possibly imagine that coal, or any other fossil fuel, could be *wanting* at the rate we need to produce it in the 21st century, *regardless* of the price.

³⁶In fact, it is more difficult to extract crude oil than coal. After we finish exploiting an oilfield, 1/2 to 2/3 of the oil present initially in the rock interstices is left there forever. Also, producing conventional crude oil and natural gas at the incredibly high rates required by the world economy will soon be *physically* impossible, regardless of their prices. Therefore substitution of crude oil by natural gas will be impossible. As pointed out by Dr. RONNEBERG, economists simply do not comprehend this looming calamity.

³⁷With the notable exception of GEORGESCU-ROEGEN (1971).

³⁸Natural soil fertilizers, soil-bonding humus, interstitial water, etc.

³⁹Unless the soil is a part of an ecosystem that recycles all mass, see Section 10.

⁴⁰In the 1930’s, the young physicists became so confused by statistical thermodynamics that a famous thermodynamicist, PERCY WILLIAMS BRIDGMAN (1955), wrote an essay on the impossibility of constructing a machine that “shall violate the second law of thermodynamics on a scale large enough to be commercially profitable (pp. 236-268).”

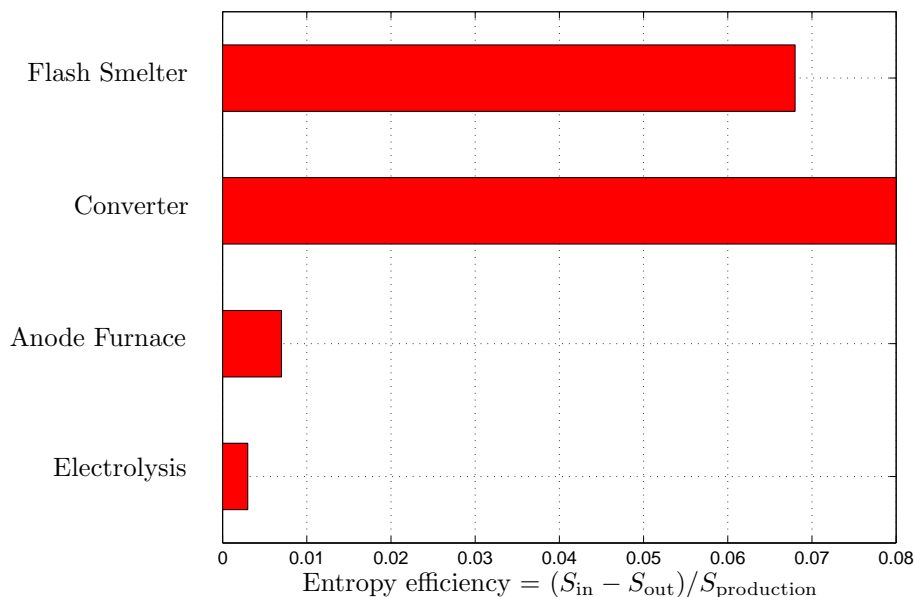


Figure 23: The Second Law efficiency of copper production is incredibly small. The steps in sorting the copper atoms are: Ore Concentration 2% \rightarrow 30% Cu (not shown); Smelter 30% \rightarrow 63% Cu; Converter 63% \rightarrow 99.1% Cu; Anode furnace 99.1% \rightarrow 99.55% Cu; and Electrolysis 99.55% \rightarrow 99.99% Cu. Source: STEFAN GÖSSLING, *Entropy Production as a Measure for Resource Use*, University of Hamburg, 2001.

MAXWELL demon⁴¹, we have merely sorted the copper atoms from all others, but in order to achieve this end-result we have used up irrevocably a greater amount of low entropy than the difference between the entropy of the copper metal and that of the copper ore. In view of **Figure 23**, it would be a great mistake to compare just the latter two entropies and exclaim: “Lo! Man has created low entropy from high!” (Georgescu-Roegen, 1971). This claim, in effect, is made by all those who say that copper can be manufactured *sustainably*. The copper scrap recycling programs are successful, only because scrap copper (and aluminum) consumes less free energy than any other way of reconstituting metallic copper⁴². Nevertheless, insofar as fossil energy is used, by recycling we only postpone the inevitable exhaustion of low entropy in the environment.

6 Economic Activity

Economists have a tendency to view the economic process as a *closed system*, while ignoring the *continuous inflow* of low entropy from the environment⁴³. From a physical point of view, the economic process is entropic; it neither consumes nor creates mass or energy, but only transforms low entropy to high. To make things worse, the parallel entropy generation process in the environment is spontaneous, and goes on by itself without human intervention.

⁴¹J. CLERK MAXWELL (2001) imagined a tiny demon posted near a microscopic swinging door separating two gases *A*, and *B* of equal temperature. The demon is instructed to open and close the door “so as to pass only the swifter molecules from *A* to *B*, and only the slower molecules from *B* to *A*.” Clearly this demon can make the gas in *B* hotter and in *A* cooler. Therefore, Maxwell’s demon creates low entropy – or does he?

⁴²According to STEFAN GÖSSLING, entropy generation per ton of copper produced from ore is 52 MJ/K, and only 12 MJ/K for copper produced from scrap; for reference, see the caption of Figure 23.

⁴³GEORGESCU-ROGEN, (Georgescu-Roegen, 1971), Chapters IX and X.

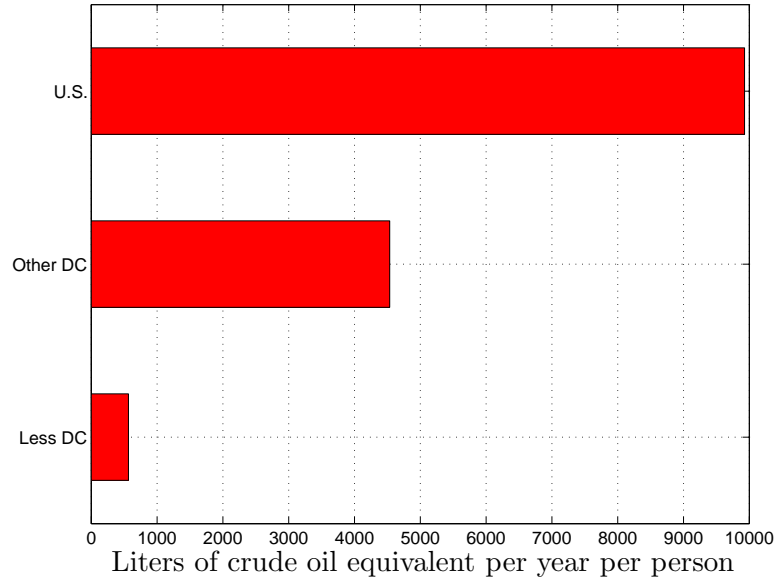


Figure 24: The 2001 per capita energy consumption in the U.S., other Developed Countries (DC), and the less Developed Countries. Source: The U.S. DOE Energy Information Agency.

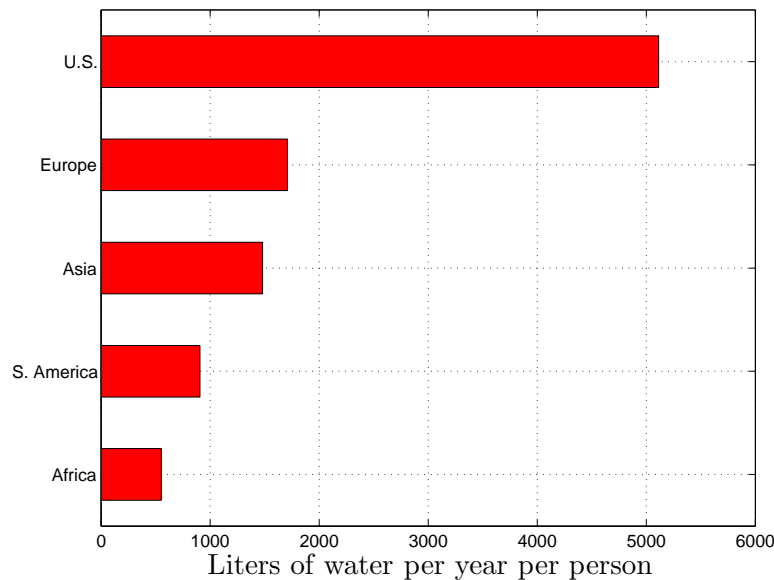


Figure 25: The 1990 per capita total (personal + industrial) water consumption in the U.S., and elsewhere. Source: Water Quality Association, 151 Naperville Road Lisle, IL 60532-1088, USA.

The material production process, in contrast, depends on the intervention of humans, who like the MAXWELL demon, sort and direct environmental entropy according to the process rules⁴⁴. This sorting activity is not a part of natural environmental processes and creates high entropy, i.e., waste, at a (much) faster rate than the biological life processes. *From a purely material point of view, the economic process always transforms low entropy into waste.* So what could be the justification for

⁴⁴When watching an SUV commercial, I often see a monstrous truck carelessly damaging a low-entropy fragile ecosystem, a pristine meadow, river bed, or an alpine mountain slope. Thus the SUV commercials are a good metaphor for the interactions of the present-day economics with the environment.

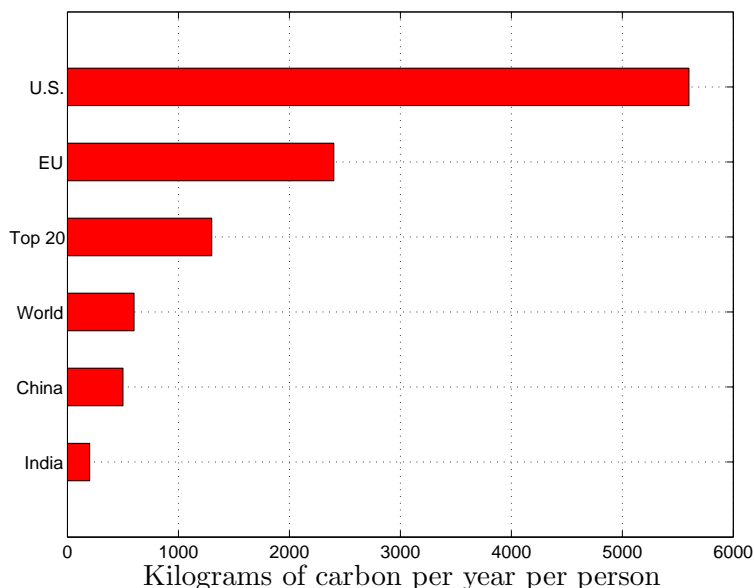


Figure 26: The 1999 per capita carbon emission estimates in the U.S., and elsewhere. Source: World Resources Institute, U.S. EIA.

economic activity? As described by GEORGESCU-ROEGEN, the true output of an economic process is not merely waste, but the *enjoyment of life*. It is not a coincidence that the very country, which on July 4, 1776 declared: “We hold these truths to be self-evident, that all men are created equal, that they are endowed by their Creator with certain unalienable Rights, that among these are Life, Liberty and the *pursuit of Happiness*,” uses over twice as much free energy per capita than any other country on the earth, see **Figures 24- 26**. In general, abundant free energy equals enjoyment of life. The converse statement is as true in real life as it is in Logic, see **Figure 27**.

7 Agriculture

The following statement is made surprisingly often: “Properly used, [the plants on the earth] can by their reproductive powers supply us indefinitely with the food, the wood, and the other natural products we require.⁴⁵” Even though the intensity of sunlight that reaches the earth has not changed appreciably over the human scale of Time, the apparent dominance of solar energy in agricultural production should not obscure the importance of the *entropic soil degradation* by continuous cultivation. Soil degradation can be severe over a human life span. Even the earliest farmers knew that manuring a soil does not remove its degradation, and to farm always meant to mine, in part, the soil. Clean water is necessary for agriculture. Water is inevitably polluted by agricultural waste; therefore, water too is mined.

It would be a mistake to believe that the practice of fertilizing soil can defeat the Entropy Law and transform food production into an everlasting process. Life feeds not only on sunlight but also on the low entropy of an *ecosystem*⁴⁶.

With time, draft animals, oxen, buffalo and horse, were replaced by machines. A tractor is made of iron, other metals, oil and coal, and it feeds on oil. The natural manure fertilizer from farm

⁴⁵(Cépède et al., 1964), p. 309.

⁴⁶A process that cycles living organisms. Only the solar energy and waste heat flow across the ecosystem boundary, everything else is recycled, cf. Section 10.



Figure 27: After the power blackout, people poured onto 8th Avenue, outside the Port Authority, unable to leave the city. The New York Times, August 14, 2003.

animals was replaced by the chemical fertilizers manufactured from methane, coal, oil, iron, copper, and the earth minerals. The importance of this switch should be self-evident: the main source of low entropy feeding agriculture has been switched from the practically *infinite* solar energy to the very *finite stock of minerals* in the earth's crust. Of course, this switch limits how long and how many people can be fed by the earth. It is no longer the practically unlimited stock of the energy in the sun that limits our survival, but the meager stock of natural minerals mined from the earth crust. If M is the *accessible fraction*⁴⁷ of this stock, and r is the average rate of its use, then $M = rt$, where t is the corresponding duration of human civilization that depends on the crustal minerals. Depending on the mineral, and its rate of use, this time can be estimated as several tens of years (high quality crude oil), through many hundreds of years (coal), to thousands of years (uranium); so much for sustainable development.

8 Industrial Production

Every car or appliance produced today means fewer human lives in the future. Every styrofoam cup, cell phone or TV set in a landfill now means fewer resources and happy humans later. In industry, just as in agriculture, the price of technological progress has been a shift from the abundant source of low entropy – the sun –, to the earth's mineral resources. Any time these resources are wasted, we shorten the survival time of humanity. Confronted in the distant future with exhaustion of mineral resources, mankind will be tempted to retrace its steps; however, in view of the Entropy Law, it is impossible. Human evolution is *irrevocable* and *cannot be retraced*.

⁴⁷This fraction may be very small. For example, the vast majority of carbon on the earth is bound in the carbonate rocks. No one in their right mind would dream of extracting this carbon. See also footnote 36.

Because of the Entropy Law, the large-scale industrial production of fossil fuels (e.g., ethanol) from the industrial plants *manufactured* (Kimbrell, 2003) by modern agriculture, will only hasten the depletion of mineral resources and the eventual demise of our civilization.

9 Waste

The ever-increasing rate of economic activity generates ever more chemical waste. For the earth as whole, this waste cannot be disposed of⁴⁸. Unless we use more free energy to process it, toxic waste once produced is here to stay. This free energy can only be used once, and will always diminish our future standard of living. Waste cleanup, collection and recycling have begun to interfere with our life and pockets⁴⁹, and are no longer hidden from our industrial civilization. The ever-accumulating garbage and the toxic by-products of agriculture and industry are the living proof of the Entropy Law in action.

The process of concentrating copper atoms, whose entropy efficiency is shown in Figure 23, is analogous to the process of fishing out the contaminant molecules dispersed in a huge volume of water and rock. The entropy efficiency of any contaminant cleanup process must be lower than that of the copper production process.

10 Sustainability

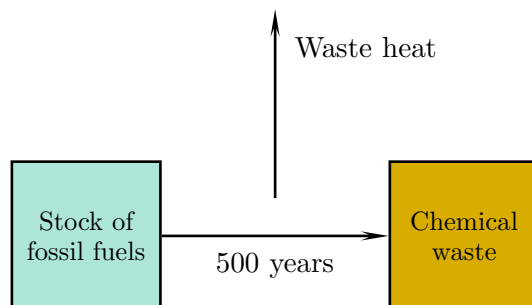


Figure 28: A linear process of converting a stock of fossil fuels into waste matter and heat cannot be sustainable. The waste heat is exported to the universe, but the chemical waste accumulates. To replenish some of the fossil fuel stock, it will take another 50-400 million years of photosynthesis, burial, and entrapment.

The opposite of as yet undefined *sustainability* is *irreversibility*, and I define it first after MAX KARL ERNST LUDWIG PLANCK (1858-1947):

Definition 3 (Irreversibility) A process is *irreversible* if it can in *no* way be reversed, all other processes are *reversible*. It is impossible, even with the assistance of all agents in nature, to restore

⁴⁸Therefore hiding our waste by injecting it into the oceans, aquifers, or burying it in shallow graves - landfills - will always come back to haunt us.

⁴⁹NYT, August 20, 2003: “Crews have completed the removal of more than 12 metric tons of weapons-grade plutonium from the Rocky Flats (CO) nuclear weapons site as part of a \$7 billion cleanup effort that is expected to be finished in 2006, 12 *years ahead of schedule*. The plant manufactured plutonium triggers for 40 years until it was closed in 1989 for safety violations. The plutonium has been shipped to South Carolina; lower-grade nuclear waste will be sent to New Mexico. The 6,000-acre Rocky Flats site is expected to become a wildlife refuge.” In other words, the cost of picking and sorting the plutonium waste is \$583/g, compared with \$11/g to buy gold.

everywhere the exact initial state when the irreversible process has once taken place (Planck, 1926). \square

Corollary 1 *From the definition above, a linear process that converts the low entropy of fossil fuels into waste is irreversible and cannot be sustainable.* \square

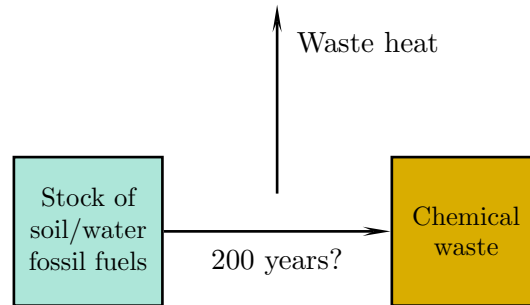


Figure 29: Current industrial agriculture is another example of a linear process, which by definition cannot be sustainable.

In a linear process, see **Figure 28** and **29**, a finite stock of fossil fuels is rapidly depleted and burned to serve as a collective heat source for all heat engines employed by our civilization, see **Figure 31a**. In addition, the atmosphere, which acts as a heat sink, becomes polluted by chemical waste from combustion (chemical entropy), as well as by waste heat (thermal entropy). The earth can only export thermal entropy through its atmosphere, see **Appendix A**. In addition to the atmosphere, the earth, which is the system in **Appendix B**, also accumulates chemical entropy. As a result, the linear fossil fuel process accumulates chemical entropy in the earth and the atmosphere, and irreversibly degrades our planet on a time scale of our civilization, measured in hundreds of years.

In contrast, a *cyclic* ecosystem can be sustainable, see **Figure 30**. A natural cycle uses the sun as its source of energy and low entropy, and it expels only waste heat into the atmosphere and, ultimately, into the universe, see **Figure 31b**. Most importantly, *all materials* involved in an ecosystem are *recycled*, and when the natural cycle is completed, only waste heat, or thermal entropy is generated.

In order to discuss the existence and constraints on *sustainability*, I need first to define it.

Definition 4 (Sustainability) A *cyclic* process is *sustainable* if and only if

1. It is capable of being *sustained*, i.e., maintained without interruption, weakening or loss of quality “forever,” and
2. The environment on which this process feeds and to which it expels its waste is also sustained “forever.”

