

Glass Transition Behavior of Aqueous Solution of Sugar-Based Surfactants

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1. Introduction

Since the end of previous century, the role of petroleum as a raw material of synthetic surfactant gradually deflated due to the reasons such as decreasing of the relative abundance of petroleum, leading to soared prices of petroleum and increasing of carbon dioxide emission by heavy utilization of petroleum. Instead, the industries concerning in the surfactants and detergents are focusing on the utilization of biobased feedstocks, intermediates and products. Under these circumstances, the biobased surfactants derived from carbohydrate or sugar are highlighted.

Sugar-based surfactants commonly used for household products are frequently applied in foods, cosmetics and pharmaceutical industrial region (Rybinski, von W. & Hill . K. (1998). Hill, K. & Rhode O. (1999). Drummond, C. J.; Fong, C.; Krodkiewska, I.; Boyd, B. J. & Baker, I. J. A. (2003). Hill, K. & LeHen-Ferrenbach, C. (2007).). They are less toxic, highly biodegradable, and able to be readily formulated with other components. And it is well known that their representative nature is that they have ability to aggregate in an aqueous solution as well as conventional surfactants (Warr, G. G.; Drummond, C. J.; Grieser, F.; Ninham, B. W. & Evans, D. F. (1986). Auvray, X.; Petipas, C. & Anthore, R. (1995). Söderberg, I.; Drummond, C. J.; Furlong, D. N.; Godkin, S. & Matthews, B. (1995). Hoffmann, B. & Platz, G. (2001). Kocherbitov, V. & Söderman, O. (2003). Imura, T.; Hikosaka, Y.; Worakitkanchanakul, W.; Sakai, H.; Abe, M.; Konishi, M.; Minamikawa, H. & Kitamoto, D. (2007). Hato, M.; Minamikawa, H. & Kato T. (2007).). The morphology of the aggregate extends over ranges from the isotropic micelle solution to the liquid crystal such as hexagonal, cubic, lamella and sponge phases.

Numerous phase diagrams of the amphiphiles, which describe the aggregative behavior of the compound, are exhibited in terms of concentration and temperature. We are able to see those of the anionic, cationic and nonionic surfactant, but the diagram under 0 °C especially in the frozen state was not reported so much. Among such studies, cationic surfactant, octyl trimethylammonium bromide is reported to be able to lower the freezing point of ice effectively due to the presence of their ionic head group (Fukada, K.; Matsuzaka, Y.; Fujii, M.; Kato, T. & Seimiya, T. (1998).). Similarly, nonionic surfactant such as polyoxyethylene

glycol decyl ($C_{10}E_m$; $m = 4-8$) and dodecyl ($C_{12}E_m$; $m = 5, 6$ and 8) ether were reported to crystallize ice below -11 °C or -4.5 °C, respectively (Andersson, B. & Olofsson, G. (1987). Nibu, Y.; Suemori, T. & Inoue T. (1997). Nibu, Y. & Inoue, T. (1998a, 1998b). Zheng, L. Q.; Suzuki, M. & Inoue, T. (2002). Zheng, L.; Suzuki, M.; Inoue, T. & Lindman, B. (2002).). Contrary to this, the phase diagram of sugar-based surfactant seems to be uncompleted particularly under supercooled conditions under 0 °C.

Some nonionic surfactants were not used as a curative agent but a plasticizer because they showed the glass transition temperature (T_g) at low temperature region (Jensen, R. E.; O'Brien, E.; Wang, J.; Bryant, J.; Ward, T. C.; James, L. T. & Lewis, D. A. (1998); Amim, J.; Kawano, Y. & Petri, D. F. S. (2009).). Tween 40, poly(oxyethylene) sorbitan monopalmitate which have 20 EO units in the molecule was reported to possess T_g at -61 °C. Triton X-100 showed its T_g at -59 °C. Ethylene oxide surfactant such as hexahydrofarnesyl ethylene oxide surfactants (EO = 1 - 8) exhibited their T_g at low temperatures below -80 °C (Fong, C.; Weerawardena, A.; Sagnella, S. M.; Mulet, X.; Krodkiewska, I.; Chong, J. & Drummond, C. J. (2011).). In addition to this, there is a report that says sugar derivatives containing a hydrophobic group are applicable as a plasticizer. Gill stated that when such a sugar derivative was added to the corresponding free sugar, T_g of the mixture tended to lower than those of the free sugar system (Gill, I. & Valivety, R. (2000a, 2000b).). Here, the sugar given hydrophobicity worked as a plasticizer for a free sugar. On the other hand, when the other component which possessed much lower T_g than that of the sugar derivative was mixed in the system, the existing sugar derivative did not necessarily work as a plasticizer.

Although it had been scarcely studied about the glass-forming property of sugar-based surfactants, but nowadays, much attention is being denoted to their interesting characteristics. It has been reported that *n*-alkyl glycosides such as α -D-glucosides, β -D-maltosides, β -D-maltotrioxide and sucrose fatty acid esters formed a glass state under anhydrous conditions (Hoffmann, B.; Milius, W.; Voss, G.; Wunschel, M.; van Smaalen, S.; Diele, S. & Platz G. (2000). Kocherbitov, V. & Söderman, O. (2004). Ericsson, C. A.; Ericsson, L. C.; Kocherbitov, V.; Söderman, O. & Ulvenlund, S. (2005). Ericsson, C. A.; Ericsson, L. C. & Ulvenlund, S. (2005). Szűts, A.; Pallagi, E.; Regdon, G. Jr; Aigner, Z.; Szabó-Révész, P. (2007).). Their T_g increased from -12.4 °C of *n*-heptyl α -D-glucopyranoside to 100 °C of *n*-dodecyl β -D-maltotrioxide in proportional to the number of saccharide unit. Thus, T_g of the sugar based surfactants are much higher than that of the other nonionic surfactants as mentioned above. That is, sugar-based surfactants possess a remarkable glass forming ability comparing to another type of surfactant. Because T_g of anhydrous sugar-based surfactant existed almost above the freezing point of water, 0 °C, therefore, we expected that the behavior and ability of making glass state of the aqueous sugar-based surfactant solution can be readily observed without ice freezing if the cooling was conducted rapidly.

