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THERMAL ANALYSIS STUDY OF ANTIMONY SULFIDES

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Lt Charles K. Kelley
Lubrication Branch
Fuels and Lubrication Division

July 1989

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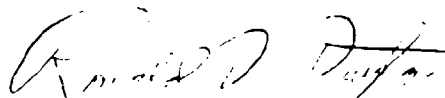
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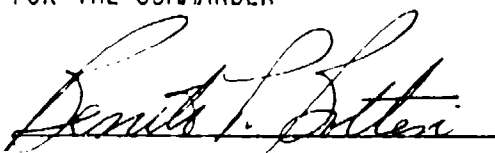


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PREFACE

This technical report was prepared by the Lubrication Branch, Fuels and Lubrication Division, Aero Propulsion and Power Laboratory (APPL), Wright Research and Development Center (WRDC), Air Force Systems Command, Wright-Patterson Air Force Base, Ohio 45433-6563. The work herein was accomplished under Project 3048, Task 304806, Work Unit 30480626, "Turbine Engine Lubricant Research," during the period of August 1987 to February 1989 with Lt Charles K. Kelley as Project Engineer. Special acknowledgment is given to Dr Dennis Flentge of Cedarville College for performing the extractions, and to Messrs Edward Pitzer and Richard Homer of WRDC/POXX for providing expert technical support in the areas of FTIR spectroscopy and thermal analysis.

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SECTION I INTRODUCTION

Antimony sulfides have received attention in such areas as semiconductors, paints, polymers and lubrication. They exist in a number of different stoichiometry - crystalline Sb_2S_3 , amorphous Sb_2S_3 , Sb_2S_4 and Sb_2S_5 . Crystalline Sb_2S_3 , stibnite, is a black material consisting of ribbon-like $(Sb_4S_6)_n$ chains. Several studies on the electrical conductivity of Sb_2S_3 have been carried out in relation to usage in semiconductor systems (1 & 2). The amorphous form of Sb_2S_3 is orange in color and has been said to contain excess sulfur. The structure of orange Sb_2S_3 , somewhat ambiguous, was examined in a recent study (3) by x-ray powder diffraction and found not to be entirely amorphous. Thermogravimetric studies on amorphous Sb_2S_3 , reveal a weight loss between 130 and 190 °C. This loss has been attributed to both removal of water and sulfur, and to sulfur oxidation (4). A transition from orange to black Sb_2S_3 occurs when the amorphous form is heated under nitrogen to 270 °C. Treatment with dilute hydrochloric acid will also transform the orange form to black.

Antimony tetrasulfide, brick-red in color and amorphous in structure (6) has been studied as an extreme pressure and an antiwear additive to greases (7). It has also been looked at as a

bonded solid lubricant (8). Thermally, Sb_2S_4 is reported to melt at $510\text{ }^\circ\text{C}$ and when heated for extended periods of time in nitrogen at $525\text{ }^\circ\text{C}$, give off sulfur to form Sb_2S_3 .

The pentasulfide is reddish orange in color. As with amorphous Sb_2S_3 , the pentasulfide contains excess elemental sulfur. Actually, there is some doubt as to the existence of the pentasulfide. Long et al., (5) examined stibnite, amorphous Sb_2S_3 , and Sb_2S_5 by Mossbauer spectroscopy. Their results indicated that in all cases, antimony was present in the +3 oxidation state. Also, there was little difference in the isomer shifts or the widths at half maximum between amorphous Sb_2S_3 and Sb_2S_5 .

Our laboratory has been involved in a study of tribological properties of solid lubricant compacts made from mixtures of various antimony sulfides and molybdenum disulfide (9). Preliminary results have indicated that the antimony sulfides can improve the tribological performance of MoS_2 ; however, the beneficial effects of the sulfides show a temperature dependence. This result has stimulated the following investigation of the thermal and oxidative properties of Sb_2S_3 , Sb_2S_4 , and Sb_2S_5 .

SECTION II EXPERIMENTAL

1. Chemicals: Antimony trisulfide and antimony pentasulfide were obtained from Fisher Scientific (reagent grade) and used as received. Antimony tetrasulfide was obtained from Dr J. King of Penwalt Corp and also used as received. The oxides of antimony--trioxide, tetra-oxide and penta-oxide were from Alpha Chemicals (ultrapure 99.99 percent). The dimorphs of the trioxide were obtained by the following method. Samples of the trioxide were placed in a platinum cylinder, sealed in quartz under argon (0.2 atm) and heated to a specific temperature for 2 hours then rapidly cooled to room temperature. To obtain the cubic form, the sample was heated to 590 °C and the orthorhombic 620 °C. Reagent grade sulfur was purchased from Fisher Scientific. All materials were sieved through a 30-micron filter to ensure uniform particle size.

2. Infrared Spectroscopy: Infrared data were collected on a Nicolet 740 spectrometer. Samples were run as polyethylene disks. Each sample was run for a total of 64 scans at 4 cm⁻¹ resolution.

3. Thermal analysis: Thermal experiments were performed on a Dupont 2100 thermal analysis system consisting of thermogravimetric analysis (TGA), differential scanning calorimetry (DSC), and high temperature differential thermal

analysis (DTA). In TGA experiments, sample sizes were kept between 20 and 25 mg, and run in platinum pans under dynamic atmospheres (200 ml/min). DSC was performed on samples ranging from 3 to 4 mg for Sb_2S_5 , Sb_2S_4 , and the antimony oxides. Due to the increased density of Sb_2S_3 , sample sizes were between 5 and 6 mg to insure complete coverage of the sample pan. Samples were run in aluminum hermetically sealed pans with inverted lids. Pinholes were placed in the lids to allow equilibration with the DSC environment. Flow rates were kept high (250 ml/min) to purge the system of corrosive gases formed during heating. Samples run under nitrogen were allowed to equilibrate for a minimum of 15 minutes. For DTA experiments, alumina sample cups were used with sample sizes kept between 10 and 20 mg. In all thermal experiments, unless otherwise stated, temperature scan rates were 20 °C per minute.

SECTION III
RESULTS AND DISCUSSION

1.0 Antimony trisulfide (Sb_2S_3)

Antimony trisulfide is a black solid, thermally stable to 575°C under an inert atmosphere as evidenced by TGA. (See Figure 1.) At temperatures above 575°C , Sb_2S_3 undergoes a reaction with the TGA's platinum sample pan resulting in a rapid weight loss. When Sb_2S_3 is heated in the presence of air, a series of reactions occur. Four distinct areas of weight change are observed: weight losses between 300 and 425°C ; 425 and 500°C ; 525 and 600°C ; and a slight increase from 600 to 625°C .

Examining the oxidation via high temperature DTA, one finds at least two exotherms - 446 and 604°C . (See Figure 2.) The 446°C peak is asymmetric and has a shoulder at 400°C indicating the presence of multiple overlapping exotherms. This exotherm lies in the temperature region as the first two areas of weight loss (300 and 500°C) in the TGA. The exotherm at 604°C coincides with the third weight loss and the small weight gain in the TGA. This peak agrees with the literature value for the oxidation of Sb_2O_3 to Sb_2O_4 (10). On going from Sb_2O_3 to Sb_2O_4 , one would expect a weight gain in the TGA; however, a loss is observed followed by a small gain and then stabilization. This is explained by a combination of Sb_2O_3 sublimation and oxidation which forms a layer of Sb_2O_4 retarding further oxidation and

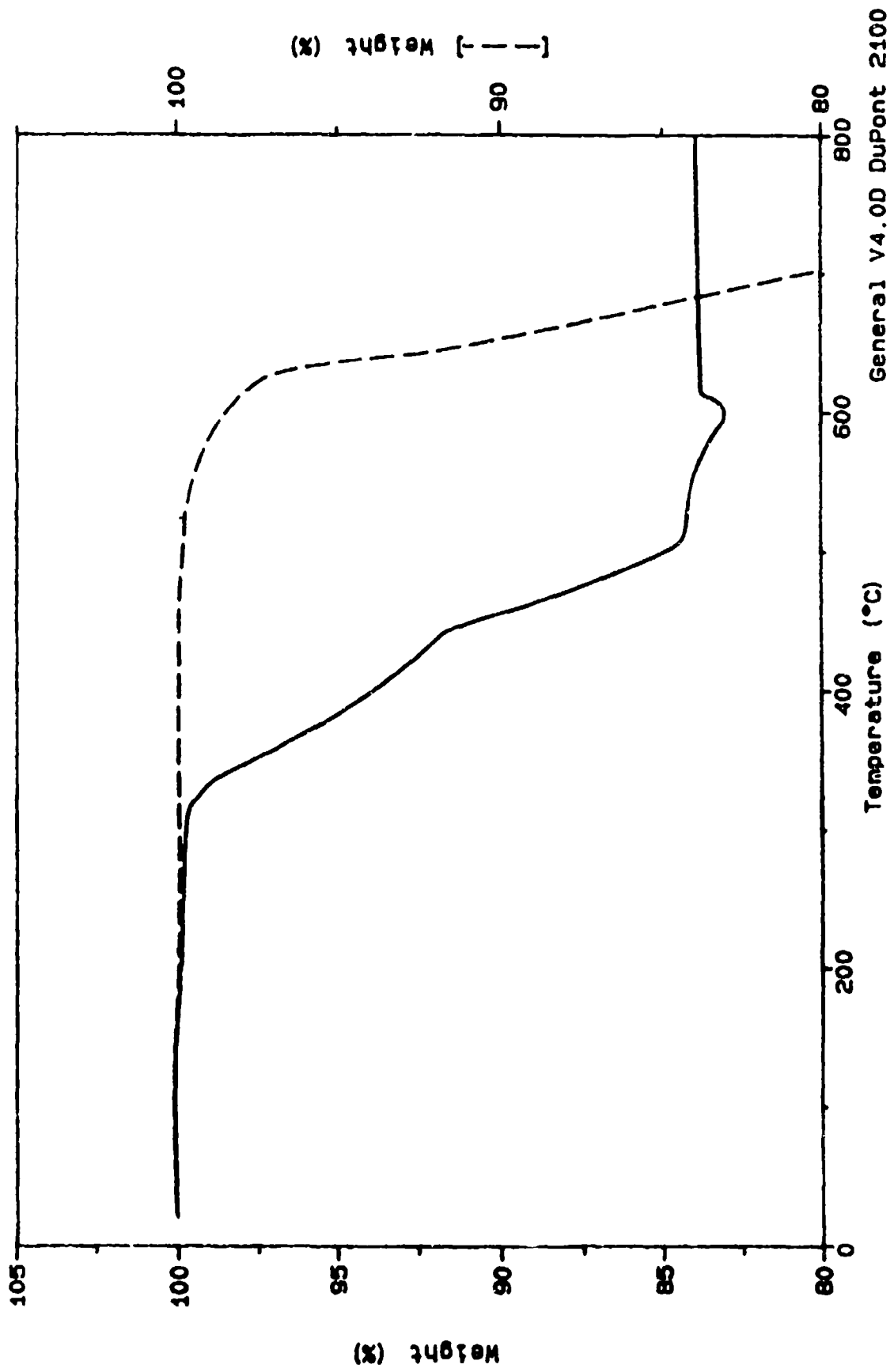


Figure 1. TGA of Sb₂S₃ (—) Air; (---) N₂.

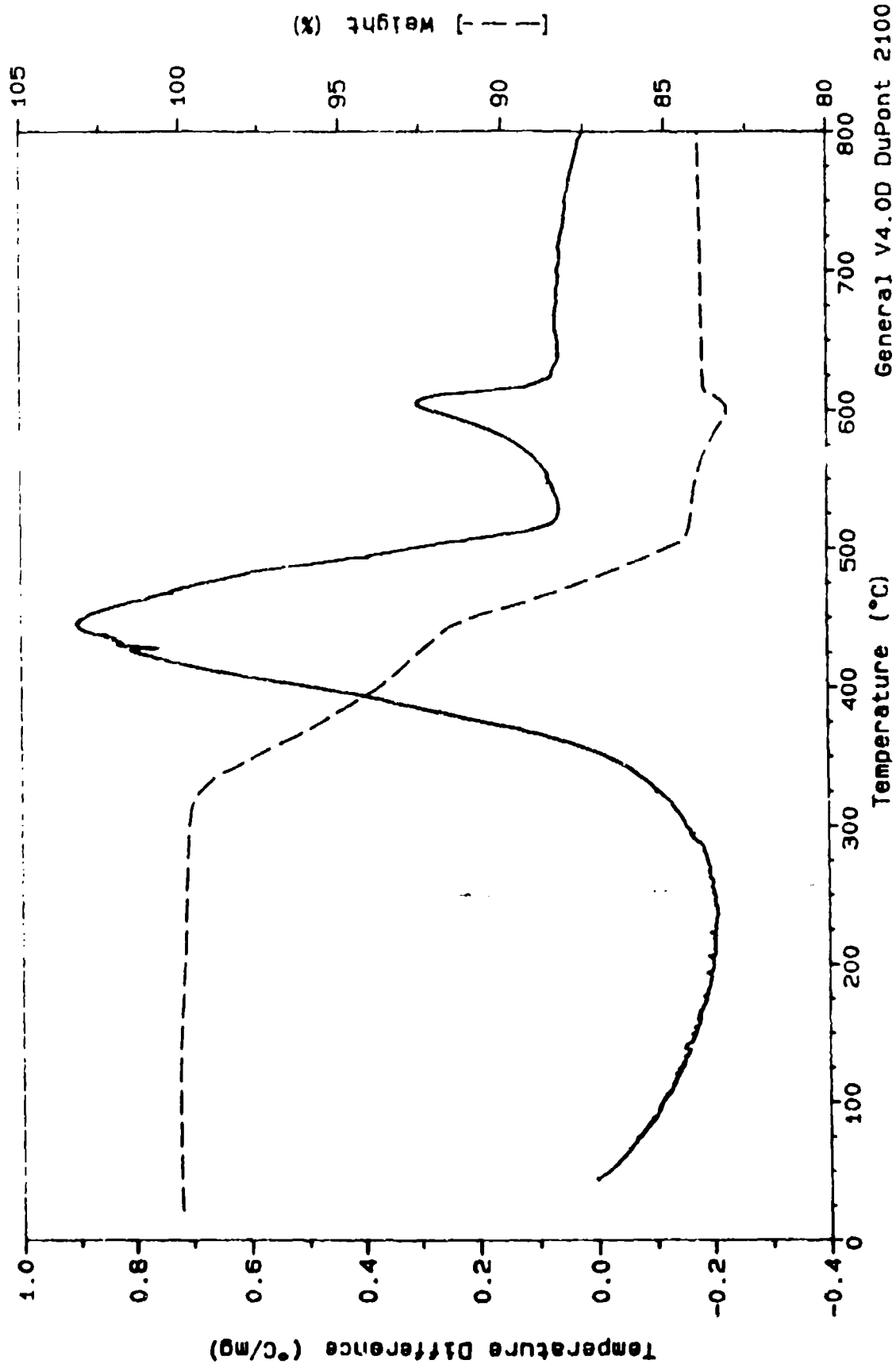


Figure 2. DTA of Sb_2S_3 (—) Air; TGA (---) Air.

sublimation. Centers (11) has shown that the oxidation of cubic Sb_2O_3 occurs by a sublimation controlled mechanism.

Assignment of the 604 °C exotherm to the oxidation of Sb_2O_3 to Sb_2O_4 indicates that the previous exotherm at 400 °C and the weight losses between 300 and 500 °C must correspond to the oxidation of Sb_2S_3 to Sb_2O_3 . The complex nature of both the TGA and the DTA curves points to a multistep process. When the oxidation is examined by DSC - a more sensitive technique at lower temperatures than the DTA (Figure 3), two peaks are observed: 434 °C with shoulders at 360 °C; 407 °C; and 473 °C with shoulders at 484 °C and 495 °C. When a smaller sample size is used (1.8 mg), the peaks are further resolved. (See Figure 4.) The first exotherm shifts to 394 °C with a shoulder at 356 °C, and the higher temperature peak splits into two - 463 and 476 °C with a shoulder at 491 °C.

In an attempt to better understand the oxidation of Sb_2S_3 , a series of isothermal TGAs (3h in air or nitrogen) were run at various temperatures corresponding to different portions of the DSC/DTA curves. The products were then examined by fourier transform infrared (FTIR) spectroscopy. Figure 5 shows the far IR spectra of Sb_2S_3 and Sb_2S_3 thermally stressed at 275 °C in N_2 . Sb_2S_3 has characteristic bands at 336, 274, 242, and 138 cm^{-1} . Thermal stressing results in no observable changes to the IR spectrum.

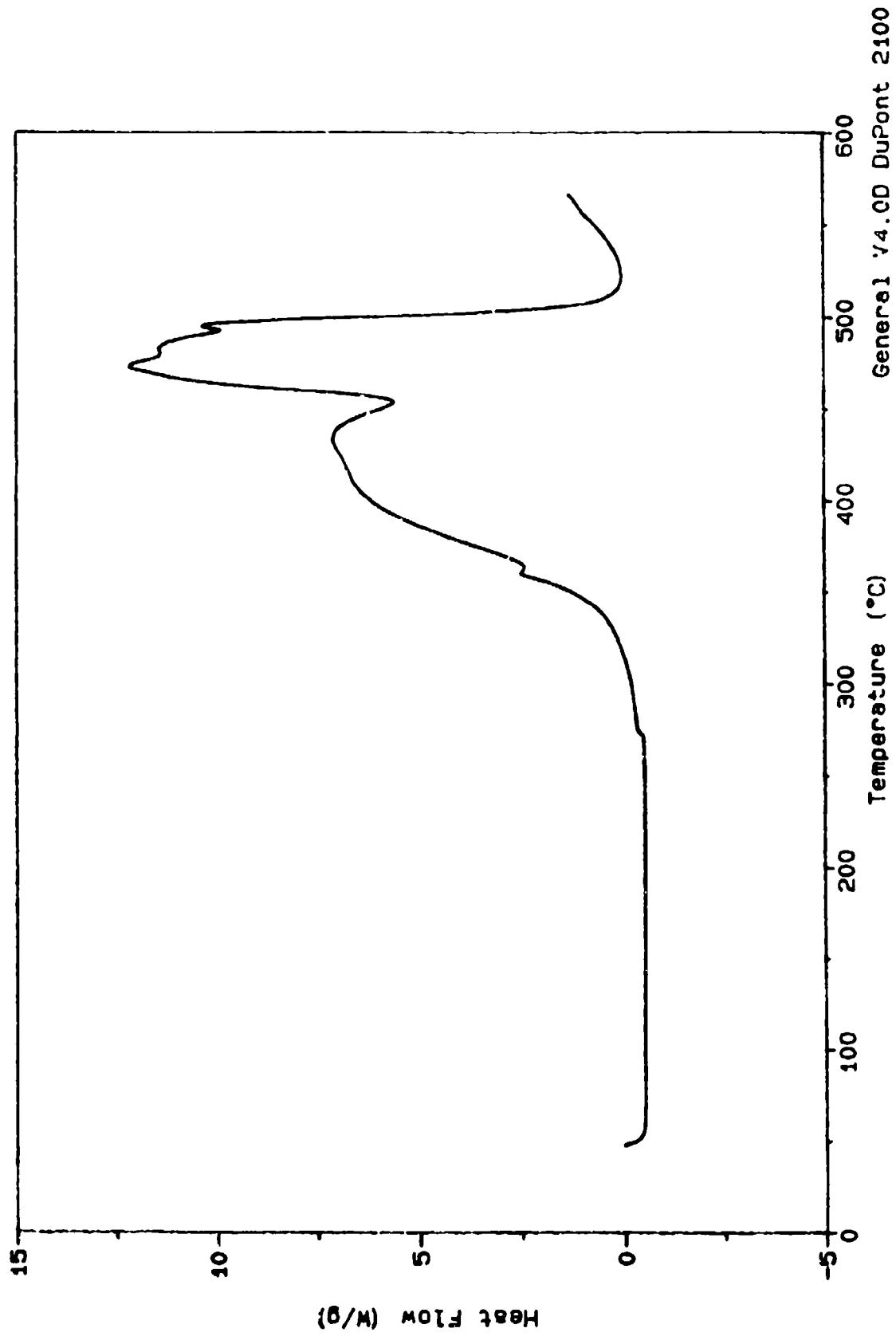


Figure 3. DSC of Sb_2S_3 in Air.

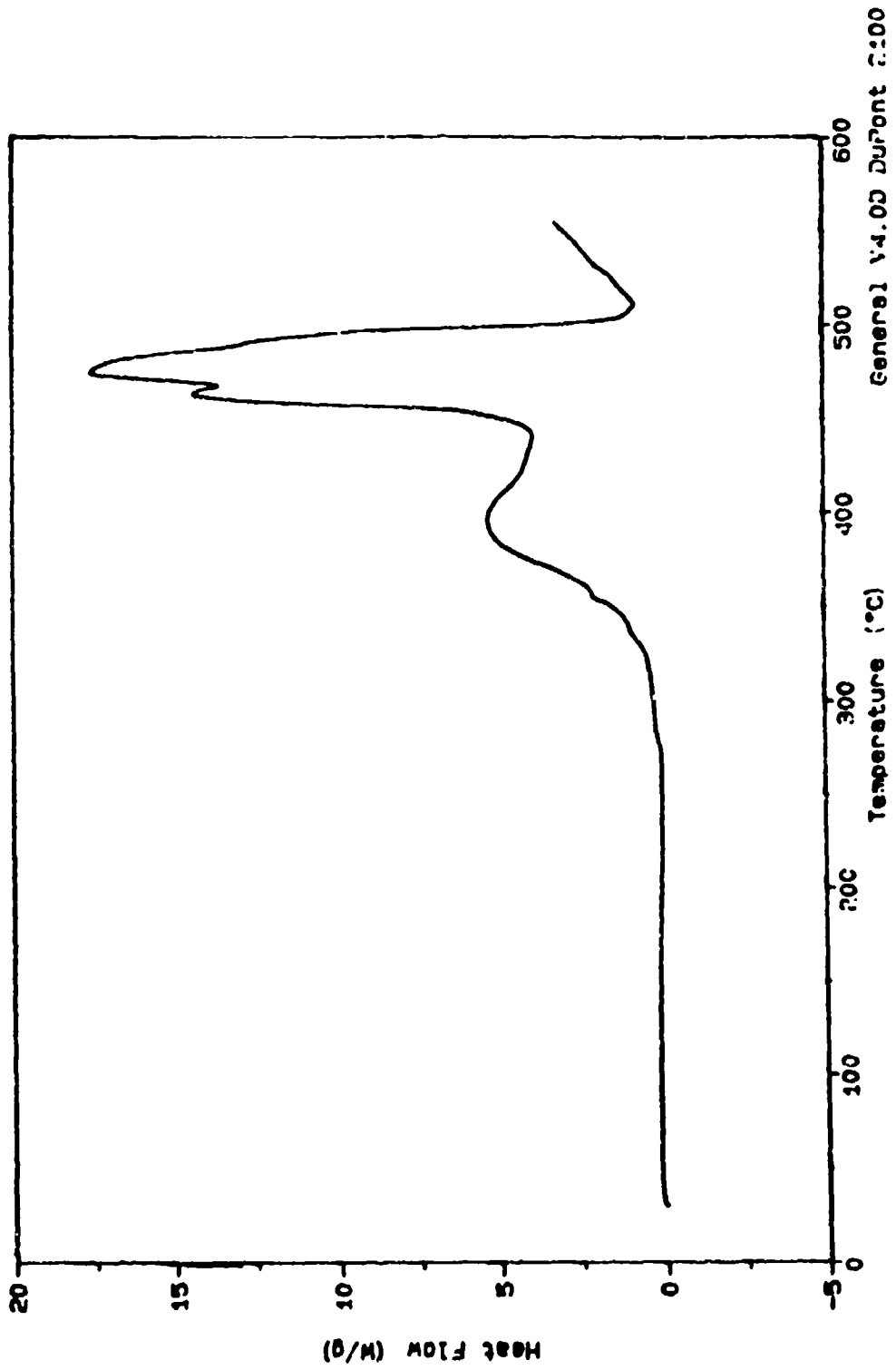


Figure 4. DSC of 1.8 mg Sample of Sb_2S_3 .

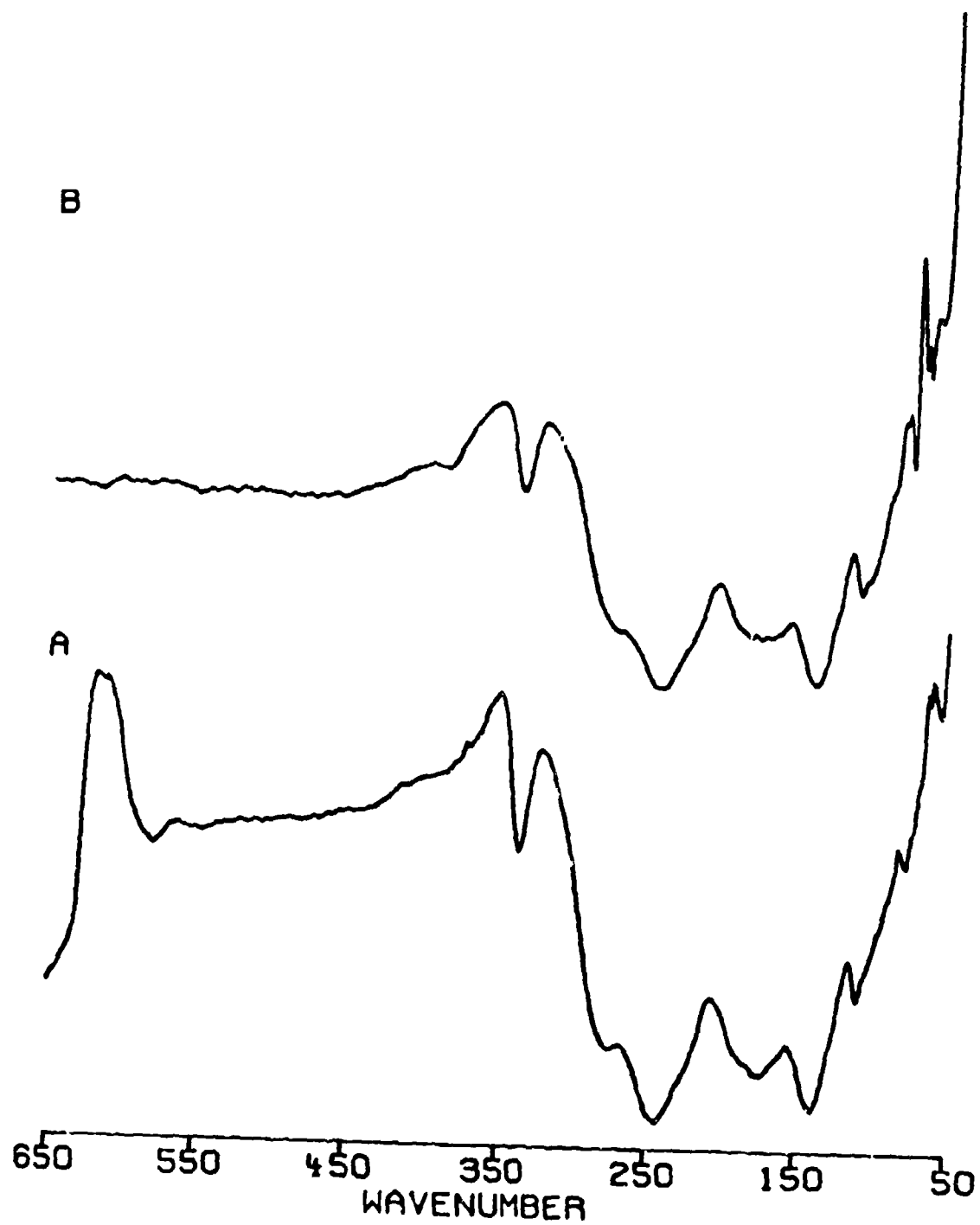


Figure 5. Far IR of Sb_2S_3 (A) Virgin; (B) Heated at 275°C in N_2 .

When Sb_2S_3 is heated in air at 400 and 480 °C, a number of changes occur. (See Figure 6.) The characteristic bands of Sb_2S_3 disappear and are replaced with bands at 384, 259, and 180 cm^{-1} . Heating at 525 °C (Figure 7) produces the same spectrum as the samples heated at 400 and 480 °C; however, heating at 800 °C results in the production of a new material showing bands at 529, 443, 365, and 287 cm^{-1} .

Comparing the spectra of the oxidation products in Figures 6 & 7 with those of the known antimony oxides - cubic and orthorhombic Sb_2O_3 , Sb_2O_4 , and Sb_2O_5 (Figures 8 & 9), one finds that oxidation at temperatures below the 504 °C exotherm leads to the formation of cubic Sb_2O_3 and temperatures above 600 °C forms Sb_2O_4 .

Examination of these samples in the mid IR, reveals an interesting band at about 1100 cm^{-1} for the samples heated in air at temperatures below 600 °C. (See Figures 10 & 11.) This band is absent in virgin Sb_2S_3 and the sample heated to 800 °C. It is also absent in the nitrogen heated Sb_2S_3 . (See Figure 12.) In addition, the band is not seen in any of the oxides. (See Figures 13 & 14.) This band is characteristic of sulfate containing compounds (12) and suggests that Sb_2S_3 oxidation goes through a sulfate containing intermediate.

Results of a DSC experiment carried out under nitrogen are shown in Figure 15. A broad endotherm at 510 °C with a shoulder at 491 °C (corresponding to the melting of Sb_2S_3) is observed.

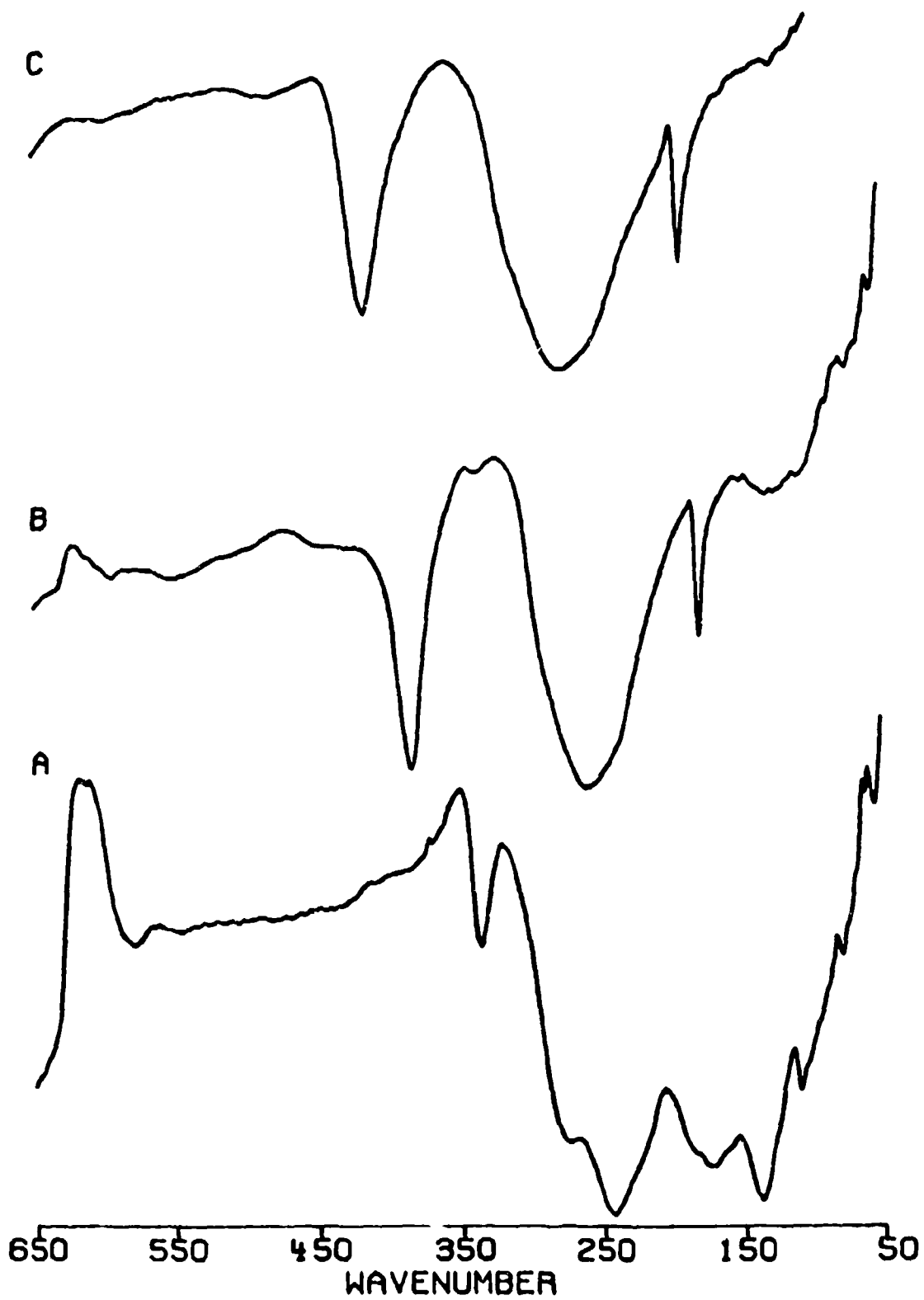


Figure 6. Far IR of Oxidized Sb_2S_3 (A) Virgin; (B) 400°C in Air;
(C) 480°C in Air.