Corollary 2 *A cyclic process, which is also “sustainable,” must not reject chemicals into the environment, i.e., its net mass production must be “close” to zero “forever.”*

Corollary 3 *A sustainable cyclic process must not reject heat into the environment at a rate that is too high for the earth to export this heat to the universe; otherwise, the environment properties will change.*

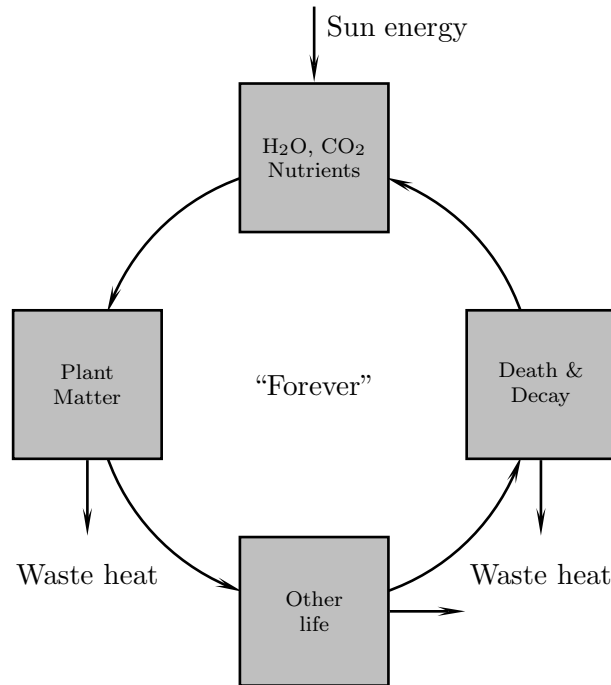


Figure 30: An ecosystem transforms the sun energy (low thermal entropy) into waste heat (high thermal entropy). The waste heat is continuously exported to the universe. *Everything* else is completely reused, or recycled.

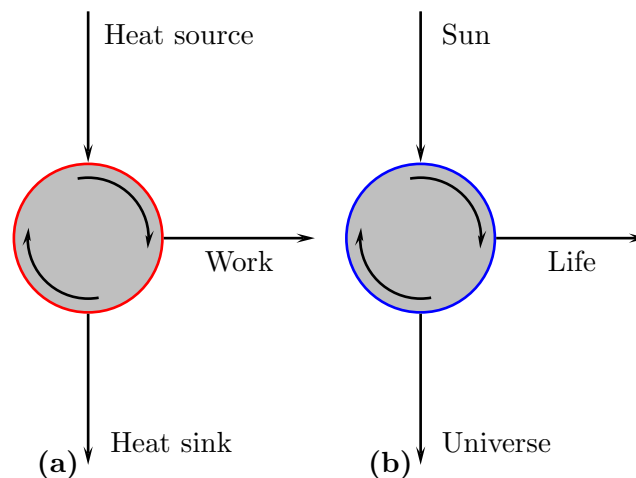


Figure 31: Thermodynamic cycles: (a) A heat engine, and (b) An ecosystem.

10.1 The Earth is an Open System to Heat Flow

Attributes (1)-(2) of a sustainable cyclic process would be a *thermodynamic contradiction* if the earth were approximately a closed system with respect to the infrared radiation (heat). These two attributes would then make a sustainable process a *perpetual machine of the second kind*. Luckily for us, the earth can be treated as an open system with respect to visible and infrared light, and a

sustainable cyclic process may generate thermal entropy at a rate per unit area of the earth surface (*specific entropy rate* or *flux*) which is no more than the average flux of entropy export from the earth to the universe, j_S^E , calculated from Eq. (34) in **Appendix A**, minus the specific rate of entropy generation in the atmosphere due to export of the solar energy, calculated from Eq. (39) in **Appendix A**.

To quantify sustainability, I first assert that a cyclic process always converts all forms of entropy to thermal entropy. Thermal entropy is the ultimate waste from all “sustainable” cyclic processes on the earth, and it should be used for comparisons. Second, per unit area of the earth, we know that (1) the *always positive* specific rate of thermal entropy generation due to everything happening on the earth is $\sigma > 0$, (2) the rate of increase of the specific thermal entropy of the atmosphere due to all these happenings is $\sigma_a > 0$, and (3) the specific rate of thermal entropy generation due to the energy transport from the earth to the universe is $\sigma_t > 0$. Then, for cyclic processes, *strong sustainability* can be defined mathematically, see **Appendix B**, as

$$\sigma + \sigma_a \leq j_S^E - \sigma_t \quad (12)$$

at every point on the earth, and at all times.

Over an arbitrary time interval $[\tau_1, \tau_2]$, we can write the global condition of sustainability of all cyclic processes on the earth as (Eq. (55) in Appendix B)

$$\underbrace{[S_a(\tau_2) - S_a(\tau_1)]}_{\text{Increase of atmospheric entropy}} + \underbrace{[S(\tau_2) - S(\tau_1)]}_{\text{Increase of earth entropy}} - S^E(\tau_1, \tau_2) + S_t(\tau_1, \tau_2) \leq 0 \quad (13)$$

where $S^E(\tau_1, \tau_2)$ is the total thermal entropy exported by the earth over the time interval $[\tau_1, \tau_2]$, and S_t is the corresponding thermal entropy generation in the atmosphere due to the solar energy export.

10.2 Conclusions

The immediate observations from the above discussion are:

- To the extent that humans use 80-90% of fossil and nuclear energy to run the heat engines that power the global economy, our civilization is 80-90% unsustainable.
- If the atmosphere dissipates more energy due to the increased greenhouse gas loading by human (and natural) activities (σ_t increases), there will be less room for all other human activities to remain sustainable. STAHL estimates, (1996), Table 1, the specific entropy generation rate in the atmosphere to be $\sigma_t = 0.2 \text{ W/K-m}^2$. So only 1 W/K-m^2 of thermal entropy generation is left to all human and other natural activities, see **Appendix A**.
- Only energy generation directly from the sun, sun-driven wind, and water can be sustainable.
- Burning or extracting large quantities of wood or green matter requires chemical fertilization and cannot be sustainable to the extent that growing plants mines fossil fuels as well as low entropy from the soil, see Section 7.
- Industrial agriculture can never be sustainable because it relies on the irreversible burning and chemical transformations of fossil fuels, see Section 7.

*Thanks to the human heart by which we live,
Thanks to its tenderness, its joys, and fears,
To me the meanest flower that blows can give
Thoughts that do often lie too deep for tears.*

— WILLIAM WORDSWORTH
Intimations of Immortality from Recollections of Early Childhood (1803-1806)

*In thermodynamics, you have to get to the point where you understand
what any damn fool means no matter what he says*

— PROF. A. R. GORDON
(*Lecture Notes*, Department of Chemistry, University of Toronto)

Part III

Sustainability of Corn-Ethanol Cycle

1 Introduction

As pointed out in Part II, our standard of life is maintained by the exploitation of natural resources that have accumulated in the earth over millions of years. A natural resource whose chemical composition *differs most* from the dead states of the elements comprising it is *most valuable*.

This part is devoted to the life-cycle analysis of industrial corn and the ethanol biofuel produced from it. In my analysis, I will move along the trail charted in the brilliant, albeit incomplete, paper by three Canadian scientists, RICHARD BERTHIAUME, CHRISTIAN BOUCHARD, and MARC A. ROSEN (Berthiaume et al., 2001). There will be, however, important differences. I define the industrial corn-ethanol system differently, include more inputs (which are more carefully estimated), and do not require the carbon and water cycles to close.

2 Available Free Energy

For example, relative to a *datum environment* ($T_0 = 25^0\text{C}$, $p_0 = 1\text{ atm}$), the quality of heat rejected by a process depends on temperature:

$$\begin{aligned} 1\text{J of heat at } 500^0\text{C} &= 0.614\text{ J of work} \\ 1\text{J of heat at } 50^0\text{C} &= 0.077\text{ J of work} \end{aligned} \tag{14}$$

GIBBS, GUOY, STODOLA, and KEENAN's available free energy, see **Appendix B**, is known in chemical and engineering thermodynamics as *exergy*, and is denoted with the symbol B . The concept of exergy and its cumulative consumption in an industrial process has been significantly advanced by the distinguished Polish thermodynamicist, JAN SZARGUT, a professor at the Chemical Engineering Department of my *Alma Mater*, the Silesian Technical University, Gliwice, Poland. His monograph, published in English with DAVID R. MORRIS and FRANK R. STEWARD in 1988 (Szargut et al., 1988) was based on two decades of development and industrial applications of life-cycle analysis. Today, this monograph is still the most comprehensive source of exergy concepts, data, and examples.

2.1 Introduction to Exergy

An easy to understand definition of exergy was proposed by LUDWIG RIEKERT (1975).

Definition 5 Exergy, B , is equal to the shaft work or electrical energy necessary to produce a material in its specified state from materials common in the environment in a reversible way, heat being exchanged with the environment at constant temperature T_0 . \square

We distinguish the *potential* exergy, B_p , *kinetic* exergy, B_k , *physical* exergy B_{ph} , and *chemical* exergy, B_{ch} :

- **Physical exergy**, B_{ph} is the work obtainable by a reversible physical process from its initial state (T, P) to the environment state (T_0, p_0) .

- **Chemical exergy**, B_{ch} , is the work obtained by taking a substance at the pressure and temperature of the environment to the state of thermodynamic equilibrium with the datum levels of components of the environment.
- **Thermal exergy**, B_{th} , is the sum of physical and chemical exergies:

$$B_{th} = B_{ph} + B_{ch} \quad (15)$$

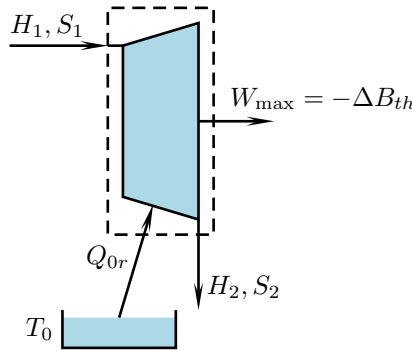


Figure 32: Exergy balance in an isothermal, ideal flow machine. The maximum possible shaft work from this machine is equal to the negative change of thermal exergy.

2.2 Change of B_{th} between Two States

Consider an ideal (reversible) flow machine, see **Figure 32**. An exergy carrier with enthalpy H_1 , and entropy S_1 enters the machine. After physical and/or chemical changes, the effluent has enthalpy H_2 , and entropy S_2 . Heat is transferred between the environment and the working fluid at the ambient temperature T_0 . The first and second law of thermodynamics are simply:

$$\begin{aligned} W_{\max} &= B_{th1} - B_{th2} = H_1 - H_2 + Q_{0r} \quad (\text{I Law}) \\ S_2 - S_1 - \frac{Q_{0r}}{T_0} &= 0 \quad (\text{II Law}) \\ B_{th1} - B_{th2} &= -\Delta B_{th} = H_1 - H_2 - T_0(S_1 - S_2) \quad (\text{I+II Law}) \end{aligned} \quad (16)$$

Physical exergy can be calculated immediately from Eq. (16)

$$B_{ph} = H - H_0 - T_0(S - S_0) = H_{ph} - T_0 S_{ph} \quad (17)$$

2.3 An Industrial Flow Process

Consider now an industrial steady-state flow process, which can occur in a heat engine, corn field, or ethanol plant, see **Figure 33**. The input to this irreversible process is an exergy carrier with the enthalpy H_1 , and entropy S_1 . The process is also supplied with the quantity of heat Q_1 from the source having temperature $T_1 > T_0$. The process effluent has enthalpy H_2 , and entropy S_2 . The rejected amount of heat Q_0 is transferred to the environment. The useful outcome of the process

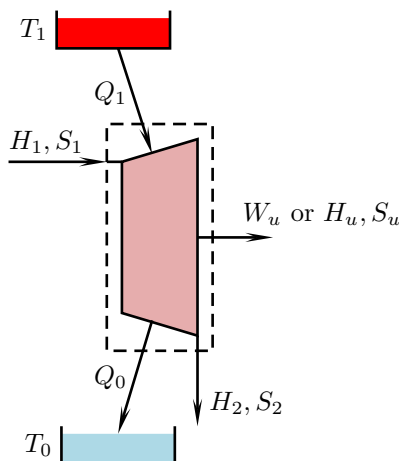


Figure 33: Exergy balance in ideal (reversible) and real (irreversible) nonisothermal industrial process.

can be mechanical work W_u or a chemical product having parameters H_u and S_u . The effect of irreversibility is studied by comparing the industrial process with a reversible process with the same inflow and outflow parameters, and the same amount of driving heat. The only difference between these two processes is the amount of heat rejected to the environment. For the reversible process this heat is Q_{0r} , and for the irreversible one, it is Q_0 .

The first and second law balances for the two processes are:

$$\begin{aligned} H_u &= H_1 - H_2 + Q_1 - Q_0 && \text{Real process} \\ H_{ur} &= H_1 - H_2 + Q_1 - Q_{0r} && \text{Reversible process} \\ H_{ur} - H_u &= Q_0 - Q_{0r} \end{aligned} \quad (18)$$

The increased useful effect of the reversible process causes the amount of rejected heat to be smaller than that in the industrial process, $Q_{0r} < Q_0$. The sum of all entropy increases in the industrial process is

$$\sum \Delta S = -\frac{Q_1}{T_1} + S_2 - S_1 + \frac{Q_0}{T_0} + S_u > 0 \quad (19)$$

while that in the reversible process is

$$\sum \Delta S_r = -\frac{Q_1}{T_1} + S_2 - S_1 + \frac{Q_{0r}}{T_0} + S_{ur} \equiv 0 \quad (20)$$

From Eqs. (19) and (20) it follows that

$$Q_0 - Q_{0r} = T_0(\sum \Delta S + S_{ur} - S_u) \quad (21)$$

and we obtain

$$\underbrace{(H_{ur} - T_0 S_{ur})}_{B_{th,r}} - \underbrace{(H_u - T_0 S_u)}_{B_{th}} \equiv \delta B = T_0 \sum \Delta S \quad (22)$$

The left hand-side of Eq. (22) represents the difference of the useful thermal exergy in the reversible and industrial process, δB . It therefore represents the *exergy loss* due to the irreversibility of the industrial process under consideration. We have recovered, again, the famous GUOY-STODOLA law, also derived in **Appendix B**.

2.4 Cumulative Exergy Consumption (CExC)

All steps of a production process leading from natural resources taken from the environment to the final product result in *exergy losses* or *exergy consumption*.

Definition 6 (CExC) The cumulative exergy consumption (CExC) is the sum of the exergy of all natural resources in all the steps of a production process. \square

The problem of cumulative *energy* consumption (CEnC), discussed in Part I, is better known, but calculation of CExC is more informative because it accounts for the exergy of non-energetic raw materials (soil, water, air, minerals) extracted from the environment.

3 The Ideal and Real Corn-Ethanol Cycle

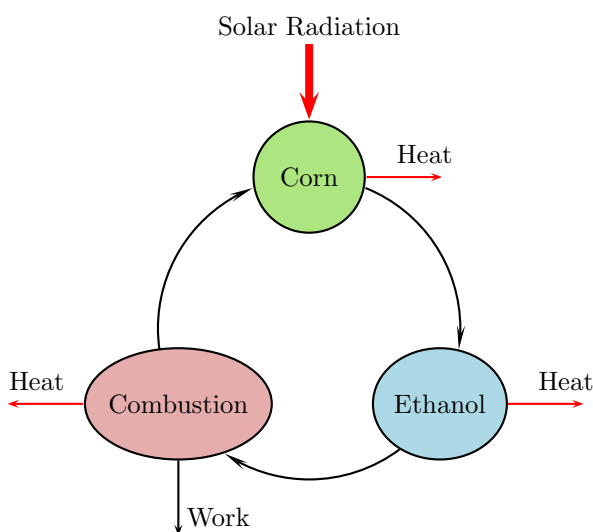


Figure 34: The ideal corn-ethanol cycle. Adapted from Figure 2 in BERTHIAUME et al. (2001).

Ideally, see **Figure 34**, the corn-ethanol system and cycle consist of three parts: (1) Sustainable corn farming, (2) Sustainable ethanol production, and (3) Ethanol combustion to produce useful work.

The cycle is driven *only* by solar energy, and all its chemical by-products are *fully* recycled. Only the low quality heat is rejected by the ideal corn-ethanol cycle into the environment, and this heat is exported through the atmosphere into the universe. All carbon dioxide is recycled, and so is all water. This low-rate ideal cycle *cannot* deliver the massive quantities of ethanol fuel from (bi)-annual corn crops, see **Figure 35**.

Remark 6 Between 1866 and 1939 (NASS, 2004b), the average yield of corn in the U.S. hovered around 26 ± 3 bushels per acre, or 1600 kg/ha, 1/5 of the average yield today. I will assume that 1600 kg/ha is the *almost* sustained corn yield using manuring, composting, crop rotation, and other not quite sustainable field practices. Note that between 1906 and 1937, the average corn yield declined, most likely due to the progressing soil deterioration⁵⁰, cf. Section 7. \square

⁵⁰PIMENTEL (2004c) observes: “Between 1900 and 1938, the early [U.S.] farmers were probably mining the soil of nutrients and soil erosion was quite severe. At that time, most farmers kept livestock and were applying manure to

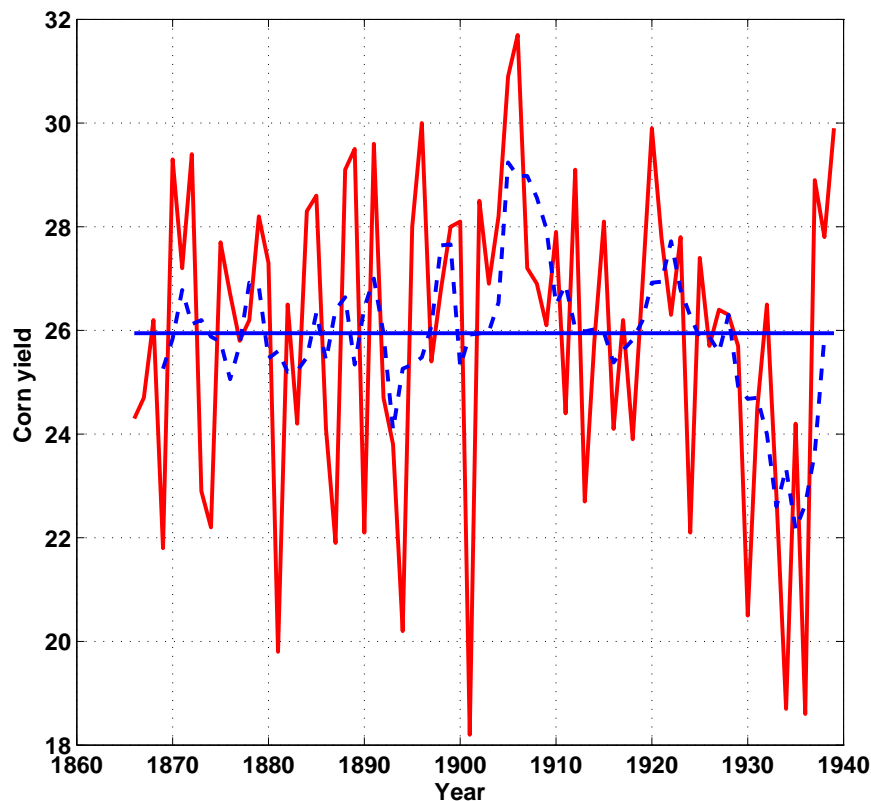


Figure 35: Between 1866 and 1939, the average yield of corn in the U.S. was 26 ± 3 bushels/acre, or 1600 kg of moist corn grain per hectare. The broken line is the 5-year moving central average of annual reports. Between 1906 and 1937, the average yield declined in general due to soil mining. Source: USDA (NASS, 2004b).