Recently, authors studied the vitrification or glassification of the aqueous solution of sugar-based surfactant, which must be associated with the specific function under freezing state (Ogawa, S. & Osanai, S. (2007). Ogawa, S.; Asakura, K. & Osanai, S. (2010).). In this chapter, we would like to elucidate some aspects of the aqueous solution of sugar-based surfactant under supercooling, where the simple primary phase transition such as gelation and

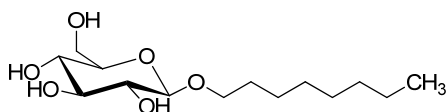
crystallization is not a key topic, but the vitrification plays an important role. We would like to indicate the some characteristics of the sugar-based surfactant in an aqueous solution under low temperature. The basic behavior of these surfactants solution will be focused on the following items.

- The glass transition of the aqueous solution of sugar-based surfactant under low temperature with forming thermotropic and lyotropic liquid crystalline phases.
- Correlation between the glass transition and the protective effect against freezing.

The information obtained from this chapter would be valuable to the researchers who engage the low temperature technologies.

2. Glass transition behavior of octyl β -D-glucoside/water binary mixtures

Octyl β -D-glucoside (G8Glu: Scheme 1) is one of the representative sugar-based surfactants. Although there are many reports on phase behavior of C8Glu/water binary system (Boyd, B. J.; Drummond, C. J.; Krodkiewska, I. & Grieser, F. (2000). Nilsson, F.; Söderman, O. & Johansson, I. (1996). Häntzschel, D.; Schulte, J.; Enders, S. & Quitzsch, K. (1999). Dörfler, H.-D. & Göpfert, A. (1999). Bonicelli, M. G.; Ceccaroni, G. F. & La Mesa, C. (1998). Sakya, P., Seddon, J. M. & Templer, R. H. (1994). Loewenstein, A. & Igner, D. (1991). Kocherbitov, V.; Söderman O. & Wadsö, L. (2002).), no report was presented on its vitrification behavior under the low temperature. In this section, we introduce the glass transition behavior of octyl β -D-glucoside/water binary mixture within a wide concentration range under the conditions without ice formation (Ogawa, S.; Asakura, K. & Osanai, S. (2010).).



Scheme 1. Chemical structure of octyl β -D-glucoside (C8Glu).

C8Glu was synthesized as described in the literature, with a little modification (Bryan, M. C.; Plettenburg, O.; Sears, P.; Rabuka, D.; Wacowich-Sgarbi, S. & Wong, C.-H. (2002).).

2.1 Thermal behavior of C8Glu/water binary system

Fig. 1 shows a typical DSC chart which illustrates the glass transition behavior of C8Glu/water mixture. Each sample with various concentrations was homogenized by heating until 120 °C prior to the measurement. The sample was rapidly cooled to -120 °C at -10 °C/min and then heated at the rate of 10 °C/min. As Fig. 1 shows, when the concentration of C8Glu was greater than ca. 80 wt%, no ice was produced during cooling, and the glass transition was observed during the heating process. Occurrence of the glass transition was confirmed by the discontinuity of the heat capacity as indicated by solid line arrows in Fig. 1. In the concentration range from ca. 80 to 82 wt% for C8Glu, the ice was formed by the devitrification and thawed in the heating process (Fig. 1(a)). Devitrification was defined as the solidification phenomena after the temperature exceeded T_g in the heating process.

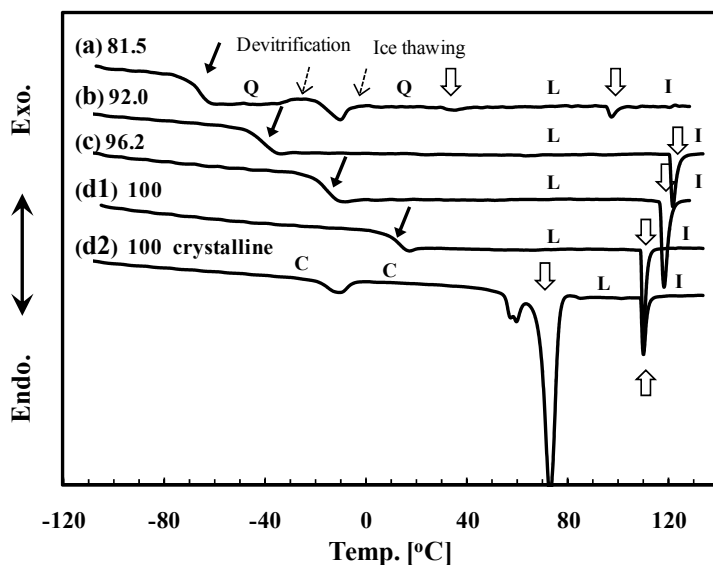


Fig. 1. Typical DSC thermograms of heating process at 10 °C/min of C8Glu/water mixture.