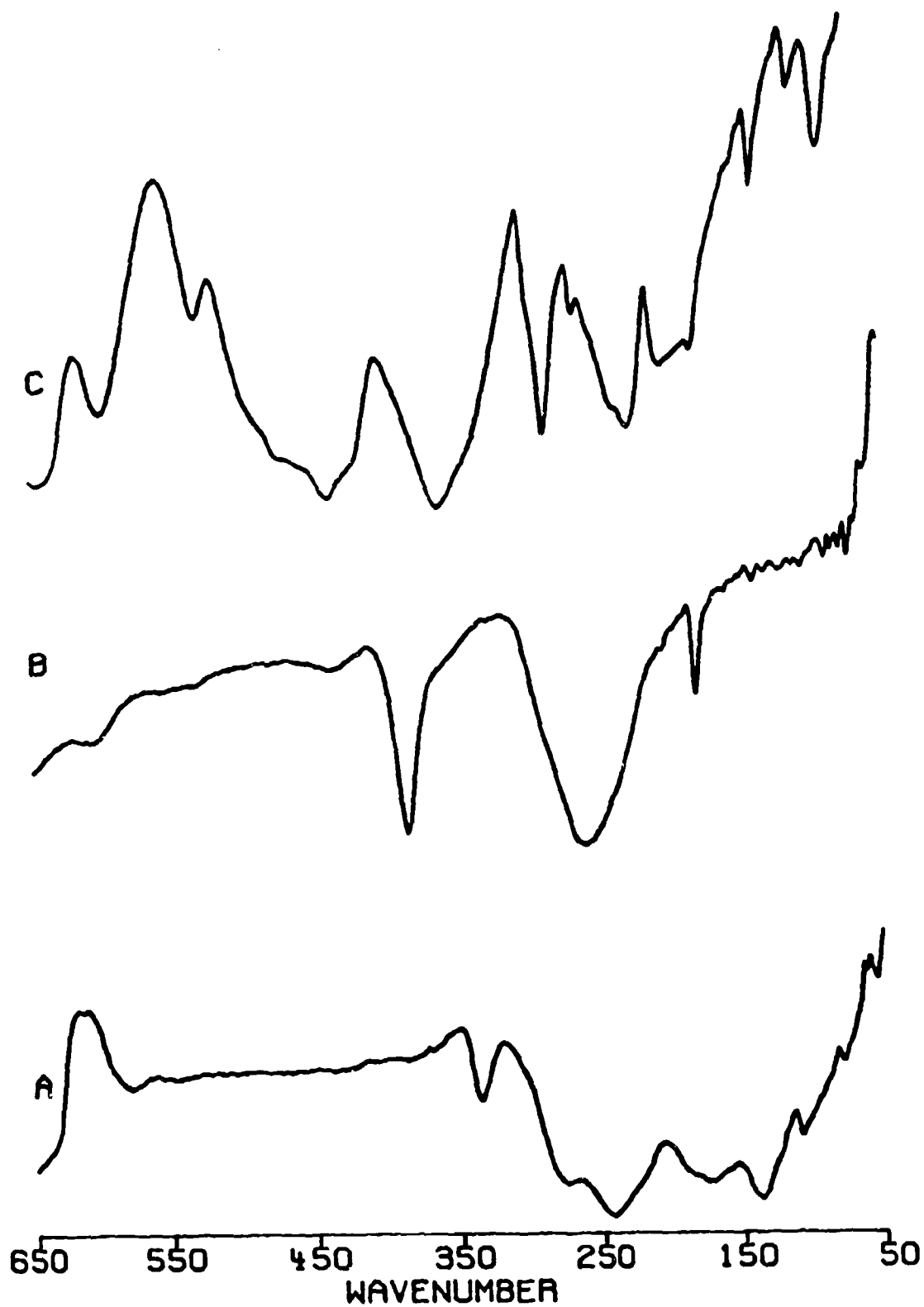


Figure 7. Far IR of Oxidized Sb_2S_3 (A) Virgin; (B) 525°C in Air;
(C) 800°C in Air.

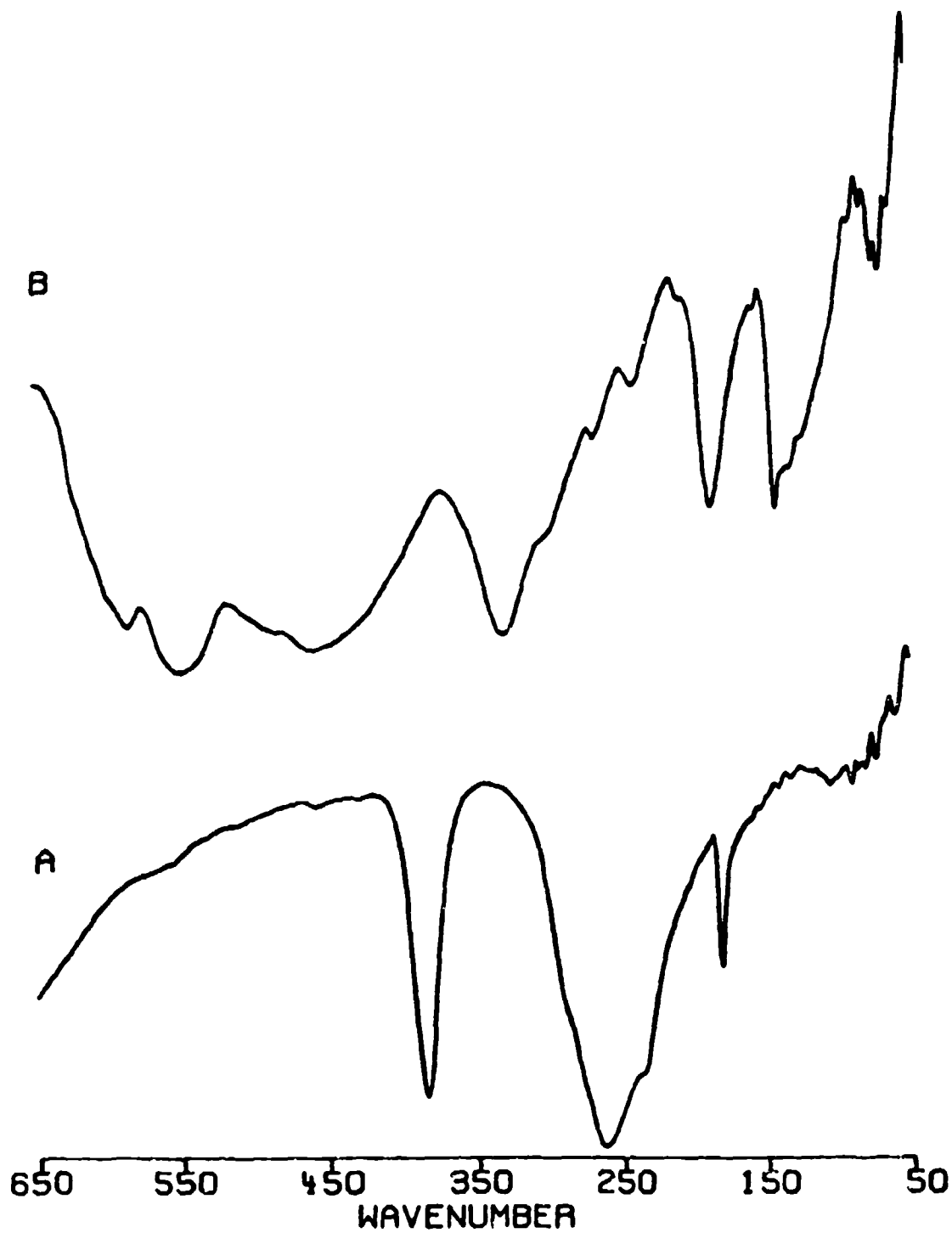


Figure 8. Far IR of Antimony Oxides (A) Cubic Sb₂O₃; (B) Orthorhombic Sb₂O₃.

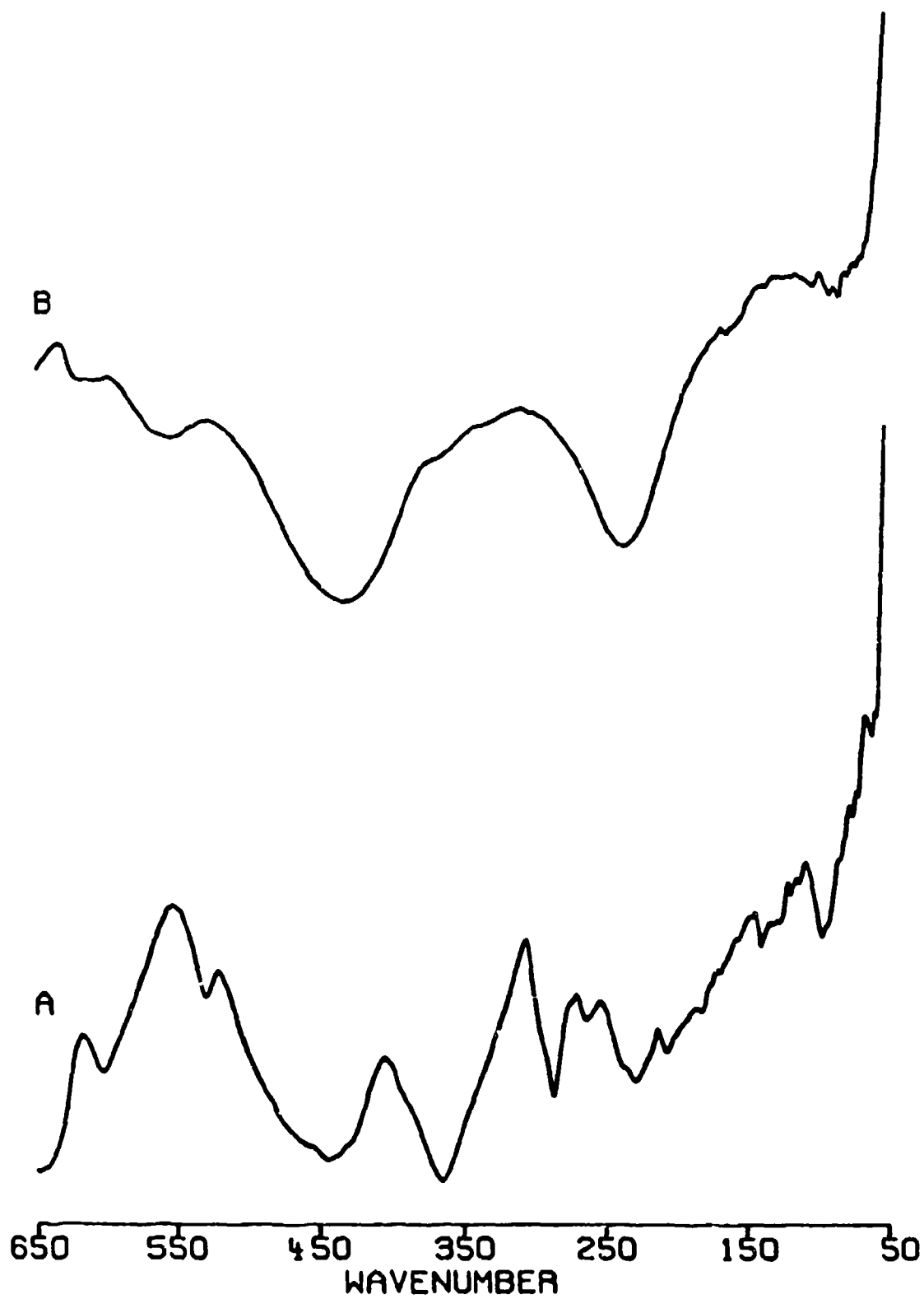


Figure 9. Far IR of Antimony Oxides (A) Sb_2O_4 ; (B) Sb_2O_5 .

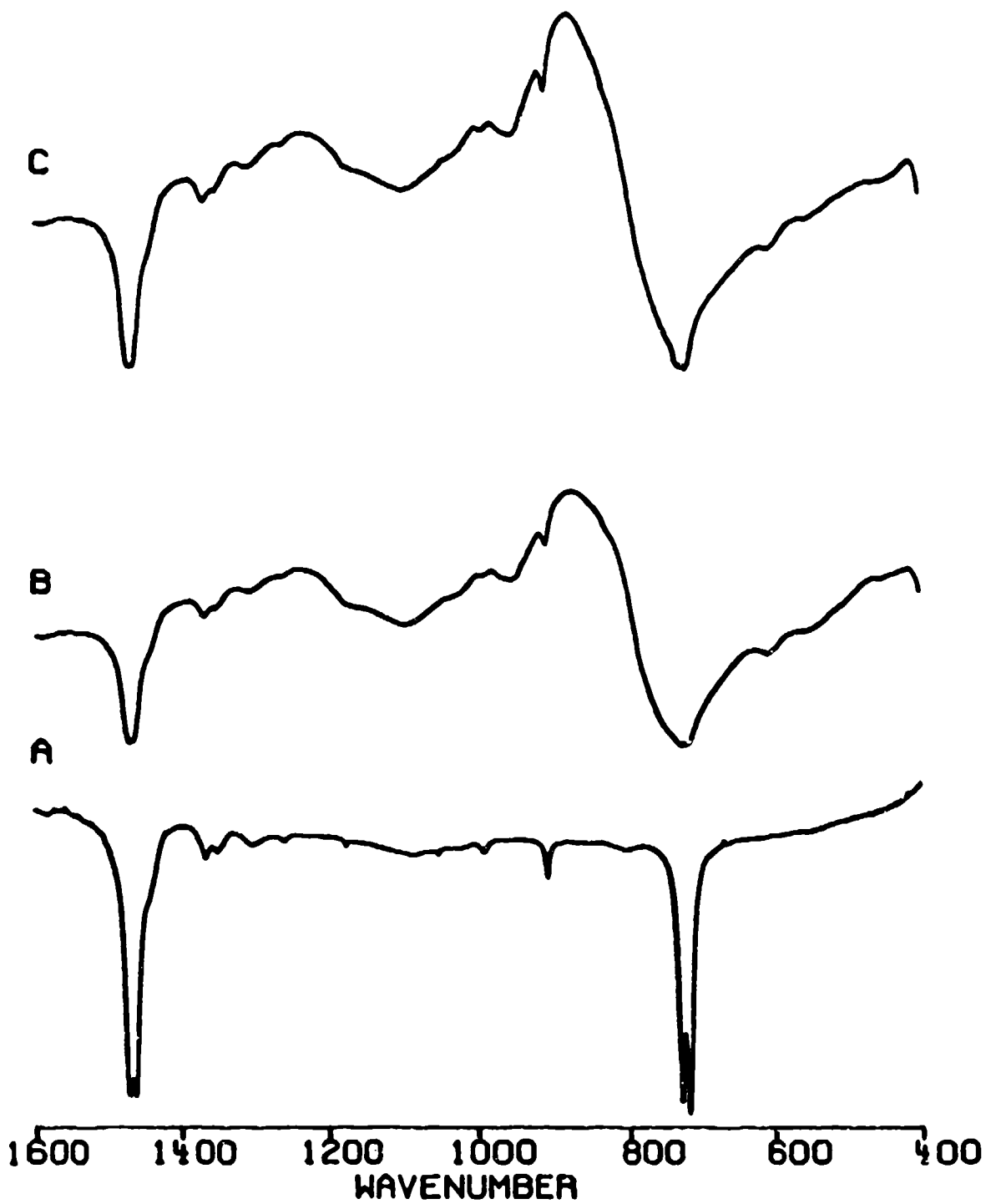


Figure 10. Mid IR of Antimony Trisulfide (A) Virgin; (B) 400°C in Air; (C) 430°C in Air.

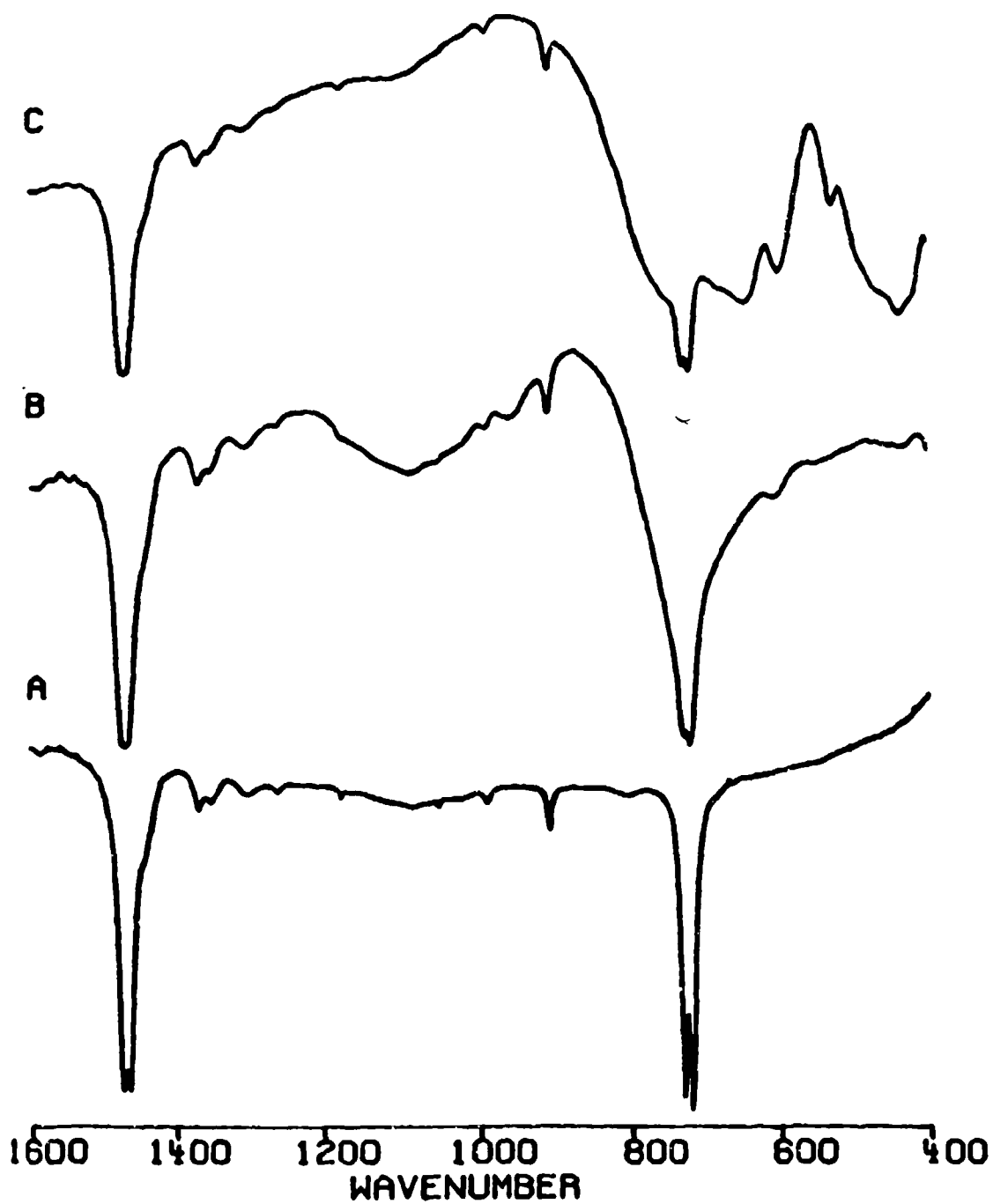


Figure 11. Mid IR of Antimony Trisulfide (A) Virgin; (B) 525°C in Air; (C) 800°C in Air.

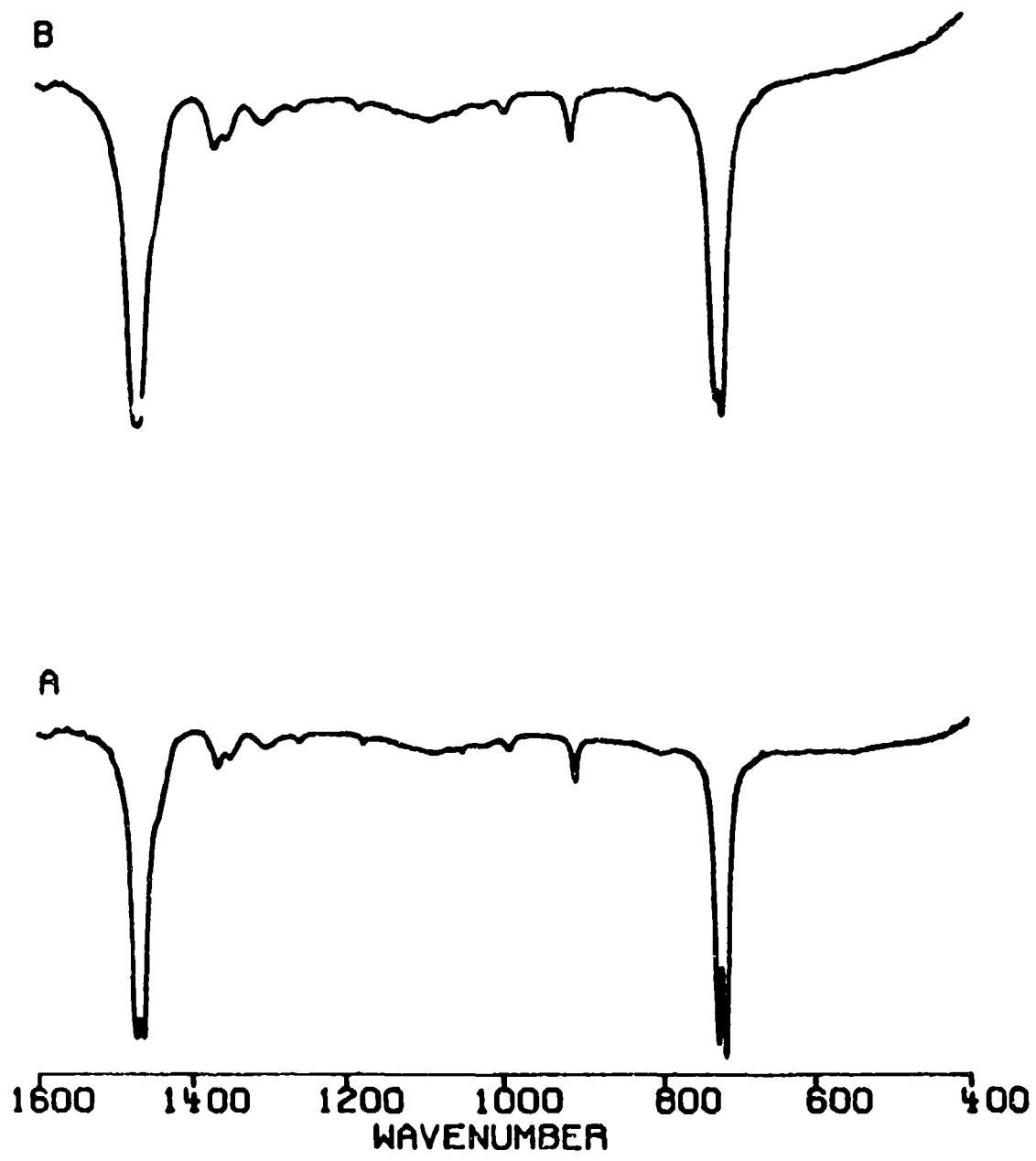


Figure 12. Mid IR of Antimony Trisulfide (A) Virgin; (B) 275°C in N₂.

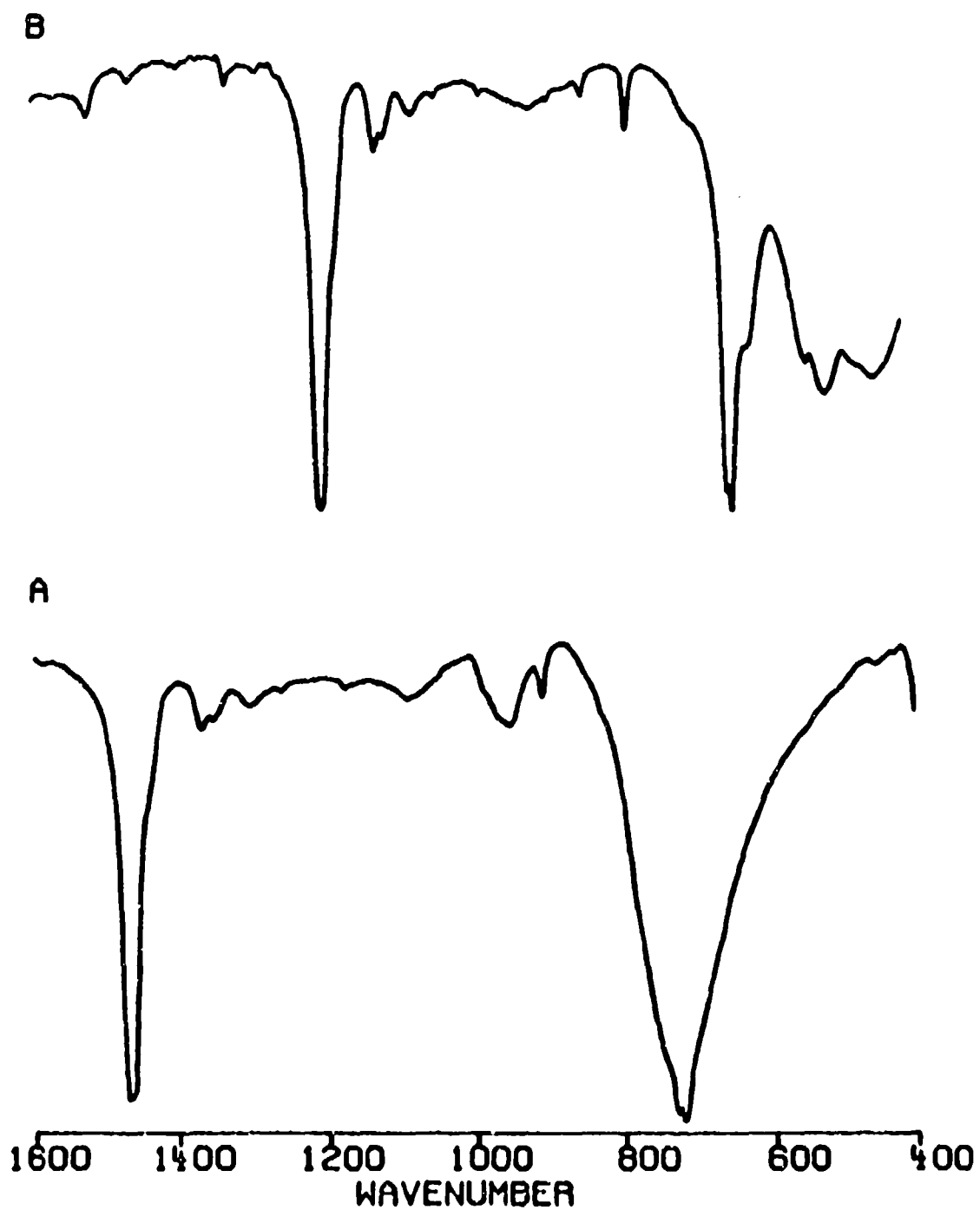


Figure 13. Mid IR of Antimony Oxides (A) Cubic Sb₂O₃;
(B) Orthorhombic Sb₂O₃.

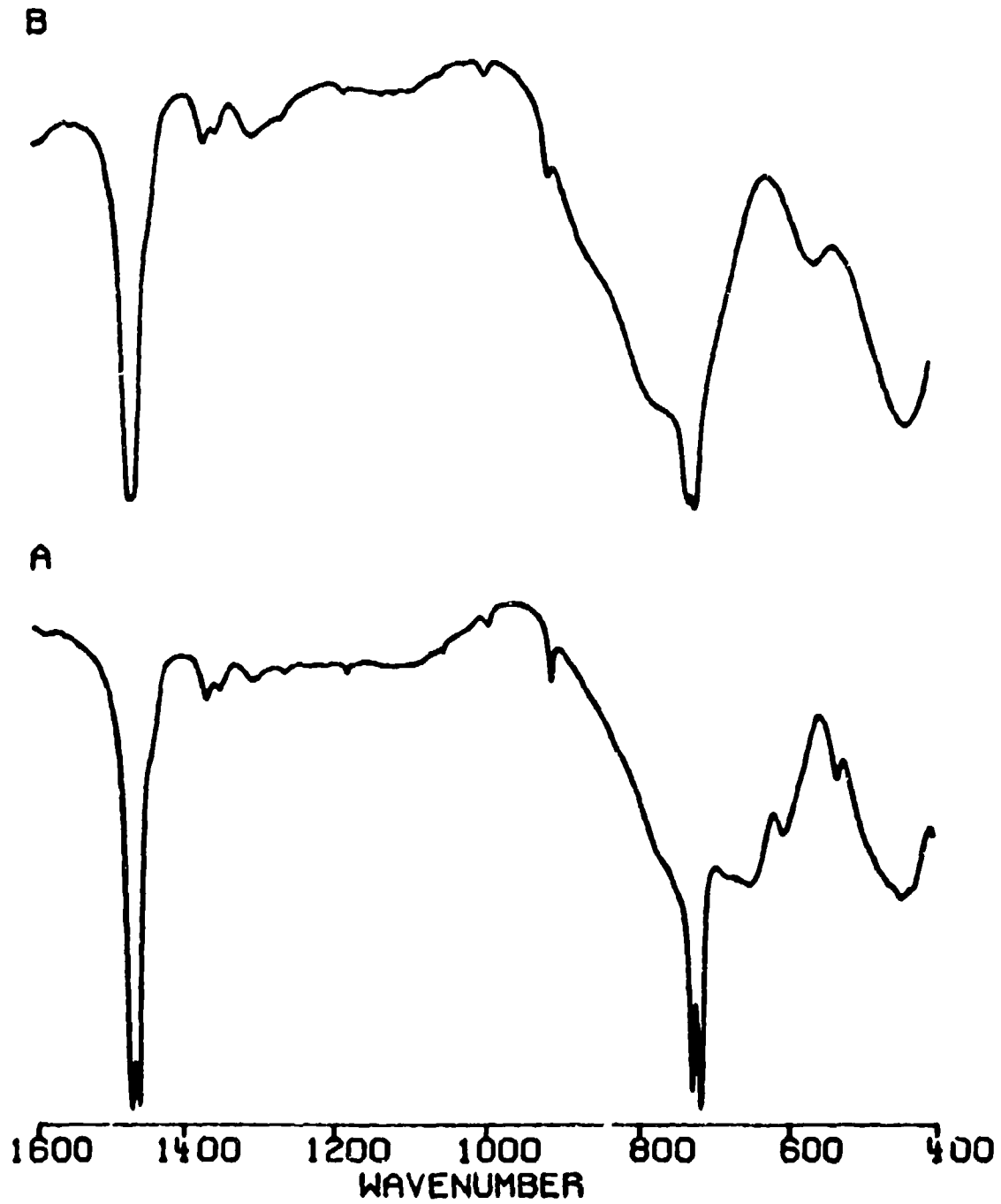


Figure 14. Mid IR of Antimony Oxides (A) Sb_2O_4 ; (B) Sb_2O_5 .