Corollary 4 *From Remark 6 it follows that without the fossil fuel-derived fertilizers, the corn produced today in the U.S. would require at 140 – 180 million hectares, close to the entire cropland area in the U.S (187 million hectares total, 122 million hectares harvested in 2002, U.S. Census). There would not be enough animals to manure the fields, and low entropy in the soil would be exhausted within some 30 years. The same conclusion follows (perhaps with a different time scale) for any other industrial source of biomass. In short, no biofuel derived from plants is sustainable. □*

Instead of relying on the current solar energy, we use the ancient solar energy in the form of fossil fuels to *accelerate* the ideal corn-ethanol cycle, see **Figure 36**. With the massive infusion of crude oil, natural gas, coal, and their products, as well as fertilizers, pesticides and herbicides, we greatly increase the corn yield and the rate of ethanol production, but we also produce massive amounts of chemical waste. In addition, we deplete soil by removing too much plant matter from the fields, and deplete groundwater whenever there is not enough rain. The industrial corn-ethanol cycle relies on the linear processes of mining fossil fuels, soil, water, and air, and in view of Part II, it *cannot*

the land. Most did not apply the manure until spring, when more than half of the nitrogen had escaped into the atmosphere. Most of the corn was grown in rotation. If the corn were grown after a legume crop, some nitrogen would be available to the corn crop. Although manure was being applied, it probably was applied in the neighborhood of only 5 tonnes of stored manure per hectare, and thus would have only about 12 to 15 kg/ha of nitrogen. For a corn yield of 1600 kg/ha the manure application only would provide a minimum amount of nitrogen.”

be sustainable. All published statements to the contrary, e.g., (Wang et al., 1997; Shapouri et al., 2002a; Deluga et al., 2004; Sheehan et al., 2004), and many others, are scientifically indefensible. Below, I will quantify just how unsustainable the industrial corn-ethanol cycle is.

4 System Boundary

As required by thermodynamics, see Figure 34, I define the *system* as the topsoil of the corn fields, the corn plants, the corn-processing ethanol refineries, and the ethanol-burning combustion engines. The system *inputs* are the manufactured hybrid corn seeds, solar energy, fossil fuels, field chemicals, earth minerals (muriate of potash, phosphates, calcinated lime, etc.), electricity, machines, air (oxygen and carbon dioxide), and the atmospheric, surface and subsurface water. The system *outputs* are heat, oxygen O_2 , carbon dioxide CO_2 , carbon monoxide CO , gaseous ammonia NH_3 , nitrogen oxides NO_x , sulfur oxides SO_x , methane CH_4 , water as liquid and vapor, and a multitude of organic and inorganic water, soil and air contaminants.

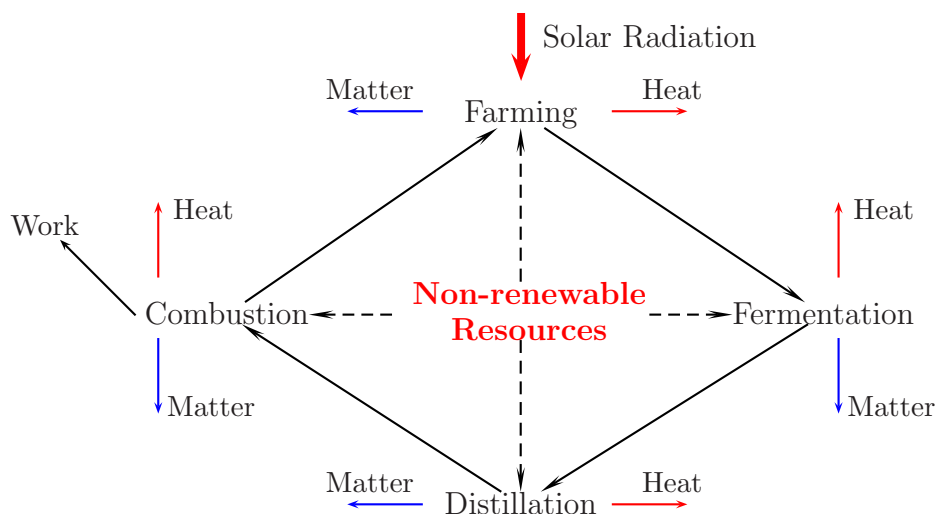


Figure 36: The industrial corn-ethanol cycle. The nonrenewable resources (NRR's) at the center of the cycle are: crude oil, methane, coal, electricity from fossil fuels, earth minerals, soil nutrients, groundwater, etc. Adapted from Figure 3 in BERTHIAUME et al. (2001).

5 The Carbon Cycle

The simplified carbon cycle *inside* the industrial corn-ethanol cycle is depicted in **Figure 37**. Consistently with the discussion in Part I, I require that all organic carbon be returned from the ethanol refineries to the fields. Therefore, the *internal* carbon cycle is *closed*. The corn plants bind CO_2 from the air, which is then released back into the atmosphere by burning the ethanol fuel produced by the cycle, and by decomposing the stalk, roots and leaves of corn plants, as well as all those parts of corn grain that were not used to produce the ethanol.

Remark 7 Most of the biomass must be returned to the field⁵¹ to preserve topsoil. This requirement puts severe restrictions of the various schemes of converting biomass (rice straw, corn straw, tree parts, whole plants, etc.) to biofuels. □

⁵¹Not necessarily the same field.

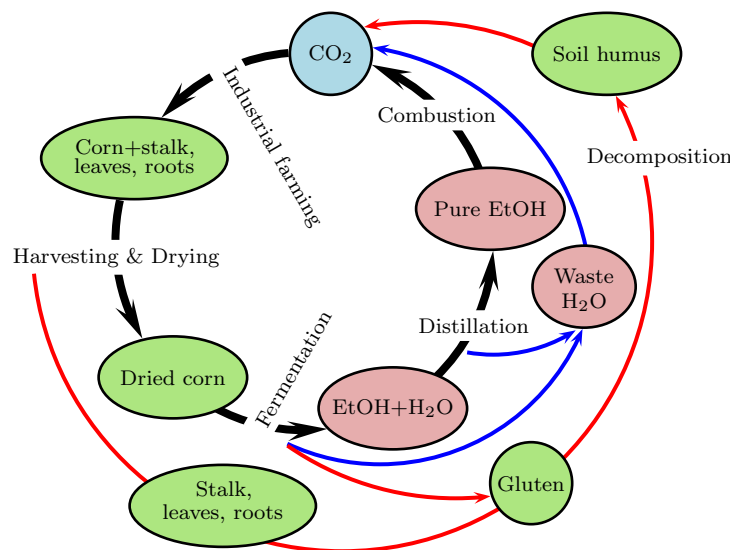


Figure 37: The internal carbon cycle in the industrial corn-ethanol cycle can be closed only by recycling most of the corn-plant and corn-grain components. Adapted from Figure 6 in BERTHIAUME et al. (2001).

5.1 Net CO₂ Emissions

As shown in Figure 36, our corn field-ethanol plant-combustion engine system uses fossil fuels as inputs, and outputs *their* combustion products into the environment. Therefore the industrial corn-ethanol cycle generates extra CO₂ and other greenhouse gases, which will all be translated into equivalent CO₂ for simplicity.

The question now is as follows: Does the industrial corn-ethanol cycle generate more equivalent CO₂ from its fossil fuel inputs than the gaseous emissions from replacing the cycle's ethanol with gasoline, methane or diesel fuel? To make this comparison fair, I will account for the cumulative exergy consumption in production of the fossil fuels by adding another 15% to their calorific values, in agreement with SZARGUT et al. (1988) and SHEEHAN et al. (1998).

By asking and answering this question, I seek to dispel common misconceptions about the industrially-manufactured biofuels, best summarized by the following quotation:

About 70 million barrels⁵² of ethanol are included in annual U.S. gasoline consumption after 1992. Because ethanol is a biofuel, the carbon it contains should not be counted as an emission. Hence, carbon from ethanol is *deducted* from transportation gasoline consumption⁵³.

EIA is right, but then the CO₂ emissions associated with the consumption of non-renewable resources in the industrial ethanol-corn cycle should be *added* to the transportation gasoline consumption. The question now is: What is the net balance?

To answer this question, I will use the EIA and the European Fertilizer Manufacturers' Association (EFMA) data on the specific carbon dioxide emissions from the fossil fuel inputs into the industrial corn-ethanol cycle, see **Table 19**. These specific emissions, in kg of CO₂ per MJ in a fossil fuel, will be multiplied by the respective energy input fluxes in MJ ha⁻¹ crop⁻¹, established

⁵²About 11 giga liters.

⁵³The U.S. Energy Information Agency (EIA) (2002), Appendix A, page A3, my italics.

in Part I. Electricity is treated differently, and its specific CO₂ emissions account for the average U.S. efficiency of conversion of thermal energy into electricity.

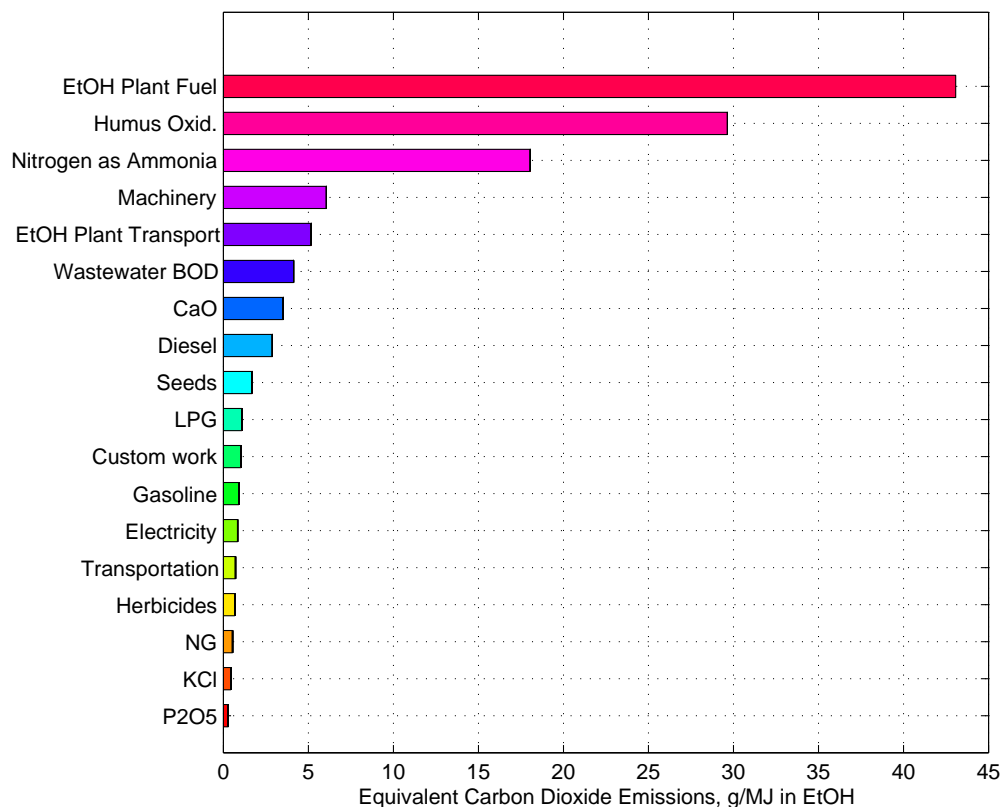


Figure 38: Equivalent CO₂ emissions from each major non-renewable resource consumed by the industrial Corn-EtOH cycle.

To convert the NO_x emissions from the industrial corn-ethanol cycle to the equivalent CO₂ emissions, I will follow the guidelines of EIA (2002), the Intergovernmental Panel on Climate Change (IPCC) (1997), and EFMA (Biermann et al., 1999):

- 1.25% of applied N fertilizer escapes into the air as N₂O.
- 30% of applied N escapes from the field, and 2.5% of that quantity is converted to N₂O in surface water.
- 10% of applied N escapes as NH₃ into the air, and 1% of that becomes N₂O.
- Nitrous oxide is 300 times more potent as a greenhouse gas (GHG) than CO₂.
- An average ammonia plant emits 0.03 kg N₂O/kg N in nitric acid, which is used to make ammonium nitrate.

The equivalent CO₂ emissions from the fertilized corn fields are then ~950 kg/ha. The equivalent CO₂ emissions from the production of ammonium nitrate are $\sim 150 \times 0.03 \times 300 \times 63/80 = \sim 1000$ kg/ha⁵⁴.

⁵⁴One may dispute this last number to the extent that ammonium nitrate is not used as fertilizer.

The CO₂ emissions resulting from electricity use in the removal of Biological Oxygen Demand (BOD) in wastewater from ethanol refineries, cf. Section 6, are also included.

In addition I have made the following assumptions:

- Emissions from humus oxidation in soil eroded by wind are included with the following assumptions: (a) topsoil contains 4% of humus (Stevenson, 1982), (b) humus contains 50% of C by weight (Stevenson, 1982), wind-eroded soil is enriched in humus by $(1.3 + 5)/2 \approx 3$ on average (Allison, 1973), (c) wind erodes only 1 mm of topsoil per year, or 10 tonnes/ha-yr (Pimentel, 2006), and (d) the airborne humus is completely oxidized.
- To make the comparisons meaningful, methane, gasoline, and diesel fuel are now charged with incremental 17% (2% more than before) emissions for their recovery, transportation, and processing.

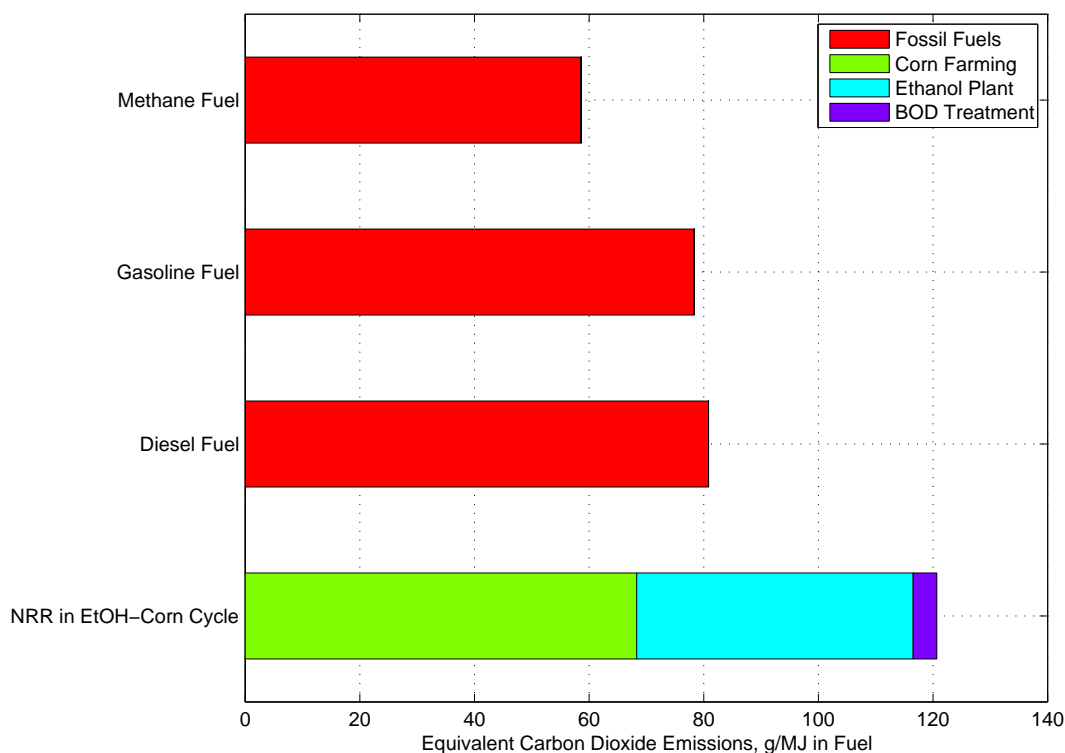


Figure 39: The total equivalent CO₂ emissions from the consumption of nonrenewable resources by the industrial corn-ethanol cycle. The CO₂ emissions from the energy-equivalent amounts of methane, gasoline and diesel fuel were increased by 17% to account for their recovery, transport, and refinement.

5.2 Conclusions

The results of my calculations, shown in **Figures 38** and **39**, lead to the following conclusions:

1. According to my estimates, 1 ha of industrial corn-for-ethanol generates 8955 kg of equivalent CO₂ from the fossil fuel inputs and humus oxidation.

2. If the amount of gasoline with the energy content of 117% of 2498 kg EtOH/ha obtained on average from corn were burned, it would generate 5817 kg of CO₂.
3. The net equivalent CO₂ emissions are then 8955 – 5817 ≈ 3100 kg/ha *above* those of gasoline.
4. Today, the industrial corn-ethanol cycle generates on 4.9 million hectares about 15.4 million metric tonnes of equivalent CO₂/year *over and above* the energy-equivalent gasoline.
5. To satisfy 10% of U.S. fuel consumption, the additional equivalent CO₂ emissions will be about 127 million metric tonnes per year.

6 Water Cycle

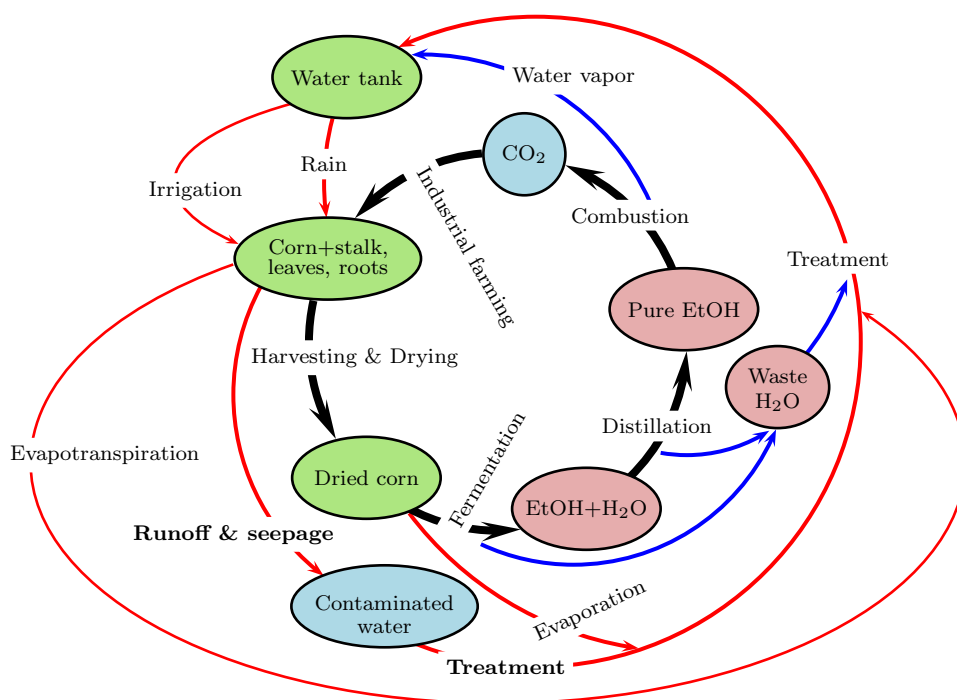


Figure 40: The water cycle in corn-ethanol production. Compared with the water requirement by corn plants, the feed water requirement of an ethanol-producing plant is small, and the link between the tank and the fermentation stage was not drawn. Adapted from Figure 7 in BERTHIAUME et al. (2001).

The water cycle can be idealized by bringing all necessary water to an imaginary “water tank” at the top of **Figure 40**. The water is then dispensed from this tank as rain, irrigation water, and ethanol plant feed water. Water is used by the corn plants, but it also evaporates from the fields, runs off as contaminated surface water, and seeps back into the groundwater, contaminating the underlying aquifers as well. Additional water is evaporated during drying of the harvested corn in NG/LPG driers.