The concentration of the sample is expressed in wt %. White arrows indicate the phase transition between the liquid crystalline phases or from the liquid crystalline phases to an isotropic solution or melt. Solid line arrows indicate the glass transition, as mentioned above. I: isotropic solution, H: hexagonal phase, Q: cubic phase, L: lamellar phase, C: crystalline phase. *Apparatus*; DSC 60 (SHIMAZU Co. Ltd.) equipped with a cooling accessory was used throughout the measurement. *Sample preparation*; Samples of an aqueous solution were prepared from C8Glu and the prescribed amount of water. The sample was prepared as following two methods. *Method A*; A dilute aqueous C8Glu solution (ca. 35 wt%) in an aluminium pan was directly concentrated by drying over phosphorous pentoxide at an ambient temperature. *Method B*; A prescribed amount of water was absorbed under a humid atmosphere or added directly to C8Glu that was free from water. The anhydrous C8Glu was prepared by placing the sample on a hot stage at 125 °C for 50 min under a N₂ atmosphere to remove any water.

G8Glu/water binary system gave various kinds of liquid crystalline (LC) phase, such as hexagonal (H), cubic (Q), and lamella (L) phases and crystalline phase (C) according to its concentration and temperature. Fig. 1(b) indicates that L phase existed after the glass transition took place at around -40 °C during a heating process and it changed into isotropic liquid at 120 °C. In other words, the glass transition did not occur in crystalline and isotropic liquid phases but in a LC phase during the cooling process. A detailed comparison of (d1) with (d2) in Fig.1 clearly demonstrated that the phase transition from glass to lamella occurred in the LC phase, because the peak due to the transformation from crystalline to lamella LC at 70 °C was not observed in (d1) chart.

Observation by the polarizing optical microscopy (POM) gave a consistent result with these findings mentioned above. The sample was rapidly cooled to -100 °C at -10 °C/min and

then heated at the rate of 10 °C/min. Fig. 2 shows the POM images of the C8Glu sample solution. Its conditions are shown in a legend of the figure. As can be seen from Fig. 2a and 2b, the L texture exhibited oily streaks at both temperatures above and below T_g . It indicated that L phase texture was maintained above or below T_g . It presented unambiguous evidence that this sample changed from liquid crystal to a glassy phase with holding its lamella structure, that is, the “glassy liquid crystals”. The basic concept of the glassy liquid crystal was introduced in references. (Yoshioka, H.; Sorai, M. & Suga, H. (1983). Kocherbitov, V. & Söderman, O. (2004).

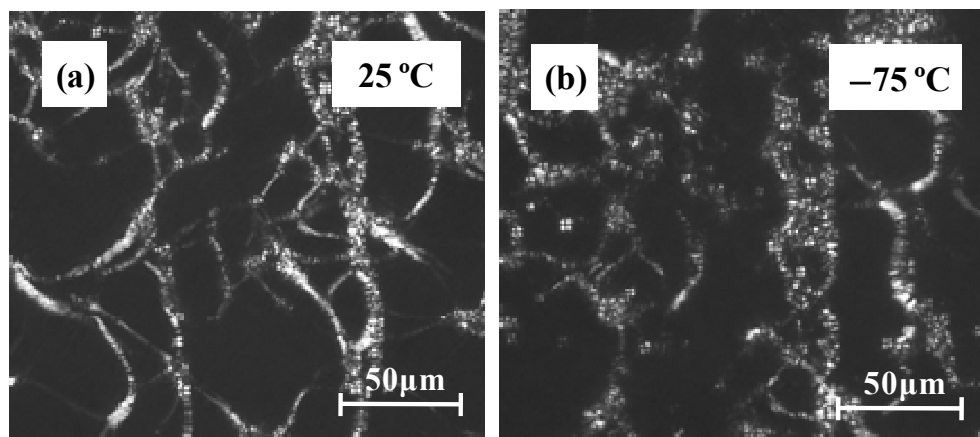


Fig. 2. POM photographs of 92.5 wt% of C8Glu/water mixture above and below T_g of -47 °C.

Apparatus; Polarizing microscopy (BH-51, Olympus) equipped with a heating/cooling stage was used for observation. *Sample preparation;* Sample solutions were prepared in a similar manner to the DSC measurement. Sample was observed through a thin specimen sandwiched between a slide plate and cover-glass plate.

2.2 Concentration-temperature phase diagram with T_g curve

Fig. 3 is a phase diagram of the C8Glu/water binary system at concentrations of more than 50 wt%. The diagram was constructed on the basis of experimental results obtained from DSC thermograms and POM photographs. Interested readers are able to refer detail methods for the determination of LC phases from numerous references as mentioned at the beginning of this section.

In this diagram, T_g curve, the ice nucleation temperature curve (INC), and devitrification temperature curve (DC) are depicted. Although INC and DC curves are variable parameter according to the rate of nucleation, they are useful to understand the dynamic behavior of the system. As shown in Fig. 3, as the concentration of C8Glu increases, the lyotropic aggregates change from an isotropic solution (Micelle solution: M) to the liquid crystal phase, such as, H, Q, and L phases at 0 °C. When the concentration of C8Glu was lower than 80%, the INC was clearly recognized. It meant that the crystallization preferentially occurred below this temperature before the vitrification took place.

