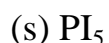
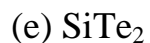
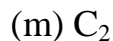
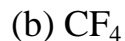
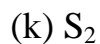


Name: _____

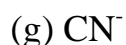
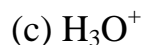
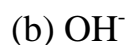
Period: ____

Bonding and Lewis Structures

(1) Draw Lewis structures for the following compounds.



(2) Determine the total number of electrons in each of the following polyatomic ions. Draw Lewis structures and calculate the formal charge on each atom.



(3) Draw all possible resonance structures for the polyatomic ion. Calculate the formal charge on each atom.



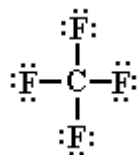
Answers

(1)

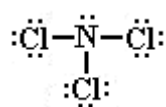
(a)



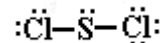
(b)



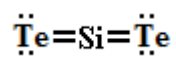
(c)



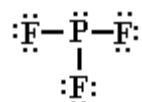
(d)



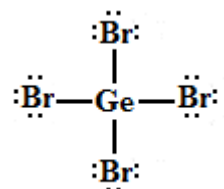
(e)



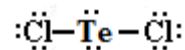
(f)



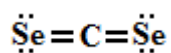
(g)



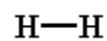
(h)



(i)



(j)



(k)



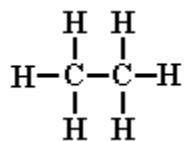
(l)



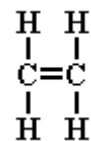
(m)



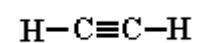
(n)



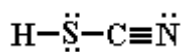
(o)



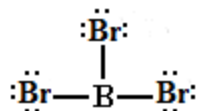
(p)



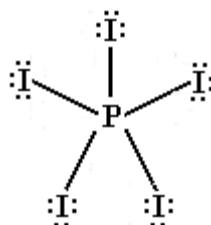
(q)



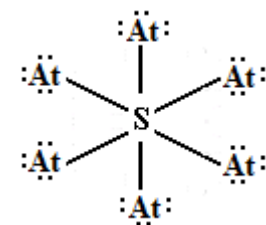
(r)



(s)

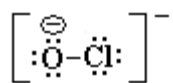


(t)

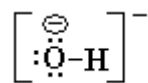


(2)

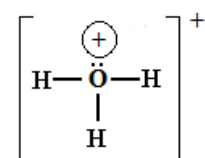
(a) 14 e⁻



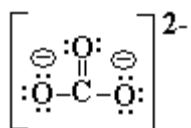
(b) 8 e⁻



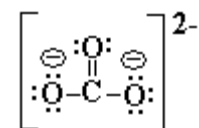
(c) 8 e⁻



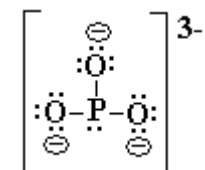
(d) 24 e⁻



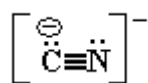
(e) 24 e⁻



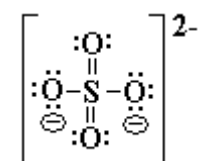
(f) 26 e⁻



(g) 10 e⁻

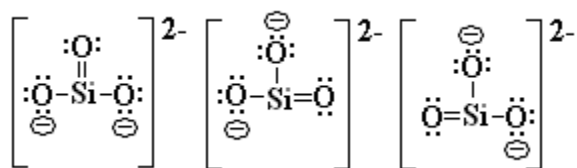


(h) 32 e⁻



(3)

(a)



(b)