To estimate the average water requirement by corn plants, I will use PIMENTEL’s (2003) average estimate of ~100 cm of water per crop⁵⁵. This translates into 10×10^6 L/ha-crop. About 15%

⁵⁵The rule of thumb is that corn needs 1000 – 1800 kilograms of water per kilogram of grain produced. This

of the water demand (USDA, 1998), 8.1×10^5 L/ha-crop, comes from pumping groundwater and surface water to irrigate corn. For comparison, the average water requirement by corn in Nebraska is 53-71 cm, and roughly 50% of corn acreage is irrigated there (Benham, 1998). Nebraska sits on top of the most prolific part of the giant High Plains aquifer (USGS, 2003; Rosenberg et al., 1999). Water mining from the High Plains aquifer continues in Texas, Kansas, Oklahoma, New Mexico and Colorado, while the water levels in Nebraska and Wyoming have stabilized or increased, see USGS (2003) and the references therein.

To estimate the water inputs into a wet-milling ethanol plant I will use the data published in the *Corn Chemistry and Technology Handbook* (White and Johnson, 2003) and by PIMENTEL (2003). According to WHITE and JOHNSON (2003), one needs the following amounts of process water per 1 liter of ethanol:

- 10-12 liters in corn fractionation (p. 450).
- 20-25 liters in glucose fermentation (p. 697).
- The total amount is 30-37 liters of clean process water per 1 liter of ethanol, or 38-46 liters of water per kilogram of ethanol.
- The average 5-day biochemical oxygen demand (BOD) of waste water is 1000-2000 mg/L.

According to PIMENTEL (2003), 159 liters of water are needed to produce 1 liter of 95% ethanol, which translates to 190 liters of water per kg of 100% ethanol. Twelve liters of waste-water are generated per 1 liter of ethanol from the separation of industrial beer with 8 percent of ethanol by volume. The wastewater BOD is 18000-37000 mg/L, 10-20 times higher than the previous estimate. This estimate makes sense because the beer wastewater contains dead yeast and unreacted glucose. After mixing with cleaner process water, the overall BOD must go down. BERTHIAUME et al. (2001), erroneously use 160 L of wastewater per L of EtOH, with the BOD of 20000 mg/L.

For the time being, I will go along with 46 L water/kg EtOH with the BOD⁵⁶ of 2000 mg/L. Thus, a wet milling ethanol plant needs on average

$$46 \frac{\text{L}}{\text{kg EtOH}} \times 2499 \frac{\text{kg EtOH}}{\text{ha-crop}} = 0.11 \times 10^6 \frac{\text{L}}{\text{ha-crop}}, \quad (23)$$

of water, and it generates 230 kg of BOD/ha-crop. This amount of process water is 100 times smaller than the amount of water needed to grow corn in the field.

Remark 8 The industrial corn-ethanol cycle needs about 10×10^6 liters of water per hectare and per crop. This amount of water must be deposited in the tank. Depending on the weather and location, some or most of this water may have to come from mining an aquifer or surface water. \square

Since the volume of process water passing through the ethanol plant is two orders of magnitude smaller than that passing through the field, and all of the process water is readily accessible in holding tanks, ponds, etc., problems with cleaning it up are *insignificant* when compared with the field runoff water.

rule translates into 86-150 cm of water requirement, see, e.g., http://www.jica.go.jp/english/publication/network/net_vol18/02.html.

⁵⁶A Google search yields BOD values between 600 and 9200 mg/L in the wastewater streams from wet- and dry milling ethanol refineries that use different biomass feeds.

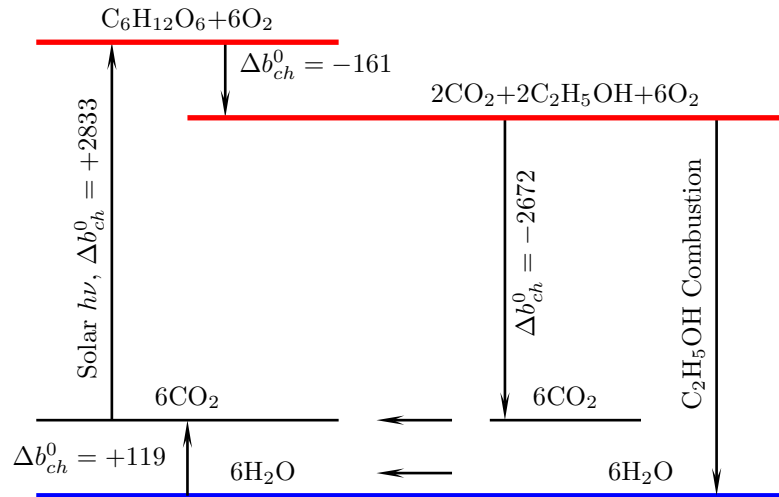


Figure 41: Exergy diagram of the ideal CO₂-Glucose-EtOH cycle.

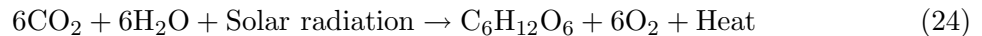
Remark 9 Industrial agriculture uses a huge land area, and it mines and contaminates huge amounts of soil, water, and air. The environmental damage it causes is much more widespread and more difficult to reign in than that from the highly-concentrated industrial sources. In addition, industrial agriculture invades and destroys large ecosystems. In other words, twenty-first century industrial agriculture poses a more acute threat to life on the earth, than the nineteenth century smoke stacks ever did⁵⁷. □

7 Exergy Analysis of the Ideal Corn Ethanol Cycle

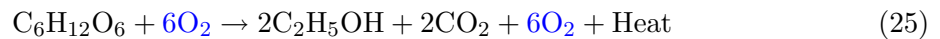
7.1 Chemistry of the CO₂-Glucose-EtOH Cycle

The ideal CO₂-Glucose-EtOH cycle consists of three steps:

Step 1 Photosynthesis of glucose from atmospheric CO₂:



Step 2 Production of ethanol from glucose:



Step 3 Combustion of ethanol:



The compounds that appear on both sides of the chemical reactions in Steps 2 and 3 do not participate in these reactions, but appear to close the cycle. The chemical exergies of all compounds are listed in **Table 20**. The chemical exergies of the products of each step of the cycle are listed in **Table 21**, and the exergy flow is depicted in **Figure 41**.

⁵⁷For an in-depth analysis of the deadly industrial agriculture, see KIMBRELL et al. (2002).

7.1.1 The Maximum Cycle Output per Unit Mass of Corn

To calculate the maximum possible energy output from the ideal CO₂-Glucose-EtOH cycle, I made the following assumptions:

- Dry corn is 66% glucose (100% hydrolyzed starch) by mass.
- Starch is converted into ~100% ethanol with 0% losses.
- Corn delivered to an ethanol plant is 15% water.
- The net chemical exergy of an ideal ethanol cycle per kg of moist corn grain is

$$15.74 \frac{\text{MJ}}{\text{kg glucose}} \times 0.66 \times \frac{180 \text{ kg glucose}}{162 \text{ kg dry corn}} \times 0.85 \frac{\text{kg dry corn}}{\text{kg wet corn}} = 9.81 \frac{\text{MJ}}{\text{kg wet corn}} \quad (27)$$

- At ~8600 kg of moist corn per hectare⁵⁸, the chemical exergy from an **industrial** ethanol cycle is

$$9.81 \frac{\text{MJ}}{\text{kg wet corn}} \times 8590 \frac{\text{kg wet corn}}{\text{ha}} = \mathbf{84.4 \text{ GJ/ha}} \quad (28)$$

- With 91.5% overall conversion efficiency of starch into 100% ethanol, the chemical exergy is

$$84.4 \text{ GJ/ha} \times 0.915 = \mathbf{77.1 \text{ GJ/ha}} \quad (29)$$

Remark 10 The output of the industrial CO₂-Glucose-EtOH cycle is the chemical exergy of ethanol equal to 77.1 GJ/ha-crop. This exergy can be transformed into useful work (e.g., shaft work or electricity) by different devices. □

For example, the efficiency of an excellent internal combustion engine is 35% (usually it is 20%)

$$W_u = 77.1 \times 0.35 = 27.0 \text{ GJ/ha} \quad (30)$$

For fuel cell/electric motor vehicles the efficiency of conversion of chemical exergy to shaft work is higher. Suppose that we could reform ethanol to hydrogen, and use a fuel cell with 60% efficiency⁵⁹ (Deluga et al., 2004)

$$W_u = 77.1 \times 0.60 = 46.3 \text{ GJ/ha} \quad (31)$$

to obtain electricity, and shaft work.

8 Exergy Analysis of the Modified Ideal Corn-Ethanol Cycle

Now let us look at the useful exergy production in the modified ideal corn-ethanol-**hydrogen** cycle discussed by DELUGA et al. (2004). This cycle is essentially the same as the cycle described in Eqs. (24 – 26). The only difference is in Step 3, which is moderately endothermic, and lowers the cycle efficiency by one percent, see **Figure 42**. The chemical exergies of the products of each step of the modified cycle are listed in **Table 22**.

Step 3, reforming ethanol to hydrogen, is a catalytic variant of water-shift reaction:

⁵⁸See Remark 6. Without synthetic fertilizers, I would have to use 1500 kg/ha-crop of corn as the average yield.

⁵⁹As shown by (Bossel, 2003a) and in **Appendix D**, a 60%-efficient PEM fuel cell-powered car cannot exist! The real-life efficiency of such a car is about 38%. The reckless promoters of a hydrogen economy neglect to mention this important downward correction.

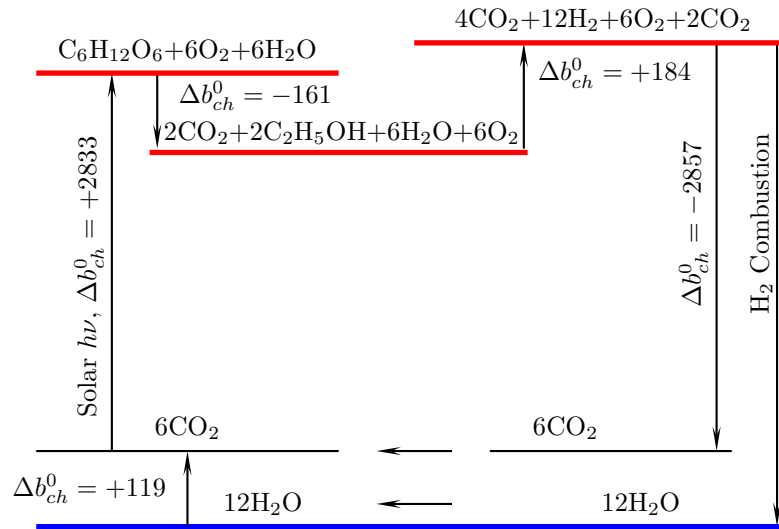
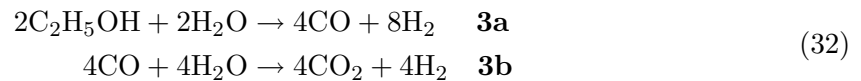


Figure 42: Exergy diagram of the ideal CO_2 -Glucose-EtOH- H_2 cycle.

Steps 3ab Ethanol oxidation to CO , and then CO_2 :



Note that DELUGA et al. (2004) use a somewhat different stoichiometry with 10H_2 and $10\text{H}_2\text{O}$, but this difference is insignificant.

Remark 11 The ideal CO_2 -Glucose-Ethanol- H_2 cycle discussed in DELUGA et al. (2004) produces practically the same amount of useful chemical exergy as the ideal CO_2 -Glucose-Ethanol cycle. Therefore, all conclusions pertinent to the latter hold for the former. \square

9 Resource Consumption and Waste Generation in the Industrial Corn-Ethanol Cycle

Now I will focus my attention on the industrial corn-ethanol cycle depicted in Figure 36. In contrast to the sun-driven ideal cycle, the industrial cycle relies heavily on fossil energy. Therefore, a part W_R of the useful work W_u , must be diverted to restore the non-renewable resources depleted by the cycle, see **Figure 43**. As long as the useful work exceeds the restoration work, $W_u > W_R$, the industrial corn-ethanol cycle is beneficial, otherwise it is indefensible.

Remark 12 The depletion of concentrated natural resources is *irrevocable*, cf. Part II. Without causing changes in the environment, we cannot remanufacture the depleted amounts of oil, methane and coal in a reversible process, and put these fuels back into their deposits. Therefore, the reversible restoration work calculation provides the *lowest estimate* of the degree to which the irreversible industrial corn-ethanol cycle is also unsustainable. \square

From Definition 6, it follows directly that the *minimum* restoration work is equal to the sum of the cumulative exergy consumption (CExC) by all the processes that convert natural resources into

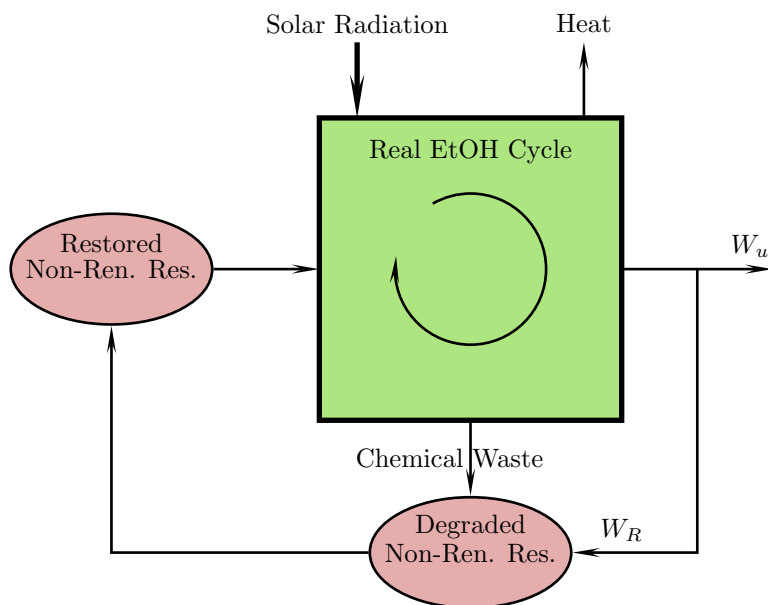


Figure 43: A part, W_R , of the useful work, W_u , from the industrial corn-ethanol cycle is diverted to “undo” mining of the environment by this cycle. If $W_u > W_R$, there is net benefit from the ethanol biofuel, otherwise its use should be stopped. Adapted from Figure 4 in BERTHIAUME et al. (2001).

inputs of the industrial corn-ethanol cycle. The specific CExC for each such input is listed in **Table 23**. For example, the CExC by the production of ammonium nitrate starts from natural gas in the subsurface and nitrogen in the atmosphere, and continues by summing up all the intermediate steps. Even though the chemical exergy of ammonium nitrate is only 10.51 MJ/kgN, the cumulative exergy consumption to produce it is almost 10 times larger. The CExC by electricity generation is based on the average exergy efficiency of ten large modern power stations in the U.S. (Gill, 1984; Termuehlen and Emsperger, 2003).

Remark 13 Currently, the fossil fuels are mined from very concentrated deposits and upgraded in very large and efficient plants. Therefore, their CExC’s are relatively small. As the rich crude oil and natural gas deposits are depleted, in part to fuel extravagant and politically-driven projects such as corn-ethanol in the U.S., the CExC by gasoline, diesel fuel, and clean natural gas will increase dramatically. I will address this looming problem elsewhere. \square

The total CExC for each input is the product of the specific value in Table 23 (in MJ/kg) multiplied by the flux (in kg/ha-crop) estimated in Part I. The CExC by cleanup of the water contaminated in the industrial ethanol cycle was not yet considered, and will be discussed now.

9.1 Cleanup of BOD in Ethanol Plant Wastewater

The CExC by the cleanup of ethanol plant wastewater is the product of the total biological oxygen demand (BOD) in the wastewater stream in kg O₂/ha-crop, and the specific CExC by BOD removal in MJ/kg O₂. As discussed in Section 6, the total BOD is

$$\text{BOD} = 46 \frac{\text{L water}}{\text{kg EtOH}} \times 0.002 \frac{\text{kg O}_2}{\text{L water}} \times 2499 \frac{\text{kg EtOH}}{\text{ha-crop}} = 230 \frac{\text{kg O}_2}{\text{ha-crop}} \quad (33)$$

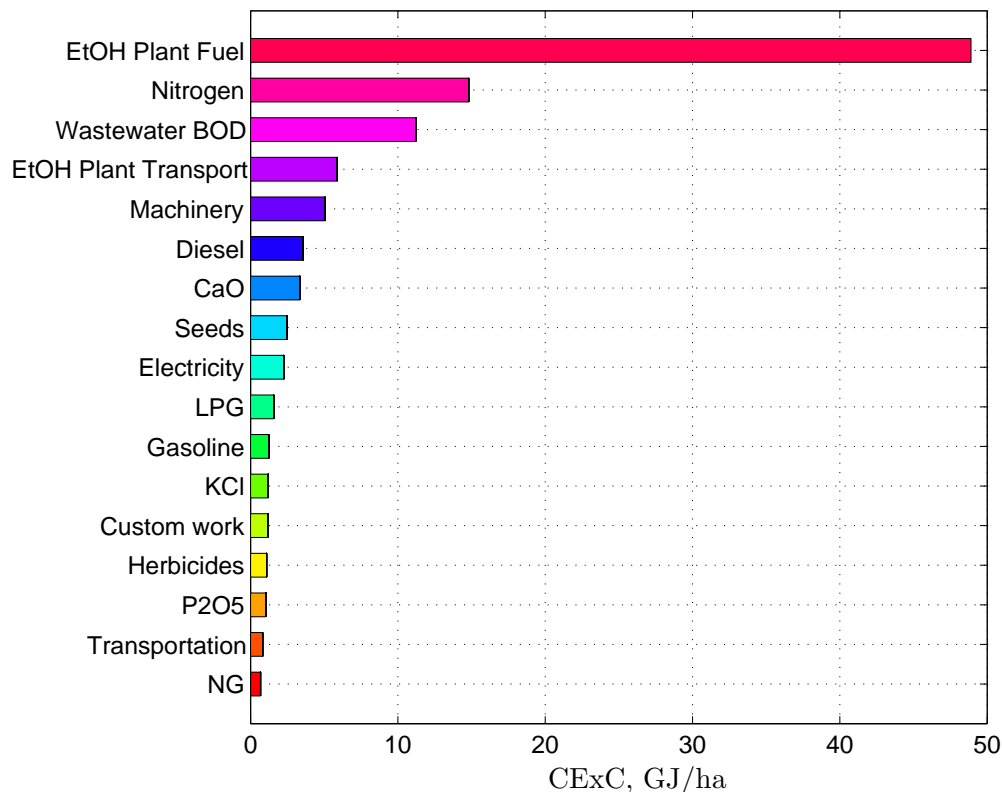


Figure 44: The specific cumulative exergy consumption (CExC) by each major non-renewable resource input to the industrial corn-ethanol cycle.

After BERTHIAUME et al. (Berthiaume et al., 2001), I will base the specific CE_{ExC} of BOD on a survey of energy consumption in municipal wastewater treatment facilities in the Canadian province of Québec. According to BLAIS et al. (1995), an average electricity consumption of 4.13 kWh per kg of BOD removal was observed there.

I have used these estimates in the CO₂ emission calculations shown in Figures 38 and 39.

9.2 Cleanup of Contaminated Field Runoff Water

Analysis of the cumulative exergy consumption in the cleanup of contaminated agricultural water that

- Seeps into the aquifers, causing, e.g., the omnipresent nitrate contamination of groundwater in the Corn Belt (Patzek et al., 2005), and
- Runs off to streams, rivers, into the Mississippi River, and to the Gulf of Mexico, generating a large anoxic zone there,

is the subject of future research and will not be discussed here.

The Second Law of thermodynamics puts a high price on concentrating and removing dilute contamination, especially from the subsurface. Therefore the restoration work of cleaning up the field runoff water will be *several orders* of magnitude larger than that of the ethanol plant runoff water.