A lot of reference showed phase diagrams of C8Glu/water system that expressed the existence of crystal phase or gel phase in the concentrated region over 90 wt%. But in our study, the crystal phase did not appear in the same concentration at cooling rate $-10\text{ }^{\circ}\text{C}/\text{min}$, between $-120\text{ }^{\circ}\text{C}$ and $120\text{ }^{\circ}\text{C}$. Instead of that, the glass transition was observed over ca. 80 wt% concentration. That was referred to T_g curve in Fig. 3. The glass transition temperature, T_g , shifted to higher, as the concentration of C8Glu moved to higher. It means that C8Glu did not work as a plasticizer but as a curative agent in an aqueous solution.

Comparing the phase diagram above and below T_g curve in Fig. 3, we are able to understand that the glassy phase was formed by cooling both of Q and L phases. It could therefore be presumed that, the formation of Q and L types of glassy LC phase occurred below T_g curve. Even if temperature crossed the phase boundary between the Q and L phases, there was no discontinuity in T_g line. It suggested that the difference among the liquid crystalline structures was not a decisive factor for the determination of T_g in the aqueous solution.

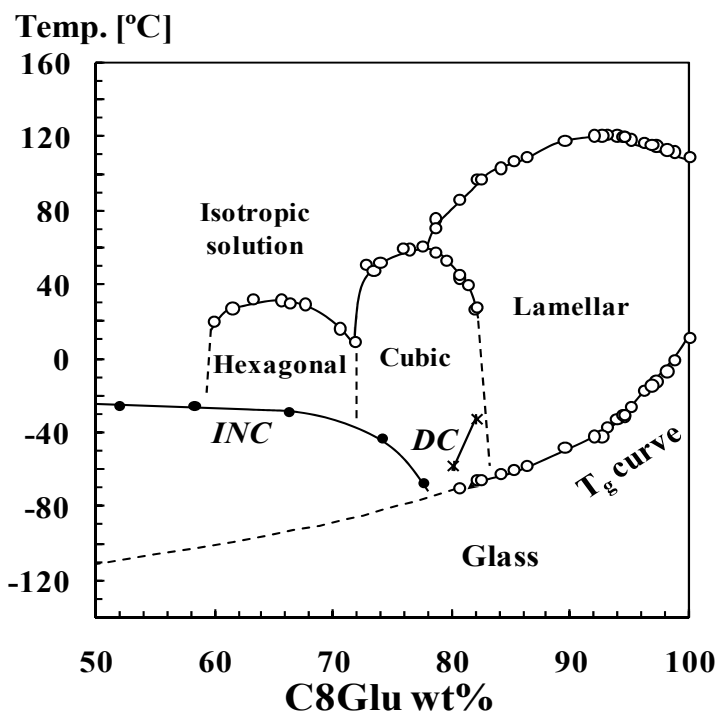


Fig. 3. Phase diagram of C8Glu/water mixture from 50 to 100 wt % C8Glu concentration including T_g curve, ice nucleation curve (INC) and devitrification curve (DC).

The dotted lines are predicted one. The phase transition temperatures were determined by the intersection of the baseline and tangent to the end of the endothermic peak of the DSC chart on heating. T_g was determined as the temperatures corresponding to half of the magnitude of the heat capacity change (ΔC_p) at T_g (Blond, G.; Simatos, D.; Catté, M.; Dussap,

C. G.; Gros, J. B. (1997).). The ice crystallization temperatures were defined as the average of five measurements.

2.3 Comparison of glass transition behavior with predicted curve

Couchman and Karasz presented a model that predicts T_g of a mixture employing classical thermodynamics (Couchman, P. R. & Karasz, F. E. (1978). Couchman, P. R. (1978).). This model treated the glass transition as if it was equivalent to an Ehrenfest second order transition. The "original" Couchman-Karasz (C-K) formula is given below:

$$\ln T_g = \frac{x_1 \ln T_{g1} + k x_2 \ln T_{g2}}{x_1 + k x_2} \quad (1)$$

where the subscripts 1 and 2 denote components 1 (C8Glu) and 2 (pure water), respectively. The symbols x_1 and x_2 represent the mole fractions of the corresponding C8Glu and pure water, respectively. T_g is the glass transition temperature of the mixture under consideration; k is a constant defined as $\Delta C_{p2}/\Delta C_{p1}$. ΔC_{p1} and ΔC_{p2} are the ΔC_p at T_g of component 1, pure solute and component 2, pure water, respectively. Eq. 1 is often "modified" to the following general form:

$$T_g = \frac{x_1 T_{g1} + k x_2 T_{g2}}{x_1 + k x_2} \quad (2)$$

The suitability of these two equations ("original" and "modified" C-K equations) was discussed by comparison with the actual experimental measurements for T_g of the C8Glu/water mixture.

Fig. 4 shows the T_g -prediction curves obtained from the "original" and "modified" C-K equations, using $\Delta C_{p1} = 142.2$ J/mol K at $T_{g1} = 284.4$ K (11.2 °C) for the amorphous C8Glu as obtained from our experimental results. At the same time, the experimental values were indicated in the corresponding figures. Values of $\Delta C_{p2} = 35.0$ J/mol K and $T_{g2} = 135$ K (-138.2 °C) for the pure water were taken from the literatures (Sugisaki, M.; Suga, H. & Seki, S. (1968). Rasmussen, D. H. & MacKenzie, A. P. (1971).). The experiment indicated that the analysis using the "original" (Eq. 1) gave a relatively good agreement with the experimental result over the entire concentration studied. By contrast, the predicted T_g obtained from the "modified" (Eq. 2) (dotted line) was not in accord with the experimental finding. By considering these results, we can state the "original" C-K equation would give a much better prediction than the "modified" one for T_g determination in mixtures. Couchman stated that if the assumption that $T_{g1}/T_{g2} \approx 1$ was applicable, the results obtained using the "modified" C-K would be valid (Couchman, P. R. & Karasz, F. E. (1978).). In our system, T_{g1} of anhydrous C8Glu is 284.4 K, and giving a T_{g1}/T_{g2} for C8Glu of 2.11. This value is far from the unity that is appropriate for the "modified" C-K equation. We guessed that the T_g ratio between two components which form the mixture must be approximately one if we want to employ the 'modified' C-K equation to predict T_g .