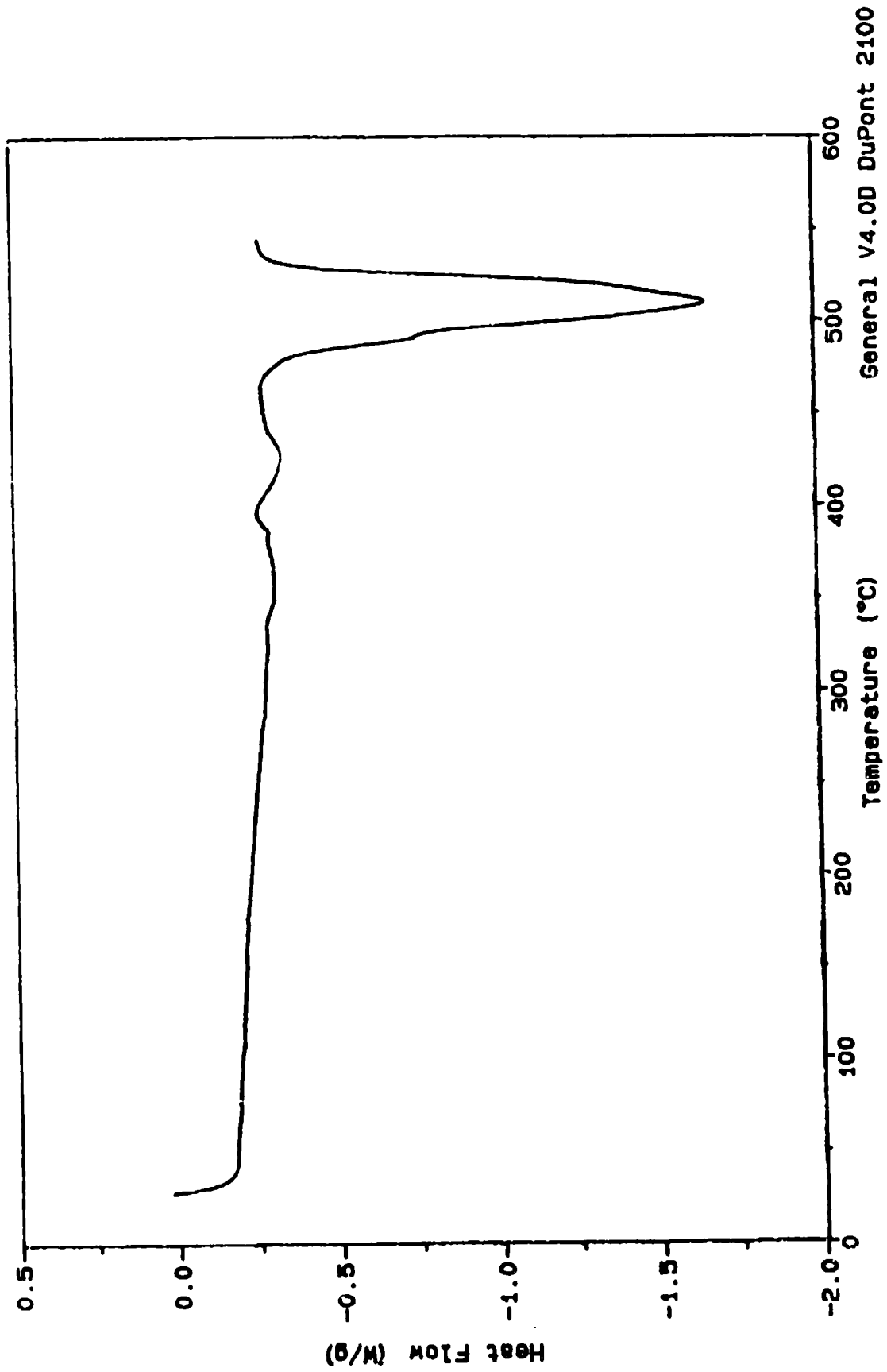
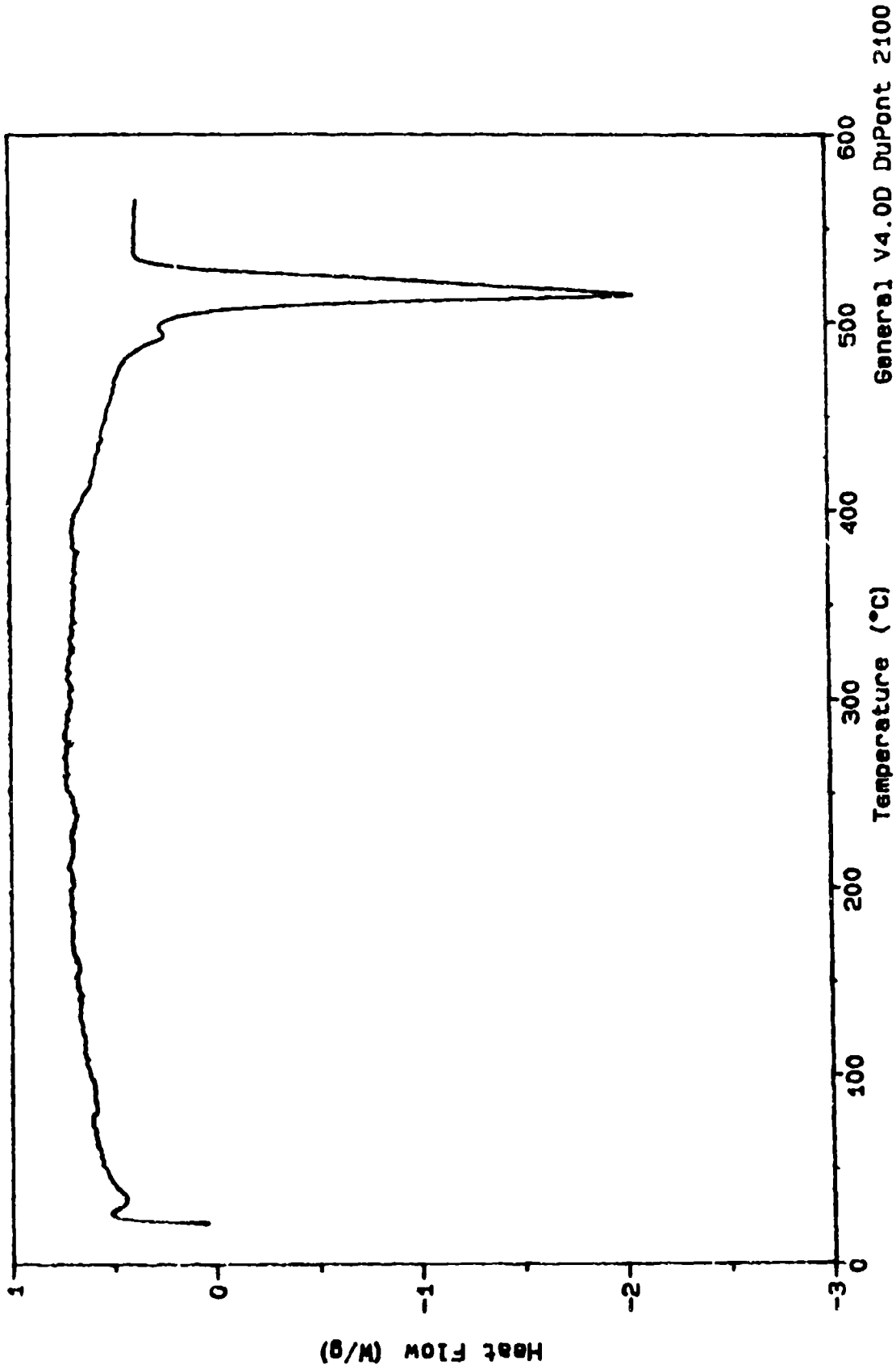


Figure 15. DSC of Sb_2S_3 in N_2 .

This is lower than the literature value of 566 °C (13) and somewhat broader than one would expect for a melting. One possible explanation may be a reaction with the sample pan similar to that seen in TGA experiments. To test this, a sample was run in a gold pan. (See Figure 16.) The endotherm was sharper; however, the shoulder at 491 °C remained - indicating some reaction with the aluminum pan. Melting endotherms in general are dependent on factors such as particle size and structure. The presence of small amounts of amorphous Sb_2S_3 , impurities, thermal decompositions and/or reaction with the sample pan could contribute to broadening of melting endotherms.

When Sb_2S_3 is run in the DSC to temperatures above the melt, rapidly cooled to room temperature and rerun, the following curve is obtained. (See Figure 17.) The rapidly cooled Sb_2S_3 has a glass transition at 232 °C followed by a heat of crystallization (69.1 J/g) at 318 °C. The melting endotherm splits into two distinct peaks - 497 and 514 °C with a shoulder at 522 °C. Repeating the cooling procedure and rerunning in air (Figure 18), yields the same glass transition and crystallization followed by oxidation. The two peaks at 488 and 507 °C are the result of a combination of the oxidation exotherms and the melting of unoxidized Sb_2S_3 . (See Figure 19.) When Sb_2S_3 is allowed to cool slowly from a melt, it solidifies into a crystalline form as evidenced from the lack of a glass transition or crystallization exotherm. (See Figure 20.)



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Figure 16. DSC of Sb₂S₃ in N₂, Cold Sample Pan.

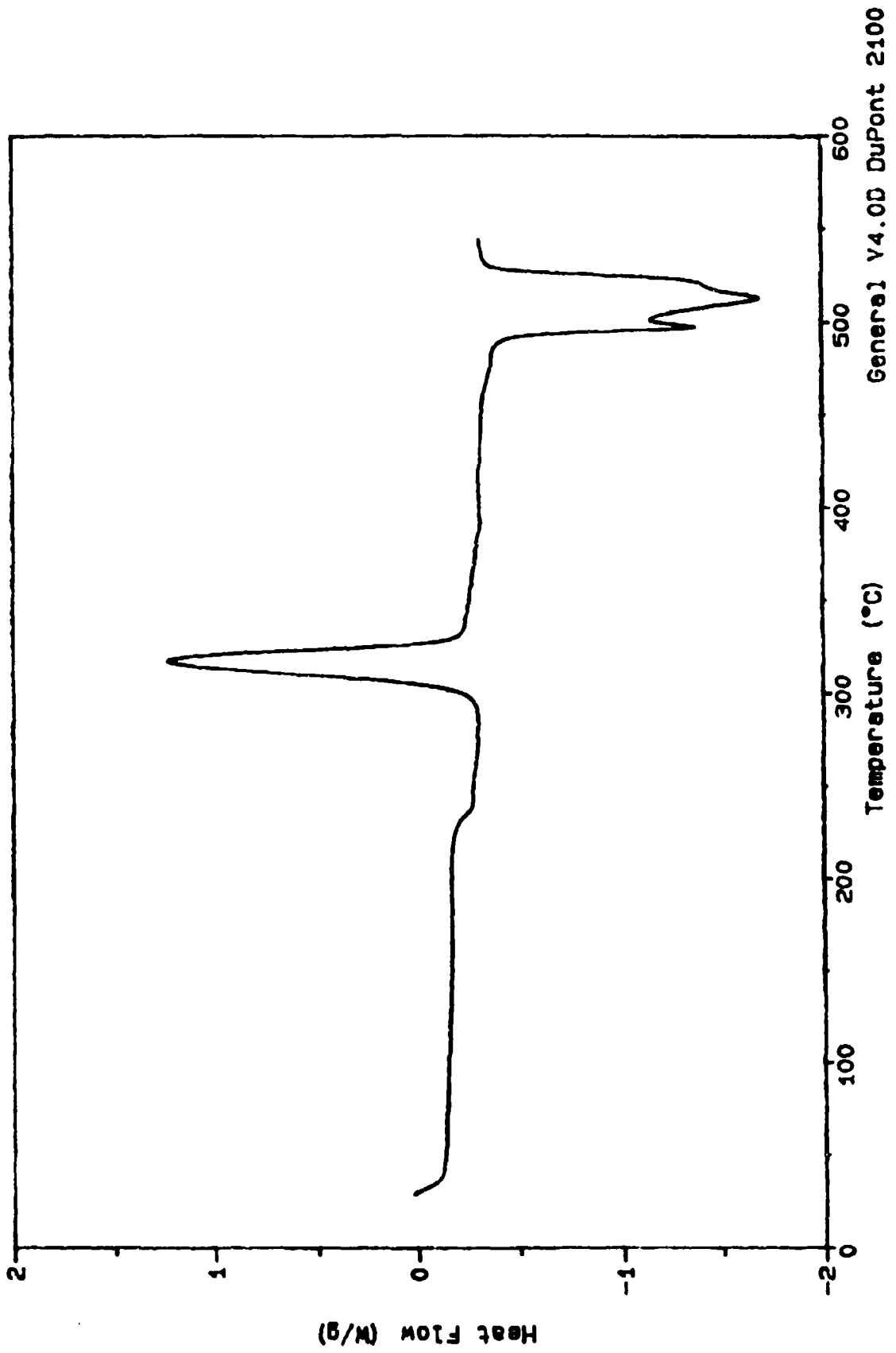


Figure 17. DSC of Amorphous Sb_2S_3 , (Rapidly Cooled from Melt), in N_2 .

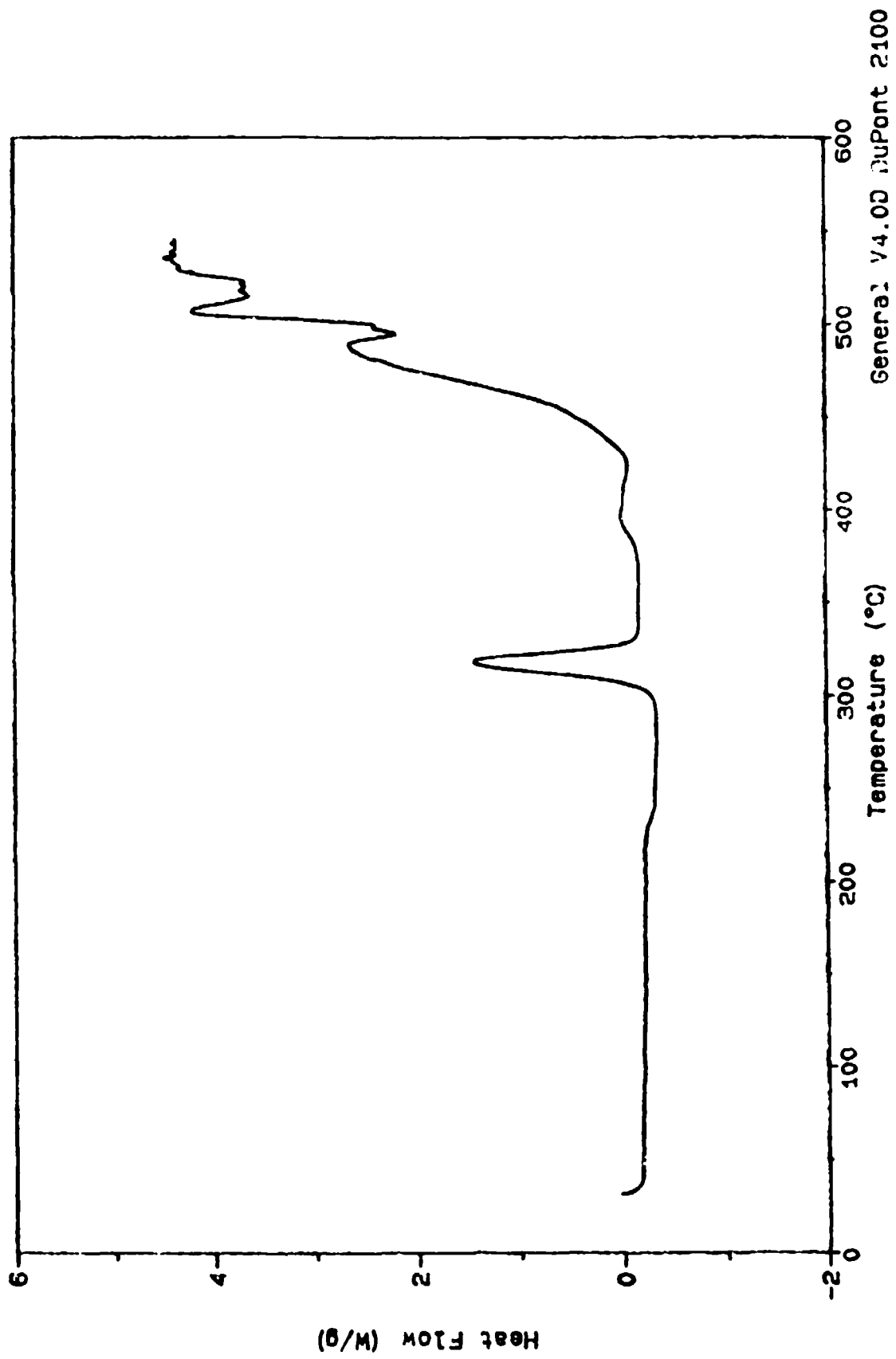


Figure 18. DSC of Amorphous Sb_2S_3 in Air.

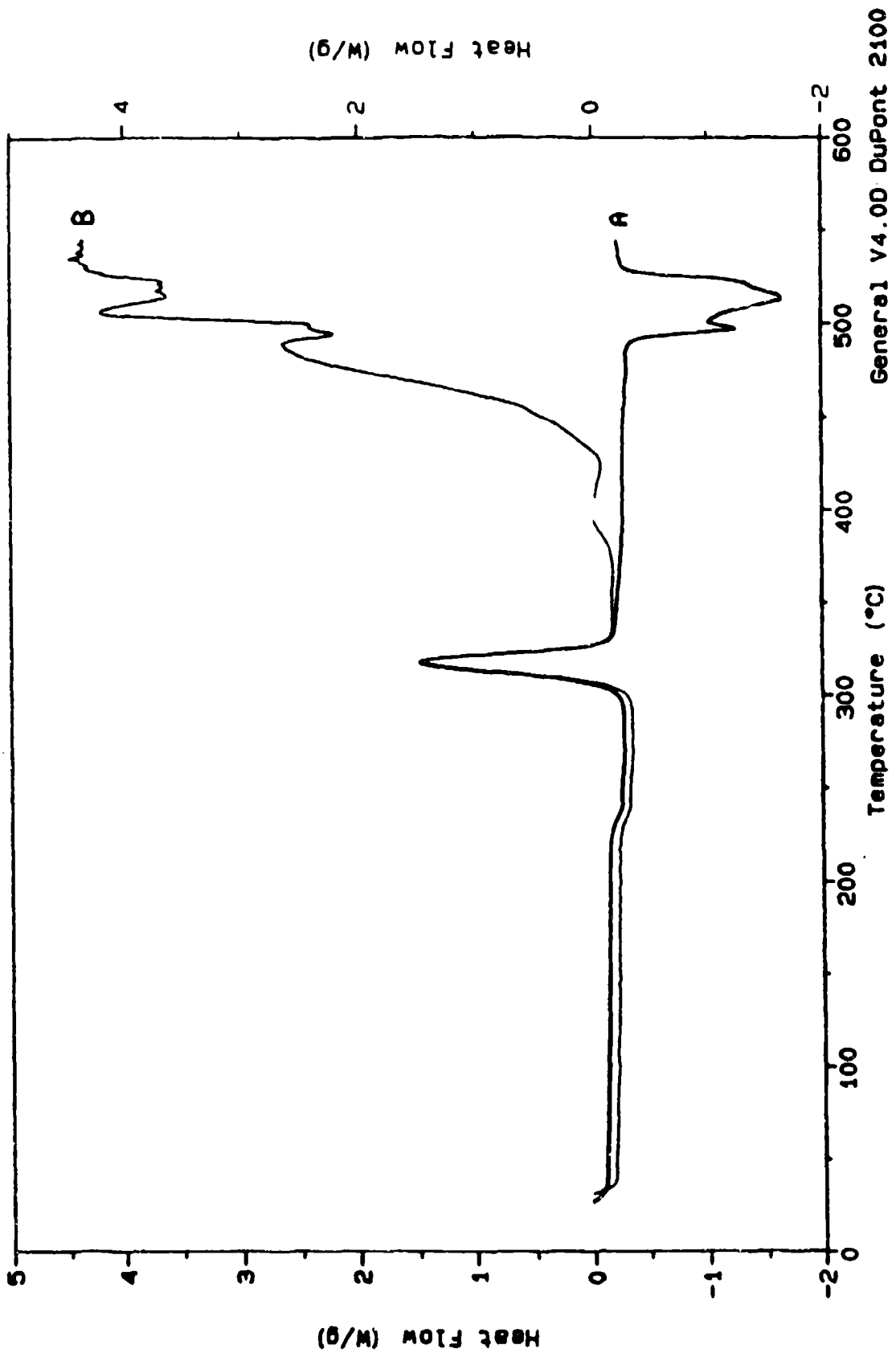


Figure 19. DSC Comparison of Amorphous Sb_2S_3 Air vs N_2 , (A) N_2 ; (B) Air.

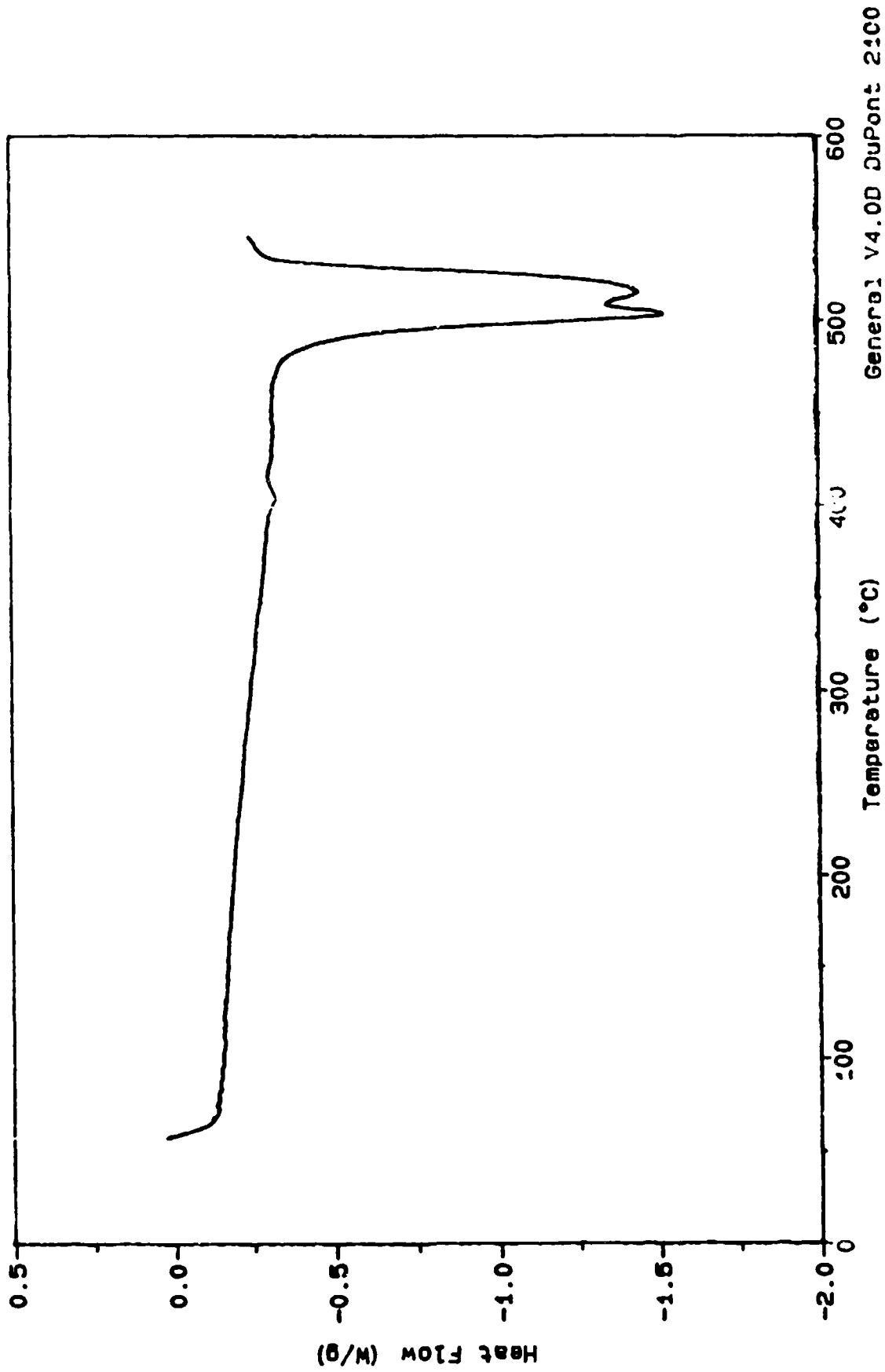


Figure 20. DSC of Crystalline Sb_2S_3 in Air (Slow Cooled from Melt).

Since the following sections will deal with antimony sulfides of higher sulfur to antimony ratios (Sb_2S_4 and Sb_2S_5), the effects of elemental sulfur on Sb_2S_3 's thermal and oxidative properties were examined. Figure 21 shows the DSC curve of Sb_2S_3 mixed with enough elemental sulfur to equal that found in Sb_2S_4 . Under nitrogen, the melting of the alpha and beta forms of sulfur can be seen (105 and 116 °C respectively) followed by volatilization at 287 °C. The added sulfur does have an effect on the Sb_2S_3 , as demonstrated by the increase in temperature of the melt to 525 °C.

When the Sb_2S_3 with sulfur is run in air, different results are obtained. As shown in Figure 22, the presence of sulfur causes some changes in the 350 to 525 °C region. There appears to be more contribution from the 434 °C exotherm and less from the one at 473 °C. In addition, a new exotherm is observed at 272 °C. This area corresponds to the oxidation of sulfur. (See Figure 23.) The intensity changes in the remaining two exotherms may result from a reaction between Sb_2S_3 and SO_2 generated by oxidation of the added sulfur to produce more of the sulfate containing species identified in the mid IR spectra of air stressed Sb_2S_3 samples.

2.0 Antimony tetrasulfide (Sb_2S_4)

Antimony tetrasulfide does not possess the same thermal stability as the trisulfide. Heating Sb_2S_4 in the TGA under a nitrogen atmosphere results in a 8.7 percent weight loss between

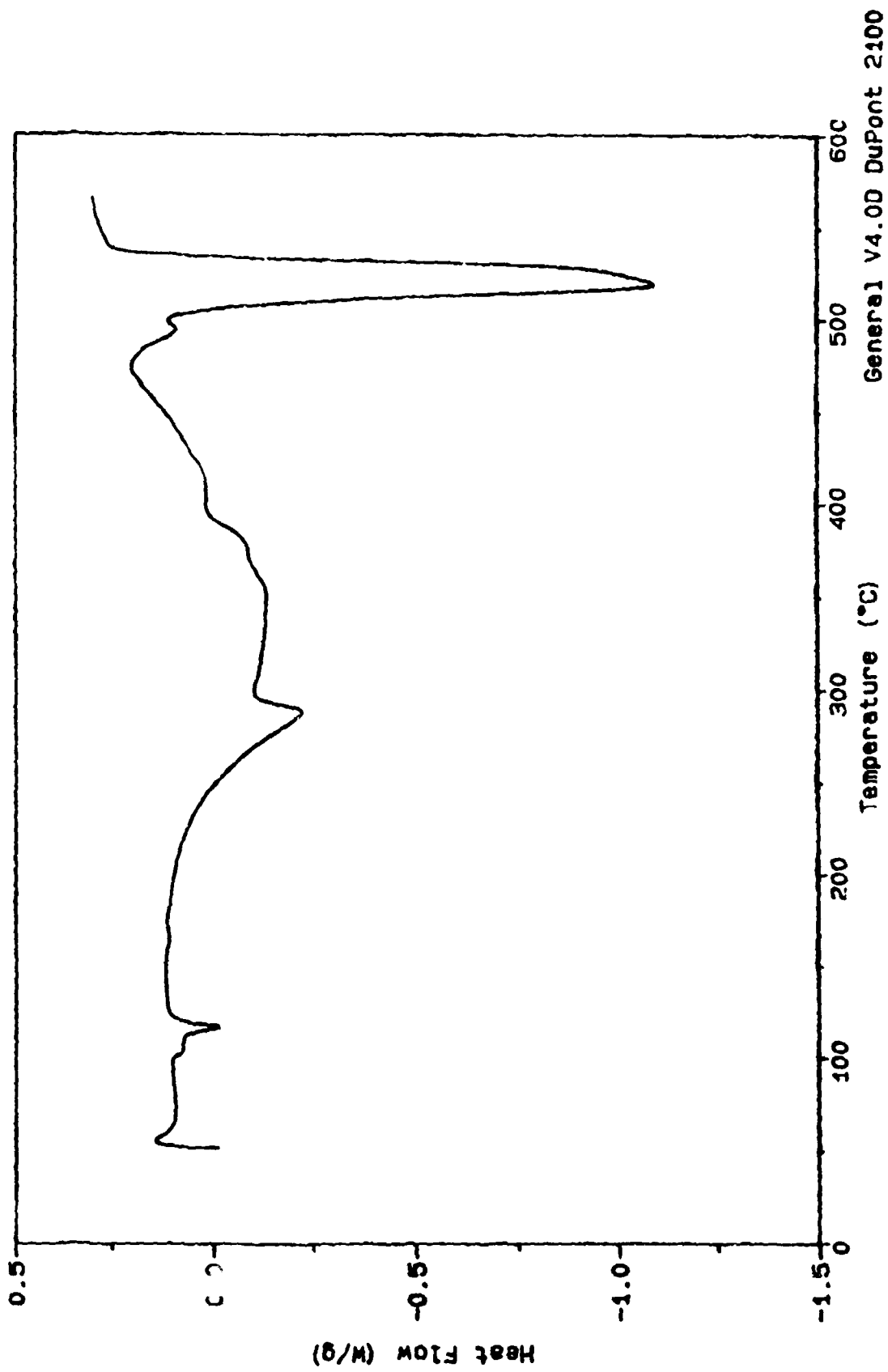


Figure 21. DSC of Sb_2S_3 and Sulfur in N_2 .