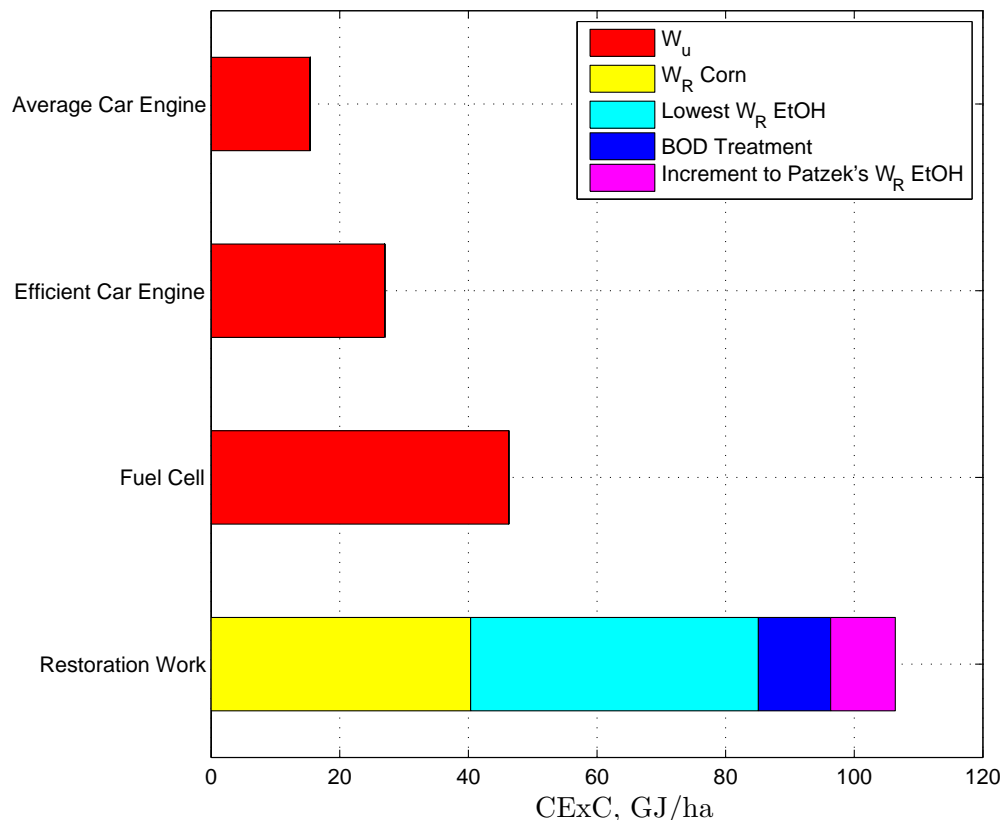


Figure 45: The minimum cumulative exergy consumption by the industrial corn-ethanol cycle and its maximum useful work, W_u . For comparison, the cycle's ethanol is burned in an average car engine, an efficient car engine, and in an ideal fuel cell. This comparison demonstrates that the industrial corn-ethanol cycle is unsustainable by a factor of 2.3-7. No adjustment of process parameters I can think of will change this terrible situation. Note that the W_R EtOH bar is the *lowest* fossil energy use (34 000 Btu/gal+0.75 kWh/gal advertised in 2004 by ICM, 310 North First Street, Colwich, Kansas 67030, www.icm-inc.com). The rightmost purple bar shows the difference between the ICM value and my estimate of fossil energy use to produce anhydrous ethanol.

Better fertilization practices (Worrell et al., 1995), and artificial wetlands (Horne, 1991; Horne and Gregg, 1993; Horne et al., 1994; Horne, 1994) around the corn fields could significantly help in containing and removing the pervasive contamination these fields generate.

9.3 Restoration Work

Now we are ready to estimate the restoration work of the non-renewable resources mined by the industrial corn-ethanol cycle. The results for each major input are shown in **Figure 44**. The three main sources of exergy consumption are ethanol plant fuel, nitrogen fertilizer, and the removal of biological oxygen demand in the ethanol plant wastewater. Note that I have not yet included the potentially huge restoration work of the High Plains aquifer (which underlies many of the Corn Belt states), other aquifers, the numerous streams and rivers which drain the field wastewater, the Mississippi River, and the Gulf of Mexico at the Mississippi River mouth.

The bottom-line comparison is shown in **Figure 45**. Here I compare the as yet incomplete

cumulative exergy consumption by the industrial corn-ethanol cycle with the cycle's maximum useful work performed by three different machines. This comparison reveals that the corn-ethanol cycle consumes 2.3-7 times more exergy than it replaces.

The lowest deficiency of the cycle, by a factor of 2.3, is realized by employing an *imaginary* 60%-efficient Proton-Exchange Membrane (PEM) fuel cell to power a car. As shown in **Appendix D** and elsewhere, such a cell simply cannot exist (Bossel, 2003b). *Real* fuel cells are 2-3 orders more expensive than a car engine, 10 times less reliable, and may never be mass-produced (Keith and Farrell, 2003; Dresselhaus et al., 2003; Bossel et al., 2003; Davis et al., 2002). A 35%-efficient internal combustion engine produces 4 times less useful work than the restoration work, and today's average car engine produces 7 times less work. Note that if one uses the lowest advertised (but not necessarily true) value of fossil fuel consumption to produce anhydrous ethanol, these factors are 2, 3, and 6 respectively.

Remark 14 No matter how efficient the engine is that transforms the industrial corn-ethanol cycle's output into shaft work, the cycle remains utterly unsustainable and unattractive as a source of fossil fuel. □

10 Conclusions

- Excluding the restoration work of decontaminating aquifers, rivers, and the Gulf of Mexico, the *minimum* cumulative exergy consumption in restoring the environment polluted and depleted by the industrial corn-ethanol cycle is over 7 times higher than the *maximum* shaft work of a car engine burning the cycle's ethanol.
- This unfavorable ratio decreases to ~ 4 , when an efficient internal combustion engine is used to burn the ethanol, and to 2.3 when an imaginary hydrogen fuel cell is used.
- The industrial corn cycle is not renewable, and is *unsustainable* by a wide margin (at least 2.3 – 7 times).
- No process changes can make this cycle more viable.
- The annual corn-ethanol biofuel production is a human assault on geologic processes and the geologic time scale, and it can *never* work.
- The *limiting factors*, nutrient-rich humus and water that carries the dissolved nutrients to plant roots are augmented by chemicals obtained in the linear, irreversible fossil fuel-based processes.
- Over the last fifty years, corn yield has grown five-fold, mostly because of the steep increases in fertilization rate of corn fields.
- Sunlight is *not* a limiting factor, and could be used to great benefit if we were in less of a hurry, cf. **Appendix C**.

*In vain, through every changeful year,
Did Nature lead him as before;
A primrose by a river's brim
A yellow primrose was to him,
And it was nothing more.*

— WILLIAM WORDSWORTH, *Peter Bell* (1819)

The outside world is something independent from man, something absolute, and the quest for the laws which apply to this absolute appeared to me as the most sublime scientific pursuit in life

— MAX PLANCK
Scientific Autobiography, and Other Papers (1949)

Part IV

Other Problems with Corn-Ethanol

1 Introduction

The results obtained in Part I and III can now be used to estimate the annual subsidies of the U.S. corn-ethanol biofuel. These subsidies fall into two categories: (1) the monetary subsidies by the U.S. Federal Government and State Governments to corn and ethanol producers, and (2) the non-monetary contributions of the U.S. population and environment polluted and mined by the industrial corn-ethanol cycle.

The first type of subsidy is easy to estimate:

- The U.S. pays its corn farmers \$10 billion a year in subsidies⁶⁰ (13% of corn acreage is now devoted to ethanol).
- The federal excise taxes per gallon of fuel⁶¹: \$0.184 for gasoline and \$0.132 for EtOH-10 (10 vol% ethanol).
- Relative to gasoline, the federal tax breaks per 1 gallon of ethanol are

$$10 \times (\$0.184 - \$0.132) = \$0.52$$

- The state excise taxes per gallon of fuel vary widely and their volume-weighted average is difficult to estimate:
 - Gasoline \$0.075 in Georgia, up to \$0.36 in Connecticut
 - EtOH-10 \$0 in Arizona, up to \$0.2535 in West Virginia
 - Minnesota offers a \$0.20 tax subsidy on EtOH-10. The state's 13 ethanol refineries have received up to \$3 million a year⁶².
- I will assume that the average state subsidy is only \$0.10/gallon of ethanol. It is likely that the state subsidies of corn growers, ethanol producers, and EtOH-10 excise tax breaks are several times higher than my estimate.

The second type of subsidy can only be estimated by comparing the maximum useful output of the corn-ethanol cycle with the minimum restoration work necessary to undo the cycle's negative impact on the environment.

2 First-Law View of Corn-Ethanol Production in 2004

In February 2004, the U.S. ethanol industry set an all-time monthly production record of 212,000 barrels per day (b/d) (12.3 GL/year), according to data released by the U.S. Energy Information Administration (EIA), the Renewable Fuels Association (RFA) reported. The previous all-time record was 211 000 b/d in the previous month. Production was up 25 percent compared to last February when 169 000 b/d (9.8 GL/year) of ethanol were produced.

⁶⁰Corn subsidies, OXFAM.

⁶¹Excise Taxes, REUTERS, Washington, Nov. 15, 2003.

⁶²CAT LAZAROFF - ENS, 3 Oct., 2002.

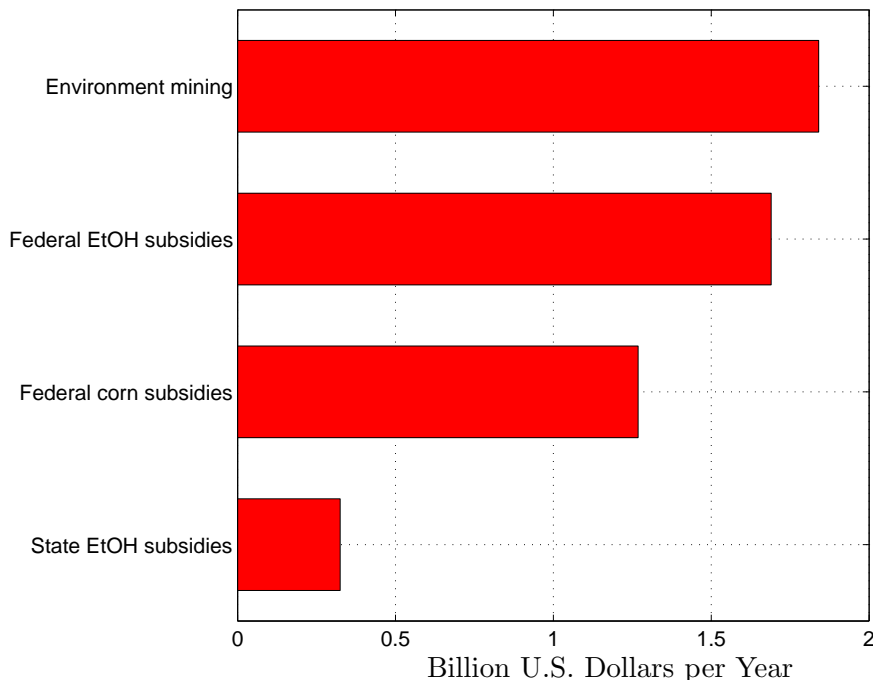


Figure 46: The U.S. environment and taxpayers subsidize corn growers and ethanol producers. The U.S. rural population, soil, water, air, plants, fish and wildlife pay the most.

“U.S. ethanol producers are doing everything they can to add much needed volume and octane to the U.S. gasoline market, thereby helping to hold down gasoline prices,” said RFA President BOB DINNEEN. The ethanol industry is expected to produce more than 3.3 billion gallons in 2004, up from 2.81 billion gallons in 2003. Currently, 75 ethanol refineries have the capacity to produce more than 3.2 billion gallons annually. *Thirteen* additional plants under construction will add 500 million gallons of annual production capacity, RFA said.

The highlights of the U.S. corn-ethanol production in 2004 are listed in **Table 24**. Note that the 12.3 GL/yr of corn-ethanol replace only 9.2 GL/yr of gasoline equivalent (GE), and require 10.1 GL/yr of GE to produce. Corn-ethanol brings *no* energy savings and *no* lessening of the U.S. energy dependency on foreign crude oil, natural gas, and liquified petroleum gas. The opposite happens, we import more methane, LPG, and crude oil. We then burn these fuels to produce corn-ethanol and, finally, we burn the ethanol in our cars, causing extensive environmental damage at each stage of the industrial corn-ethanol cycle. Also note that the ethanol and corn tax-subsidies projected for 2004, will be 10 times higher than the total political contributions of Agribusiness over the last 14 years⁶³.

3 Second-Law View of Corn-Ethanol Production in 2004

Second Law exergy analysis brings into the picture the significant contribution of the environment to corporate profits. Depending on which thermodynamic device transforms the chemical exergy

⁶³According to the Center for Responsive Politics, <http://www.opensecrets.org/industries/indus.asp?Ind=A>, between 1990 and 2005, Agribusiness paid \$355 534 846 to elected officials, 31% to Democrats and 69% to Republicans. The peak donations occurred in the 2000 and 2002 election cycles, exceeding 59 and 54 million dollars, respectively. The 2004 election cycle donations exceeded 53 million dollars.

of corn-ethanol into useful work, the difference between the minimum restoration work and the maximum useful work by the cycle will vary. This difference in GJ/ha, can be translated roughly into \$/ha, and into the cumulative environmental cost of the industrial corn-ethanol cycle. The hidden cost of mining the environment by the industrial corn-ethanol cycle is real, but rarely mentioned. According to the RFA President BOB DINNEEN, ethanol displaces imported crude oil. Therefore, to arrive at an estimate of corn-ethanol's environmental costs, I assume that the total exergy deficit will be "paid" with the imported crude oil, whose price in the first half of 2004 was close to \$35/barrel. I also assume that 1 barrel of oil is 136 kg of 35⁰ API oil, with the specific energy content of 45 MJ/kg. The results are listed in **Table 25**.

In 2004, the environment has contributed an estimated 1.8 billion dollars per year by being continuously and irrevocably damaged and depleted. This huge gift to the corporate coffers from the U.S. rural population, soil, water, air, plants, and wildlife is as real as the federal tax subsidies. The 60%-efficient fuel cell car does not exist now, or in the future (Bossel, 2003b; Patzek and Pimentel, 2006). But even if in the next 20-years we were to replace all existing cars with efficient fuel cell cars, the environment's contribution would still be \$1.2 billion per year. If 10% of fuel consumption in the U.S. were supplied by corn-ethanol, the annual contribution from the environment would be \$12 billion.

Again, my current estimate should be viewed as the *lowest* bound on the environmental costs for two reasons: (1) The true restoration work is irreversible and significantly larger than the reversible restoration work, and (2) I have not yet calculated the minimum reversible work of restoring surface and ground water, and soil contaminated by the corn field runoff water.

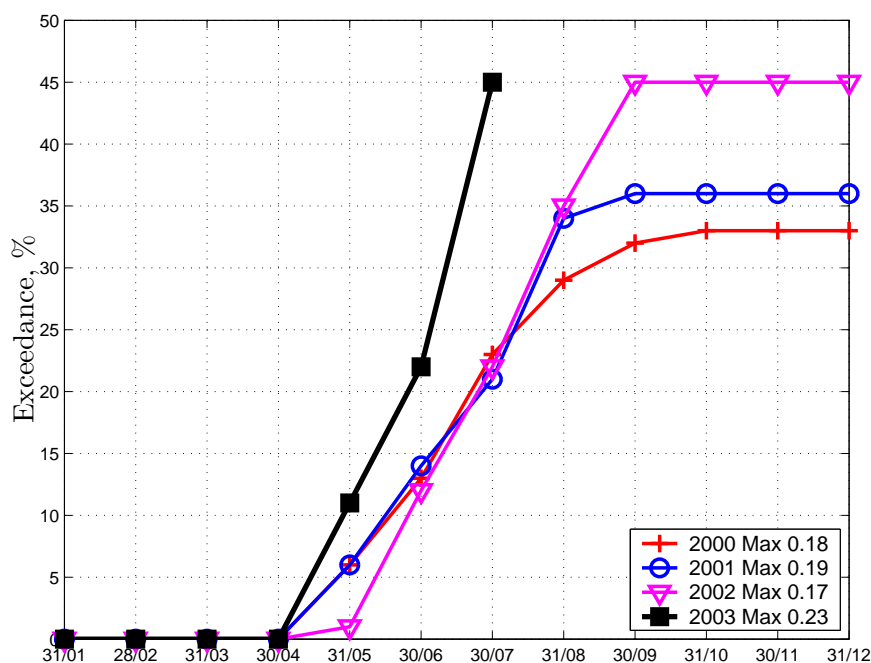


Figure 47: The cumulative one-hour exceedances of maximum legal ozone level in Southern California. Source: CAL HODGE, President of A 2nd Opinion, Inc.

All the subsidies to corn growers, ethanol producers and distributors are compared in **Figure 46**.

4 Public Health Problems

The stated goal of adding ethanol from corn to gasoline was to help in cleaning the air we breathe and lessen the U.S. dependence on foreign oil. The opposite is achieved. Air becomes more polluted, and as much oil and more methane are burned as without the corn-ethanol. At the same time, additional health hazards are created by the agricultural chemicals, fertilizers, insecticides and herbicides, and by the waste water streams.

For example, in 2002, twelve Minnesota ethanol refineries agreed to spend \$2 million per plant, pay penalties of \$29,000-\$39,000, and limit the following air emissions⁶⁴

- Volatile organic compounds by 2400 - 4000 tons per year,
- Carbon monoxide emissions by 2000 tons per year,
- Nitrogen oxides emissions by 180 tons per year,
- Particulate matter by 450 tons per year,
- Other hazardous air pollutants by 250 tons per year.

Ethanol-in-gasoline seriously pollutes the air (Hodge, 2002). The reactivity of the combined exhaust and evaporative emissions using the ethanol-blended reformulated gasoline is estimated to be about 17% larger than those using the MTBE-blended reformulated gasoline (NRC, 1999). Ethanol does reduce the carbon monoxide emissions, but increases those of nitrogen oxides (NO_x), acetaldehyde, and peroxy-acetyl-nitrate (PAN) (Rice et al., 1999). The negative effects of using gasoline-ethanol blends are clearly seen in Southern California, where ozone levels in the air exceeded the one-hour legal limits more often, see **Figure 47**. By 2003, over 70% of gasoline produced in Southern California was blended with ethanol.

In 2004, the California Air Quality Board completed a study (Hancock, 2005) in which the fuel systems of several vehicles were tested for diurnal evaporative permeation emissions with fuels containing MTBE and EtOH. The results were applied to the existing fleet in the South Coast Air Basin and Sacramento. Their analysis showed a 17-ton-per-day (ton/d) increase in the South Coast Air Basin on an ozone episodic day, an increase of 14% for evaporative emissions. For the Sacramento Metropolitan area, the increase in evaporative emissions due to ethanol was estimated to be 18% and 2.4 ton/d. In an earlier draft report, the California Air Quality Board concluded (ARB, 2005) that in the South Coast Air Basin alone, the *removal* of ethanol oxygenate from the reformulated gasoline CaRFG3 would decrease hydrocarbon emissions by 27.4 ton/d, NO_x emissions by 6.7 ton/d, and increase CO emissions by 1.4 tons/d. So much for the cleaner air...