$$\Delta C_p [J / g K] = \Delta F / (W_{sample} \times scan\ rate) \quad (3)$$

$$\Delta C_p [J / mol K] = \Delta C_p [J / g K] (MW_1 \times x_1 + MW_2 \times x_2) \quad (4)$$

where W_{sample} is the sample weight in a sealed pan and ΔF is the heat flow change (Fig. 4b). MW_1 and MW_2 are the molar weight of C8Glu [MW_1 : 292.19], and pure water [MW_2 : 18.02].

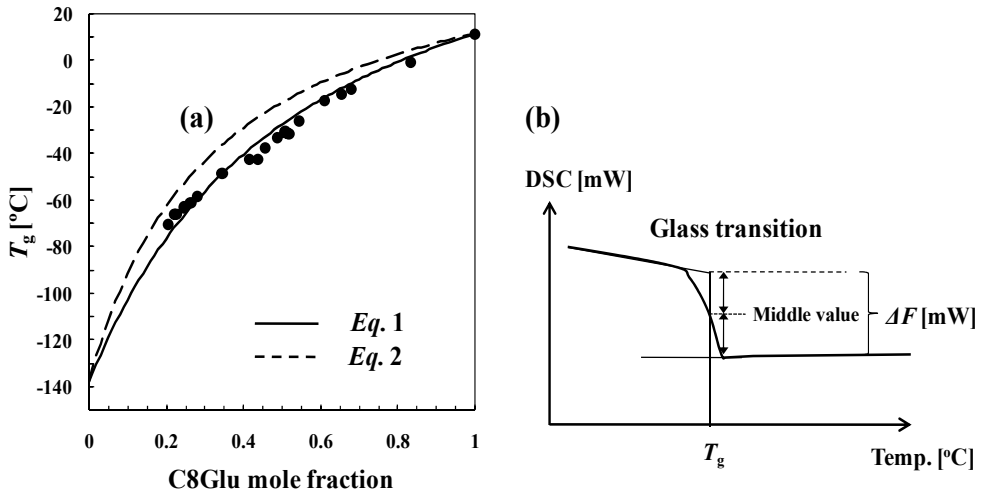


Fig. 4. (a) Comparison of experimental T_g of C8Glu/water mixture obtained by two kinds of predicted curves by C-K equations and (b) estimation of T_g and ΔF .

Each plot shows the experimental data. ΔF was determined as sketched in Fig.4(b). Using this parameter, ΔC_p was obtained as follows.

2.4 Influence of concentration on ΔC_p , ΔH and phase transition temperature T_g

Fig. 5 shows the ΔC_p curve as a function of the mole fraction of C8Glu. As mole fraction is decreasing from 1.0, there were two bending points at 0.65 and 0.40. The curve is divided into three regions (A, B, and C) according to their characteristics.

In the mixture of C8Glu/water, ΔC_p for pure water (ΔC_{p2}) obtained by extrapolation of the plots in region C to zero C8Glu content, was 35.0 J/mol K. This value is compatible with the experimental value reported by Sugisaki et al. (Sugisaki, M.; Suga, H. & Seki, S. (1968).). On the other hand, ΔC_p for pure C8Glu (ΔC_{p1}) obtained by extrapolation of the plots in region C to 1.0 C8Glu content, was about 175 J/mol K. There was an apparent difference between the extrapolated and actual experimental value, 142.2 J/mol K. By contrast, the extrapolation of plots in region A reached 0 J/mol K. These results showed that ΔC_p of the binary mixture was not predictable using a simple linear function composed of ΔC_{p1} and ΔC_{p2} .

In order to obtain information for clarification of this complex behavior of ΔC_p , the transition between lamella and isotropic liquid phases were studied further in detail.

Fig. 6 indicates the relationships of the phase transition between lamella (L) and isotropic solution (I) with C8Glu mole fraction. Enthalpy (ΔH) and temperature of the phase transition are depicted on the two vertical axes. Generally speaking, enthalpy of C8Glu solution decreased as the mole fraction of C8Glu reduced. A clear bending point was recognized at a particular

concentration with mole fraction of 0.65, which was in fair agreement with that of the first bending point shown in Fig. 5. The enthalpy of the system reached 0 J/mol by extrapolation of plots in region A in analogy with the result shown in Fig. 5. These results meant that the amount of water in region A would have no influence not only on ΔC_p but also on ΔH .

Here, we adopted a concept of the "non-continuous water" to propose a hypothesis that interprets above behavior of ΔC_p in C8Glu/water mixture.