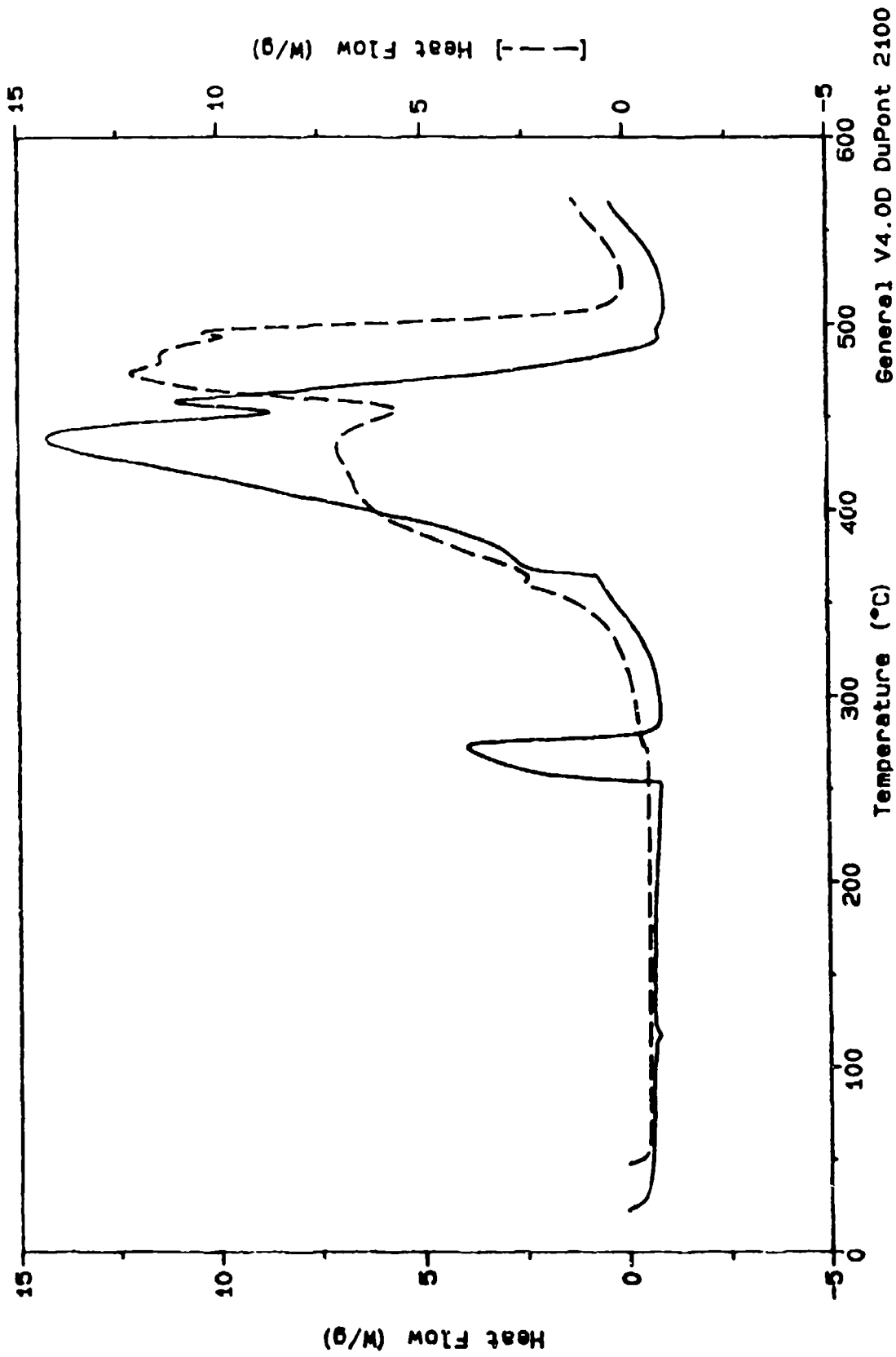


Figure 22. DSC of Sb_2S_3 and Sulfur in Air.

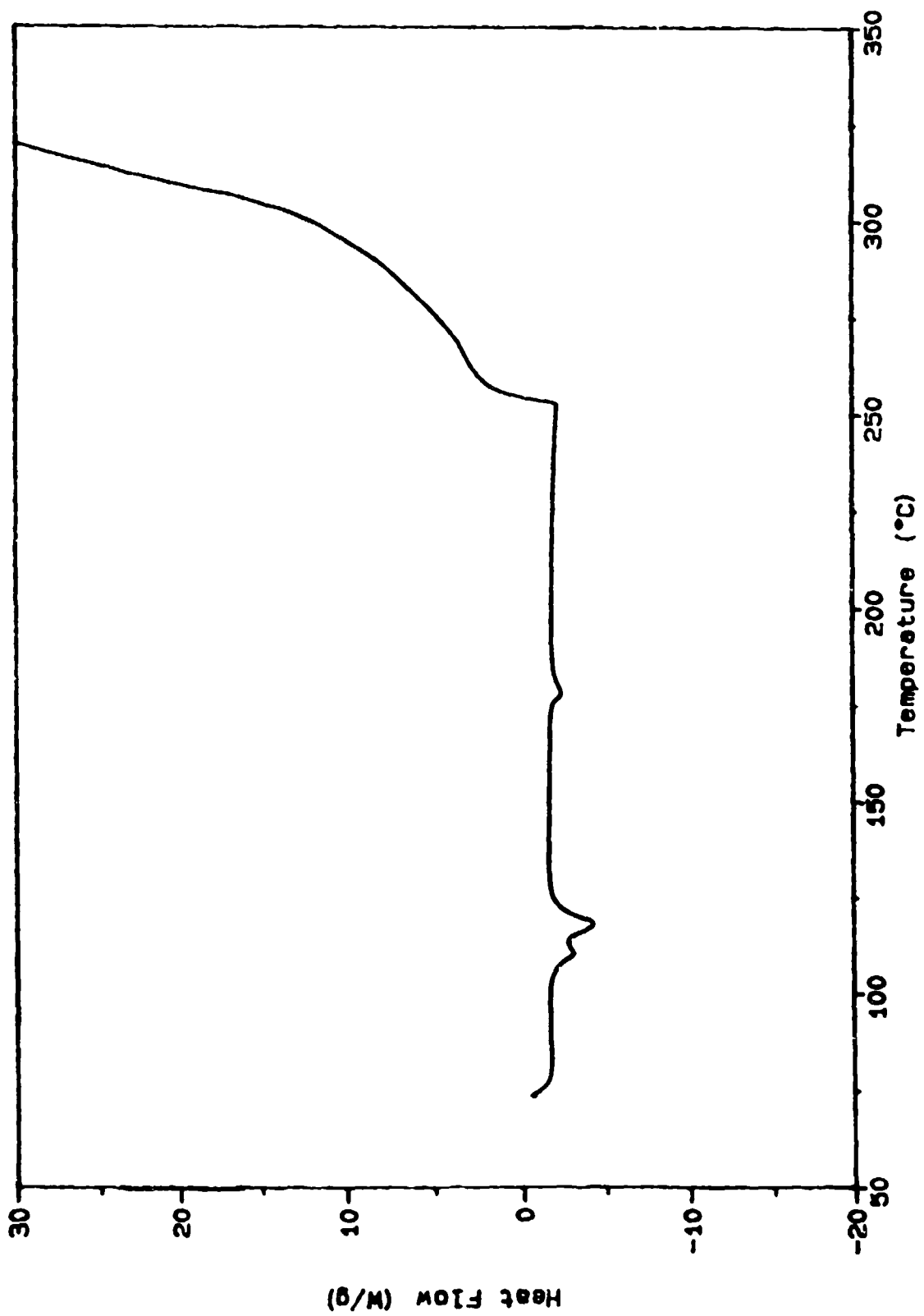


Figure 23. DSC of Sulfur in N₂.

195 and 230 °C. (See Figure 24.) After the initial loss, the sample weight remains constant to 600 °C. Although not shown, at temperatures beyond 600 °C a rapid weight loss occurs and evidence of reaction between the sample and the platinum sample pan can be seen. If heating is stopped before reaction with the platinum sample pan can occur, one sees that the Sb_2S_4 changes color from brick red to black. This color change along with the 8.7 percent weight loss, suggests that sulfur is being lost to yield Sb_2S_3 . (See Reaction 1.) The theoretical weight loss for reaction 1 is 8.6 percent.



The Far IR spectrum of the above product along with those of Sb_2S_3 and Sb_2S_4 are shown in Figure 25. Sb_2S_4 has one broadband at 280 cm^{-1} and a shoulder at 150 cm^{-1} . The sample heated under nitrogen to 275 °C, displays a spectrum almost identical to Sb_2S_3 . This is in agreement with King and Asmerom (6), who reported that prolonged heating of Sb_2S_4 at 525 °C resulted in a 8.6 percent weight loss and formation of Sb_2S_3 (determined by x-ray analysis).

Differential scanning calorimetry of Sb_2S_4 in nitrogen (Figure 26) shows several features not seen with Sb_2S_3 . There is an exotherm at 210 °C (40 J/g) with a small shoulder at 170 °C, followed by a small endotherm at 302 °C. The melting endotherm is also sharper than that observed for the trisulfide and occurs at 548 °C, which is 38 °C higher.

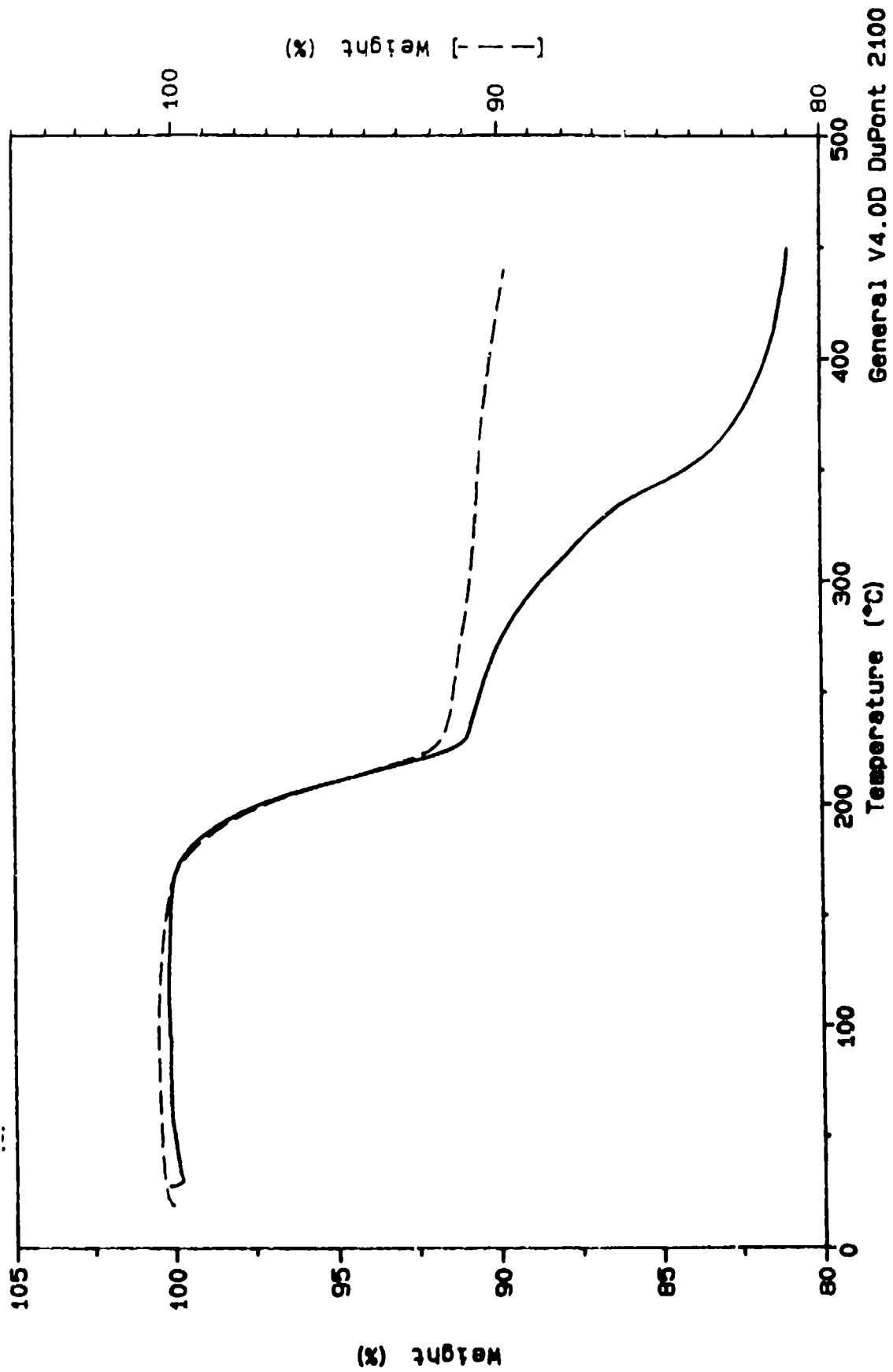


Figure 24. TGA of Sb_2S_4 (—) in Air; (----) in N_2 .

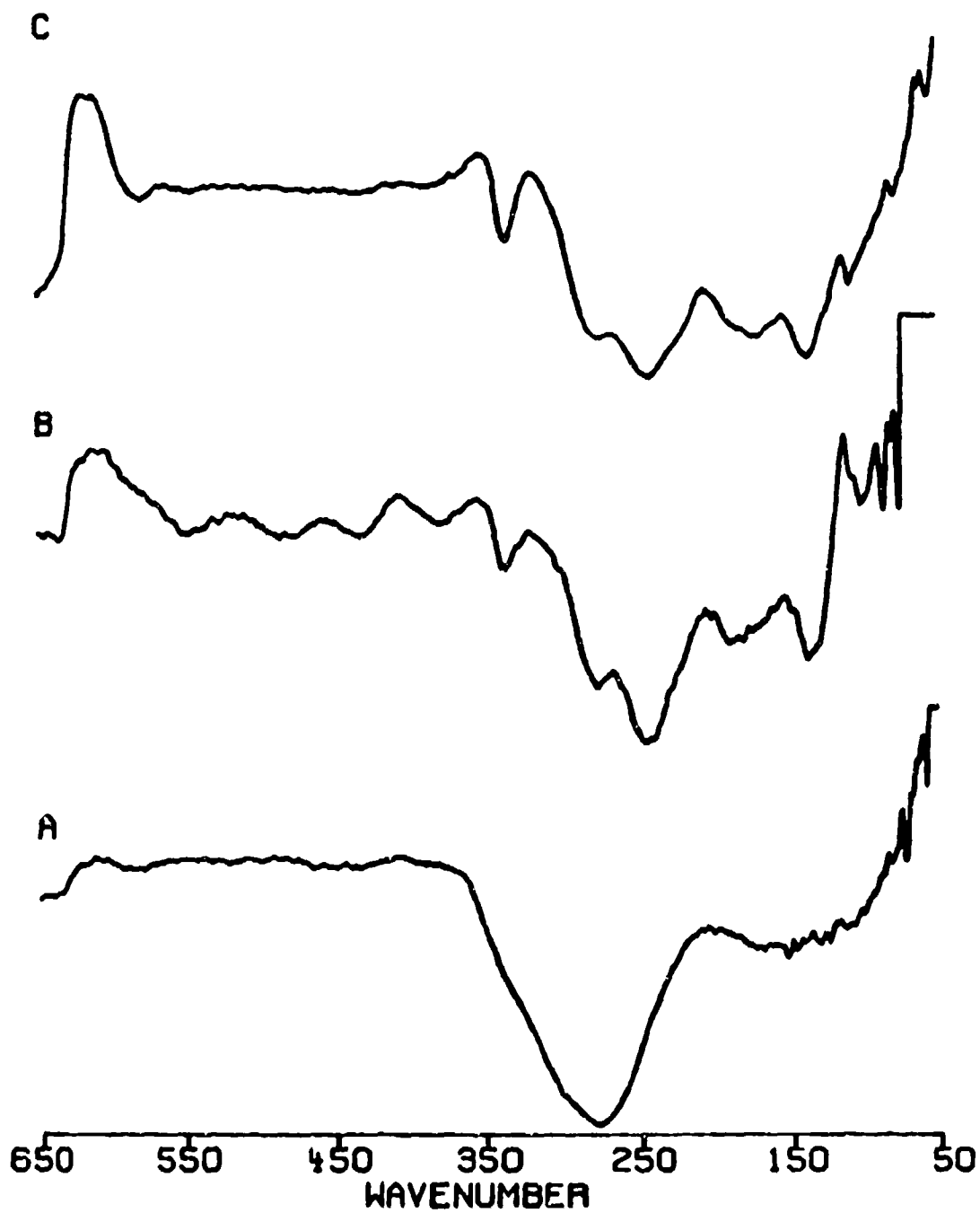


Figure 25. Far IR of Sb_2S_4 (A) Virgin; (B) Heated 275°C in N_2 ;
 (C) Sb_2S_3 (Virgin).

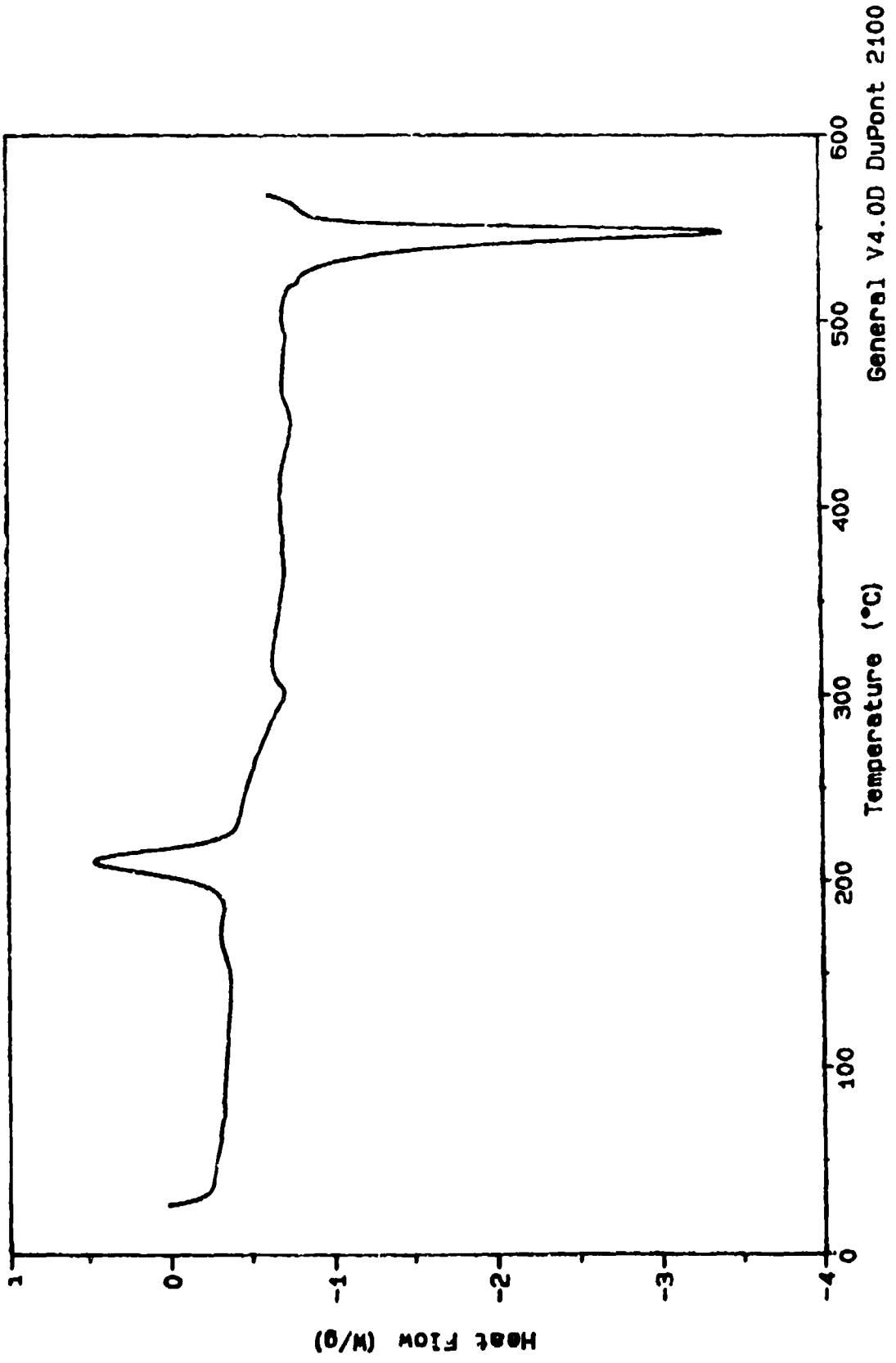


Figure 26. DSC of Sb_2S_4 in N_2 .

When DSC and TGA curves are overlaid, (Figure 27), one sees that the exotherm and the 8.7 percent weight loss occur concurrently. This exotherm is most likely associated with a structural rearrangement occurring along with the release of sulfur. Further evidence for this assignment is found by running a sample in the DSC through the 210 °C peak, rapidly cooling and rerunning. (See Figure 28.) One can see the melting of the released sulfur. (See the alpha and beta forms.)

Returning to Figure 26, the small endotherm at 302 °C can now be identified as the volatilization of released sulfur. The melting of Sb_2S_3 generated by reaction 1 (548 °C) is closer to the literature value of Sb_2S_3 . This may result from a combination of several factors - particle size and crystal structure of the Sb_2S_3 formed, side reactions and purity. Also as seen in the previous section, when elemental sulfur was added to Sb_2S_3 , melting occurred at higher temperatures. The free sulfur generated in reaction 1 may contribute in the same way to yield the higher value.

Differential thermal analysis performed in air, produces the thermogram in Figure 29. Several exotherms are observed: 212, 269, 369, and 581 °C. These are associated with reaction 1, the

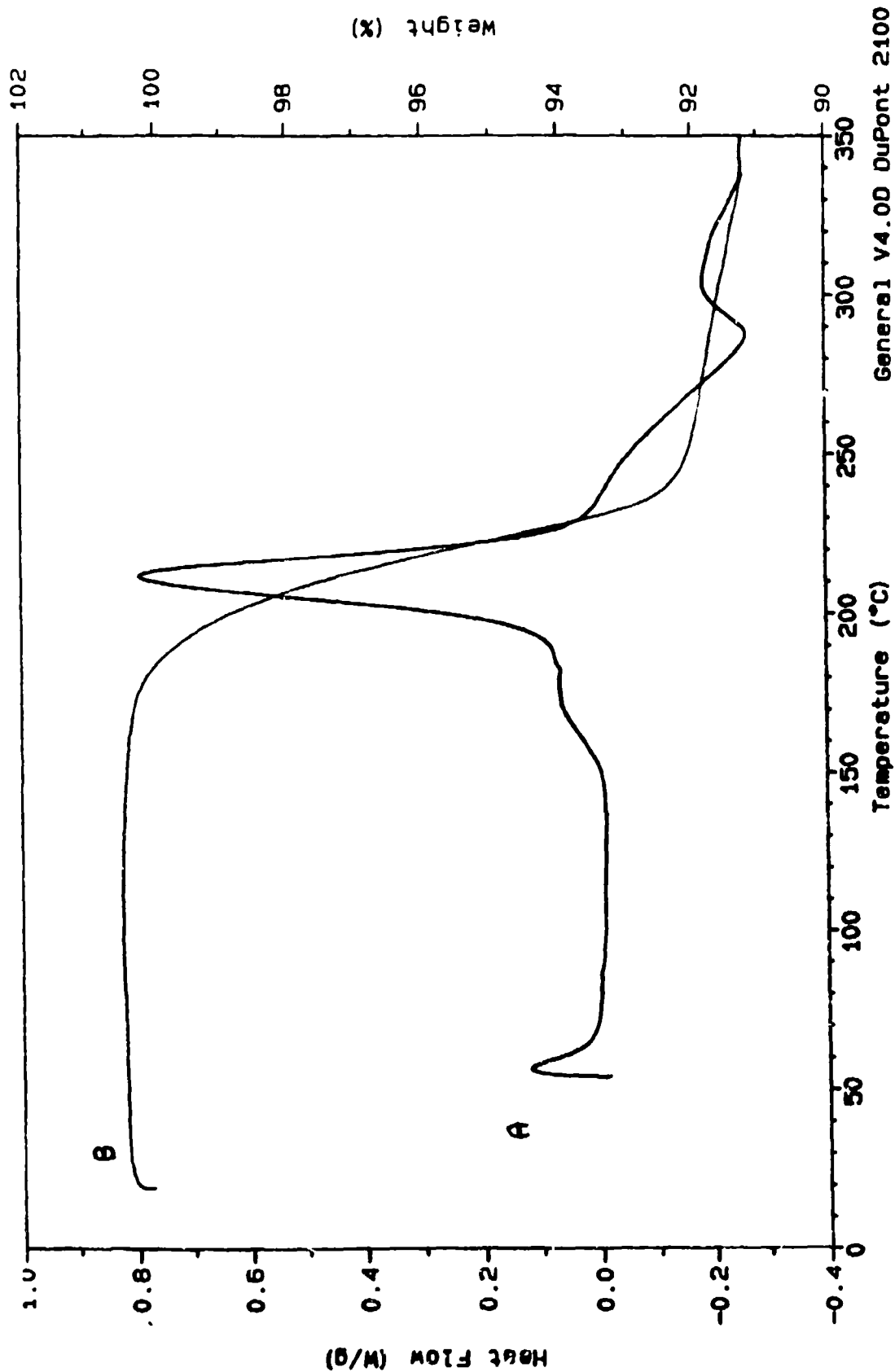


Figure 27. Comparison of (A) DSC and (B) TGA of Sb_2S_4 in N_2 .

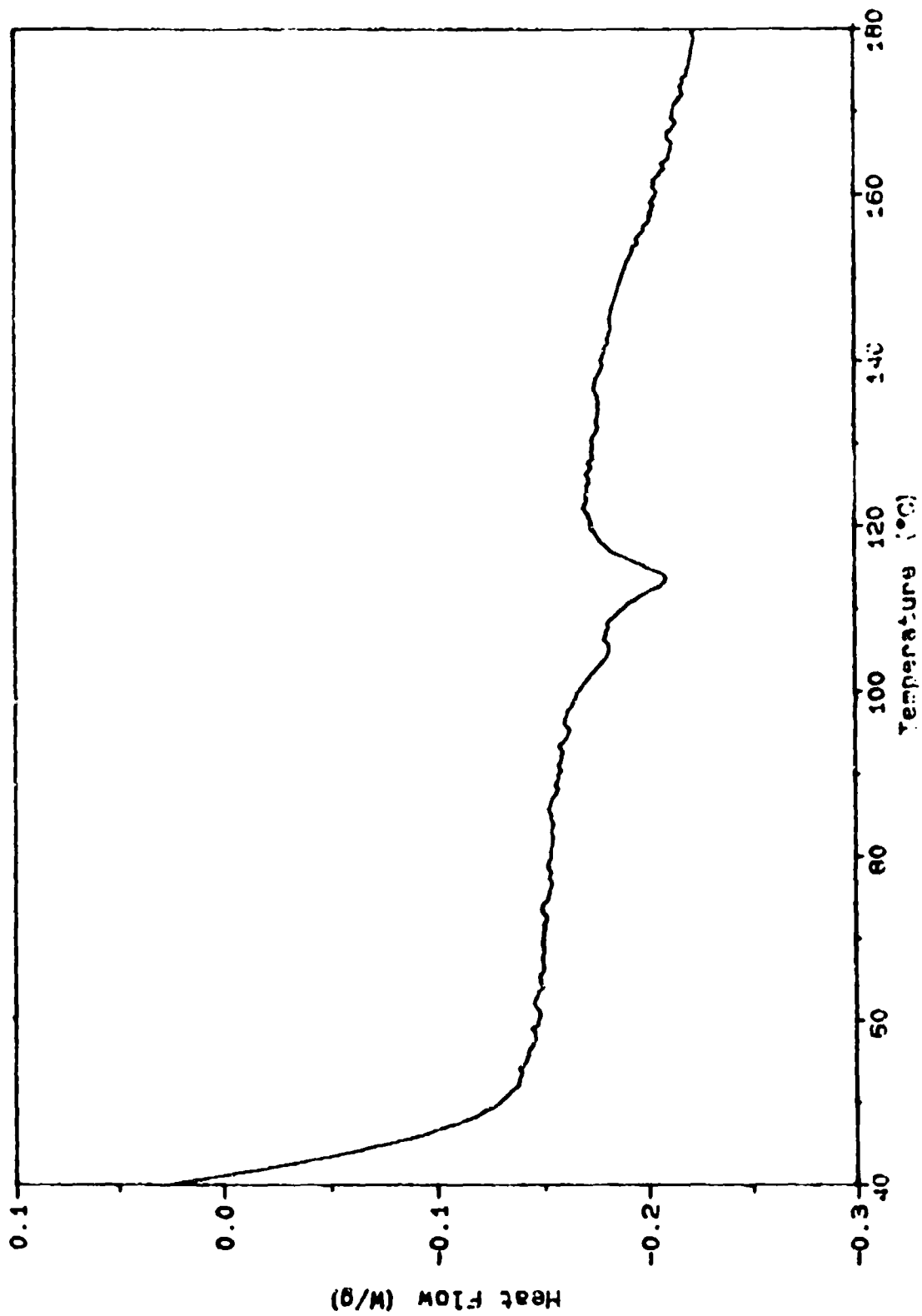


Figure 28. DSC of Released Sulfur (Sb_2S_4 previously heated to 250°C in N_2).

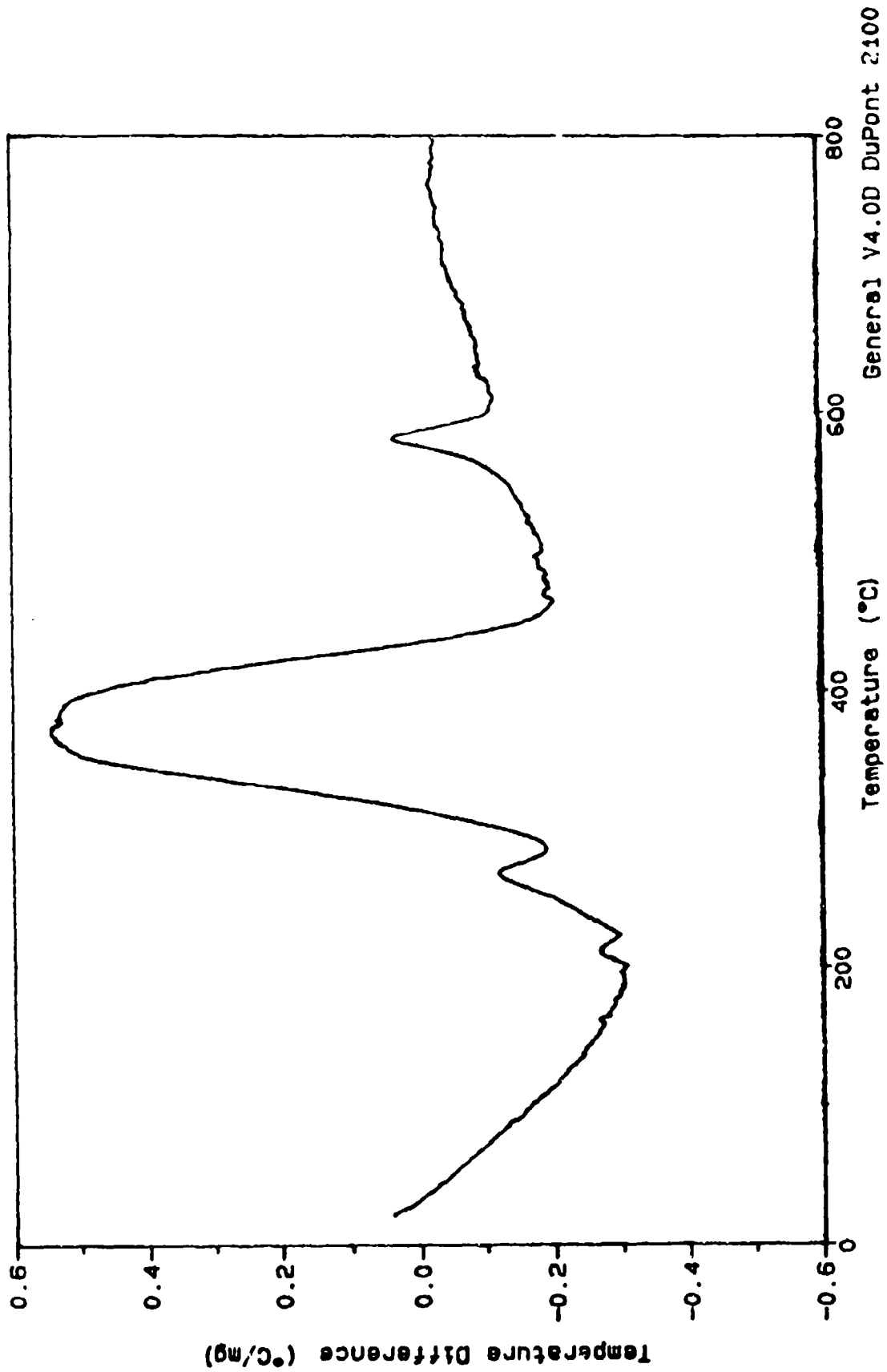


Figure 29. DSC of Sb_2S_4 in Air.

oxidation of the released sulfur, oxidation of Sb_2S_3 and finally the oxidation of Sb_2O_3 to Sb_2O_4 . (See below.)