*Friends, Romans, countrymen, lend me your ears;
I come to bury Cæsar, not to praise him.
The evil that men do lives after them;
The good is oft interred with their bones;*

— WILLIAM SHAKESPEARE
Julius Cæsar Act 3, Scene 2,
(1599)

⁶⁴CAT LAZAROFF - ENS, 3 Oct 2002.

Part V

Summary & Conclusions

The purpose of this paper was to prove beyond any reasonable doubt that the industrial corn-ethanol cycle accelerates the *irrevocable* depletion of natural resources: fossil fuels, minerals, top soil, surface and subsurface water, and air, while creating wide-spread environmental damage throughout the continental United States. My arguments relied entirely on the First and Second Law of thermodynamics, and on the Law of Mass Conservation.

I have tried to avoid political questions, but at some point one should ask how it was possible for a poor agri-industrial technology to grow so explosively in the last four years? The only plausible answer lies in politics. The recent growth of ethanol production could occur only because of the massive transfer of money from the collective pocket of the U.S. taxpayers to the transnational agricultural cartel, represented most notably by Archer Daniel Midlands Co., Cargill Inc., Monsanto Co., and A. E. Stanley Manufacturing Co. This flow of billions of dollars from the pockets of the many to the pockets of the few was accomplished by federal subsidies of corn producers, and the federal and state tax subsidies of ethanol producers. It was spearheaded by many powerful, and I would like to think, thoroughly misinformed politicians.

More ominously, as a country, we have diverted our collective attention from the most important issue of this century: energy conservation and increased reliance on the *only* renewable source of energy, the sun, and its weak derivative, the wind, see **Appendix C**. Instead, we have somewhat *accelerated* the rate of depletion of the precious natural gas and crude oil deposits, in exchange for the significantly more wide-spread pollution of water, soil and air over roughly 1/2 of the area of the United States, the incremental carbon dioxide emissions, the substandard ethanol fuel, and the continuous drain of taxpayers' money.

To make things worse, the scientific community in the U.S. seems to be preoccupied with promulgating empty illusions of a future global energy bliss brought about by the new and sexy, but inherently unsustainable technologies. The ethanol biofuel for hydrogen (Deluga et al., 2004), the fossil fuel-based "hydrogen economy" (Davis et al., 2002; Bossel et al., 2003; Dresselhaus et al., 2003; Keith and Farrell, 2003; Tromp et al., 2003; DOE, 2003), the subsurface carbon dioxide sequestration (Celia, 2002), etc., come to mind. I suggest that we – the scientists – should instead be advocating the simpler and less expensive, but painful, real solutions of the overwhelming energy problems facing the world. These solutions must involve far more energy conservation in every aspect of the U.S. economy, and the significantly increased reliance on the sun.

The philosophical, ethical, and political arguments ought to be developed further, but I will leave this task to the others, see e.g., the transcript of an excellent speech by NICHOLAS E. HOLLIS, *Ethics and Agribusiness – In Search of the New Food Security*, given in Newcastle-on-Tyne, United Kingdom, March 15, 2004. Here I will only reiterate the following:

1. The industrial corn-ethanol cycle brings *no* energy savings and *no* lessening of the U.S. energy dependency on foreign crude oil, natural gas, and liquified petroleum gas. The opposite happens, (a) we import somewhat more methane, LPG, and crude oil; (b) we burn these fossil fuels to produce corn and ethanol; and (c) we burn the corn ethanol in car engines. All three steps of this cycle increase the extent of environmental damage beyond that caused by burning the same fossil fuels directly in the cars.
2. The industrial corn-ethanol cycle generates more carbon-dioxide equivalents than would be generated by the energy-equivalent quantity of gasoline or diesel fuel penalized by 15% to

account for the cumulative use of free energy in their production. Currently these incremental emissions are about 11 million tonnes of equivalent CO₂ per year.

3. The taxpayers' subsidies of the industrial corn-ethanol cycle are estimated at \$3.5 billion in 2004. The parallel subsidies by the U.S. environment are estimated at \$2.0 billion in 2004. The latter estimate will increase manifold when the restoration costs of aquifers, streams and rivers, and the Gulf of Mexico are also included.
4. The rate of sequestration of the unlimited solar energy as organic plant matter is controlled by the availability of water in soil and the minerals dissolved in this water. For the reasons explained in this paper, water and soil nutrients are finite, easily degradable, and must be replenished by decomposing all corn plant leftovers, including the byproducts of ethanol production.
5. The mass balance of starch in dry corn grain sets the theoretical efficiency of conversion of corn to ethanol to 0.374 kg EtOH/kg dry grain (0.475 L EtOH/kg dry grain), or to 3.19 gallons EtOH/dry bushel = 2.71 gallons EtOH/equivalent bushel of corn with 15% moisture.
6. A 4-year average of the ethanol industry reported yields is 2.477 gal EtOH/wet bushel, or 91.5% of the theoretical average.
7. Therefore, the commonly accepted USDA estimate of ethanol yield, 2.682 gallons EtOH/bushel (Shapouri et al., 2002b) is inconsistent with the industry data and too high.
8. *All* considered analyses of the fossil energy inputs into the industrial corn-ethanol cycle (Wang et al., 1997; Shapouri et al., 2002a; Pimentel, 2003; Berthiaume et al., 2001; Patzek et al., 2005) have been corrected and/or amended, as detailed in Part I.
9. With the corrected estimates of the fossil fuel inputs to corn farming and ethanol production, and using the high heating values of fossil fuels, three published estimates predict that production of corn-ethanol is the fossil energy losing proposition, and two show slight surpluses, see **Figure 19**. When the low heating values are used, all published analyses show that the corn ethanol's calorific value is less than those of the fossil fuels (Patzek, 2004).
10. In view of Conclusion 4, the robust ~30% energy credits for ethanol production used in (Wang et al., 1997; Shapouri et al., 2002a) are indefensible. In Part II, the Second Law analysis of the industrial corn-ethanol cycle reaffirms this conclusion.
11. The mere energy balance of the industrial corn-ethanol cycle does not take into account the cumulative consumption of the non-fossil environmental resources: soil, water, air and minerals, and provides an incomplete picture of this cycle.
12. Because all published First Law balances of the industrial corn-ethanol cycle have been incomplete by definition, their comparison has led to different interpretations, and endless acrimonious debate, see e.g., (Patzek et al., 2005), which served little useful purpose. Worse yet, this debate has diverted our collective attention from the real problems with corn-ethanol and provided cover for the ethanol lobby.
13. Only when the Second Law of thermodynamics and the concept of available free energy, or *exergy*, are introduced (Parts II and III, Appendix A and B), a definitive analysis of the industrial corn-ethanol cycle is possible.

14. In the Second Law analysis, the *environment* is defined as anything but the top soil of the corn field, the corn plants, the ethanol-producing plants, and the devices that process the ethanol (internal combustion engines and fuel cells). The environment, therefore, is the sun, water, air, nutrients, crude oil, methane, coal, electricity, field chemicals, roads, trucks, etc., and the cold universe to which the heat generated by the cycle's ethanol is rejected. Thus, the environment defined here *cannot* be dismissed summarily as the raving of a green lunatic.
15. From the definitions of *irreversibility* and *sustainability* (Part II, Appendix A and B), it follows clearly that the industrial corn-ethanol cycle, which relies heavily on mining the environment, is irrevocable and unsustainable. The common references to the cycle sustainability, see, e.g., (Deluga et al., 2004), are scientifically indefensible.
16. In view of Conclusion 4, the recently-advertised ethanol production from "agricultural waste," i.e., from plant leaves, stems, and roots is even more unsustainable.
17. Because the industrial corn-ethanol cycle is irrevocable, the precious natural resources wasted by it disappear forever from the earth, and diminish the wellbeing of future generations.
18. The free energy available on the earth (exergy) is consumed by the industrial corn-ethanol cycle. The cumulative net consumption of the exergy is a good quantitative indicator of the degree of unsustainability of this cycle.
19. The maximum useful free energy generated by the industrial cycle, 77.1 GJ/ha as corn-ethanol, produces useful shaft work, electricity, or both.
20. When an average car engine (20% efficiency) burns corn-ethanol, it produces 15.4 GJ/ha of useful work. A very efficient car engine (35% efficiency) produces 27.0 GJ/ha of useful work. An imaginary fuel cell/electric motor car with 60% efficiency, produces 46.3 GJ/ha of useful work (see Appendix D).
21. The *maximum* useful work from the industrial corn-ethanol cycle has been compared with the *minimum* reversible work required to undo the environmental damage caused by this cycle. Of course, this environmental damage cannot be undone in reality, and the minimum reversible restoration work is only a *tool* used to quantify the damage.
22. The cumulative exergy consumption in performing the minimum reversible work of "undoing" the industrial corn-ethanol cycle, exceeds by a factor 7, 4 and 2.3, respectively, the maximum useful work from a 20%-efficient car engine, 35%-efficient car engine, and a nonexistent 60%-efficient fuel cell car.
23. No process change can improve the very unfavorable ratio of benefits from the industrial corn ethanol cycle to its environmental costs, and I have not yet included the exponentially growing costs of restoring surface and ground water, and soil contaminated by the fertilizer-, pesticide- and insecticide-rich runoff from the corn fields.
24. More attention should be devoted to harvesting the only renewable source of energy on the earth: sunlight. An inefficient solar cell that continuously operates at moderate latitude produces ~ 100 times more electricity than a corn ethanol-burning fuel cell (see Appendix C). Somewhat better, more durable and cheaper solar cells that are distributed worldwide would dramatically increase the chances of long-term survival of our civilization.

25. One hectare of solar cells placed anywhere can free 100 hectares of fertile agricultural land from industrial corn, and allow for the low-intensity, diversified, and almost sustainable agriculture.

17 Now I beseech you, brethren, mark them which cause divisions and offences contrary to the doctrine which ye have learned; and avoid them.

18 For they that are such serve not our Lord Jesus Christ, but their own belly; and by good words and fair speeches deceive the hearts of the simple.

— THE EPISTLE OF PAUL THE APOSTLE TO THE ROMANS,
(*The New Testament*)

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Subsequent to the CRPS article publication, Dr. DREW RONNEBERG of Technology and Management Services, Inc. has pointed out my error in calculating the theoretical efficiency of ethanol production from corn, kindly provided me with additional information about corn ethanol refineries, and highlighted the incompleteness of JEVONS' arguments about coal production.

A Examples of Entropy Production and Disposal

The Rate of Entropy Export by the Earth

In the simplest model, the earth is in thermal equilibrium; continuously heated by the sun's radiation, and cooled by the infrared radiation into the universe.

The *solar constant* is the power collected at the top of the earth atmosphere by a unit area (1 m²) perpendicular to the light path. This power is remarkably constant, see e.g., (Hickey et al., 1980), and equal to $\dot{q} = 1370$ Watts⁶⁵ per square meter (W/m²). The projection of the sun-lit earth hemisphere in the direction perpendicular to the sun light is πr_e^2 , where $r_e = 6371$ km is the mean volumetric earth radius, or 1/2 of the hemisphere area, and 1/4 of the earth surface area ($A_e \approx 510 \times 10^6$ km²).

The PLANCK temperature of the sun's radiation is $T_s \approx 5700$ K, and the PLANCK temperature at which the earth radiates its heat is $T_e \approx 254$ K⁶⁶. The earth reflects about 30% of the sun radiation, so its surface is reached by only 0.7 of the solar energy. Therefore, the time-averaged flux of entropy exported by the earth into the universe is

$$j_S^E = \frac{\Delta \dot{S}}{4\pi r_e^2} = \frac{4}{3} 0.7 \dot{q} \frac{1}{4} \left(\frac{1}{T_e} - \frac{1}{T_s} \right) = 1.20 \quad \text{W/K-m}^2 \quad (34)$$

where the factor 4/3 comes from the STEFAN-BOLTZMANN law⁶⁷. This estimate agrees very well with the $j_S^E = 1.2$ W/K-m² reported by Prof. ARNE STAHL (1996).

The Simplest Climate Model

The earth is in thermal equilibrium:

$$\begin{aligned} \text{Rate of energy input from the sun} &= \text{Rate of energy radiation by the earth} \\ \dot{E}_{in} = \dot{E}_{out} = P &= \text{const} \end{aligned} \quad (35)$$

Similarly to FRONDEL, OERTEL and RÜBBELKE (2002), I assume that the earth's atmosphere is a heat-transporting⁶⁸ gas layer. The surface temperature of the earth is T_0 and the PLANCK temperature of its radiation is T_e . The stationary heat flow through the atmosphere occurs at a constant rate:

$$P = k A_e (T_0 - T_e) \quad (36)$$

where A_e is the surface area of the earth. In this simple model, the steady-state rate of energy export from the earth depends on the mean temperature difference between the earth surface and the uppermost atmosphere. The overall heat transfer coefficient, k , depends on how effectively the atmosphere transports heat. With the increasing concentration of heat absorbers (the greenhouse gases) this coefficient decreases, and the earth temperature must go up.

At steady state, Second Law of thermodynamics requires the entropy flow rate to satisfy the following equation:

$$P = T_0 \dot{S}_0 = T_e \dot{S}_e \quad (37)$$

⁶⁵1 Watt = 1 Joule/second.

⁶⁶The actual temperature of the earth surface is about 34 K higher due to the greenhouse effect. Therefore there is entropy generation in the atmosphere, see the section below.

⁶⁷See SZARGUT's monograph, (1988), page 72, Eq. (2.53).

⁶⁸Heat transport through the atmosphere proceeds through turbulent convection and mixing, water evaporation and condensation, thermal conduction, and radiation.

where \dot{S}_0 is the rate of entropy change near the earth surface, and \dot{S}_e is the rate of entropy change at the outer layers of the atmosphere. Therefore, the net rate of entropy generation in the atmosphere due to energy transport is

$$\dot{S}_t = P \left(\frac{1}{T_e} - \frac{1}{T_0} \right) > 0 \quad (38)$$

Since both P and T_e are constant, it follows that as the temperature of the earth surface increases, so does the rate of entropy generation in the atmosphere. This *positive feedback* effect, or *domino effect*, will intensify the impact of human entropy production on the earth climate. The rate of entropy generation in the atmosphere per unit surface area of the earth defines the specific atmospheric entropy generation rate:

$$\sigma_t = \frac{\dot{S}_t}{A_e} \quad (39)$$

With the numbers used in Eq. (34), this specific rate is equal to 0.11 W/K-m², not bad for such a simplistic model. For comparison, STAHL (1996) reports 0.2 W/K-m².

Entropy Generated by a Human

A sedentary human needs about 2400 kcal per day in food to live and work a little. This food consumption translates roughly into an average sustained power of $\dot{Q} = 100$ W, i.e., one man = one 100 W bulb shining infrared light for 24 hours a day. This power is dissipated through work, thermal convection, radiation, excrements, etc., and appears as heat at the ambient temperature $T_0 = 273 + 15 = 288$ K. Thus the entropy generation rate is

$$\dot{S} = \frac{\dot{Q}}{T_0} = 0.35 \quad \text{W/K} \quad (40)$$

per person. At 3500 kcal per day, a quantity more representative of the U.S. feeding habits, the rate of entropy generation by one person is $\dot{S} = 0.5$ W/K, in agreement with (Stahl, 1996). This means, that if one crowds 2-3 people per square meter, they will generate more entropy than the earth can export.

Entropy Generated By Fossil Fuels

In the year 2000, the U.S. burned 83 quads (1 quad = 1 quadrillion or 10^{15} BTU⁶⁹) in fossil energy and 2.9 quads in wood, alcohol, etc. For simplicity, I will treat all this energy as the source of heat rejected⁷⁰ into the earth environment at the mean ambient temperature of $T_0 = 288$ K. The U.S. population was $N = 282$ million people in 2000. Therefore, in A.D. 2000, the per capita thermal entropy generation rate in the U.S. was

$$\dot{S} = \frac{\dot{Q}}{T_0 N} = \frac{(83 + 2.9) \times 10^{15} \times 1055/365/24/3600}{288 \times 282 \times 10^6} = 36 \quad \text{W/K-person} \quad (41)$$

My estimate of the per capita thermal entropy generation rate in the U.S. is somewhat higher than the 35 W/K-person calculated by STAHL (1996).

In the year 2000, the U.S. emitted about 1.65×10^{12} kg of carbon dioxide by burning fossil fuels (Marland et al., 2004) (1/4 of the global carbon dioxide emissions). The standard molar entropy

⁶⁹1 BTU = 1055 Joules of energy.

⁷⁰In reality, a portion of the fossil energy, probably 20-30%, generates work, resulting in less heat dumped into the environment. Therefore, this analysis provides the uppermost bound on the entropy generation rate.

of carbon dioxide is 213.8 kJ/kmol-K. Therefore, the per capita U.S. rate of generation of chemical entropy of carbon dioxide in the atmosphere:

$$\dot{S}_{\text{chem}} = \frac{1.65 \times 10^{12} \times 213.8 \times 10^3}{44 \times 365 \times 24 \times 3600 \times 282 \times 10^6} = 0.9 \text{ W/K-person} \quad (42)$$

is much smaller than the corresponding thermal entropy rate in Eq. (41). Unfortunately, this chemical entropy cannot be exported to the universe, and thus accumulates on the earth.

The whole world produced 400 quads of energy in the year 2000, therefore the U.S. used roughly 1/4 of it. The per capita entropy rate for the whole world excluding the U.S. was 6 W/K-person, or 1/6 of the U.S. rate. This means that where it really counts, in entropy generation rate, the U.S. society is *six times as wasteful* as the world at large, including Europe and Japan.

The U.S. thermal entropy generation rate may also be expressed per unit surface area of the country (9,629,100 km²), resulting in the specific anthropogenic entropy rate $\sigma + \sigma_a = 0.001$ W/K-m², or 0.1% of the net entropy flux exported from the earth. Of course, most of this entropy is generated on a small fraction of the U.S. land area. If all anthropogenic entropy were generated over 0.1% of the U.S. area⁷¹, roughly the area of Los Angeles, it would overwhelm the entropy export capacity of the earth. This calculation omits the entropy generation rate to “restore” the environment polluted chemically by the fossil fuel by-products. Nevertheless, the anthropogenic thermal entropy generation is still very small when compared with the capacity of the earth to export entropy. The chemical entropy, in contrast, continuously accumulates on the earth. Even the transport entropy in the atmosphere is generated at a specific rate 100-200 times higher than the mean generation rate of anthropogenic specific entropy in the U.S.. Note that in his interesting paper RÜBBELKE (1998) made a mistake (p. 198) by claiming that the U.S. population produced entropy at the rate of 35 W/K-m², not 35 W/K-person as calculated by STAHL (1996). The consequences of this mistake influenced erroneously his conclusions about sustainability.

What this calculation does not show, is the fast and irreversible exhaustion of the meager stock of high quality crustal materials (low entropy) that feed our civilization with the accompanying increase of chemical entropy.

⁷¹A physical impossibility!

B Availability and Irreversibility in Thermal Systems

Because the earth can export entropy by infrared radiation from the outer layers of the atmosphere, in this appendix I define any thermodynamic system on the earth as interacting with the atmosphere only. The atmosphere will be treated as being in stable dynamic equilibrium, and characterized by the constant absolute temperature T_0 , volume V_a , and the hydrostatic pressure, $p_0 = 1$ atm, at sea level. By including within the system as much surface land area, surface water, groundwater, minerals, material, plants, machinery, etc.⁷², as affected by the process of interest, one is always able to construct the system that interacts with the atmosphere only. For simplicity, chemical entropy generated by the system is not considered here.

GIBBS⁷³ showed that for any process which can occur under these circumstances the quantity defined as

$$\Phi = E + p_0V - T_0S \quad (43)$$

decreases

$$\Delta\Phi \leq 0 \quad (44)$$

where E is the total energy of the system, V its volume, S its entropy, and the increment of Φ , $\Delta\Phi$, is taken in the direction of increasing chronological Time.