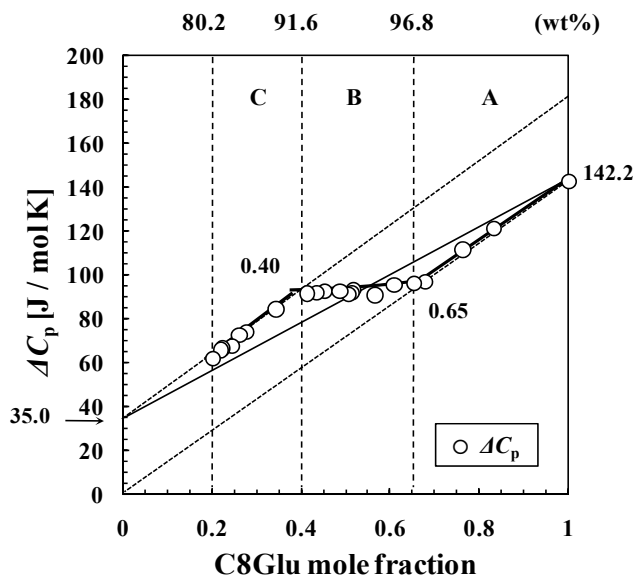


Fig. 5. ΔC_p behavior of C8Glu/water mixture with variation in concentrations.

The linear solid line connecting two ΔC_p values, 35.0 and 142.2 shows ΔC_p line when the mixing was carried out under holding ideal state.

We interpreted the behavior of C8Glu in different concentration of aqueous solution as follows: Fig. 7 shows the relationships among ΔC_p at T_g , ΔH of the phase transition and the phase transition temperature in C8Glu/water mixture systems on the basis of the experimental results. In region C, the molar ratio of water: C8Glu was 1.5 : 1 - 4 : 1, that is, molecule's number of water is larger than that of the C8Glu. The water in this region will constitute the aqueous phase keeping continuous state among a bimolecular membrane lamellar structure. It is a kind of bulk water. Reduction of water means simple decrease of the bulk water stated above. The fact that extrapolation of plots of ΔC_p in region C reached 35 J/mol K and coincided with that of pure water proved its validity.

On the other hand, in region A, the molar ratio of water : C8Glu is 1 : 2. In other words, the number of water molecule is less compared with that of C8Glu. The scarcity of water will be further signalized by consideration of their relative magnitude of the molecular bulkiness. This circumstance will not enable the water molecule to exist in continuous state. The water molecules in region A would be present in a non-continuous state with creating a new hydrogen bond among the glucoside molecules.

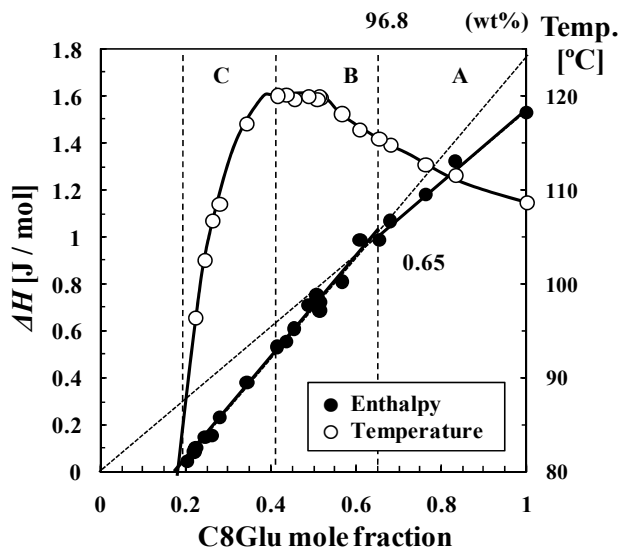


Fig. 6. Relationships of phase transition behavior between lamellar phase and isotropic solution with C8Glu fraction.

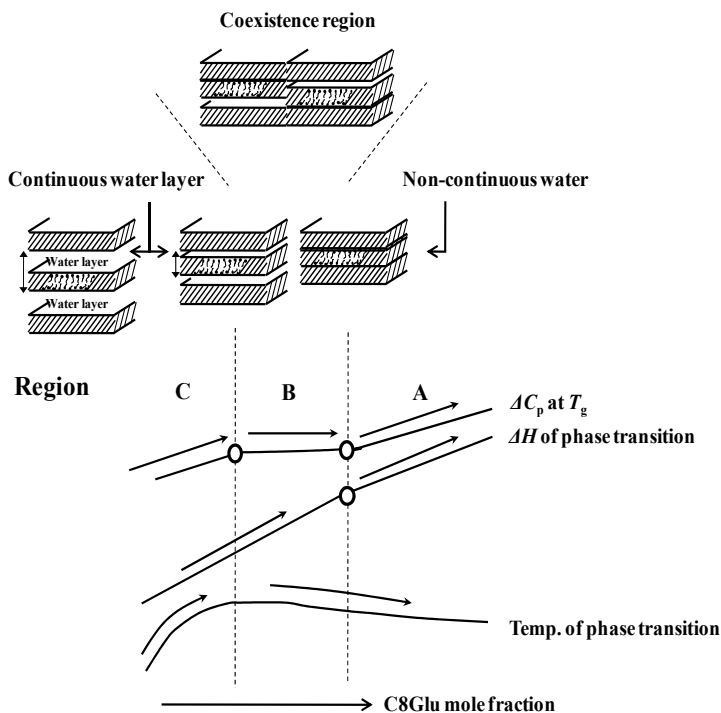


Fig. 7. Schematic figure of C8Glu/water mixture systems with variation of C8Glu concentration.

Region A: The water molecules would be present in a non-continuous state.

Region B: The water characteristic is determined by the mixed system composed of the continuous and non-continuous water existed in the region C and A.