Isothermal TGA in air followed by IR analysis reveals a process similar to that of Sb_2S_3 's oxidation. Exposure to temperatures below 581°C (temperature of the last exotherm) results in the formation of cubic Sb_2O_3 . (See Figure 30.) At 525°C , the major species is Sb_2O_3 ; however, the presence of shoulders at 442 and 286°C indicates the beginning of Sb_2O_4 formation. Above 600°C (Figure 31), an IR spectrum corresponding to Sb_2O_4 is seen. As with Sb_2S_3 , mid-IR analysis reveals the 1100 cm^{-1} band typical of a sulfate containing species, for samples heated in air at 425 , 525 , and 650°C . (See Figures 32 & 33.) Virgin Sb_2S_4 and Sb_2S_4 heated under nitrogen do not display this band. (See Figure 34.) This band is also absent in the sample heated in air to 1050°C , indicating the sulfate is not stable at elevated temperatures.

When Sb_2S_4 is heated under nitrogen in the DSC above its melting point (a temperature where Sb_2S_3 should be present as a result of reaction 1), rapidly cooled and rerun, the thermogram in Figure 35 is obtained. A glass transition at 227°C is seen along with two overlapping exotherms at 279 and 293°C . This indicates the formation of amorphous Sb_2S_3 . The two exotherms point to the possibility of two distinct structures being formed. The melting endotherm on the second run is broader than the first and is lowered to 541°C . Slow cooling from the melt and rerunning shows

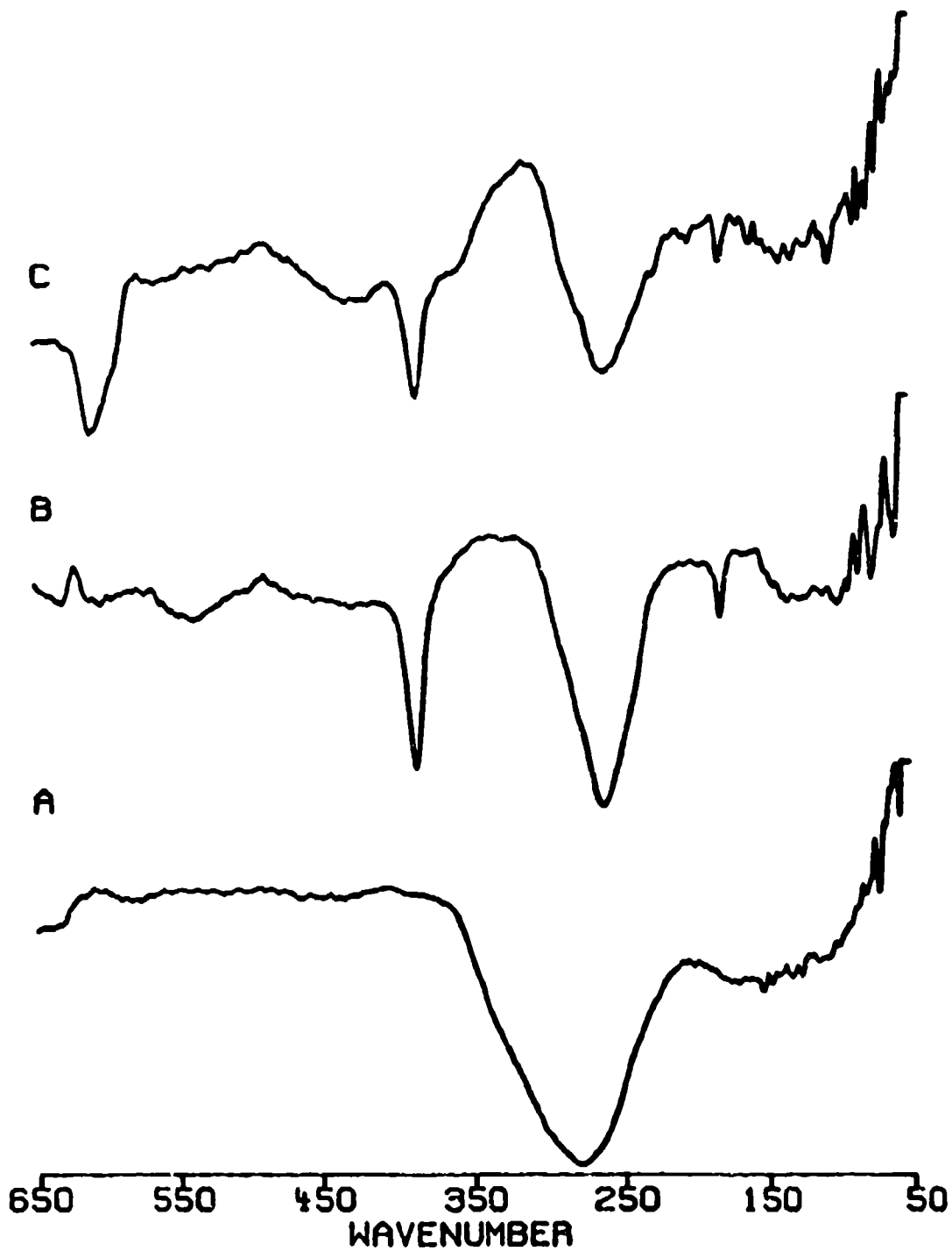


Figure 30. Far IR of Sb_2S_4 (A) Virgin; (B) 425°C in Air; (C) 525°C in Air.

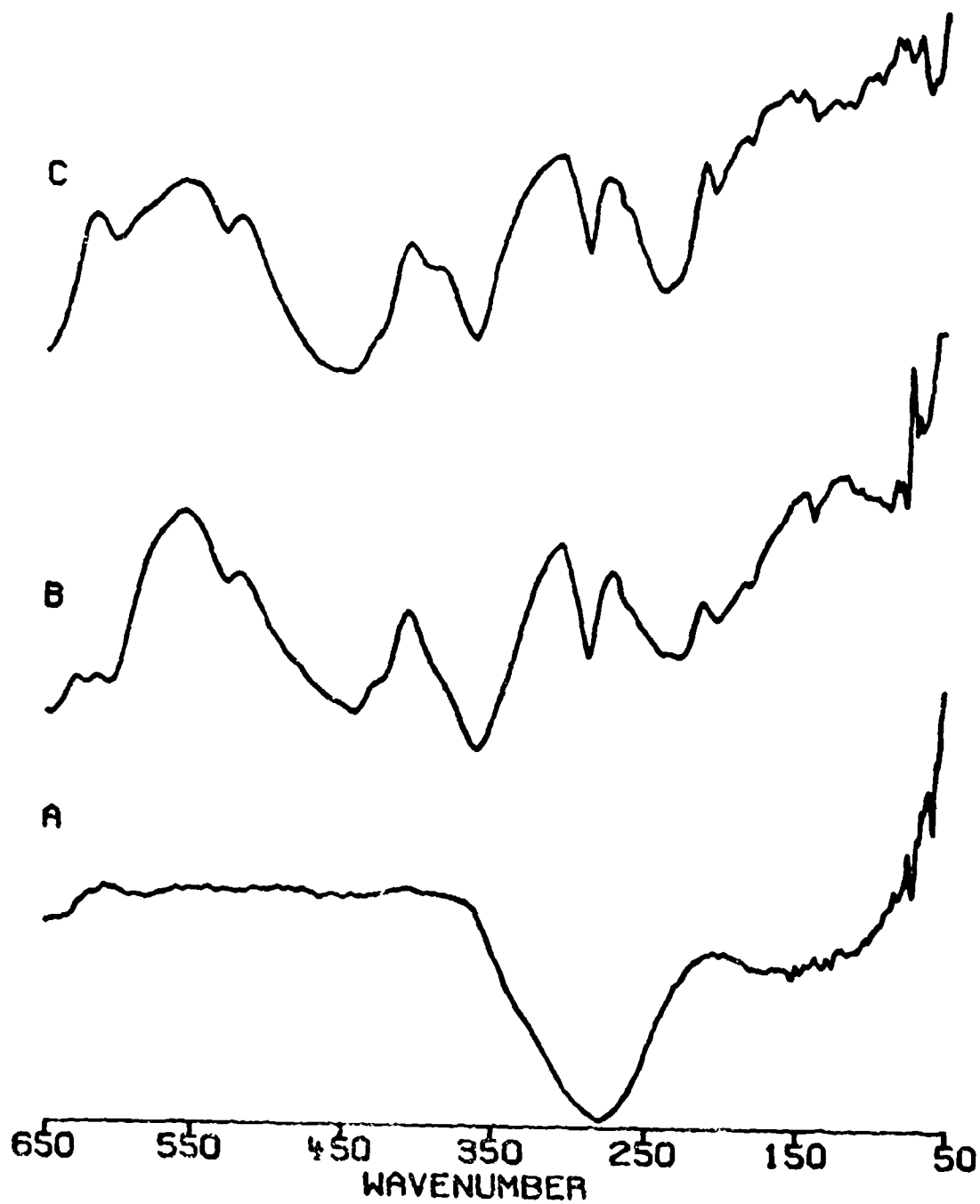


Figure 31. Far IR of Sb_2S_4 (A) Virgin; (B) 650°C in Air; (C) 1050°C in Air.



Figure 32. Mid IR of Sb_2S_4 (A) Virgin; (B) 425°C in Air; (C) 575°C in Air.



Figure 33. Mid IR of Sb_2S_4 (A) Virgin; (B) 650°C in Air; (C) 1050°C in Air.

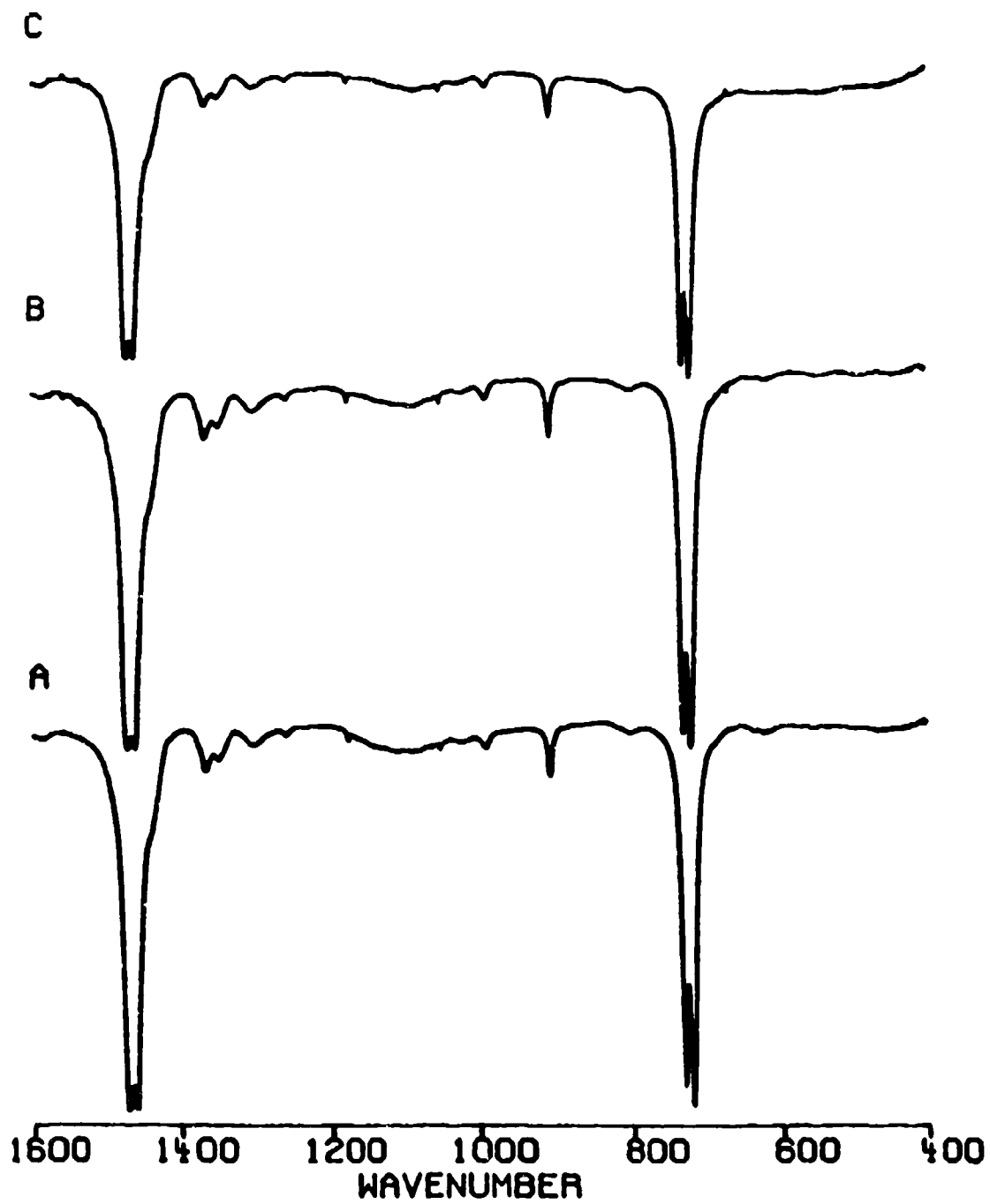


Figure 34. Mid IR of Sb_2S_4 (Virgin) (A) Virgin; (B) 230°C in N_2 ; (C) Sb_2S_3 .

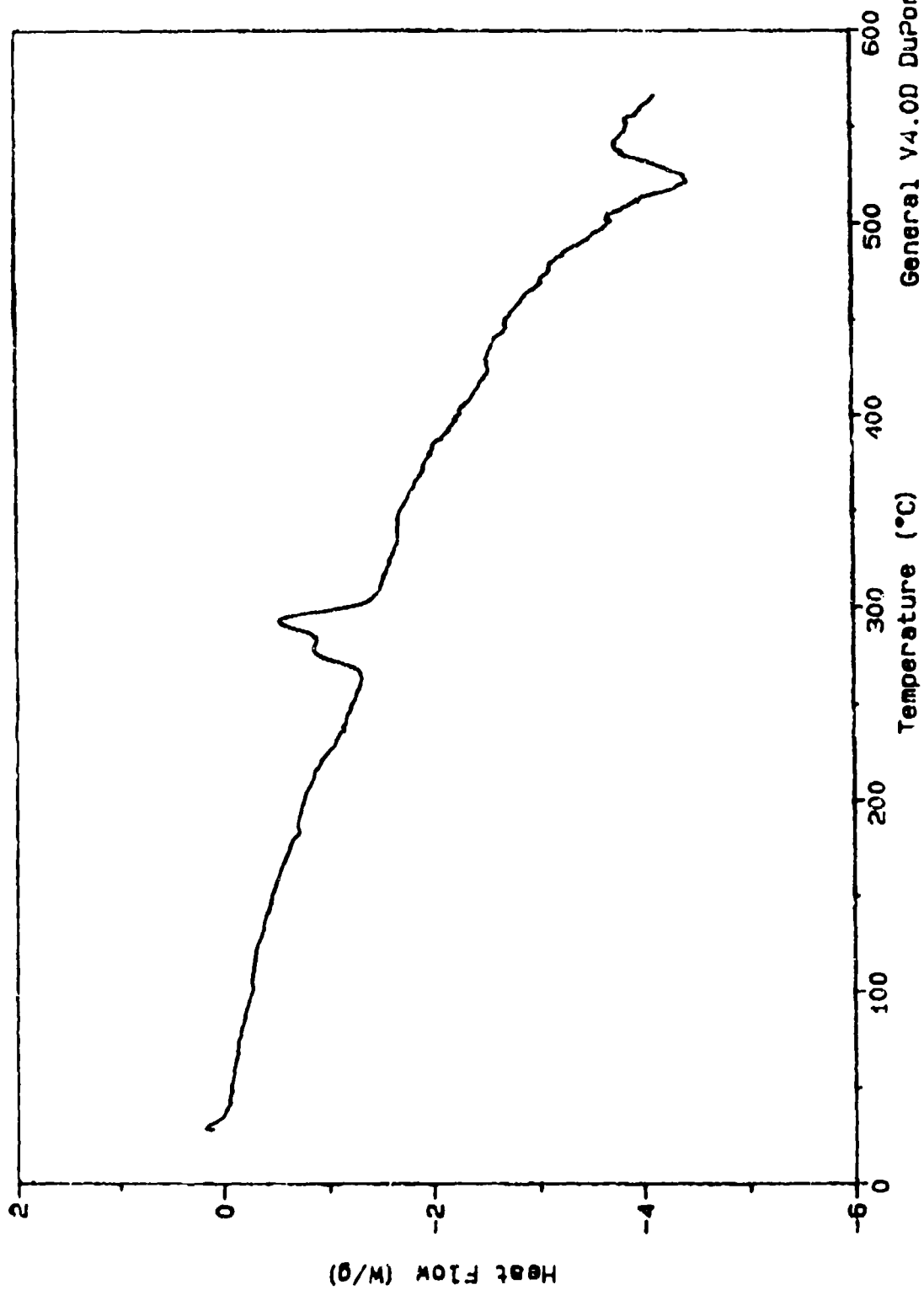


Figure 35. DSC of generated Sb_2S_3 (Sb_2S_4 rapidly cooled from melt in N_2).

no glass transitions or heats of crystallization, indicating crystalline Sb_2S_3 formed. (See Figure 36.) Melting occurs at 545°C and the endotherm is broader and less intense than in the first run - indicating there may be some decomposition or side reactions. Also a small endotherm at 450°C is unseen in any of the previous experiments.

Addition of elemental sulfur to Sb_2S_4 in an amount equal to that found in antimony pentasulfide produced some interesting effects. Figure 37 shows the results of DSC performed under anaerobic conditions on Sb_2S_4 + sulfur, along with those of Sb_2S_4 . The most pronounced effect is a lowering in temperature of the exotherm associated with the release of sulfur to form Sb_2S_3 . Unlike Sb_2S_3 , there is no increase in the melting point on the addition of elemental sulfur. When the experiment is carried out in air (Figure 38), the same shift in the 210°C exotherm is observed; however, the oxidations of sulfur and the generated Sb_2S_3 occur at slightly higher temperatures.

3.0 Antimony pentasulfide (Sb_2S_5)

As previously mentioned, antimony pentasulfide, a reddish brown solid, contains antimony in the +3 oxidation state along with considerable amounts of excess sulfur. In order to examine the effects of excess sulfur contained in Sb_2S_5 , samples were extracted with CS_2 and parallel studies were conducted with extracted and unextracted Sb_2S_5 .

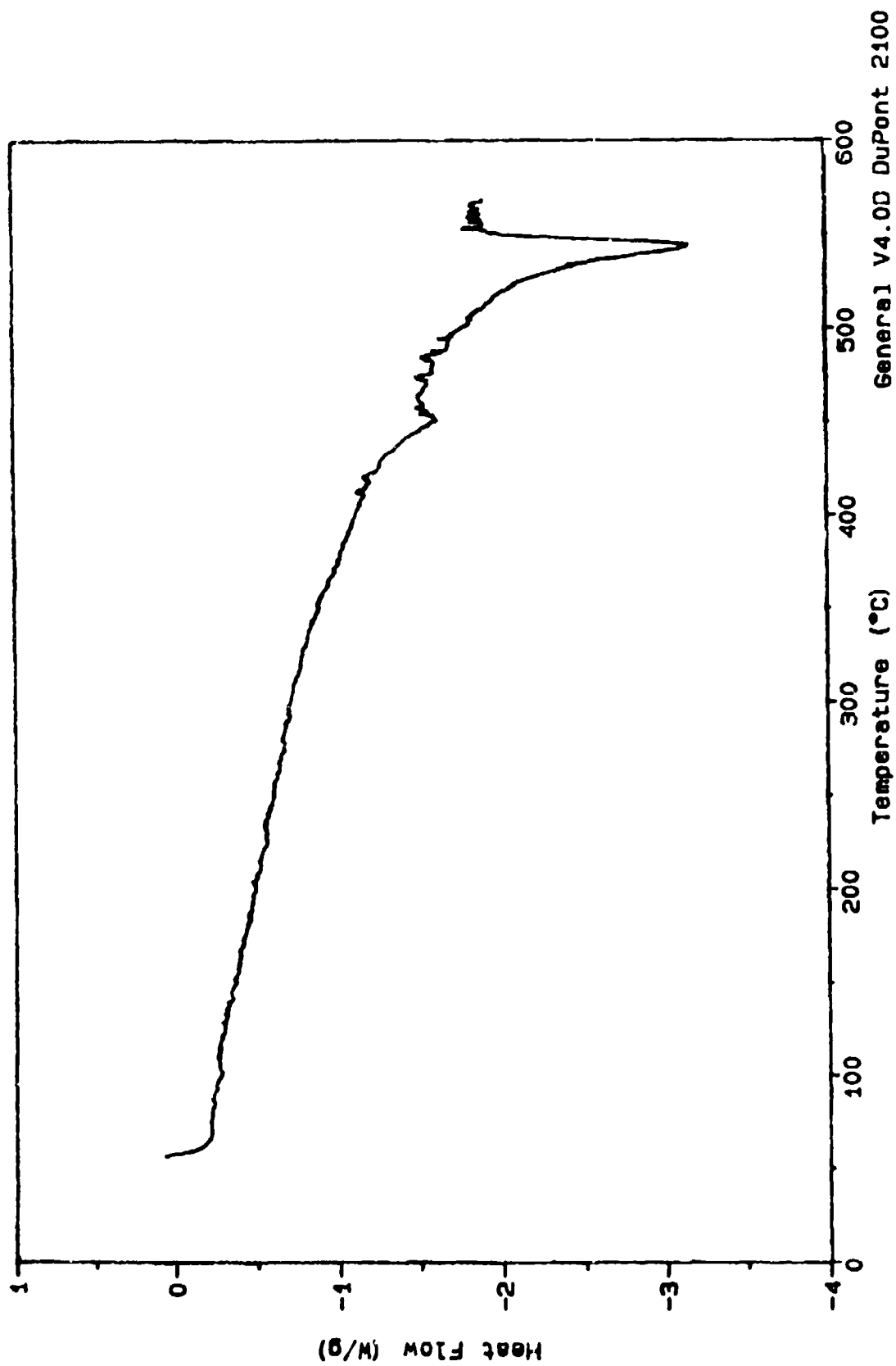


Figure 36. DSC of Generated Sb₂S₃ (Sb₂S₄ slow cooled from melt in H₂).

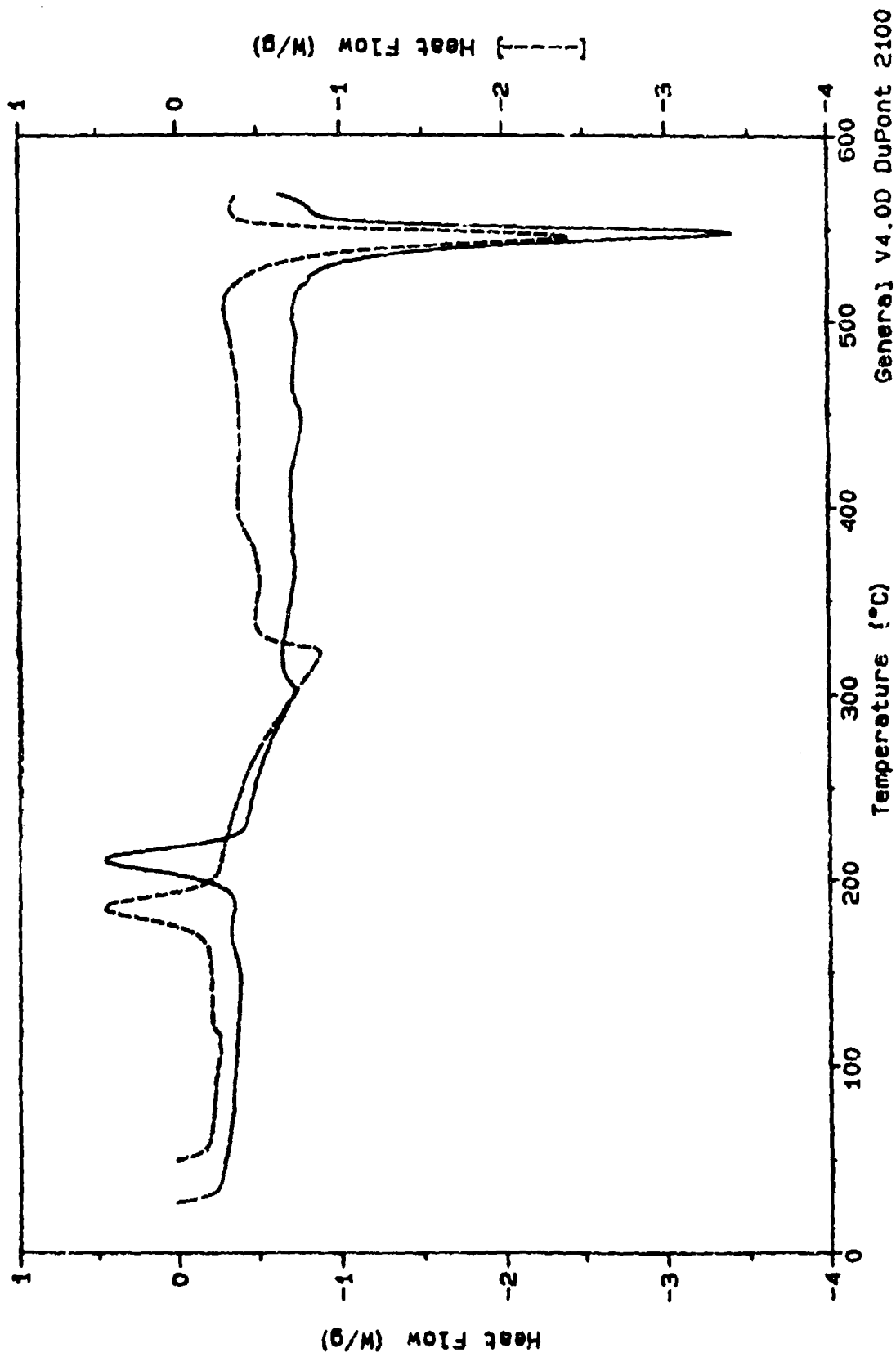


Figure 37. DSC of (—) Sb_2S_4 ; (---) Sb_2S_4 and Sulfur in N_2 .

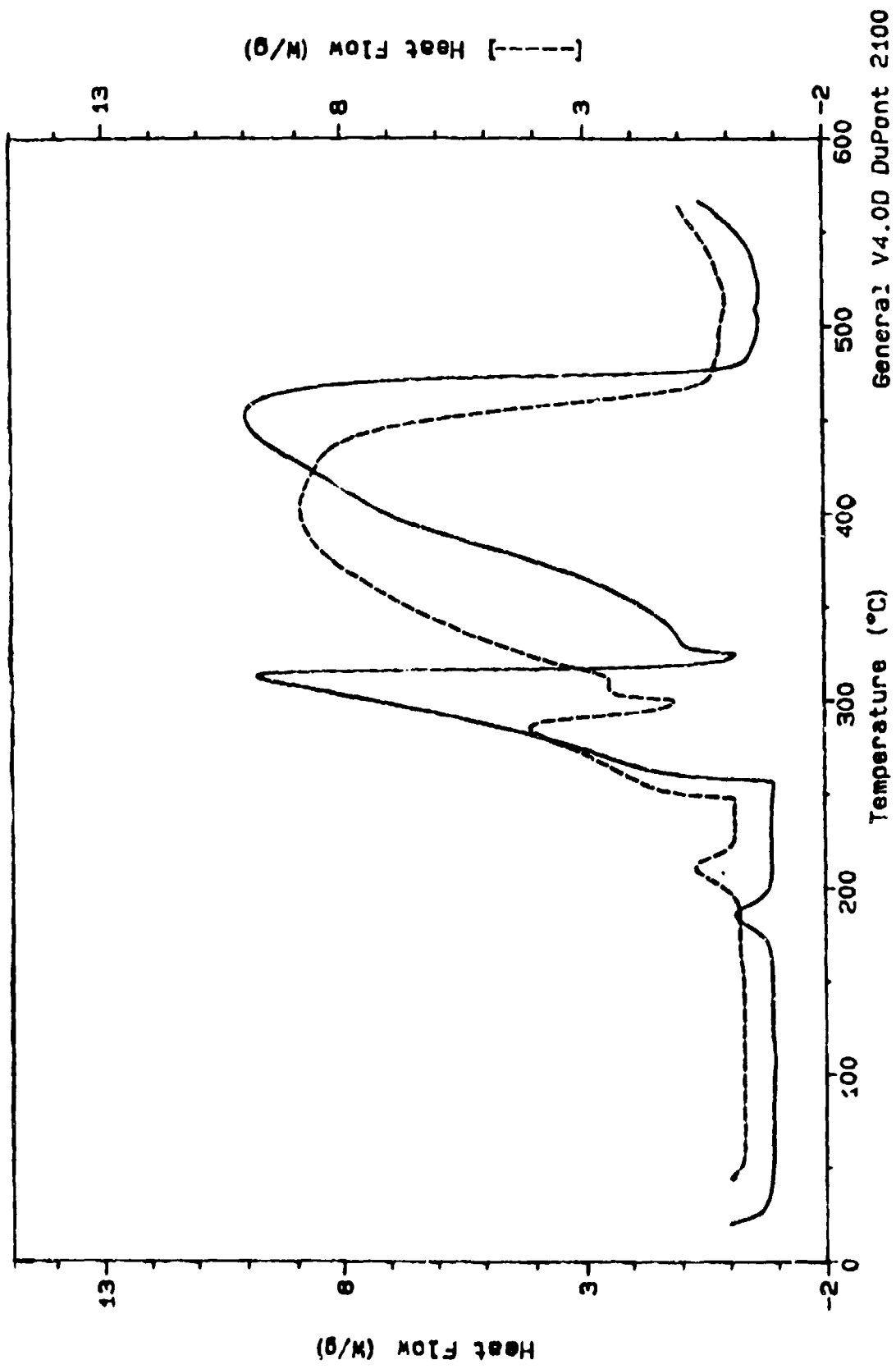


Figure 38. DSC of (—) Sb_2S_4 and Sulfur; (----) Sb_2S_4 in Air.

When extracted with CS_2 in a Soxhlet extraction apparatus, Sb_2S_5 experienced a 15 percent weight loss. This loss equalled the weight of sulfur recovered from evaporation of the CS_2 . Determination of antimony concentration by permanganate titration (14) gave 69.5 percent for extracted Sb_2S_5 and 57.59 percent for unextracted Sb_2S_5 . Theoretical values for Sb_2S_5 , and Sb_2S_3 are 60.30, and 71.69 percent respectively. The observed value for extracted Sb_2S_5 is closer to the theoretical value for Sb_2S_3 (71.9 percent) than Sb_2S_5 (60.3 percent); however, the color of the material remains reddish brown. This tends to support the theory that Sb_2S_5 is really amorphous Sb_2S_3 with excess sulfur.