The process of interest can only occur until the system pressure is uniformly hydrostatic and its uniform temperature is T_0 . Therefore, the state from which no *spontaneous* change can occur is the state in which the system has the hydrostatic pressure (p_0 at sea level) and the atmospheric temperature T_0 , and for which Φ has the smallest possible value, Φ_{\min} . If only one state of the system results in this minimum value, the system is in *stable equilibrium*. Otherwise, if there are several states corresponding to the minimum value of Φ_{\min} , the system is in *neutral equilibrium* of maximum stability.

GIBBS⁷⁴ referred to the difference

$$\Phi - \Phi_{\min}, \quad (45)$$

where Φ corresponds to the state in question, as the "... available energy of the body (*our system*) and the medium (*our surroundings*)."

JOSEPH H. KEENAN (1951) later showed that for the system undergoing change from an *earlier* state 1 to a *later* state 2, the amount of useful work W_u is

$$W_u \leq \Phi_1 - \Phi_2 \leq \Phi_1 - \Phi_{\min} \quad (46)$$

Therefore, for any state 1, the maximum possible useful work done by the system is $\Phi_1 - \Phi_{\min}$. KEENAN proposed to call this maximum value *availability*, Λ . It may be said⁷⁵ that for any system in the stable atmosphere

$$\Lambda \geq 0 \quad (47)$$

and that for the most stable state of the system

$$\Lambda = 0 \quad (48)$$

From Eq. (46) it also follows that

$$\begin{aligned} \Delta\Lambda &= \Delta\Phi \\ W_u \leq W_{u,\max} &= \Phi_1 - \Phi_{\min} = -\Delta\Phi = -\Delta\Lambda \end{aligned} \quad (49)$$

⁷²The entire globe, if necessary.

⁷³(Gibbs, 1994), p. 40.

⁷⁴(Gibbs, 1994), p. 53; my comments in italics.

⁷⁵(Keenan, 1951), Eqs. (12) and (13).

KEENAN also quantified the *irreversibility*⁷⁶, I , of a process executed by the system-atmosphere combination:

$$\begin{aligned} I &= W_{u,\max} - W_u \\ &= -\Delta\Phi - W_u \\ &= -\Delta\Lambda - W_u \end{aligned} \quad (50)$$

and showed that

$$\begin{aligned} I &= T_0\Delta S + \Delta E_a + p_0\Delta V_a \\ &= T_0\Delta(S_a + S) \end{aligned} \quad (51)$$

Of course, KEENAN's irreversibility was discovered much earlier and independently by GOUY (1889) and STODOLA (1898; 1927)⁷⁷. Regardless, the irreversibility of a process is equal to the increase of entropy of everything involved in the process multiplied by the temperature of the atmosphere.

One may use Eq. (51) in the differential form, and per unit area of the earth surface, by writing

$$\begin{aligned} \frac{1}{A_e} \frac{dI}{dt} &= T_0 \left(\frac{1}{A_e} \frac{dS_a}{dt} + \frac{1}{A_e} \frac{dS}{dt} \right) \\ \sigma_I &= T_0(\sigma_a + \sigma) \end{aligned} \quad (52)$$

where σ_I is the specific rate of irreversibility, σ_a is the specific rate of entropy increase in the atmosphere generated by the process, and σ is the specific rate of entropy increase of the system. Note that as the atmospheric temperature increases, so does the irreversibility of any process on the earth.

From Appendix A it follows that we can treat the atmosphere as an open system that exports entropy to the universe with the flux j_S^E calculated from Eq. (34). The energy transport through the atmosphere generates entropy at the specific rate of σ_t calculated from Eq. (39). Thus, we may rewrite Eq. (52) as

$$\sigma_I^{\text{net}}/T_0 = \underbrace{(\sigma_a + \sigma)}_{\text{Thermal entropy from Earth processes}} + \underbrace{(-j_S^E + \sigma_t)}_{\text{Thermal entropy from Solar processes}} \quad (53)$$

Using equation (53), we can define sustainability as

$$\begin{aligned} \sigma_I^{\text{net}} &\leq 0 \\ \sigma_a + \sigma - j_S^E + \sigma_t &\leq 0 \end{aligned} \quad (54)$$

For all places on the earth, and at all times

For *sustainability*, equation (54) requires that the net rate of increase of entropy of everything at every place on the earth and for all times be less or equal to zero! Note that the process and energy transport increase the entropy of the earth and the atmosphere, and low-temperature heat radiation decreases it.

Remark 15 As derived, Eq. (54) is quite deceiving. The anthropogenic part of the thermal entropy generation rates, $\sigma + \sigma_a$, can only be sustainable if this entropy is generated in cycles in which all process materials are completely recycled, and all chemical entropy is transformed into thermal entropy. If we rely on a finite stock of fossil energy, then even if the entropy generation rate in a process is much lower than the entropy export flux, the process is *never* sustainable. Therefore, sustainability can only be discussed in the context of *cyclic processes*. \square

⁷⁶(Keenan, 1951), Eqs. (31) - (36).

⁷⁷In particular, Volume II of STODOLA's monograph (1927), pp. 1271-1330, brings a thorough discussion of *The Highest Possible Conversion into Work on the Basis of the Second Law of Thermodynamics*.

Over an arbitrary time interval $[\tau_1, \tau_2]$, we can write the global condition of *sustainability* of all cyclic processes on the earth as

$$\underbrace{[S_a(\tau_2) - S_a(\tau_1)]}_{\substack{\text{Increase of atmospheric entropy} \\ \text{from earth processes}}} + \underbrace{[S(\tau_2) - S(\tau_1)]}_{\text{Increase of earth entropy}} - S^E(\tau_1, \tau_2) + S_t(\tau_1, \tau_2) \leq 0 \quad (55)$$

For the entire earth, and arbitrary τ_1, τ_2

where $S^E(\tau_1, \tau_2) = j_S^E A_e(\tau_2 - \tau_1)$ is the total thermal entropy exported by the earth over the time interval $[\tau_1, \tau_2]$, and $S_t(\tau_1, \tau_2) = \sigma_t A_e(\tau_2 - \tau_1)$ is the corresponding thermal entropy generated in the atmosphere by the steady-state energy transport to the universe.

C Is Economic Sustainability Possible?

I have already demonstrated that anything sustainable on the earth must involve a closed material loop (a cycle) which is driven by the sun. Solar energy is like a fine mist: it dribbles at a constant rate on every square inch of the earth's surface. We cannot increase its flow rate, it dries up overnight, and we cannot store it easily. So one may ask legitimately: is solar power enough to drive the world economy, and if so what constraints does it impose?

The time- and area-average solar energy flux in the U.S. is $\sim 200 \text{ W/m}^2$ at the earth surface. At more moderate latitudes this flux goes down to about $100\text{-}150 \text{ W/m}^2$, and in the tropics it can be as high as 300 W/m^2 . Thirty percent of the earth's surface, $153 \times 10^6 \text{ km}^2$, is land. Suppose that 1% of land mass on the earth were covered with solar cells that were 10% efficient, and received only 100 W/m^2 of solar energy. The total solar power converted by these cells to electricity would then be 1.5×10^4 gigawatts ($1\text{GW} = 10^9\text{W}$). Currently, the mean world power consumption is about 1.3×10^4 GW. Therefore covering one percent of the earth's land with solar cells would deliver power equal to the mean world power consumption. For the U.S., it would take closer to 3% of the land area, or $300,000 \text{ km}^2$ (the area of Poland) to replace the entire mean fossil and nuclear power.

There are problems, however, with these calculations:

1. With the necessary infrastructure, the surface area of the solar power plants would probably double or triple.
2. Mass production of solar cells will generate associated chemical waste (high chemical entropy, which cannot be exported by the earth).
3. There are few big transmission lines in the hot barren areas of the world.
4. The peak power consumption in the world can be several times the mean power, and power is needed at night.
5. A significant portion of the world power (over 1/3 in the U.S.) is used by automobiles.
6. Large-scale manufacturing of solar cells does not exist (Hayden, 2002).
7. At 1 U.S. dollar per watt (1/7-1/5 of the current U.S. price, given sufficient manufacturing capacity), it would take 15 trillion dollars to replace most of the fossil and nuclear energy worldwide with solar cells. This amount would probably double if massive new transmission grids and the required transformers were to be built worldwide.

The first and the last item above suggest that decentralization of power generation is necessary. With small and micro power plants, one avoids access roads, big transformers, and long transmission lines. Covering large areas with solar cells is also impossible in the mountains and forests.

The second item requires an intensive research and development effort to weigh the solar cell efficiency and reliability versus the production costs, the use of exotic rare-earth elements, and the generation of chemical waste.

The fourth item dictates a dramatic adjustment of industrial activity: (1) Most factories would have to shut down before sunset. (2) For winter and nocturnal use, energy could be stored as hydrogen from solar-driven water electrolysis. (3) The industrial world would have to slow down, and reassess its social priorities, in sharp contrast to my introductory remarks in Part II.

Solar Cells vs. Biofuels. The superiority of solar cells over *any* biofuel can be illustrated as follows. The cumulative energy balance in Part I shows that the *unsustainable* industrial corn

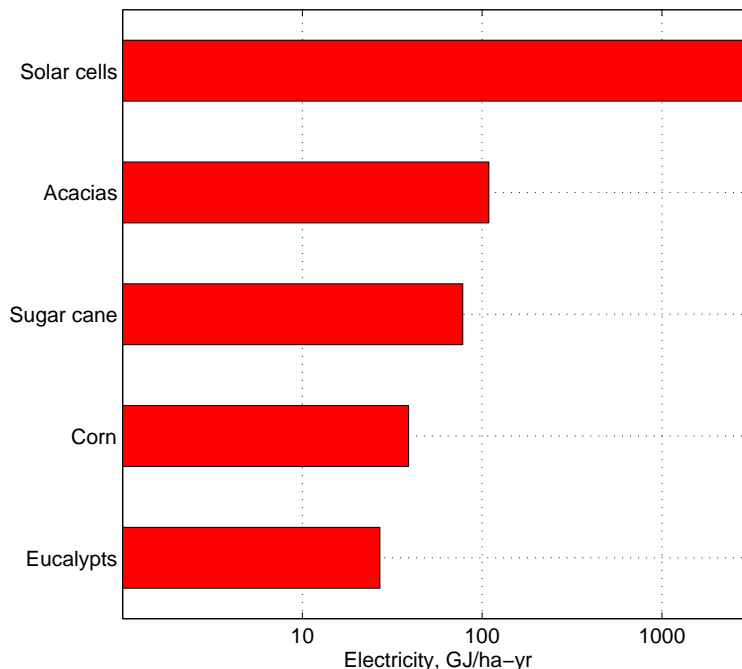


Figure 48: The amount of electricity generated in 1 year with solar cells on 1 ha dwarfs those generated with hydrogen-burning 60%-efficient fuel cells or in power stations. The hydrogen is obtained with no losses from ethanol from corn and sugarcane, and electricity is generated with the 35%-efficiency from eucalypt and acacia wood pellets. The annual electricity yields have been calculated by PATZEK & PIMENTEL (2006). The *minimum* solar cell yield follows the assumptions in this appendix. Note that the bar scale is *logarithmic*, and annually the solar cells produce ~ 100 times more electricity than corn ethanol.

farming captures about 100 GJ/ha per crop that lasts, say, 120 days. Because soybean or another legume is planted afterwards, the annualized energy yield from corn remains the same. The product, corn grain, is a fossil fuel which is later converted into another fossil fuel, ethanol, at a large expense of fossil energy. The ethanol yield is about 74 GJ/ha-yr, and the energy conversion efficiency of a fuel cell is 60%. Thus, corn can deliver $74 \times 0.60 = 44$ GJ/ha-yr as electricity. The fossil fuel inputs exceed the calorific value of the ethanol by about 2-20 GJ/ha-yr. In Part III, I showed that a fuel cell can deliver about 46 GJ/ha-yr as electricity from the industrial corn-ethanol cycle, but the restoration work is at least twice as large. During one year, an inefficient solar cell that converts only 10% of only 100 W/m² will *sustainably* produce 3000 GJ/ha of electricity, see **Figure 48**. Therefore, the solar cells described above can produce 100 times more electricity than the corn ethanol(hydrogen)-burning fuel cells. One hectare of solar cells placed anywhere can therefore free 100 hectares of fertile agricultural land from industrial corn, and allow for the low-intensity, diversified, and almost sustainable agriculture. The socially and scientifically defensible choice is clear⁷⁸.

⁷⁸Some, e.g., (Sheehan et al., 2004), argue that corn should be farmed continuously (with no crop rotation): the whole plants are harvested, the grain is processed for ethanol, and the rest is burned for energy. This energy supply scheme requires 2-10 times more field chemicals (see Part I), irreversibly depletes and damages the soil, and is even *more* unsustainable than a single corn crop (see Parts II and III).

D Efficiency of a Fuel Cell System

In their Science paper, DELUGA et al. (2004) claim the following:

... Further, combustion used for transportation has ~20% efficiency as compared with up to 60% efficiency for a fuel cell. . . The efficiency of these processes for a fuel cell suggests that it may be possible to capture >50% of the energy from photosynthesis as electricity in an economical chemical process that can be operated at large or small scales. (p. 996).

Following DELUGA et al.'s paper and common chemical engineering knowledge, I will assume the following:

1. The catalyst is made of a rare-earth metal, rhodium⁷⁹, and a Lanthanoid, cerium⁸⁰.
2. The catalytic reaction has 100% selectivity and >95% conversion efficiency. I will assume here the conversion efficiency $\eta_1 = 0.96$.
3. The ethanol-water reactant mixture for the catalytic conversion to hydrogen is ultra pure (on the order of 99.9999% pure C₂H₅OH and H₂O). Otherwise, any and all of the impurities listed in Footnotes 81 and 82, and carbon deposition, will destroy the catalyst and reactor. Therefore, one must separate all impurities from the 8 – 10% alcohol⁸¹-yeast-*rest* beer⁸² by multiple distillations, rectification, and molecular sieve exclusion (Maiorella, 1985). As described in Part II, all these processes consume irreversibly large amounts of free energy.
4. The reaction products are hydrogen, carbon dioxide, carbon monoxide, water vapor, plus whatever other impurities. The hydrogen fuel separation from all other components of reactor effluent is almost perfect. Compared with the reference hydrogen of 99.9999999% purity, even tiny amounts of impurities can cause a noticeable drop in performance of a fuel cell⁸³. Some of these impurities are CO (at 5-10 ppm), SO₂ (at 2 ppm), H₂S (at 1-2 ppm), HCHO (at 10-20 ppm), and HCOOH (at 50-100 ppm).
5. In summary, by the time we are ready to use the ethanol-derived hydrogen in a fuel-cell/electric car, we have spent a lot of free energy on picking ethanol, water, and hydrogen molecules; *much* more than on distilling the relatively dirty ethanol to be mixed with gasoline. The latter free energy expenditure is ~15 MJ/L EtOH, or >50% of the ethanol's high heating value.

⁷⁹Rhodium is a precious metal whose price is about US\$30 000/kg, 3×more expensive than gold, <http://www.kitco.com/charts/rhodium.html>.

⁸⁰The nanoparticles of cerium dioxide are called *ceria*, and cost \$250/kg, <http://www.advancedmaterials.us/58N-0801.htm>

⁸¹Alcohol dissolves a large number of substances insoluble in water and acids, such as many inorganic salts, phosphorus, sulphur, iodine, resins, essential oils, fats, coloring matters, etc. (Wright, 1994).

⁸²The beer obtained by mashing and fermenting consists of volatile components, such as water, alcohols (mostly ethyl, but also traces of methyl, propyl, butyl, amyl, and other alcohols – depending on the impurities in glucose), essential oils, and a little acetic acid; and of non-volatile substances, such as cellulose, dextrine (an intermediate product of starch hydrolysis), unaltered sugar and starch, mineral matter, lactic acid, glycerol, etc.

⁸³Japan Automotive Research Institute (JARI), DOE Hydrogen Codes and Standards, Coordinating Committee Fuel Purity Specifications Workshop, April 26, 2004, http://www.eere.energy.gov/hydrogenandfuelcells/pdfs/fuel_purity_notes.pdf

After BOSSEL (2003a), I will summarize efficiency of a Proton Exchange Membrane (PEM) fuel cell as follows. In fuel cells, gaseous hydrogen is combined with oxygen to water. This process is the reversal of the electrolysis of liquid water and should provide an open circuit voltage of 1.23 V (Volts) per cell. Because of polarization losses at the electrode interfaces the maximum voltage observed for PEM fuel cells is between 0.95 and 1.0 V. Under operating conditions the voltage is further reduced by ohmic resistance within the cell. A common fuel cell design voltage is 0.7 V. The mean cell voltage of 0.75 V may be representative for standard driving cycles. Consequently, the average energy released by reaction of a single hydrogen molecule is equivalent to the product of the charge current of two electrons and the actual voltage of only 0.75 V instead of the 1.48 V corresponding to the hydrogen high heating value⁸⁴. Therefore, in automotive applications, PEM fuel cells may reach mean voltage efficiencies of

$$\eta_2 = \frac{0.75 \text{ V}}{1.48 \text{ V}} = 0.50 \quad (56)$$

However, there are more losses to be considered. The fuel cell systems consume part of the generated electricity. Typically, automotive PEM fuel cells consume 10% or more of the rated stack power output to provide power to pumps, blowers, heaters, controllers, etc. At low power demand the fuel cell efficiency is improved, while the relative parasitic losses increase. The small-load advantages are lost by increasing parasitic losses. Let us assume optimistically that for all driving conditions the net power output of an automotive PEM fuel cell system is about $\eta_3 = 0.9$ of the power output of the fuel cell stack.

Depending on the chosen drive train technology, the DC power is converted to frequency-modulated AC or to voltage-adjusted DC, before motors can provide motion for the wheels. Energy is always lost in the electric system between fuel cell and wheels. The overall electrical efficiency of the electric drive train can hardly be better than $\eta_4 = 0.9$.

By multiplying the efficiency estimates, one obtains for the maximum possible tank-to-wheel efficiency of a hydrogen fuel cell vehicle

$$\eta = \eta_1 \eta_2 \eta_3 \eta_4 = 0.96 \times 0.50 \times 0.90 \times 0.90 = 0.38, \quad (57)$$

or 38%. This *optimistic* estimate agrees with another analysis (31-39%) (Fleischer and Ørtel, 2003), and is significantly less than the 60% used by the reckless promoters of a hydrogen economy and hydrogen fuel cell vehicles.

⁸⁴According to Faraday's Law, the standard enthalpy of combustion of hydrogen, $\Delta H_f^0 = -285.9 \text{ kJ/mol}$, can also be expressed as an electrochemical potential ("standard potential") $U^0 = -\Delta H_f^0/n_e F = 1.48 \text{ V}$ with $n_e = 2$ being the number of electrons participating in the conversion and $F = 96485 \text{ Coulomb/mol}$ the Faraday constant.

E Cumulative Exergy Consumption in Steel Production

The primary commercial iron ores in the world are hematite (Fe_2O_3) and magnetite (Fe_3O_4). Taconite, the principal iron ore mined in the United States, has a low (20 – 30%) iron content and is found in hard, fine-grained, banded iron formations. About 99% of iron ore is used in the iron and steel industry. Scrap can be considered a supplement to iron ore in the steelmaking process but is limited as a major feed material owing to inadequate supply of high-quality scrap. Alternatives, such as direct reduced iron (DRI), are also available, and their use continues to grow.