Region C: The water behaves like bulk water and constitutes the aqueous phase keeping continuous state among a bimolecular membrane lamellar structure. Reduction of water means simple decrease of the bulk water.

In region B, ΔC_p kept in a constant state irrespective of its concentration of the system. This result demonstrated the additivity of two kinds of states at the corresponding concentration of the bending points in Fig. 5. It means the behavior of water in this region would be determined by the mixed system composed of the continuous and non-continuous water existed in the region C and A, respectively.

3. Glass transition behavior of octyl β -D-glucoside/NaCl/water ternary mixtures

The aqua-system of the life organism contains a various kinds of ions and exhibits complicated buffer actions to maintain its physiological functions in a normal state. As it is cooled down, eutectic phase composed of electrolyte and ice was generated in the concentrated unfrozen phase (Mullin, J. W. (2001)). Occurrence of the eutectic would be responsible of direct causes for damages against cells and enzymes and resulted unusual pH change would become a trigger for abnormal interactions (Heber, U.; Tyankova, L. & Santariu, K. A. (1971). Mollenhauer, A.; Schmitt, J. M.; Coughlan, S. & Heber U. (1983). Han, B. & Bischof, J. C. (2004). Wang, C.-L., Teo, K. Y. & Han, B. (2008). Goel, R.; Anderson, K.; Slaton, J.; Schmidlin, F.; Vercellotti, G.; Belcher, J. & Bischof, J. C. (2009)). In actual circumstances where the life organisms are treated under extremely and mildly cool atmosphere, various kinds of cryoprotectants and lyoprotectants such as salts, amino acids, carbohydrates, artificial and natural polymers are used to stabilize these bio-tissues from the cooling damages (Heber, U.; Tyankova, L. & Santariu, K. A. (1971). Tyankova, L. (1972). Izutsu, K.; Yoshioka, S. & Kojima, S. (1995). Koshimoto, C. & Mazur, P. (2002). Chen, N. J.; Morikawa, J. & Hashimoto, T. (2005). Chen, Y.-H. & Cui, Z. (2006). Kawai, K. & Suzuki, T. (2007). Izutsu, K.; Kadoya, S.; Yomota, C.; Kawanishi, T.; Yonemochi, E. & Terada, K. (2009)).

The purpose of this section is to clarify the inhibition effect of sugar-based amphiphiles on eutectic formation in the freeze-thawing process of aqueous NaCl solution (Ogawa, S. & Osanai, S. (2007)).

3.1 Thermal behavior of sugar-based amphiphiles/NaCl/water ternary system

Fig. 8 shows DSC charts of the ternary system consist of C8Glu/NaCl/water. A solution containing C8Glu at the same concentration (C8Glu to water = 1:9 [wt%]) was mixed with various concentration of NaCl as shown in Fig. 8. Each sample was cooled to $-100\text{ }^\circ\text{C}$ at $-10\text{ }^\circ\text{C}/\text{min}$ and then heated at the rate of $3\text{ }^\circ\text{C}/\text{min}$. The peak appeared at $-21\text{ }^\circ\text{C}$ referred to the fusion peak of eutectic of $\text{NaCl} \cdot 2\text{H}_2\text{O}/\text{ice}$ and another peak at about $0\text{ }^\circ\text{C}$ was that of ice (Hvidt, A. & Borch, K. (1991)). These samples were classified into three groups according to the concentration of NaCl, Group I, II and III.

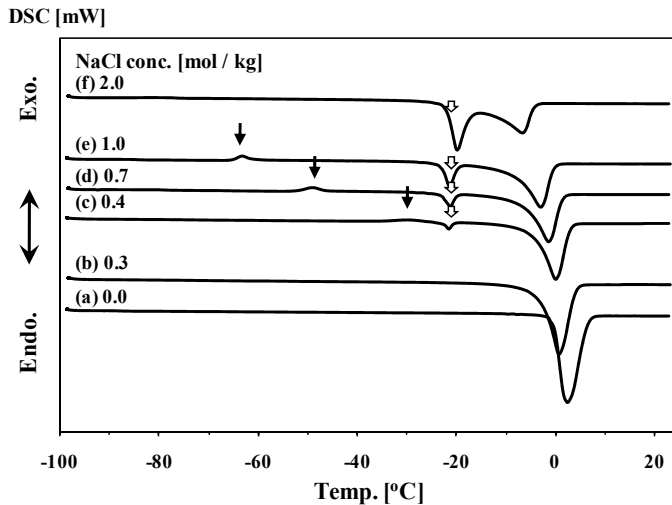


Fig. 8. DSC thermograms of C8Glu/NaCl/water systems in the thawing process at the heating rate of 3 °C/min.

The weight ratio between C8Glu and water was constant (C8Glu : water = 10 : 90 wt%). NaCl concentrations were shown in the figure. *Apparatus*; DSC 60 (SHIMAZU) was used throughout. *Sample preparation*; Each sample was prepared by dissolving NaCl and the C8Glu in a prescribed amount of water and leaving to stand for at least 2 h.

Group I: Chart (a) and (b) in Fig. 8. Their NaCl concentration was low. Only one peak due to fusion of ice was noticeable, that is, formation of the eutectic was completely restrained.

Group II: Chart (c), (d) and (e). Concentration of NaCl was moderate. The exothermic peak due to devitrification was also observed in addition to the peaks due to fusion of ice and eutectic were observed.