Far IR spectra of unextracted and extracted Sb_2S_5 , along with Sb_2S_4 are shown in Figure 39. All three display similar patterns - a broadband centered around 290 cm^{-1} with a shoulder at about 150 cm^{-1} . Band maxima are at 292, 285, and 179 cm^{-1} , respectively. These results indicate that removal of excess sulfur is producing a structure similar to Sb_2S_4 .

Thermogravimetric analysis in nitrogen (Figure 40) shows a 16 percent weight loss occurring between 193 and 239 $^{\circ}\text{C}$. Unlike Sb_2S_4 , this weight loss does not occur in a single step. The curve is asymmetric and the derivative spectrum shows two peaks. This suggests that the excess sulfur is coordinated in the antimony sulfide in two distinct ways. From 200 $^{\circ}\text{C}$ on, the sample shows the same thermal stability as Sb_2S_3 and Sb_2S_4 - ultimately

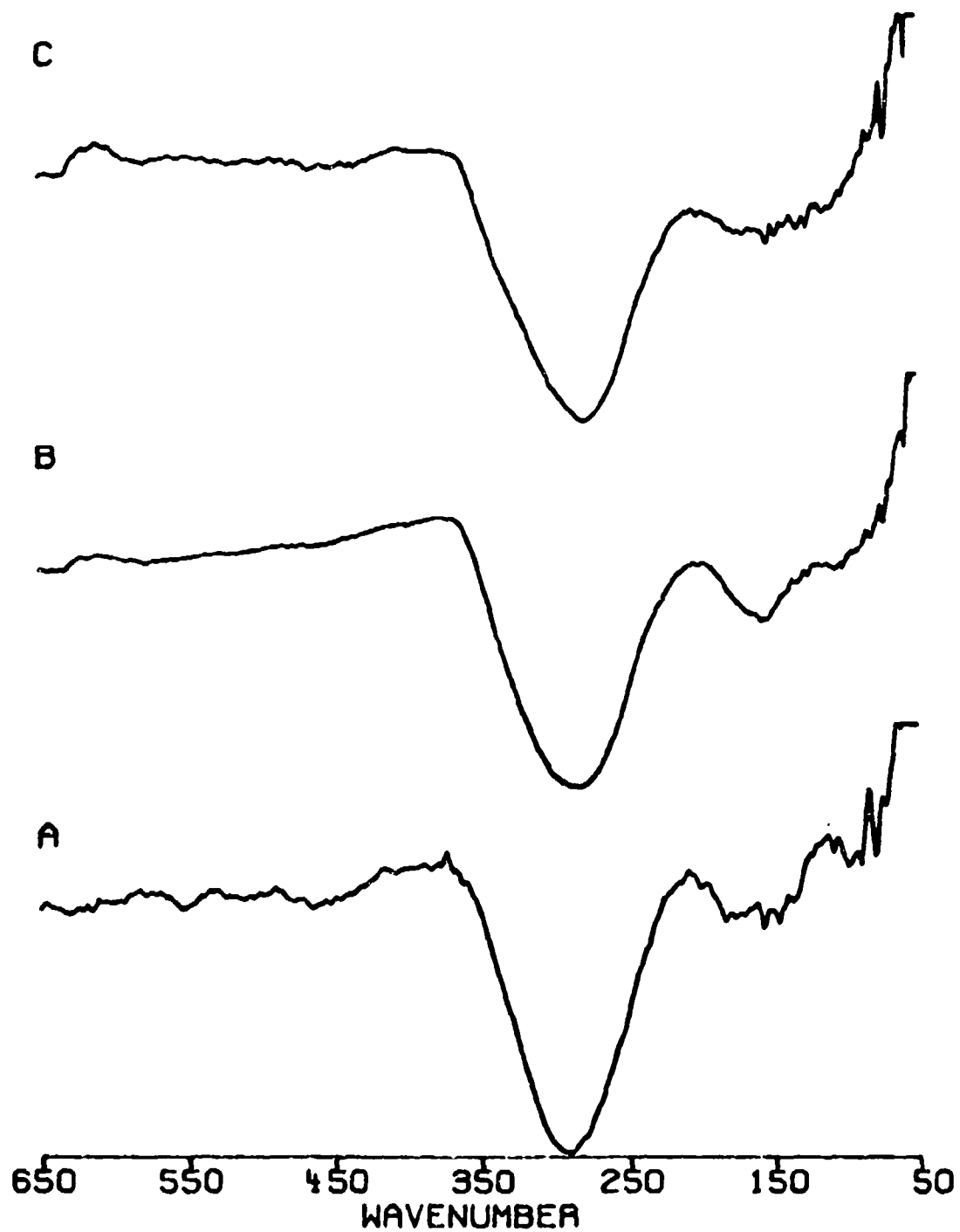


Figure 39. Far IR of Sb_2S_5 (A) Sb_2S_5 ; (B) Sb_2S_5 extracted;
(C) Sb_2S_4 .

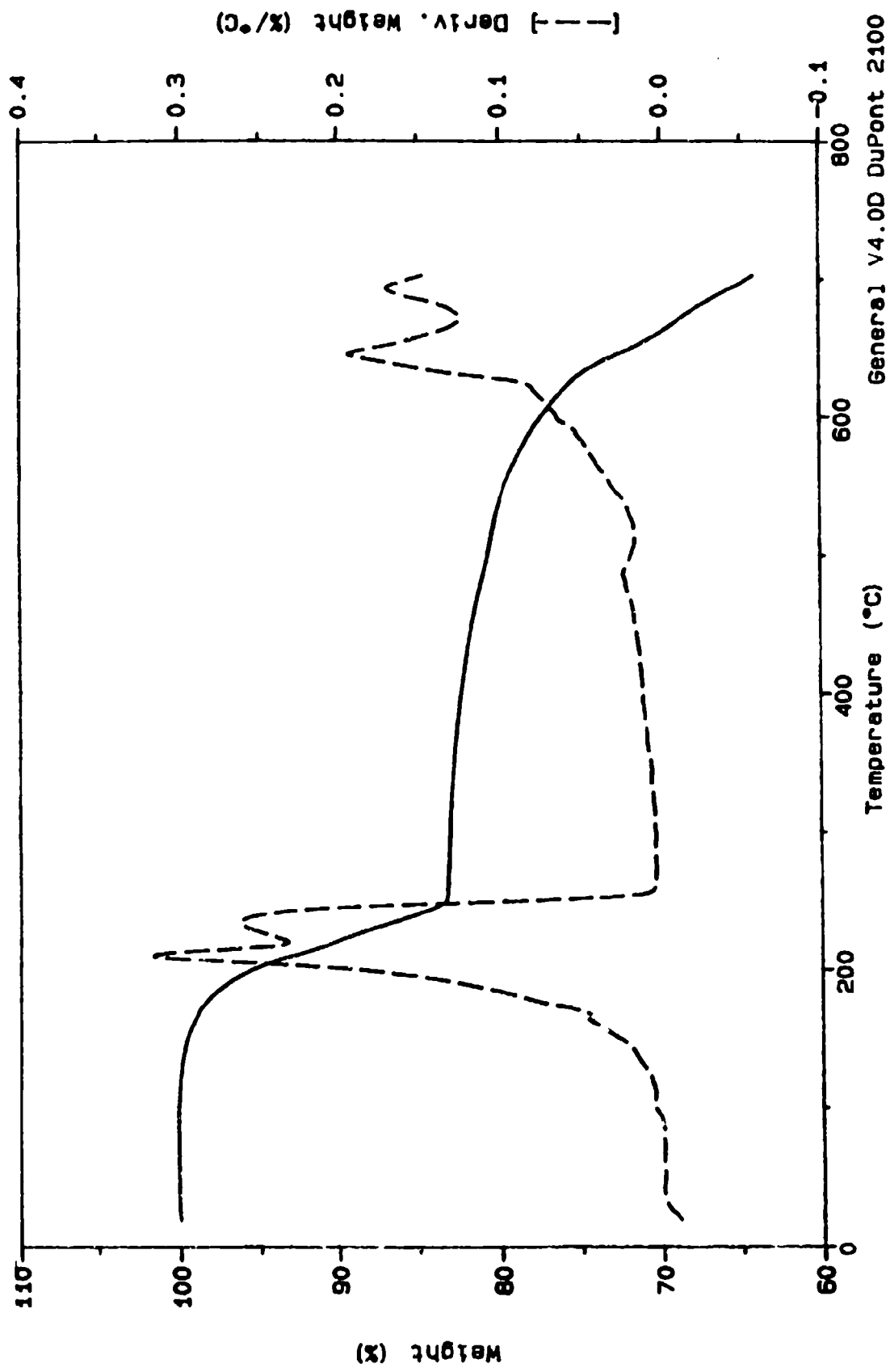


Figure 40. TGA of Sb_2S_5 (—) % wt loss; (----) 1st derivative.

reacting with the platinum pan at temperatures above 600 °C. The initial weight loss is close to that predicted for reaction 2 (16 percent vs 15.9 percent). It is also close to the amount of sulfur obtained by extraction.



Results of DSC on Sb_2S_5 under nitrogen are shown in Figure 41. Sb_2S_5 behaves in the same manner as Sb_2S_4 with the exception of the presence of excess sulfur as seen by the melting endotherm at 115 °C. (Note: only the beta form is observed.) The first exotherm occurs at 231 °C (15 J/g) followed by the melt at 553 °C which is 5 °C higher than the one for Sb_2S_4 . Again, this increase may be due to a combination of factors - generated particle size, crystallinity and/or reactions with the sample pan. When DSC and TGA curves are overlaid (Figure 42), one sees that the exotherm coincides with the second portion of the weight loss. This supports the idea that the excess sulfur is present (or bound) within the antimony sulfide in two distinct ways.

Isothermal TGA experiments support the loss of sulfur to produce Sb_2S_3 . When Sb_2S_5 is heated under nitrogen to 270 °C (just beyond the 231 °C exotherm) and the product is examined by IR, a spectrum matching that of Sb_2S_3 is observed. (See Figure 43.)

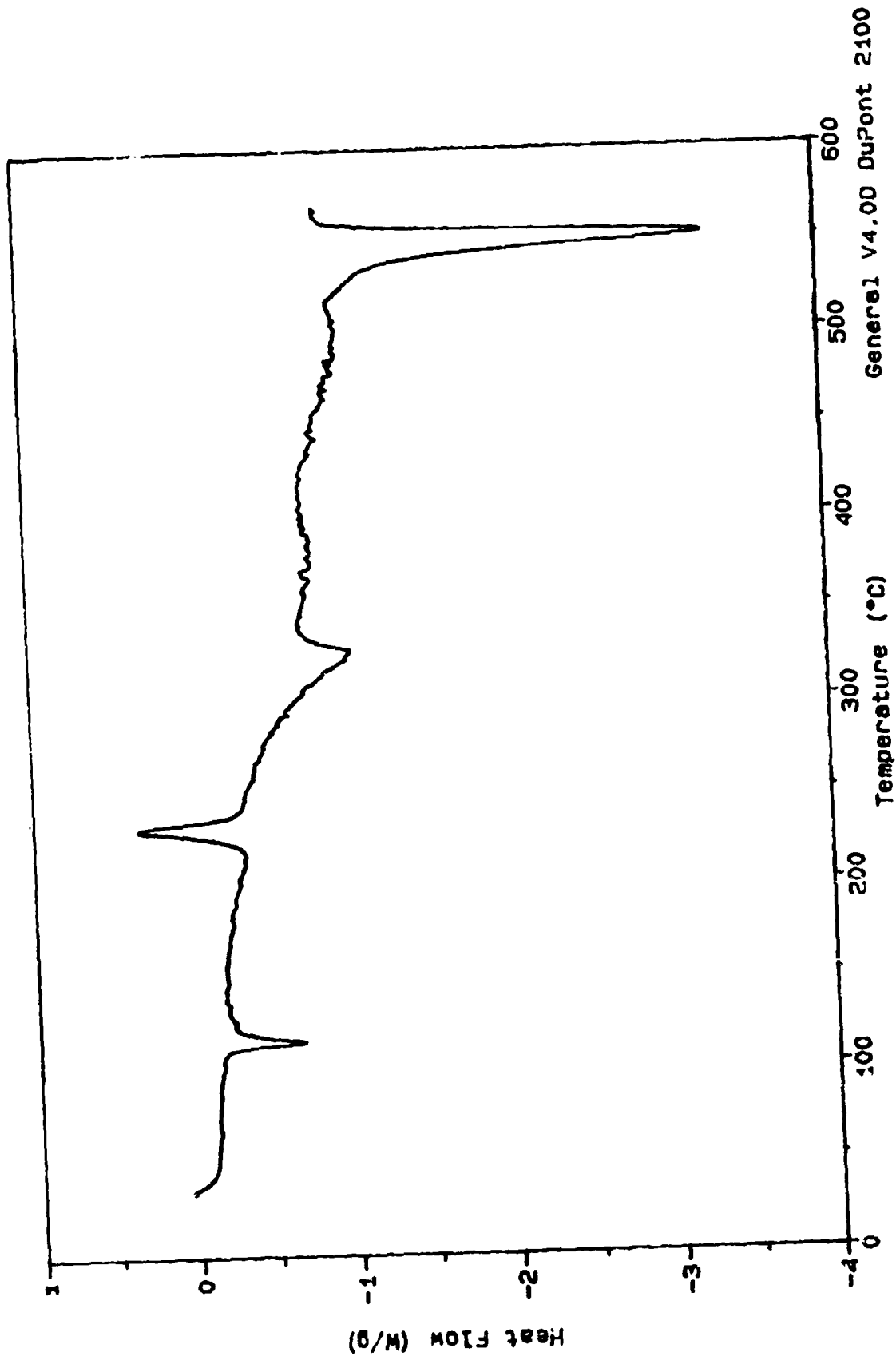


Figure 41. DSC of Sb_2S_5 in N_2 .

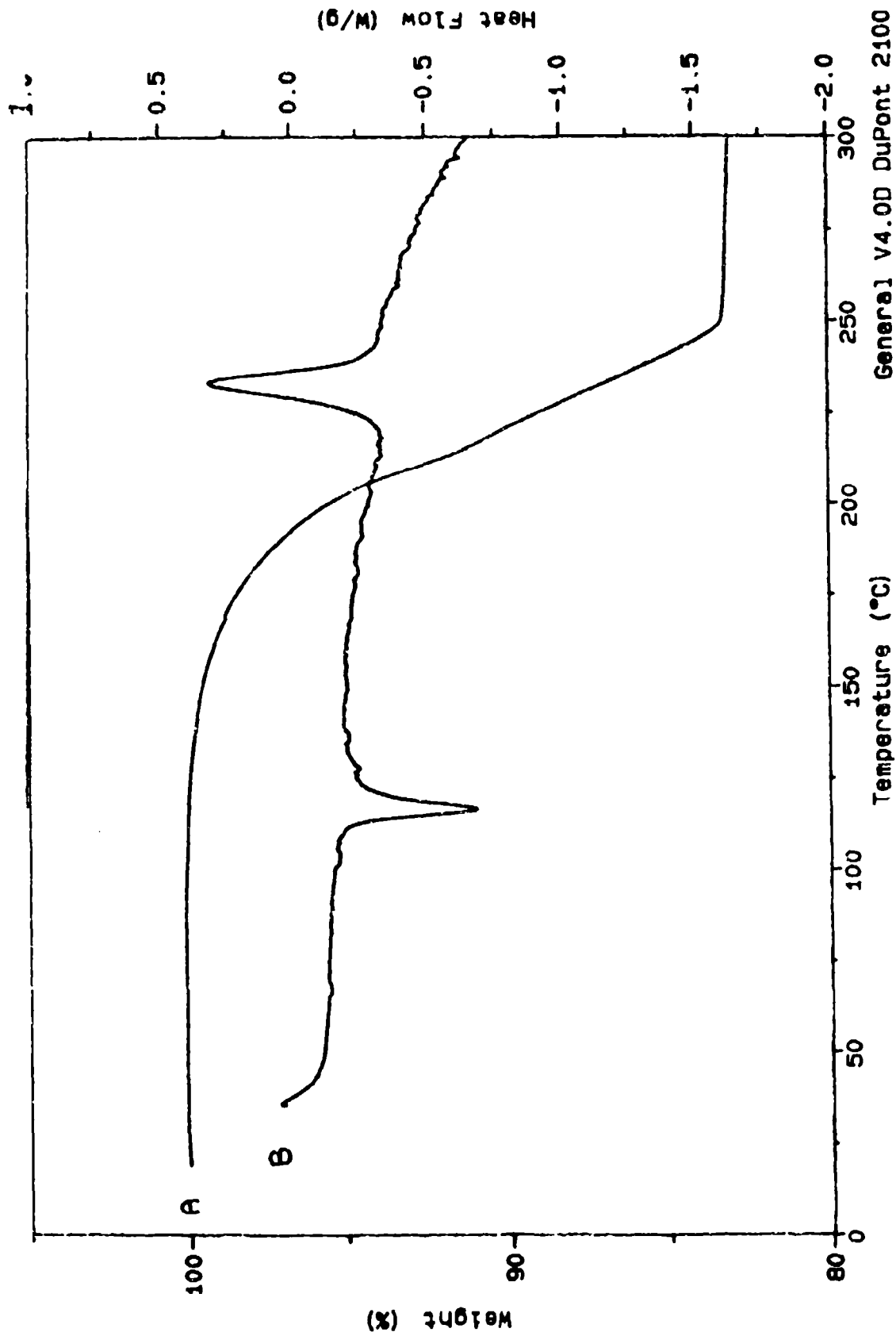


Figure 42. DSC - TGA Comparison of Sb_2S_5 (A) TGA (B) DSC. (Both in N_2).

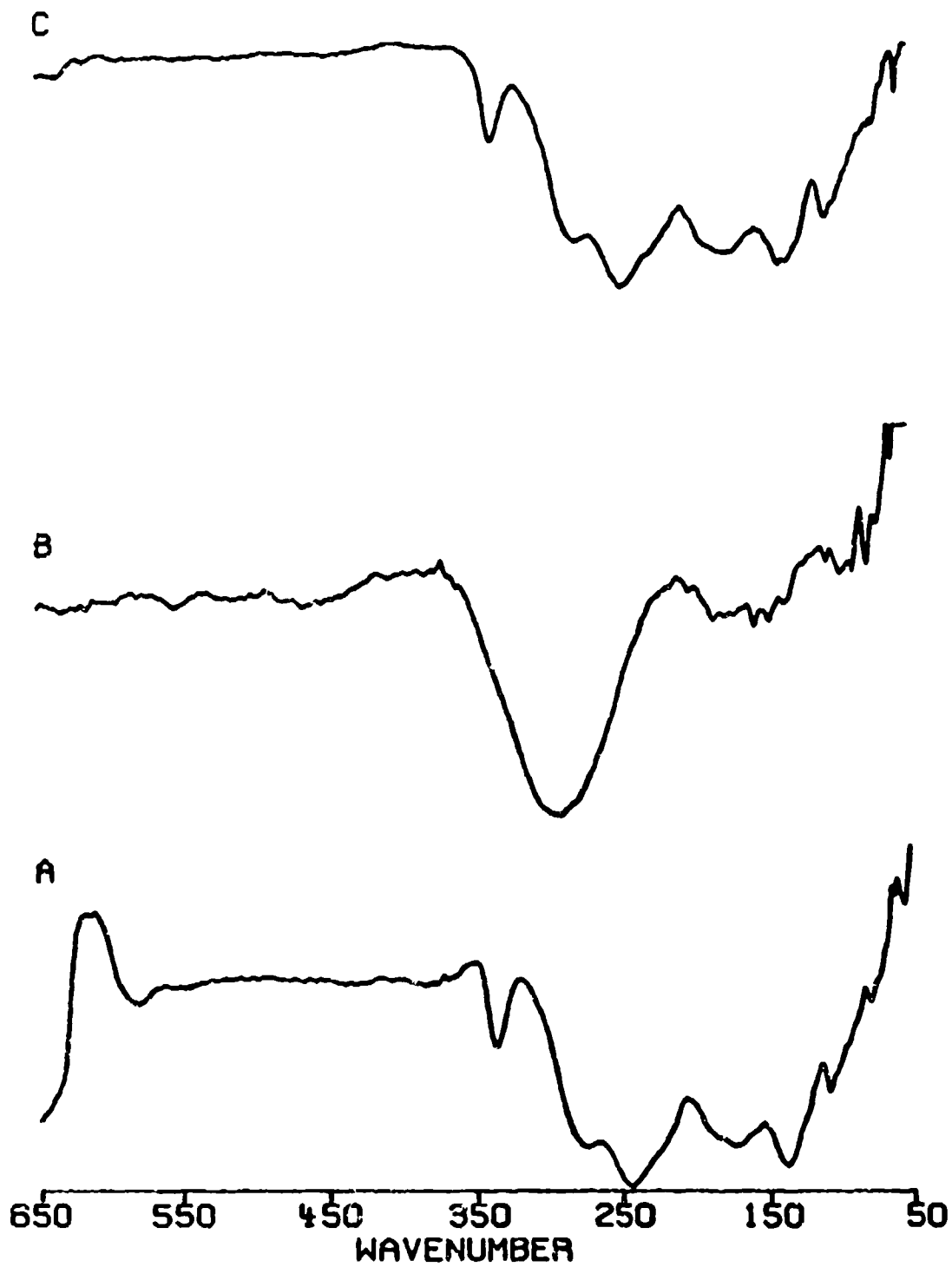


Figure 43. Far IR of Sb_2S_5 (A) Sb_2S_3 ; (B) Sb_2S_5 ; (C) Sb_2S_5 at 270°C in N_2 .

Like other antimony sulfides, the pentasulfide undergoes a complex oxidation. TGA (Figure 44) shows an initial weight loss occurring in a biphasic manner (as in nitrogen), followed by a series of incremental losses from 300 to 600 °C. This multistep oxidation can be seen in the DTA and DSC curves. (See Figures 45 & 46.) DTA results are similar to those of Sb_2S_4 ; two large asymmetric exotherms (292 and 403 °C) followed by a smaller one at 604 °C. The larger exotherms are better resolved in the DSC. The 292 °C peak is shifted to 290 °C with shoulders at 259 and 280 °C. The 403 °C exotherm remains; however, a shoulder at 333 °C is observed.

Oxidation of the pentasulfide in the TGA, followed by IR analysis produced results similar to those for Sb_2S_4 and Sb_2S_3 . Figure 47 shows the Far IR spectra of Sb_2S_5 heated in air to 400, 25 and 700 °C. At 400 °C, a combination of cubic Sb_2O_3 and Sb_2S_3 can be seen. The sample heated to 525 °C shows no signs of Sb_2S_3 , but does reveal a combination of cubic Sb_2O_3 (384 and 180 cm^{-1}) and Sb_2O_4 (434 and 605 cm^{-1}). As with Sb_2S_3 and Sb_2S_4 , temperatures above 600 °C yield a spectrum matching Sb_2O_4 . Mid IR analysis reveals the presence of the sulfate containing species common to only air heated samples. (See Figures 48 & 49.)

Heating Sb_2S_5 to a melt in the DSC, rapidly cooling, and rerunning under nitrogen produces the thermogram shown in Figure 50. A glass transition at 228 °C followed by crystallization at 271 °C (35 J/g) indicates that amorphous Sb_2S_3 formed from rapid

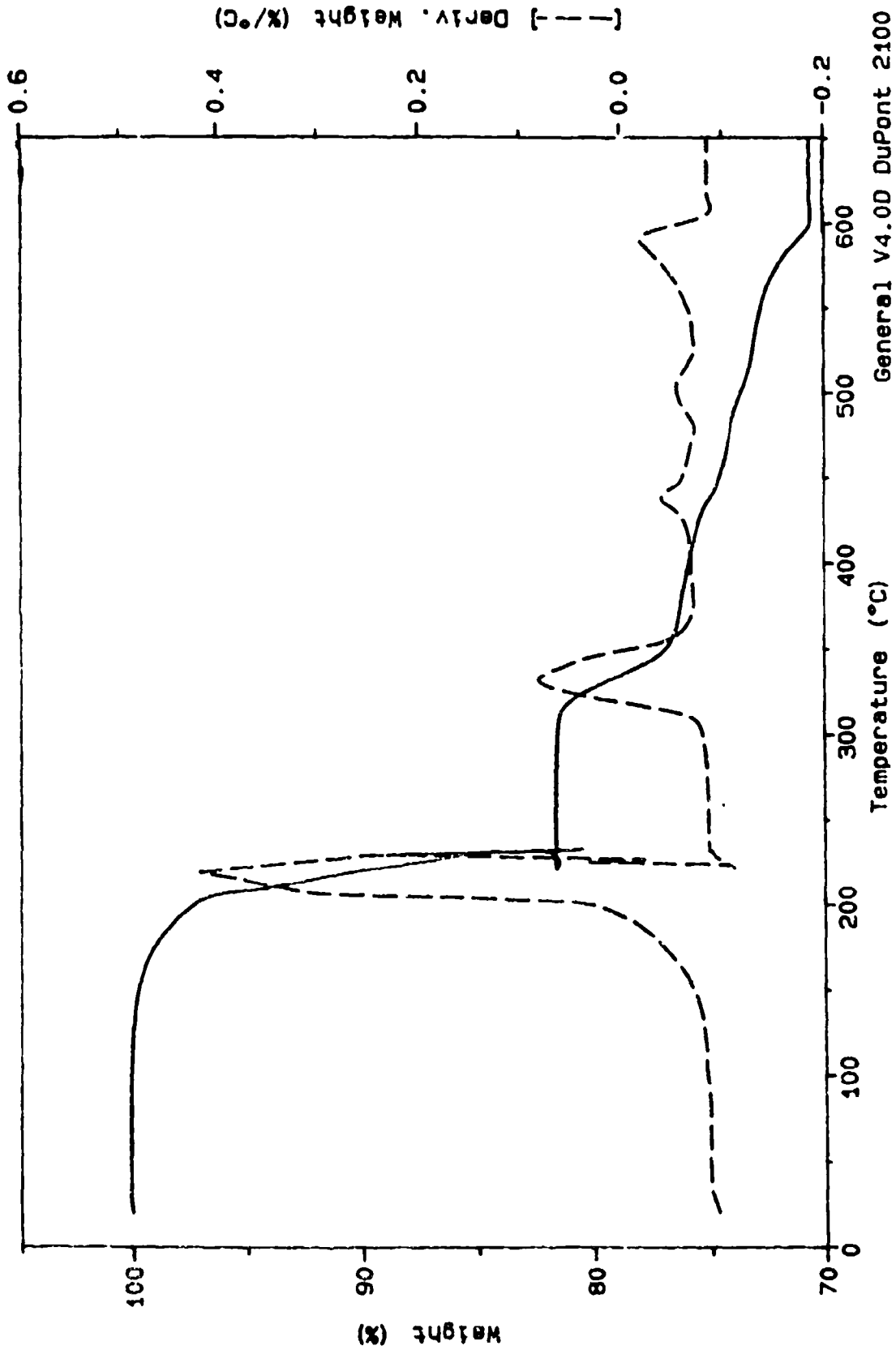


Figure 44. TGA of Sb_2S_5 in Air.

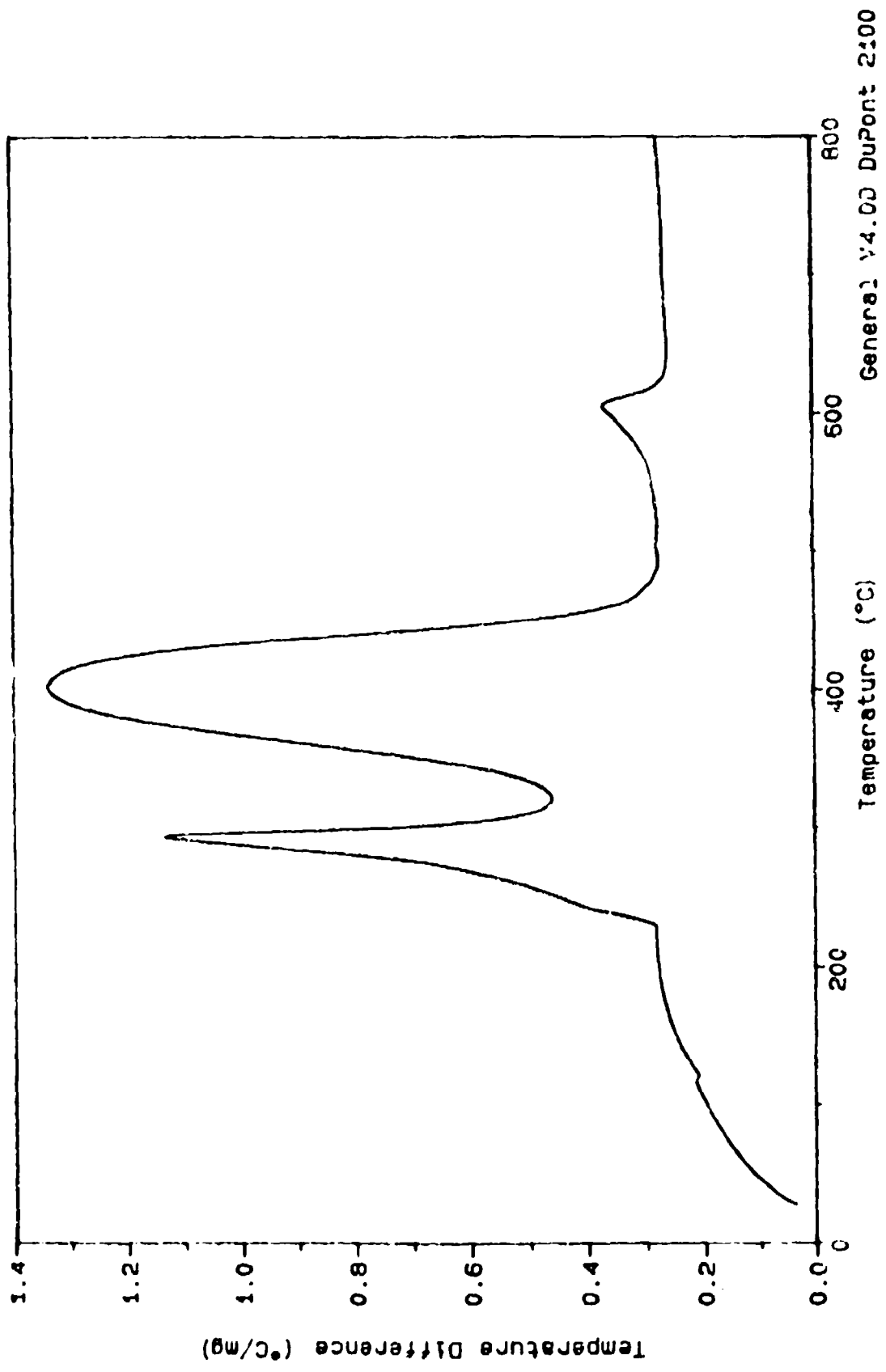


Figure 45. DTA of Sb_2S_5 in Air.