Ore is put into a blast furnace and smelted to produce molten iron, which is then converted to steel by removing most of the remaining carbon in a basic oxygen furnace (BOF). Almost all molten iron goes directly to the BOF, eliminating the molds. The blast furnace product is usually referred to as pig iron.

Iron ore consumption in the U.S. in 2003 was 61 million metric tonnes (Mt), a rise of slightly more than 1 Mt from that of 2002. There was an average of 30 blast furnaces active during 2003, up slightly from that of 2002 when the average number of blast furnaces operating was 29, the lowest since 1961. Accordingly, pig iron production at 40.6 Mt in 2003 was slightly above that of 2002, which had been the lowest since 1982. Crude steel production at 94 Mt increased by 2% compared with that of 2002. Steel demand remained constant at revised 2002 levels of 107 Mt. The large difference between ore production and steel demand is explained by examining the minimill sector and net imports of iron ore substitutes. In 2003, the minimill sector of the steel industry produced more than 50% of the crude steel in the United States. Minimills do not use iron ore as feedstock; instead they use iron and steel scrap, and some DRI in electric arc furnaces (EAF). For a more detailed summary, see JORGENSEN & KIRK (2003).

Several literature estimates of primary energy consumption in producing steel via integrated U.S. mills and BOF are listed in **Table 26**. The same-source estimates of primary energy consumption in producing steel from 100% scrap in EAF are listed in **Table 27**. The embedded energy and equivalent CO_2 emissions in SAE 1045 and SAE 15V45 steel are listed in **Table 28**.

E.1 Steel component manufacturing

As an example, the track roller shaft in a Caterpillar tractor has been investigated in (Beloff et al., 2004). This track roller is conventionally produced with medium carbon steel, typically SAE 1045 or modified SAE 5038. Due to physical and cost benefits at least within the “shape and treat” manufacturing stage of the lifecycle, the part is now produced with microalloyed steel, namely SAE 15V45.

In addition to the primary energy contained in the energy sources, exergy includes work that is available in other materials that contribute to the lifecycle. For simplicity, only iron and minerals that make up the compositions of the steels have been included in (Beloff et al., 2004) in addition to the primary energy, i.e., wastes and auxiliary materials have been ignored. With these assumptions, the primary energy embedded in the track roller is 30 and 27 MJ/kg-part, respectively, for SAE 1045 and SAE 15V45 steels. The standard chemical exergy embedded in these two steels is 6.7 MJ/kg. The incomplete cumulative exergy consumption (CEXC) is 37 and 34 MJ/kg-part respectively. The GHG emissions are about 2.7 kg CO_2 equiv./kg-part. Other, apparently more complete estimates of cumulative exergy consumption in steelmaking (Szargut et al., 1988) are listed in **Table 29**. They range from 46 to 84 MJ/kg of finished steel products.

Remark 16 The average cumulative exergy consumption per kg of steel in machinery is between 34 and 84 MJ/kg. Taking into account the extensive infrastructure this machinery uses to function,

and the concomitant environment pollution, the steel-only estimate should at least double (Wall, 1997) to 68-168 MJ/kg. □

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Part VI

Tables

Table 1: Average dry mass composition of corn grain (White and Johnson, 2003)

Component	% by mass
Starch	66
Oil	3.9
Gluten feed (21% protein)	24
Gluten meal (60% protein)	5.7
Losses	0.4

Table 2: Average application rates of corn field chemicals in 2001 (NASS, 2002)

Compound	True kg/ha	Fraction ha w/ applied	Mean kg/ha
N	148.8	0.96	142.8
P ₂ O ₅ ^a	62.5	0.79	49.4
K ₂ O ^a	93.5	0.65	60.8
Herbicides	2.54	0.98	2.49
Insecticides	1.08	0.29	0.31

^aUSDA (NASS, 2002) reports “P” and “K” but, according to Ms. BARBARA TIDWELL of the NASS/MISO Customer Service, they mean P₂O₅ and K₂O.

Table 3: Specific energy consumption and application rates of nitrogen fertilizer

Active Ingredient	Specific Energy MJ/kg	Application Rate kg/ha	Source
N	63.43	148.0	Pimentel, 2003
N	54.43	148.8	Patzek, 2004
N	43.00	140.0	Shapouri et al., 2002
N	49.06	153.0	Wang et al., 1997
N	54.43 ^a	150.0	Berthiaume et al., 2001

^a BERTHIAUME et al. (2001) do not give the specific N, P, K, Ca fertilizer energies, only the specific exergies. I have *assumed* that the specific energies of BERTHIAUME et al. are equal to my estimates, and used their reported application rates.

Table 4: Energy consumption in superphosphate production (Kongshaug, 1998)

Process	MJ/kg P ₂ O ₅
Phosphate mining	> 0.3
Apatite mining	2.9
Dihydrate process	2.5
Hemihydrate process	6.5

Table 5: Specific energy consumption and application rates of phosphorus fertilizers

Active Ingredient	Specific Energy MJ/kg	Application Rate kg/ha	Source
P ₂ O ₅	17.44	53.0	Pimentel, 2003
P ₂ O ₅	6.80	62.5	Patzek, 2004
P ₂ O ₅	4.76	54.0	Shapouri et al., 2002
P ₂ O ₅	11.40	56.0	Wang et al., 1997
P ₂ O ₅	6.80	55.0	Berthiaume et al., 2001

Table 6: Energy consumption in potassium fertilizer production (Kongshaug, 1998)

Fertilizer	K%	MJ/kg K ₂ O
Chloride	52	6.8 (additive to phosphates)
Sulphate	49	Not reported
Nitrate	45	13.5 (KNO ₃ solution evaporation) 43 MJ/kg N

Table 7: Specific energy consumption and application rates of potassium fertilizer

Active Ingredient	Specific Energy MJ/kg	Application Rate kg/ha	Source
K ₂ O	13.77	57.0	Pimentel, 2003
K ₂ O	6.80	93.5	Patzek, 2004
K ₂ O	8.71	85.0	Shapouri et al., 2002
K ₂ O	5.30	66.0	Wang et al., 1997
K ₂ O	6.80	85.0	Berthiaume et al., 2001

Table 8: Specific energy consumption and application rates of calcinated lime

Active Ingredient	Specific Energy MJ/kg	Application Rate kg/ha	Source
CaO	1.33	699.0	Pimentel, 2003
CaO	1.75	333.0 ^a	Patzek, 2004
CaO	1.70	276.0	Shapouri et al., 2002
CaO	1.70	276.0 ^b	Wang et al., 1997
CaO	1.75	270.0	Berthiaume et al., 2001

^aSHAPOURI et al's data (Table 2 in (Shapouri et al., 2002a)) with the two zero entries omitted.

^bWANG et al. (1997) does not report lime use, I have corrected their estimates by adding the lime use reported by SHAPOURI et al. (2002a).

Table 9: Specific energy consumption and application rates of herbicides

Specific Energy MJ/kg	Application Rate kg/ha	Source
422.00	2.10	Pimentel, 2003
261.00	2.54	Patzek, 2004
261.00	4.73	Shapouri et al., 2002
237.30	3.07	Wang et al., 1997

Table 10: Specific energy consumption and application rates of insecticides

Specific Energy MJ/kg	Application Rate kg/ha	Source
422.00	0.15	Pimentel, 2003
268.40	1.08	Patzek, 2004
268.40	0.22	Shapouri et al., 2002
243.00	0.22	Wang et al., 1997

Table 11: Average high and low heating values of some fossil fuels from (Castorph et al., 1999; Bossel, 2003b; Spiers, 1961). Other sources are listed in the footnotes.

Fuel	Density kg/sm ³	HHV ^a MJ/kg	LHV ^a MJ/kg	HHV ^a MJ/kg	LHV ^a MJ/kg	Source ⁱ
Gasoline	720-800	46.7 ^b	42.5 ^b	46.8	43.6	Table 339
Diesel fuel	840	45.9	43.0	45.3	42.3	Table 350
Methane	0.66 ^d	55.5 ^c	50.1 ^c	55.1(g)		Table 347
LPG ^e	0.58	50.0	46.0			
NG ^f	0.84	48.7	43.9			
Ethanol	787 ^h	29.7 ^g	26.7 ^g	29.6	26.8	Table 353
Corn grain dry		18.8 ^j				
Corn stover ^k		17.7	16.5			
Corn stalks ^l		15.8	14.8			
Corn meal ^m		16.0				
Corn oil ⁿ	909.5	39.5	38.8			

^aHHV = High Heating Value; LHV = Low Heating Value

^bAverage of gas station fuels, I choose the mean density of 740 kg/m³, The leftmost gasoline, diesel fuel, LPG, and NG data are from (Castorph et al., 1999)

^cFrom BOSSEL, Table A, (Bossel, 2003b)

^d(Lide, 1994), 6-25

^eA mixture of propane and butane, C₃H₈ and C₄H₁₀

^fAn average of natural gas compositions from Groningen, Orenb., Ekofisk, and Leman Bank

^ghttp://bioenergy.ornl.gov/papers/misc/energy_conv.html

^h(Lide, 1994), 15-46

ⁱ(Spiers, 1961)

^jBased on the mean of the values reported by SCHNEIDER & SPRAQUE (1955), p. 496, 2033 kcal/lb; and MILLER (1958), p. 639, 2059 kcal/lb. 1 thermochemical kcal = 4.184 kJ

^kThe mature corn stalks of corn from which the ears have been removed. (Domalski et al., 1987), p. 16

^l(Domalski et al., 1987), p. 17

^m*Zea mays*, 11.79% moisture (Domalski et al., 1987), p. 93

ⁿLiquid vegetable oil (Domalski et al., 1987), p. 223

Table 12: Calorific values and specific volumes of gasoline used in corn farming

Calorific Value MJ/kg	Volume L/ha	Source
56.32	56.0	Pimentel, 2003
46.70	29.0	Patzek, 2004
59.10	29.0	Shapouri et al., 2002
46.70	30.8	Wang et al., 1997

Table 13: Calorific values and specific volumes of diesel fuel used in corn farming

Calorific Value MJ/kg	Volume L/ha	Source
50.24	90.0	Pimentel, 2003
45.90	80.0	Patzek, 2004
53.97	82.1	Shapouri et al., 2002
45.90	74.3	Wang et al., 1997
45.90	71.0	Berthiaume et al., 2001

Table 14: Calorific values and specific volumes of Liquefied Petroleum Gas (LPG) used in corn farming

Calorific Value MJ/kg	Volume L/ha	Source
50.00	47.0	Patzek, 2004
49.18	59.0	Shapouri et al., 2002
50.00	28.7	Wang et al., 1997
50.00	252.0 ^a	Berthiaume et al., 2001

^a Large quantities of LPG are used to dry wet corn in Quebec.

Table 15: Calorific values and specific volumes of methane^a used in corn farming

Calorific Value MJ/kg	Volume sm ³ /ha	Source
55.50	21.3	Patzek, 2004
61.90	14.0	Shapouri et al., 2002
55.50	35.7	Wang et al., 1997

^a I will treat natural gas from the various sources as methane

Table 16: Specific electric energy inputs to corn farming

Amount kWh/ha	Source
39.9	Pimentel, 2003
191.0	Patzek, 2004
207.6	Shapouri et al., 2002
24.1	Wang et al., 1997
132.0	Berthiaume et al., 2001

Table 17: Specific energy used in transportation related to corn farming. Source: WANG et al. (1997)

Category	Plant→ Center	Center→ Mixer	Mixer→ Farm
Travel distance (km one way)	1700/ 850	80	50
Mode	barge/ rail	Class 8b truck	Class 6 truck
Energy use (kJ/kg)	71	25	53

Total energy use is 720 kJ/kg of field input materials or ~400 MJ/ha + Commute of Personnel

Table 18: Ethanol yield, specific energy inputs, and energy credits

Yield L/ha	Yield ^a GJ/ha	Transport MJ/L	Fuel MJ/L	Total Inputs MJ/L	Credits MJ/L	Source
3201	74.82	1.32	14.75	16.07	1.88	Pimentel, 2003
3175	74.22	1.74	14.45	16.19	0.00	Patzek, 2004
3137	73.32	0.44	14.45	14.89	5.89	Shapouri et al., 2002
3189	74.55	0.43	13.96	14.39	5.38	Wang et al., 1997
2406	56.24	0.00	17.18	17.18	0.00	Berthiaume et al., 2001

^aBased on High Heating Value of ethanol

Table 19: Specific CO₂ emissions

NRR	Emission	Units
Electricity	0.322	kg CO ₂ /kWh
Coal	0.089	kg CO ₂ /MJ
Gasoline	0.067	kg CO ₂ /MJ
Diesel	0.069	kg CO ₂ /MJ
LPG	0.059	kg CO ₂ /MJ
Methane	0.050	kg CO ₂ /MJ
Lime	0.785	kg CO ₂ /kg CaO
Nitrogen ^a	2.70	kg CO ₂ /kg N

Sources: Energy Information Administration (EIA, 2002) *Emissions of Greenhouse Gases in the United States 2001*, Table B1

^a G. KONGSHAUG (1999)

Table 20: The specific chemical exergies of compounds participating in the ideal corn-glucose-EtOH cycle. Source: SZARGUT et al. (1988), Tables I and II in Appendix.

Component	Standard Chemical Exergy MJ/kmol
$C_6H_{12}O_6(s)$	2928.8
$C_2H_5OH(l)$	1363.9
$CO_2(g)$	19.87
$O_2(g)$	3.97
$H_2O(l)$	0.00
$H_2(g)$	236.1

Table 21: The calculated chemical exergies of the products of steps in Eqs. (24 – 26).

Step	Product Exergy After Step MJ/kmol $C_6H_{12}O_6$
1	2952.6
2	2791.4
3	119.2
Net	2833

Net chemical exergy per kg of glucose = 15.74 MJ/kg

Exergy of matter after Step 1 per kg of glucose = 16.4 MJ/kg = 3920 kcal/kg

Cycle efficiency = 96%

Table 22: The calculated chemical exergies of the products of steps 1-3 of the ideal corn-ethanol-hydrogen cycle.

Step	Product Exergy After Step MJ/kmol $C_6H_{12}O_6$
1	2952.6
2	2791.4
3	2976.2
4	119.2
Net	2809.8

Net chemical exergy per kg of glucose = 15.61 MJ/kg

Cycle efficiency = 95%

Table 23: CExC of major non-renewable resources used in the industrial corn-ethanol cycle. Sources: Table 5.2 in SZARGUT et al. (1988), and calculations by the author.

NRR	Exergy	CExC	Units	Comments
Ammonium Nitrate	10.51	99.6	MJ/kgN	30-years old technology
Phosphate	0.1	10.1	MJ/kg P ₂ O ₅	H ₂ SO ₄ CExC
KCl	0.26	6.09	MJ/kg KCl	Sylvinite ore, 1:1 K:Na
Lime	1.96	10	MJ/kg CaO	Calcinated limestone
Herbicides	261	300	MJ/kg	1.15 exergy
Seeds	104	119	MJ/kg	1.15 exergy
Electricity	3.6	11.83	MJ/kWh	Plant eff. of 34.6%
Diesel	44.4	53.2	MJ/kg	Typical value
Gasoline	48.3	57.5	MJ/kg	Mean value
Natural Gas	50.7	57.9	MJ/kg	Typical value
LPG	48.8	61.6	MJ/kg	Autothermic cracking
Steel	7.1	45.9	MJ/kg	Process ore, blast furnace

Table 24: The First Law summary of the U.S. corn-ethanol production in 2004

29.6 million hectares	of corn harvested in the U.S.
299.67 million tonnes	of moist corn grain harvested
3.8 million hectares	of U.S. cropland growing corn for ethanol
12.7 %	of all U.S. corn is farmed for ethanol
0.435 liters	of ethanol from 1 kg of corn
12.28 GL/yr	of ethanol produced in the U.S.
3.25 billion gal/yr	of ethanol produced in the U.S.
9.21 GL GE/yr	as ethanol produced in the U.S.
10.01 GL GE/yr	burned to produce this ethanol
1.4 %	of U.S. automobile fuel from ethanol
25.9 million hectares	for 10% U.S. automobile fuel energy
\$1.69 billion/yr	in federal subsidies for ethanol
\$0.32 billion/yr	in average state subsidies for ethanol
\$1.27 billion/yr	in corn-for-ethanol price subsidies
\$3.28 billion/yr	in total ethanol subsidies

GL = Giga Liter = 10⁹ L; GE = Gasoline Equivalent

Table 25: The Second Law summary of the U.S. corn-ethanol production in 2004

\$ 1.84 billion/yr	Subsidy by the environment with 20% efficient car engine
\$ 1.59 billion/yr	Subsidy by the environment with 35% efficient car engine
\$ 1.18 billion/yr	Subsidy by the environment with 60% efficient car engine
\$ 12.69 billion/yr	10% of U.S. fuel consumption with 20% efficient car engine
\$ 10.98 billion/yr	10% of U.S. fuel consumption with 35% efficient car engine
\$ 8.13 billion/yr	10% of U.S. fuel consumption with 60% efficient car engine

Table 26: Primary energy consumption in steel production via integrated mills (Beloff et al., 2004)

Source	Processes Reported	Energy (MJ/kg-steel)	Notes
Young & Vanderburg (1994)	Cokemaking, blast furnace, BOF, rolling	27 (range 18 - 36)	Review of literature data, includes 30% scrap
Fruehan et al. (1994)	Cokemaking, ore agglomeration, blast furnace, BOF, hot rolling	31.35	Reported as "typical" energy consumption; hot rolling for flat steel
BUWAL (1990)	Ore mining, sintering, blast furnace, BOF, hot	26.94	Based on US production 1975-1990; for hot-rolled rolling coils
IDEMAT (2001)	Raw materials extraction, cokemaking, sintering, pelletizing, blast furnace, BOF, oxygen making, casting	18.85	Based on production at CORUS plant in the Netherlands, may include some scrap

Table 27: Primary energy consumption in steel production from 100% recycled scarp via electric arc furnace (EAF) minimills (Beloff et al., 2004)

Source	Processes Reported	Energy (MJ/kg-steel)	Notes
Young & Vanderburg (1994)	EAF, rolling	18 (range 9 - 27)	Review of literature data
Fruehan et al. (1994)	EAF, hot rolling	9.46	Reported as "typical" energy consumption; hot rolling for flat steel
BUWAL (1990)	EAF, hot rolling, electrode production	14.09	Based on US production 1980-1985; for hot-rolled rolled coils
IDEMAT (2001)	Recycling, EAF, casting	23.95	Based on steel alloy produced in the Netherlands

Table 28: Estimates of primary energy embedded in steel and its constituents (Beloff et al., 2004)

Component	Energy (MJ/kg)	GHG Emissions kg CO ₂ -equiv./kg	Source
Steel	22.5	2.1	Young & Vanderburg (1994)
Manganese	70	5.4	(Haasen and Cahn, 1996)
Vanadium	415	33	IDEMAT (2001)

Table 29: Estimates of cumulative exergy consumption (CExC) in steelmaking

Component	Exergy (MJ/kg)	CExC MJ/kg	Notes
Steel billets from ore in the ground	7.1	48.35	(Szargut et al., 1988)
Track roller shaft in a Caterpillar tractor 50% ore in the ground and 50% scrap	6.7	34 – 37	(Beloff et al., 2004), material, ore extraction and steelmaking, shape & treat, no accounting for wastes and utilities
Steel forgings from ore	7.1	83.75	(Szargut et al., 1988)
Steel (finished general products) from ore in the ground	7.1	45.9	(Szargut et al., 1988), ore processing, blast furnace, BOF, continuous casting and finishing
Steel (finished general products) from ore in the ground	7.1	45.9	(Szargut et al., 1988), ore processing, direct reduction, EAF, continuous casting and finishing