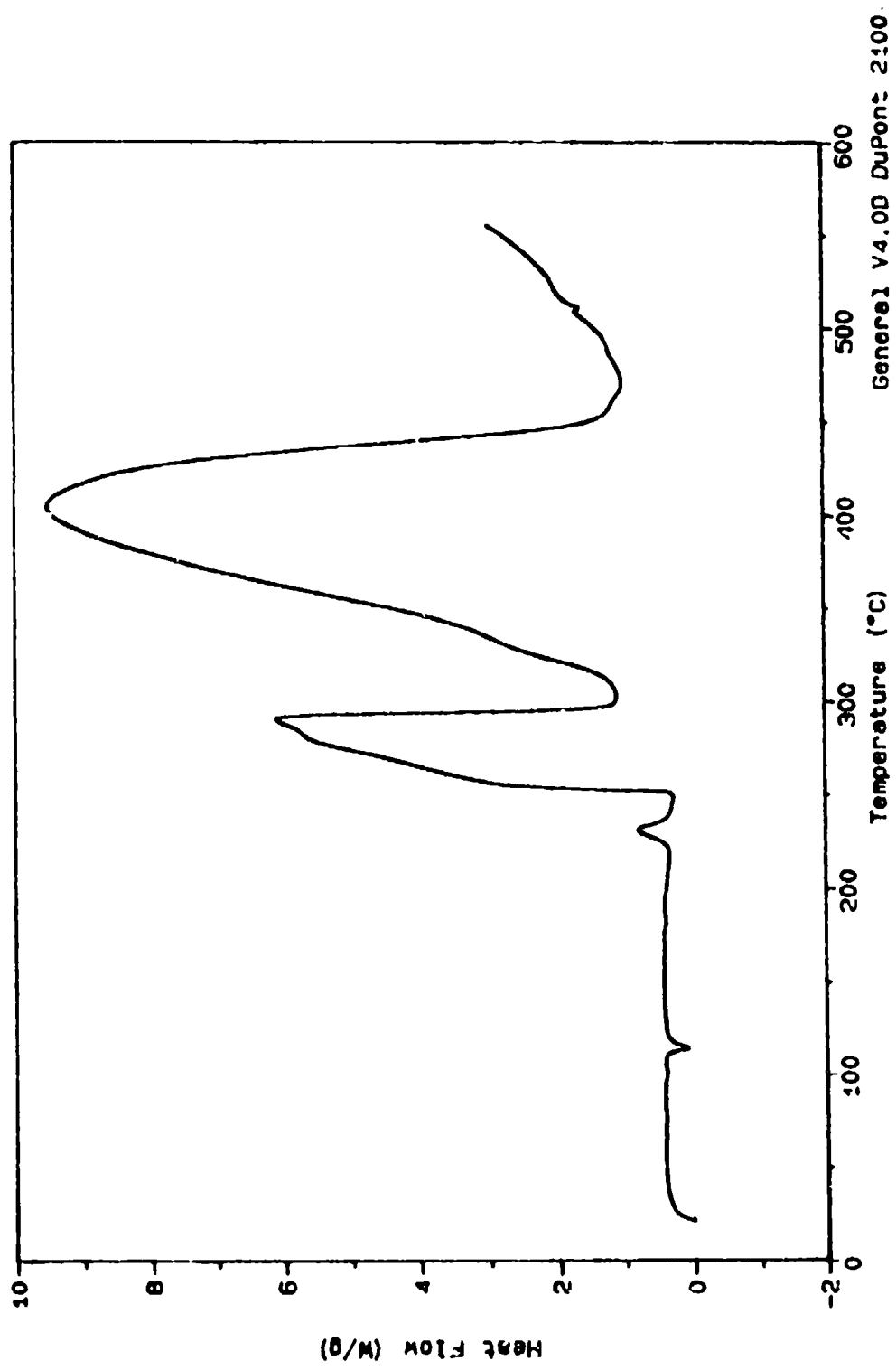


Figure 46. DSC of Sb_2S_5 in Air.

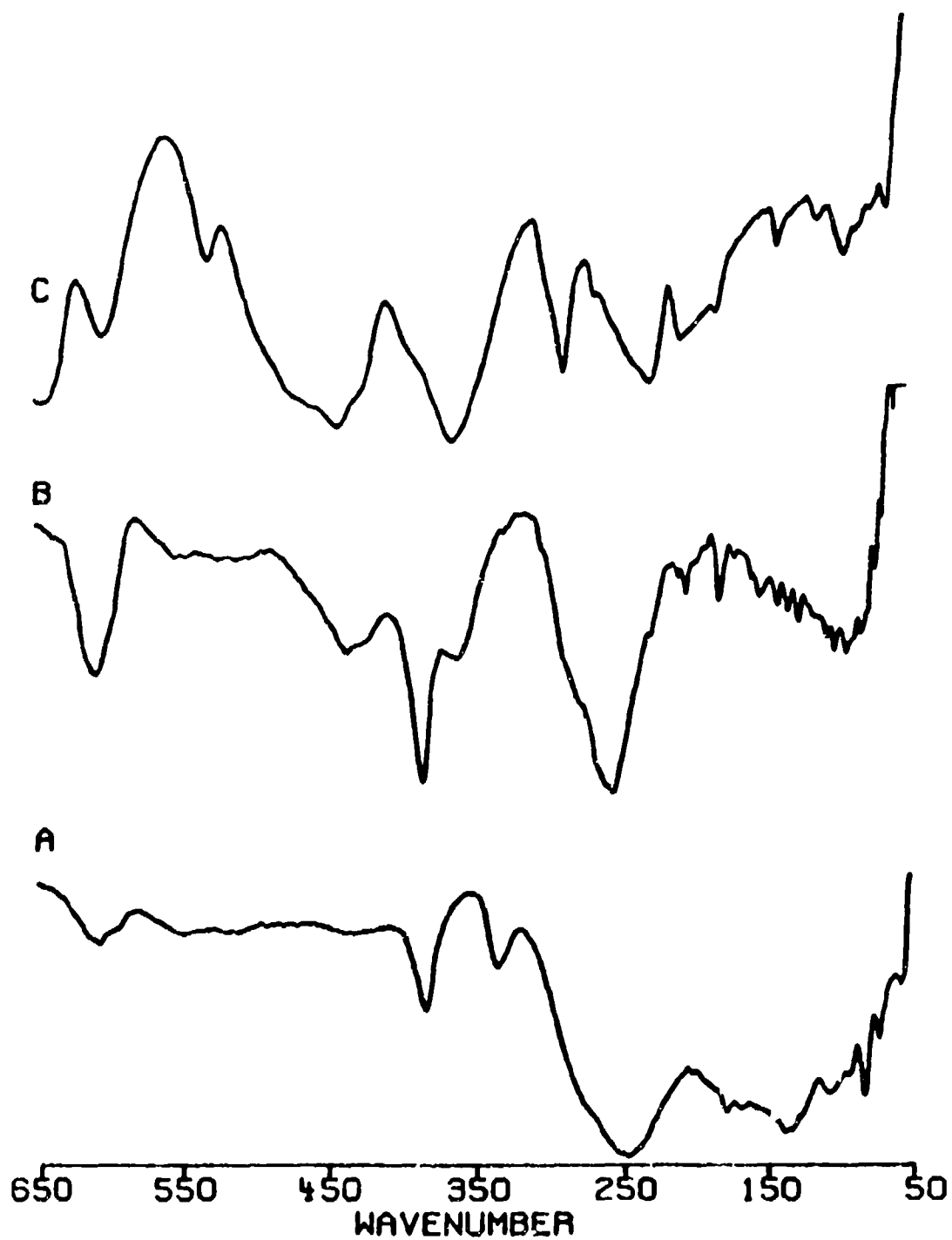


Figure 47. Far IR of Sb_2S_5 (A) 400°C in Air; (B) 525°C in Air; (C) 700°C in Air.



Figure 48. Mid IR of Sb_2S_5 (A) 400°C in Air; (B) 525°C in Air;
(C) 700°C in Air.

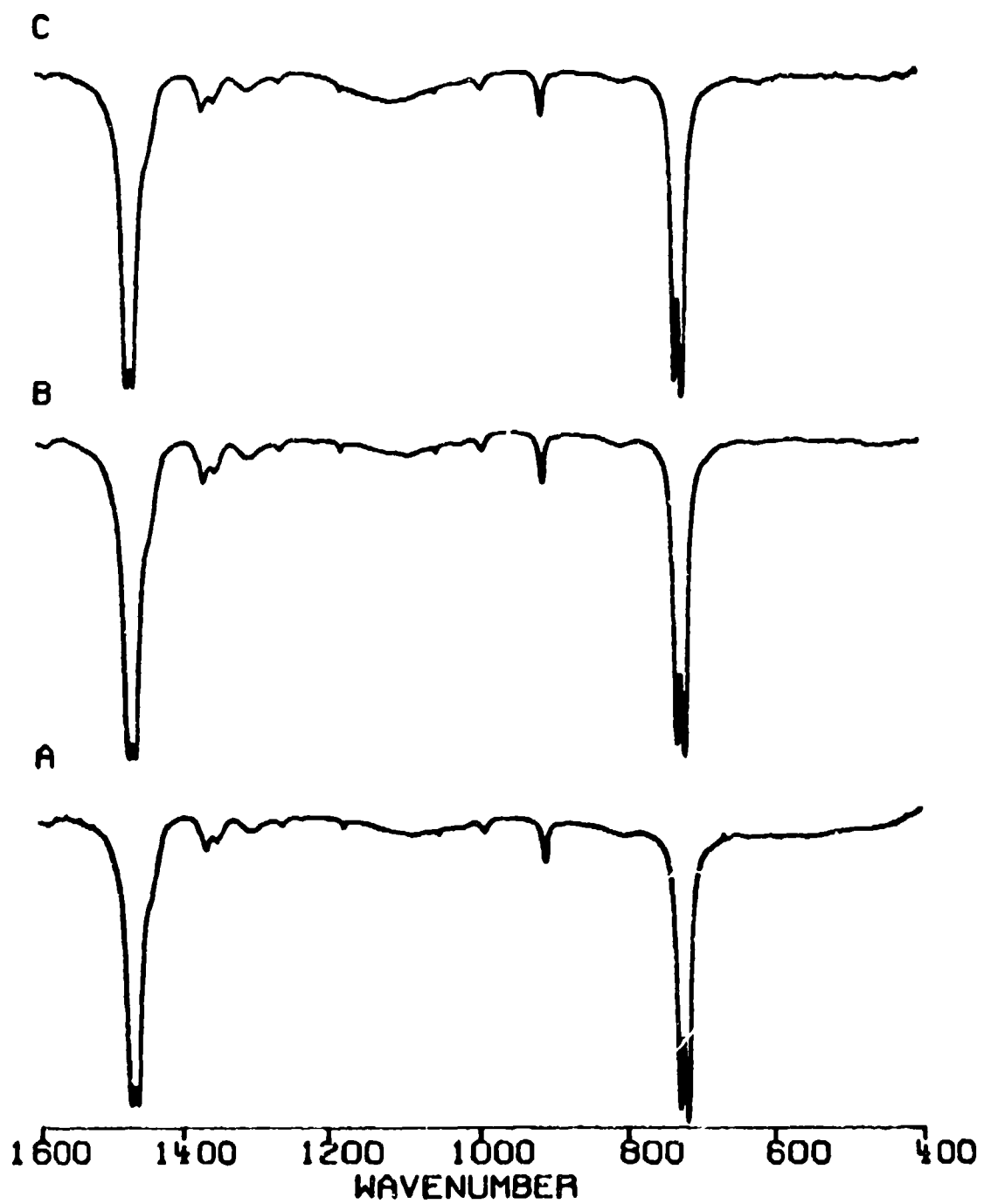


Figure 49. Mid IR of Sb_2S_5 (A) Sb_2S_3 ; (B) Sb_2S_5 ; (C) Sb_2S_5 at 275°C in N_2 .

cooling the melt. When slowly cooled, Sb_2S_3 solidifies in its crystalline form as evidenced by the lack of a glass transition and heat of crystallization. (See Figure 51.)

The thermal stability of extracted antimony pentasulfide is somewhat different from the unextracted material. In TGA experiments, (Figure 52) both under nitrogen and air, only a 2 percent weight loss is observed between 227 and 243 °C as compared to 16 percent for the unextracted material. This 2 percent is most likely due to residual sulfur not removed by extraction. After the initial loss, in nitrogen, the extracted material demonstrates the same thermal stability as Sb_2S_3 , ultimately reacting with the platinum pan at temperatures above 600 °C.

In air, extracted Sb_2S_5 undergoes a series of weight losses between 290 and 580 °C. These losses are similar in pattern to unextracted Sb_2S_5 (Figure 53), but different in magnitude. A loss of 8.9 percent occurs at 291 °C in the extracted sample - 4 percent less and 25 °C earlier than unextracted Sb_2S_5 . The total weight loss for the extracted sample is 16 percent, compared with 29 percent for unextracted Sb_2S_5 - a difference approximately equal to the amount of sulfur removed by extraction.

DSC experiments on extracted Sb_2S_5 parallel those of the unextracted Sb_2S_5 with the exception of the melting and volatilization of the excess sulfur. (See Figure 54.) On closer

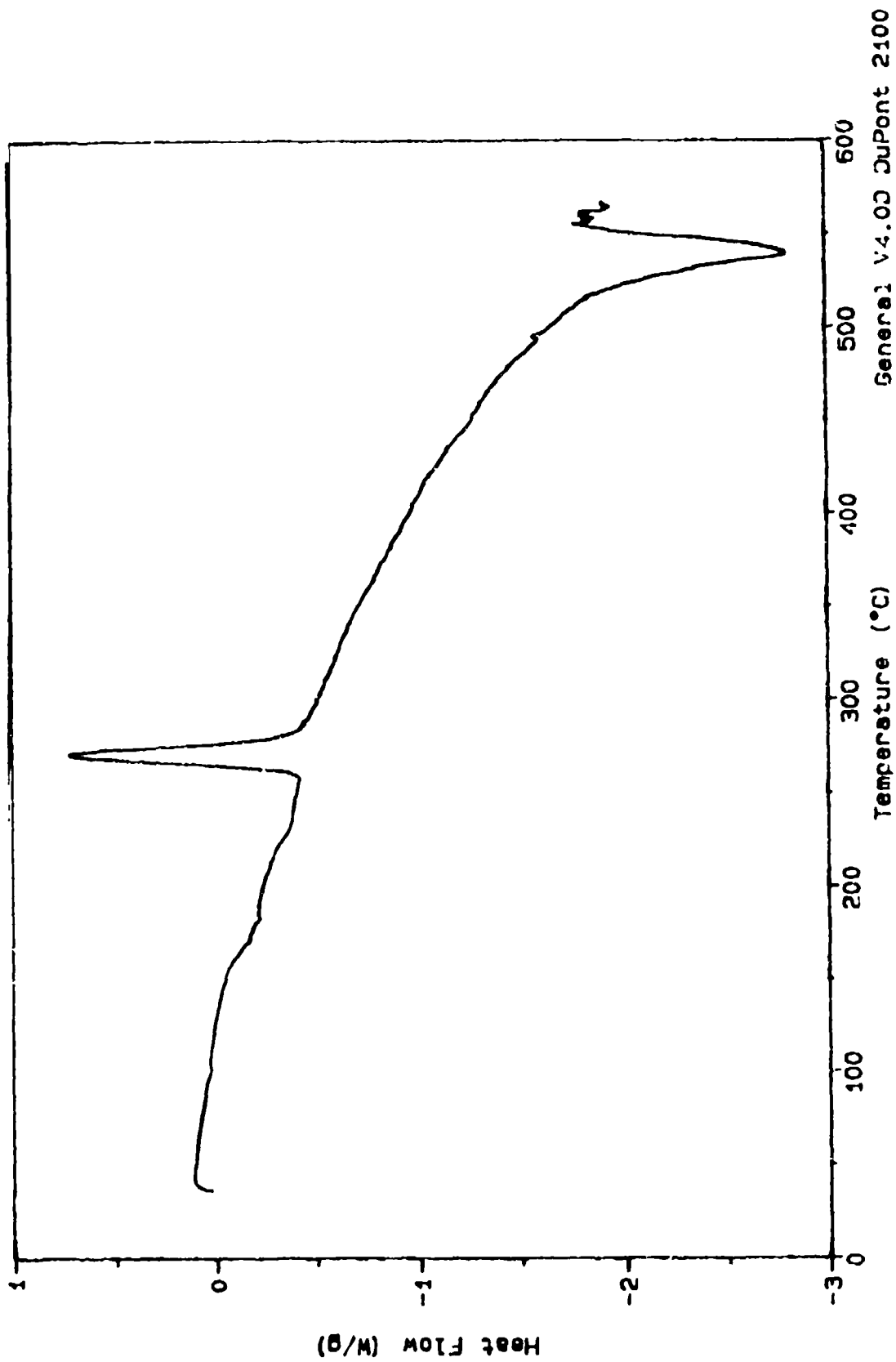


Figure 50. DSC of Generated Amorphous Sb_2S_3 (Sb_2S_5 rapidly cooled from melt in N_2).

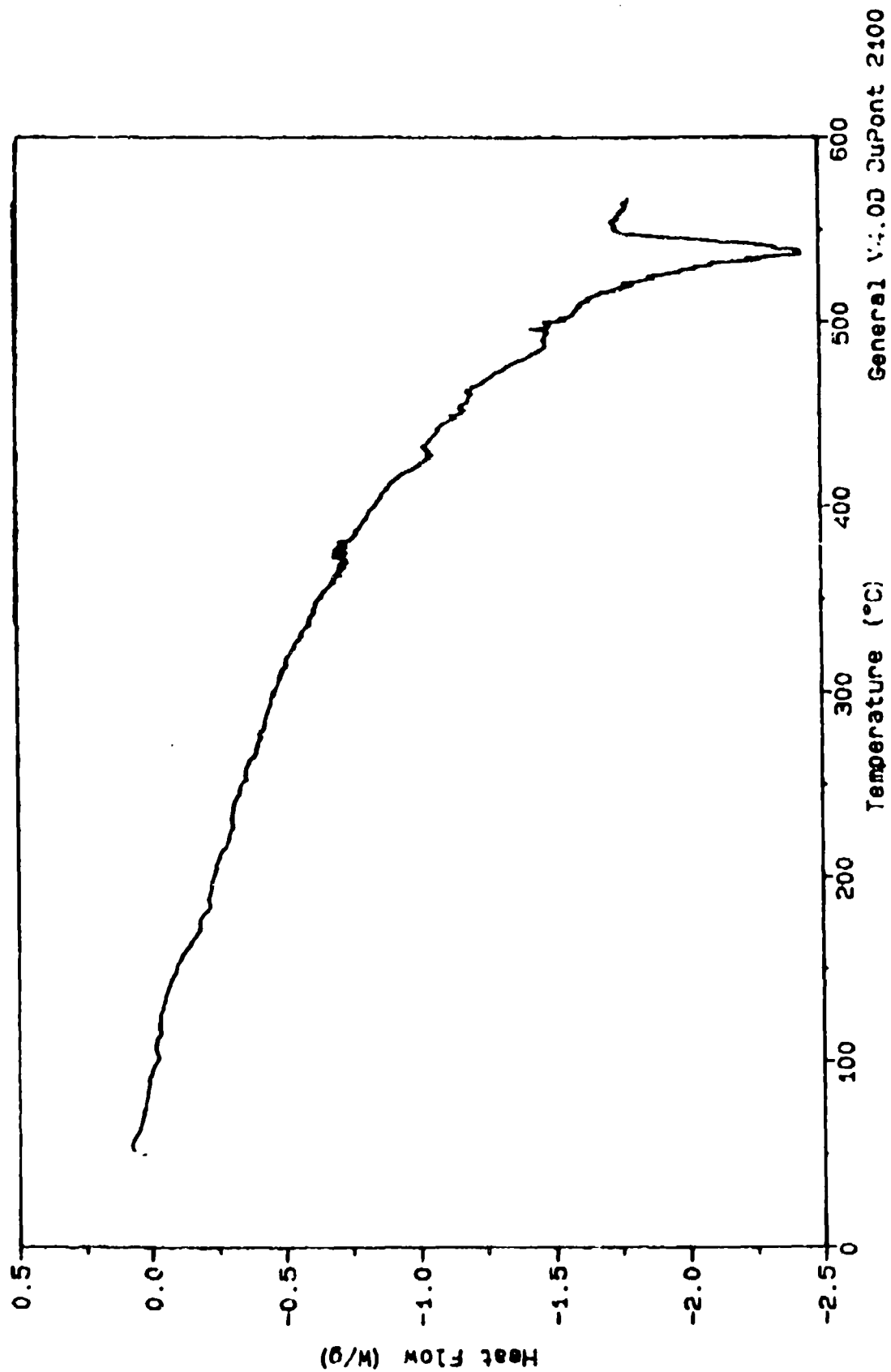


Figure 51. DSC of Generated Crystalline Sb_2S_3 (Sb_2S_3 slow cooled from melt in N_2).

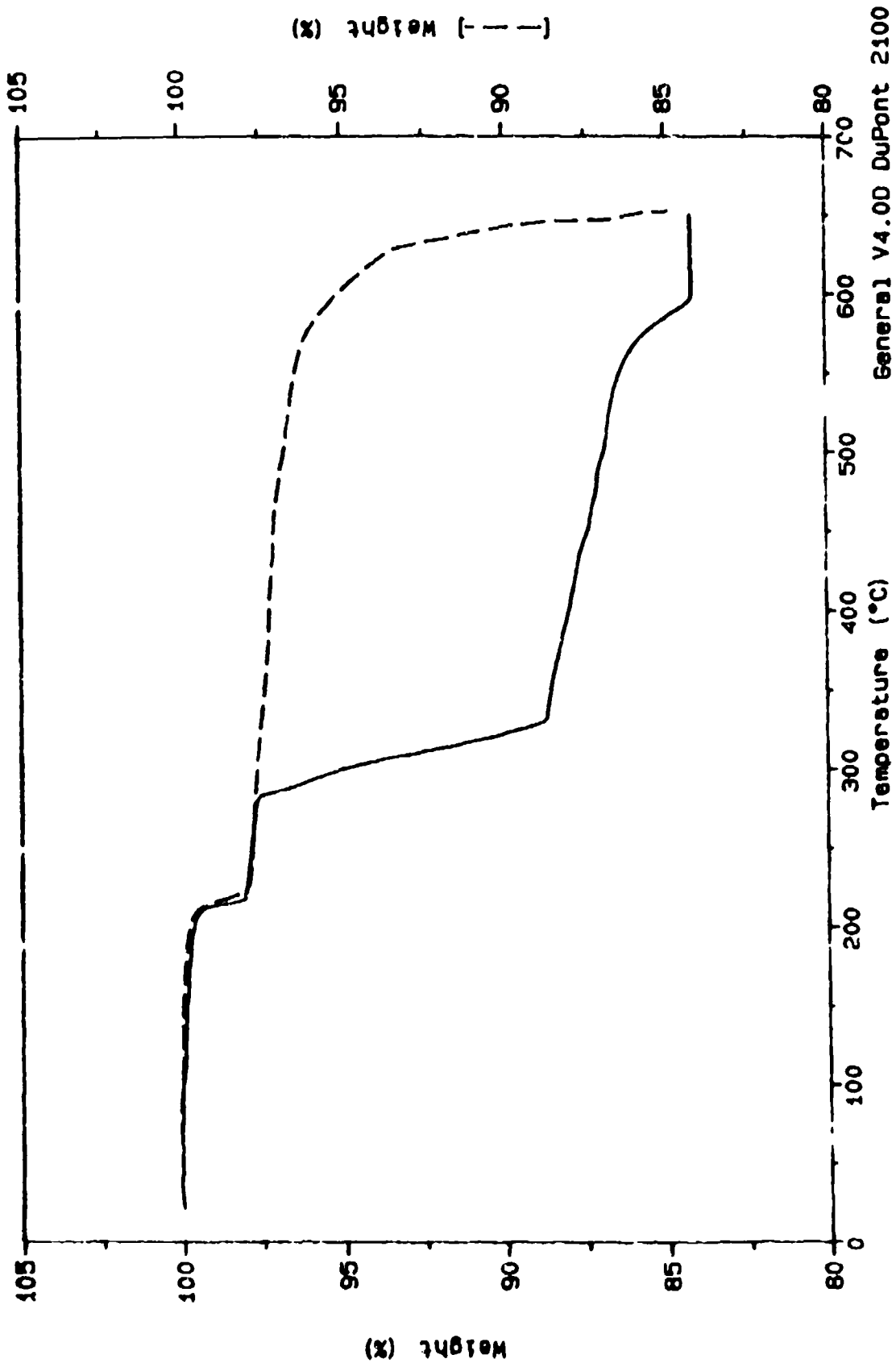


Figure 52. TGA of Extracted Sb_2S_5 (—) in Air; (----) in N_2 .

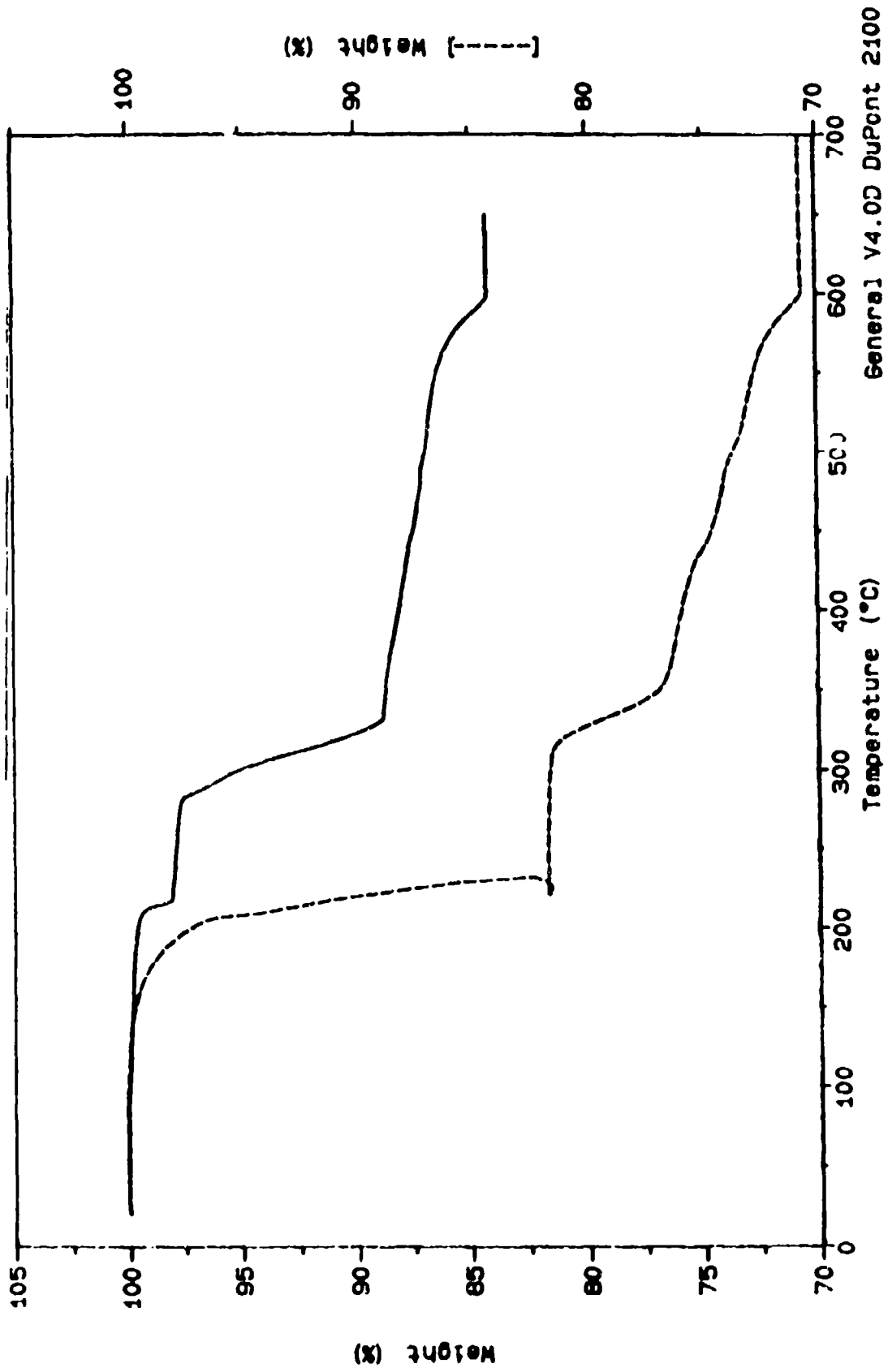


Figure 53. TGA of (-----) Extracted vs (—) Unextracted Sb_2S_5 .

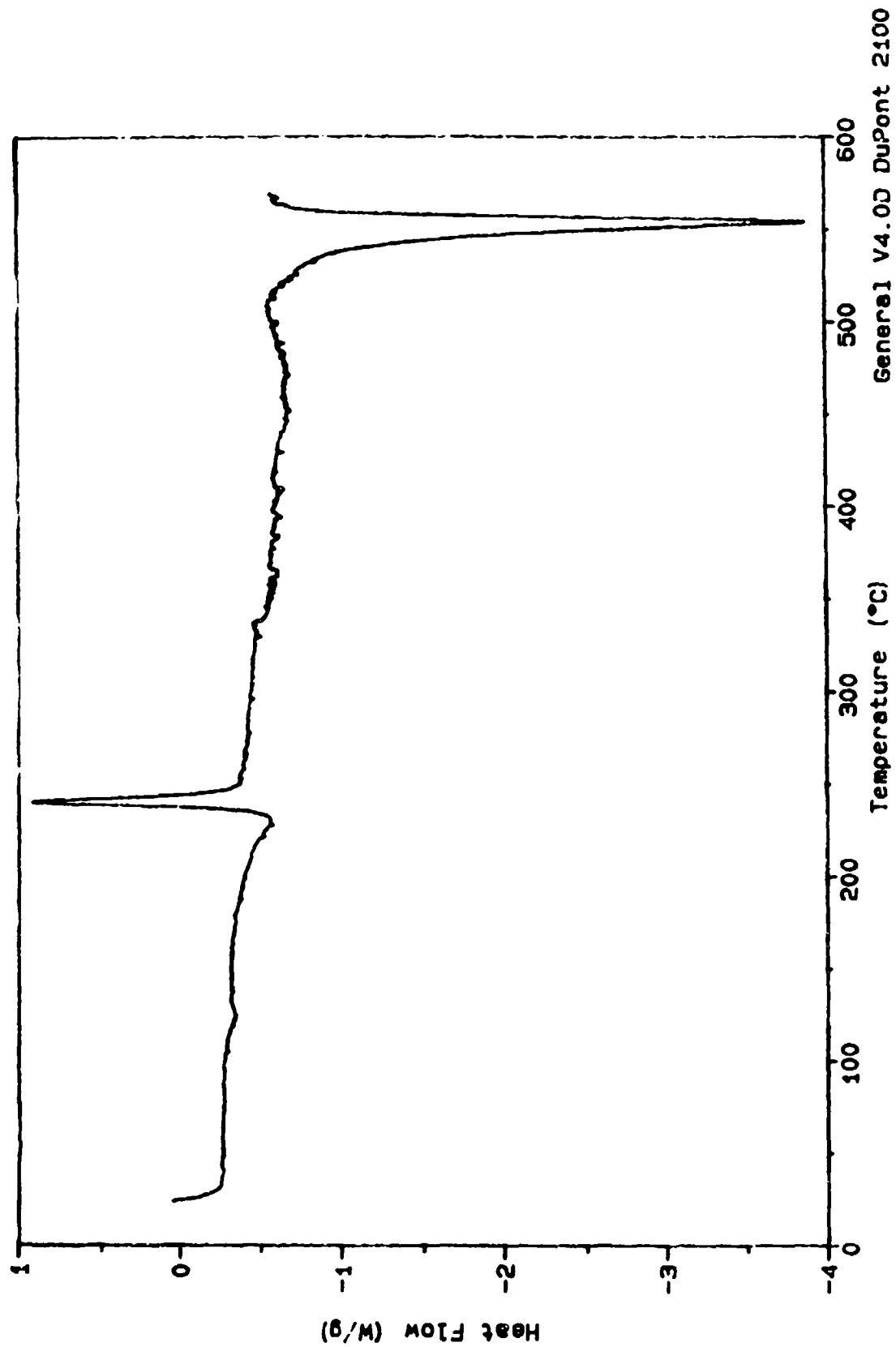


Figure 54. DSC of Generated Amorphous Sb_2S_3 (Extracted Sb_2S_5 rapidly cooled from melt).

examination, one sees that the exotherm is asymmetric and contains a contribution from an underlying endotherm. The maximum at 239 °C is 7 °C higher than that observed for unextracted Sb_2S_5 . This does not however, take into account the contribution from the endotherm, the real maxima may well be the same as that of the unextracted sample. The melting point of the generated Sb_2S_3 is unaffected by the extraction process.

Results of DSC performed on extracted Sb_2S_5 under oxidative conditions, are shown in Figure 55 along with previously reported results for the unextracted material. The initial exotherm occurs at a slightly higher temperature (241 °C), also the magnitude of the exotherm is larger - 23 J/g vs 12 J/g. The higher temperature may be the result of the extracted sulfur. The corresponding exotherm in Sb_2S_4 shifted to lower temperatures when elemental sulfur was added.

The extracted Sb_2S_5 does not display the two well defined exotherms at 292 and 403 °C, only one broadband containing several shoulders is observed. Again, this may be a result of a smaller amount of sulfur present in the extracted sample reducing the size of the exotherm at 286 °C, previously identified as the oxidation of sulfur, to a point where it appears as a shoulder on the larger exotherm.

Infrared analysis of samples of extracted Sb_2S_5 heated under oxidative conditions mirror results obtained for all the sulfides

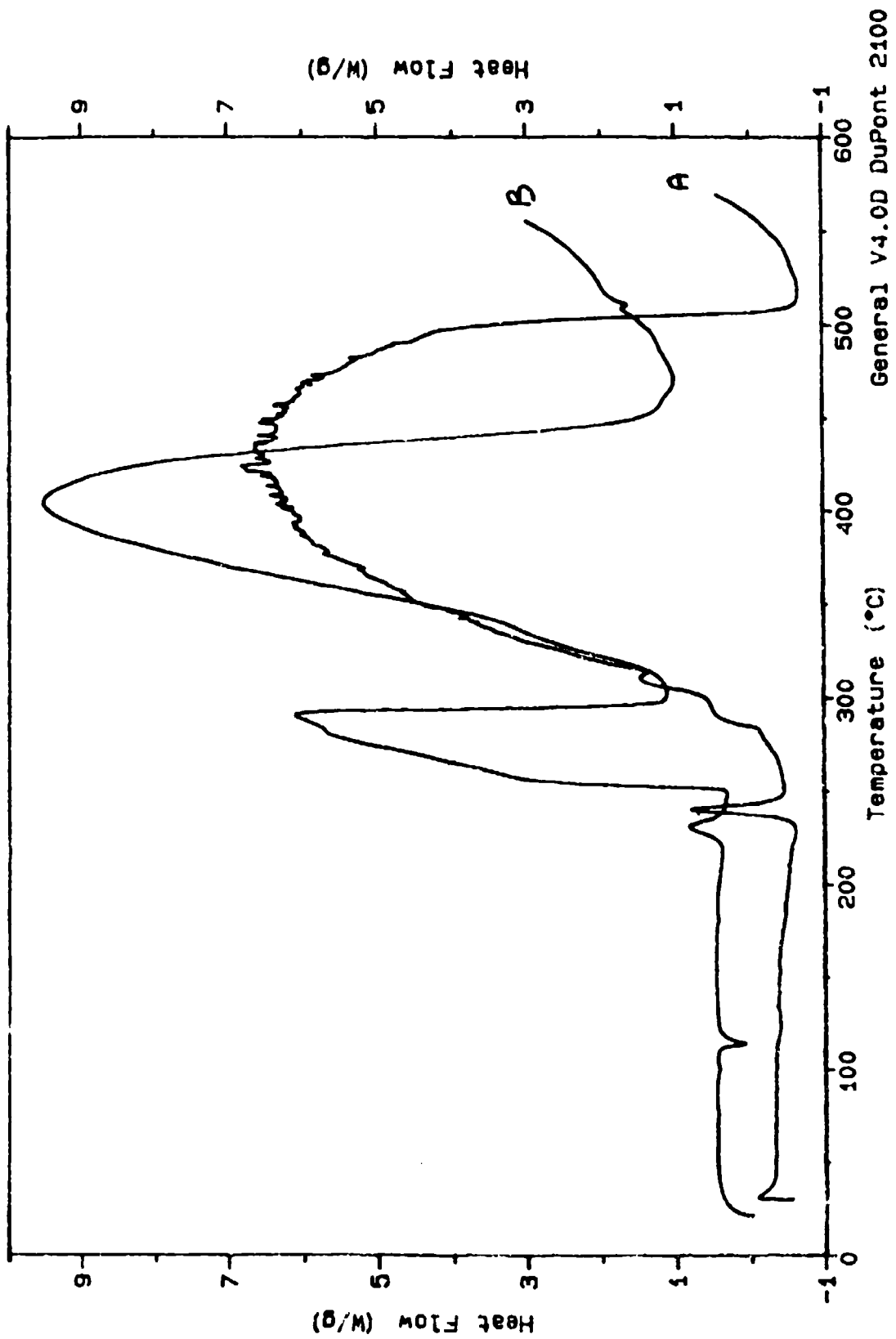


Figure 55. DSC of (A) Extracted vs (B) Unextracted Sb_2S_5 in Air.

previously examined. Shown in Figure 56, heating at 400 °C, produces bands at 384, 264, and 180 cm⁻¹ indicating formation of cubic Sb₂O₃; while at 600 °C, bands at 440, 365, 287, and 227 cm⁻¹ identify the presence of Sb₂O₄. Again, mid IR (Figure 57) reveals the presence of the characteristic 1100 cm⁻¹ band for the samples heated in air and its absence (Figure 58) in virgin and nitrogen heated samples. Finally, as with Sb₂S₄ and unextracted Sb₂S₅, heating in a inert environment produces Sb₂S₃. (See Figure 59.)

Solidification from a melt patterns the unextracted Sb₂S₅. On rapid cooling, DSC (Figure 60) shows the typical glass transition (225 °C) and crystallization exotherm (272 °C) indicating that amorphous Sb₂S₃ formed from the melt. Allowing the melt to cool slowly results in formation of crystalline Sb₂S₃ as evidenced by the lack of a glass transition and crystallization exotherm. (See Figure 61.)

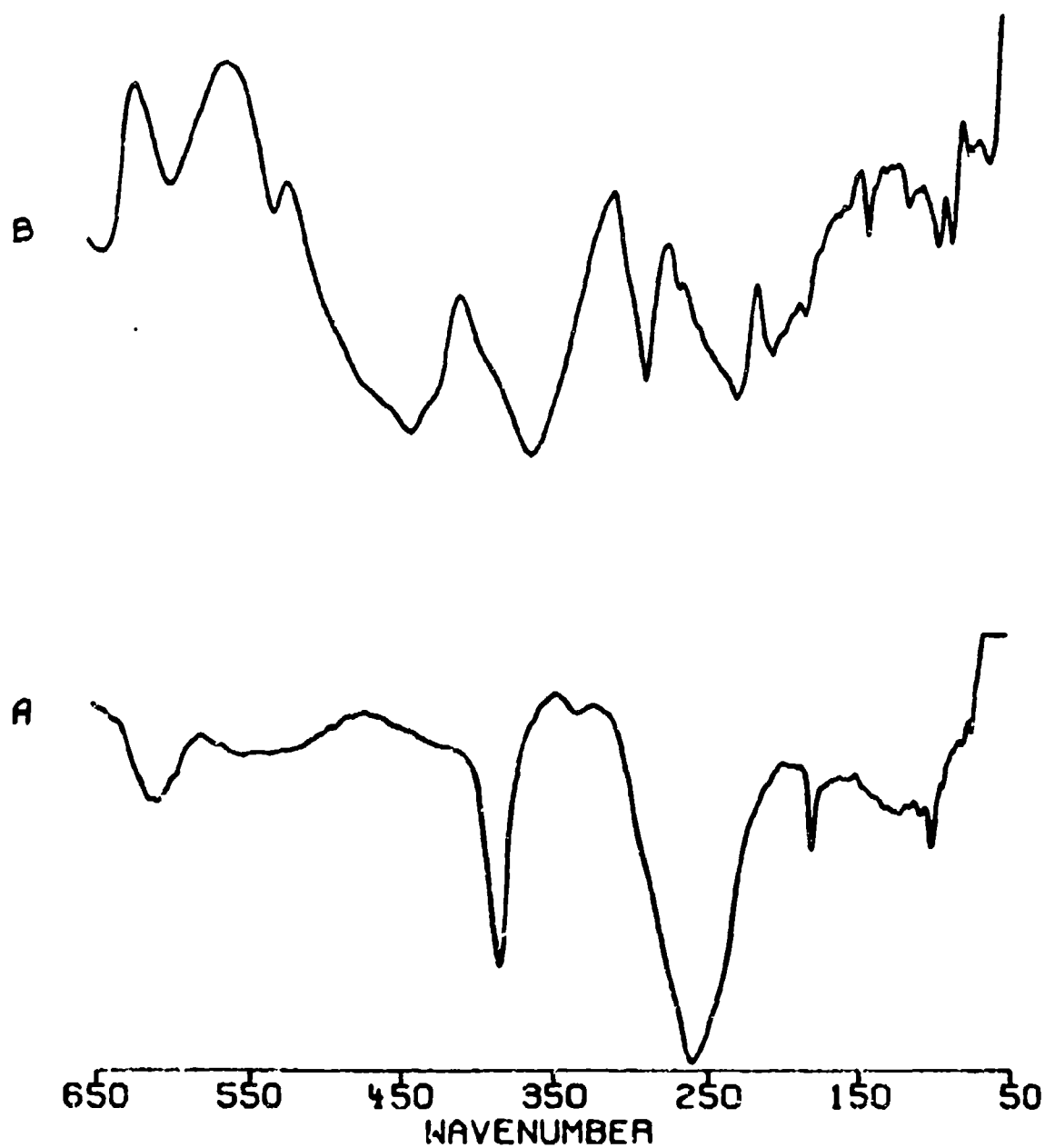


Figure 56. Far IR of Extracted Antimony Pentasulfide (A) 400°C in Air; (B) 65°C in Air.

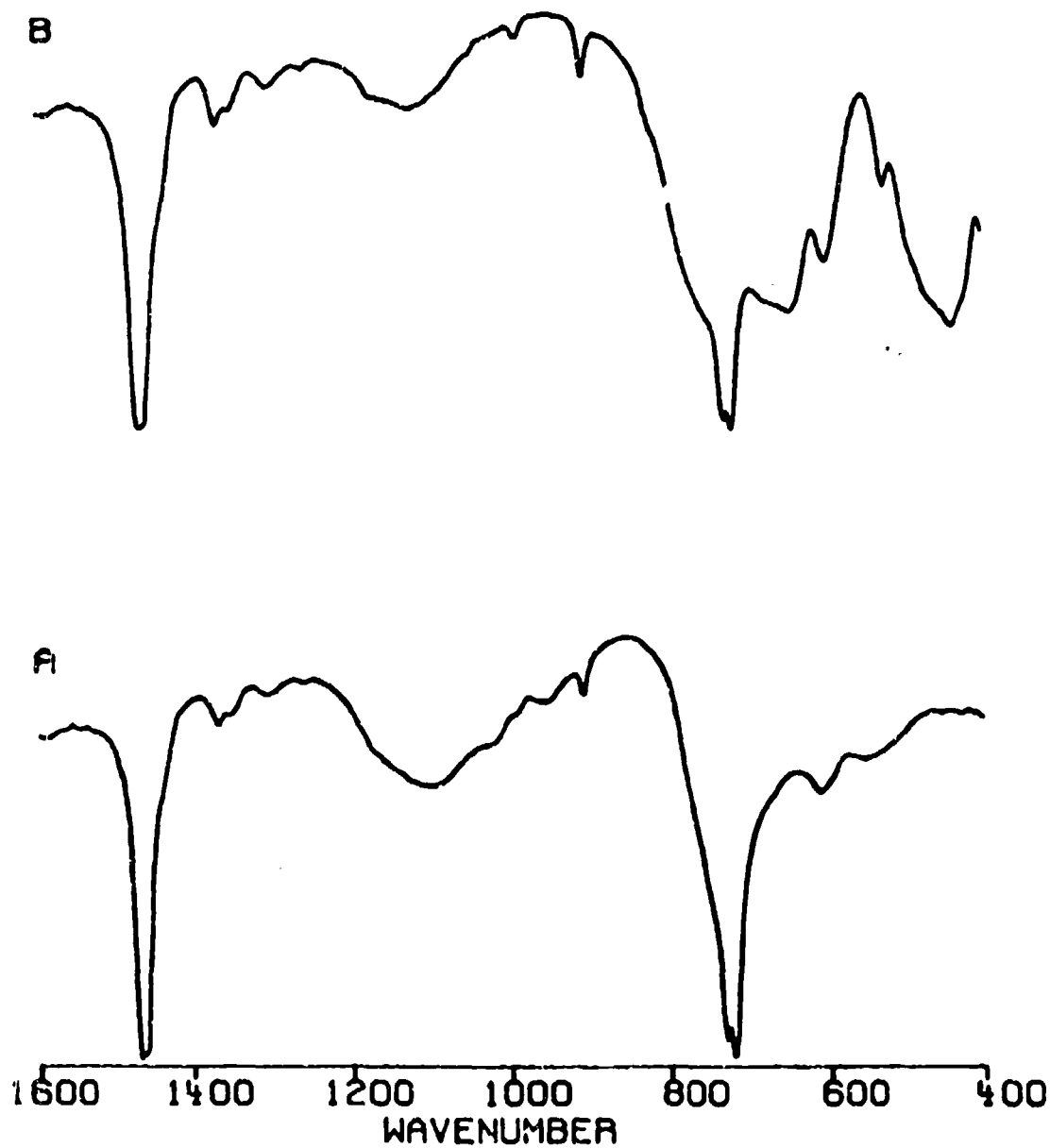


Figure 57. Mid IR of Extracted Sb_2S_5 (A) 400°C in Air; (B) 650°C in Air.

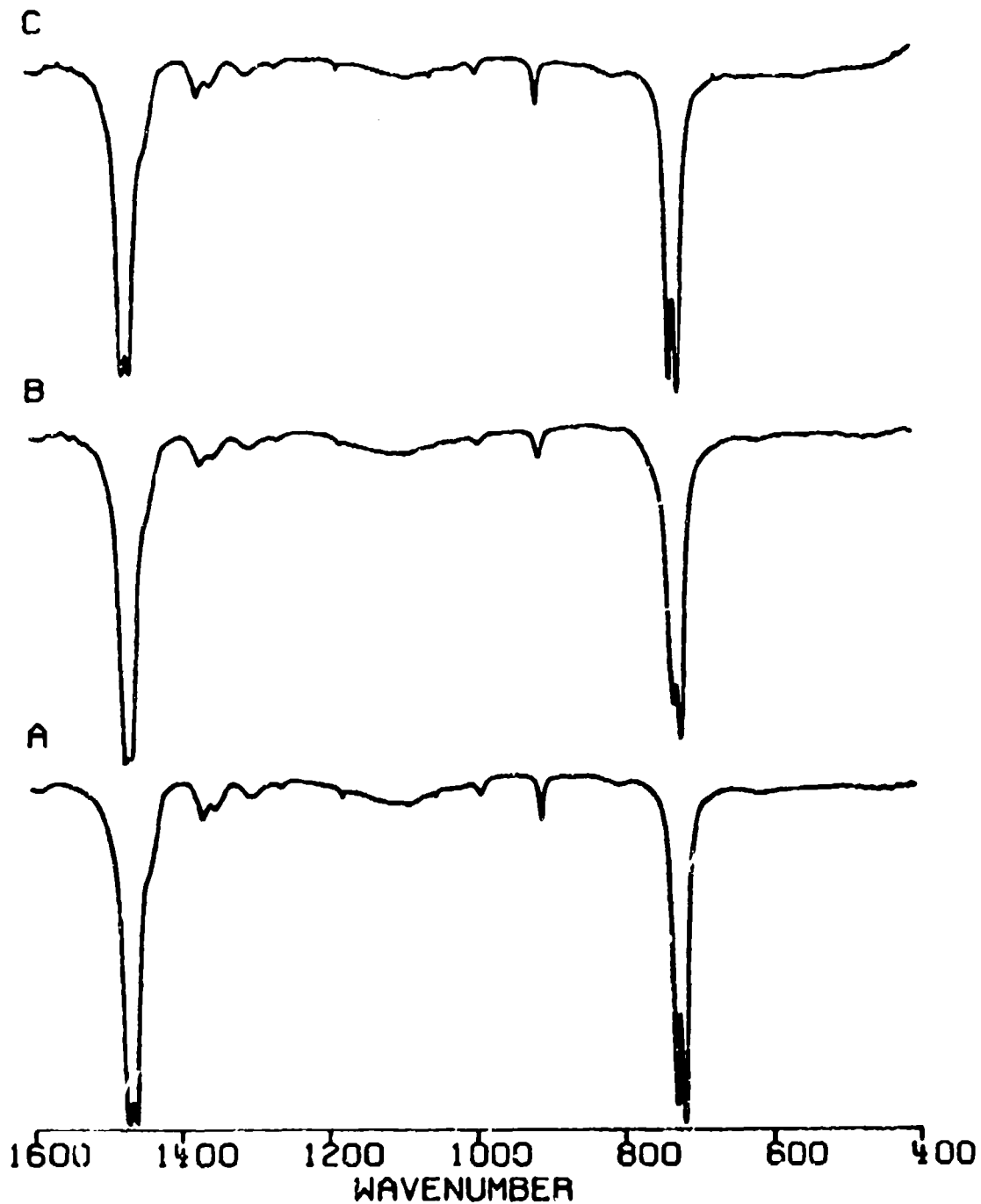


Figure 58. Mid IR of Extracted Sb_2S_5 (A) Virgin; (B) 350°C in N_2 ; (C) Sb_2S_3 .

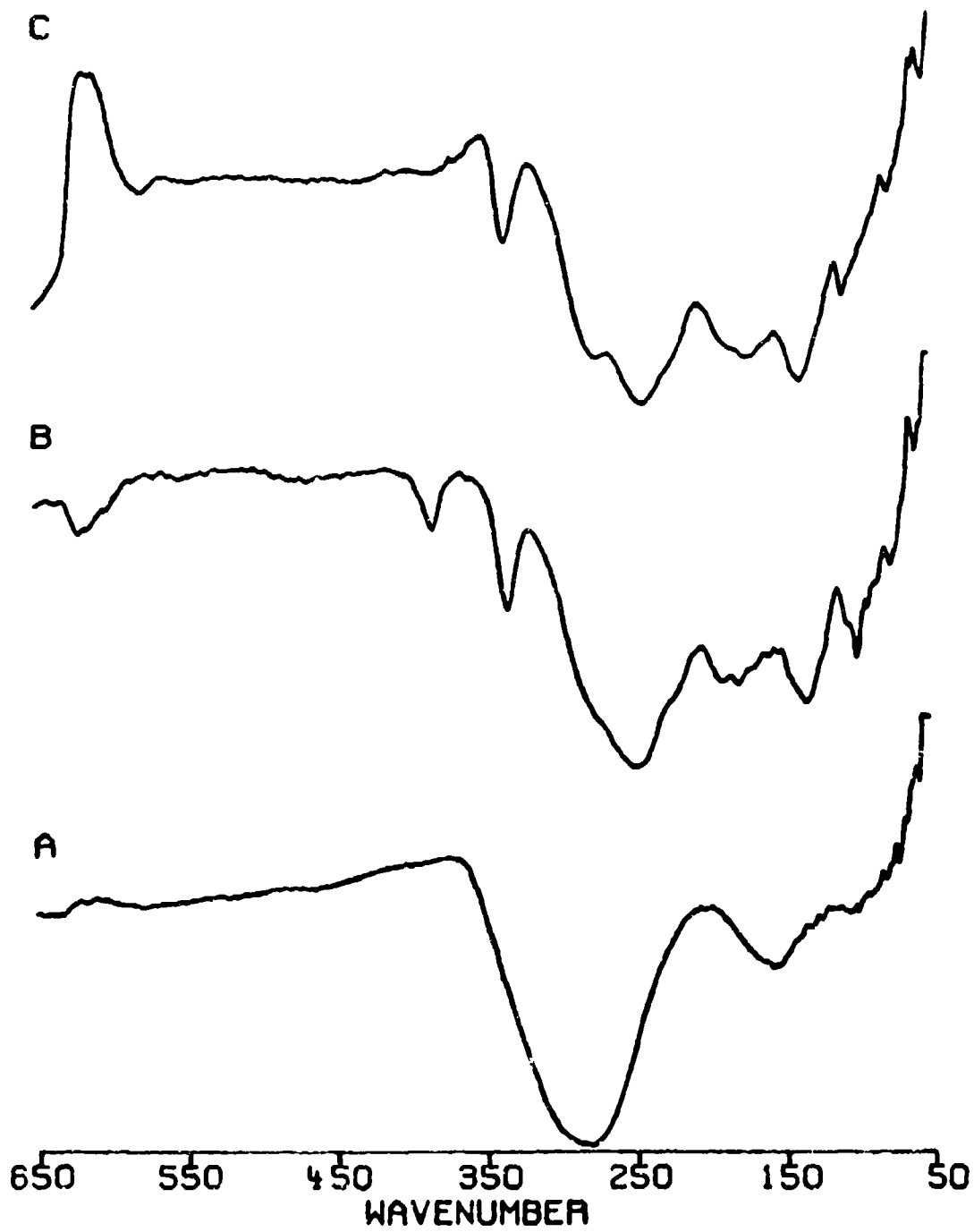


Figure 59. Far IR of Extracted Sb_2S_5 (A) Virgin; (B) 350°C in N_2 ;
 (C) Sb_2S_3 .

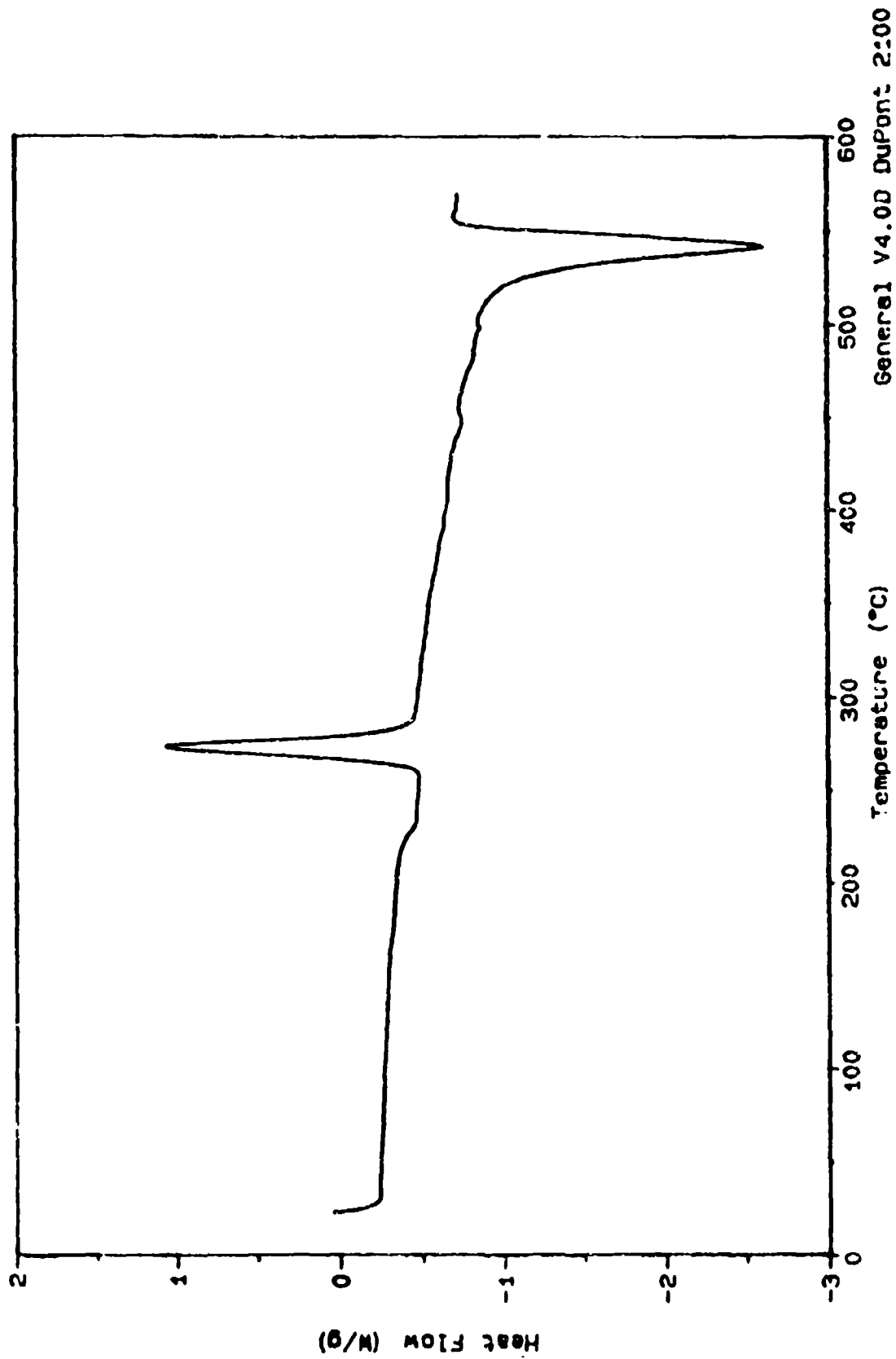


Figure 60. DSC of Generated Amorphous Sb_2S_3 (Extracted Sb_2S_5 rapidly cooled from melt in N_2).

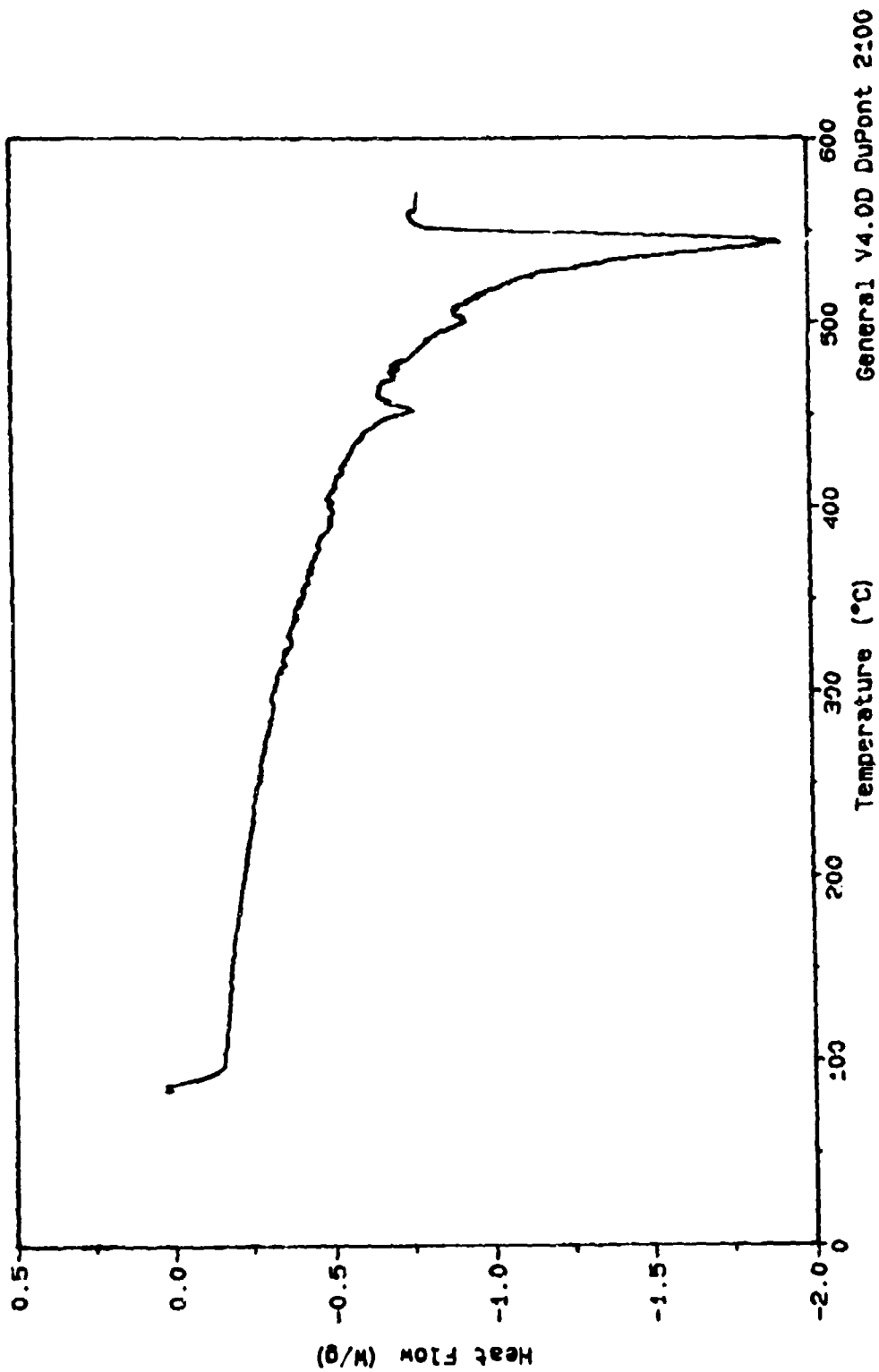


Figure 61. DSC of Generated Crystalline Sb_2S_3 (Extracted Sb_2S_3 slow cooled from melt in N_2).

SECTION IV CONCLUSION

1. Antimony trisulfide: Under nitrogen, antimony trisulfide is thermally stable to 575 °C. Sb_2S_3 oxidizes at 300 °C. The oxidation produces a sulfate containing species along with cubic Sb_2O_3 . At 604 °C, cubic Sb_2O_3 oxidizes to form Sb_2O_4 . The amorphous form of Sb_2S_3 can be created in situ in the DSC, by rapidly cooling a sample from a melt. The amorphous Sb_2S_3 undergoes a glass transition at 232 °C and a crystallization at 318 °C. Addition of elemental sulfur to Sb_2S_3 raises the observed melting point from 510 °C to 520 °C.

2. Antimony tetrasulfide: Sb_2S_4 releases sulfur between 195 and 230 °C to form Sb_2S_3 . The corresponding transition to Sb_2S_3 is seen in the DSC at 210 °C. The Sb_2S_3 generated in situ melts at 548 °C. It undergoes the same oxidative behavior as virgin Sb_2S_3 ; forming an intermediate sulfate and the cubic trioxide at temperatures below 600 °C, and the tetraoxide above 600 °C. Amorphous Sb_2S_3 generated from a rapid cool undergoes a glass transition at 227 °C and a crystallization at 279 and 293 °C. The addition of elemental sulfur lowers the temperature of the transition to Sb_2S_3 to 186 °C in the DSC.

3. Antimony pentasulfide: Sb_2S_5 contains approximately 15 percent excess sulfur - as determined by CS_2 extraction. The

excess sulfur is present within the antimony sulfide in two forms as demonstrated by the biphasic weight loss in the TGA between 193 and 239 °C. The associated change in the DSC occurs at 231 °C. The Sb_2S_3 generated in situ, melts at 553 °C. It's oxidation follows the same pattern as the tri and tetrasulfides. Amorphous Sb_2S_3 generated by a rapid cool of a melt, undergoes a glass transition at 228 °C and a crystallization at 271 °C. Extracted Sb_2S_5 has a sulfur to antimony ratio of 3:2, and undergoes a 2 percent weight loss between 227 and 243 °C. The corresponding change in the DSC occurs at 239 °C; however, this exotherm overlaps an endotherm which in effect raises the measured maximum. Removal of the excess sulfur by extraction does not effect the melting of generated Sb_2S_3 . Amorphous Sb_2S_3 created by a rapid cooling of a melt, has a glass transition at 225 °C and a crystallization at 272 °C.

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