Group III: Chart (f). Concentration of NaCl was high. Only two peaks due to fusion of the eutectic and ice were observed and the devitrification was not recognized. The exothermic peak due to devitrification was also observed in addition to the peaks due to fusion of ice and eutectic.

3.2 Analysis of enthalpy for the fusion of eutectic and ice

Fig. 9a shows the relationship between the fusion enthalpy of the eutectic and the ice under the presence and absence of C8Glu. It was examined based on the each DSC chart in Fig. 8. In Fig. 9a, two dotted lines represent the corresponding results obtained under without C8Glu, that is, the result of NaCl solution. Quantitative analysis of the two peaks was conducted as shown in Fig. 9b.

As can be seen from Fig. 9a, it was confirmed that when the amphiphilic sugar derivative, C8Glu, was not present, the fusion enthalpy of ice decreased and that of the eutectic increased linearly with the concentration of NaCl. This result was interpreted that formation of the eutectic was regulated by NaCl concentration.

On the other hand, when C8Glu was present in the system at the concentration of 10 wt% of water mass, the fusion enthalpy of eutectic was zero in a region of Group (I), and that of the ice slowly decreased compared with that in other Groups. It meant that a part of water was retained as non-freezing water, which could not be attributed to the formation of ice even below $-100\text{ }^{\circ}\text{C}$. In the Group (II) and (III) in Fig. 9a, dotted and solid two lines were depicted in parallel. It signified that formation of a definite amount of eutectic was depressed by C8Glu in the system regardless of NaCl concentration.

In this section, C8Glu clearly depicted the conception on the additive effect of amphiphilic sugar derivatives for eutectic formation. Some other sugar derivatives such as C12Raffinose, C12Sucrose, C12Maltose, C8Mannose, C8Glucose appeared in the following section also exhibited a similar behavior. From their nonspecific behavior, it was concluded that the characteristics that amphiphilic sugar derivatives possess the ability to depress the formation of eutectic was general one.

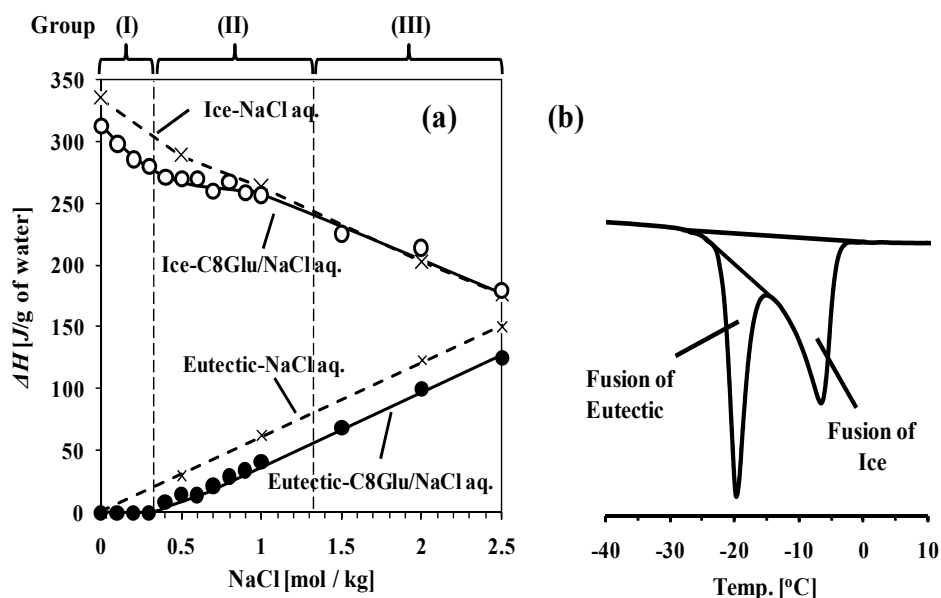


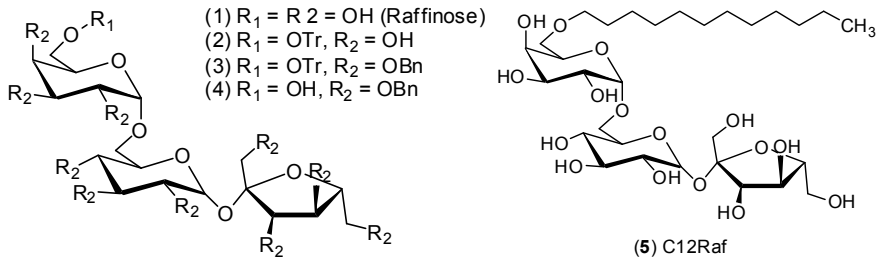
Fig. 9. Analysis of melting enthalpies in NaCl/water system with and without C8Glu. (a) Fusion enthalpies of ice (above) and eutectic (below). (b) Calculated enthalpy areas of ice fusion and eutectic fusion.

3.3 Simultaneous XRD-DSC analysis

The depression effect of another amphiphilic sugar derivative for eutectic formation was studied to clarify its mechanism in detail. Here, C12Raf was used as a specimen instead of C8Glu. Scheme 2 shows its chemical structure and synthetic route.

Fig. 10 indicates DSC thermograms of C12Raf/NaCl/water ternary systems in the thawing process. The sample of C12Raf solution was prepared in a same concentration; C12Raf to water = 1:3 [weight ratio]. The molality of two NaCl solutions were 1.0 mol/kg and 2.5

mol/kg in Fig. 10(a) and (b). The appearance of the chart in Fig. 10(a) was similar to that of C8Glu system of Group (I) stated in Fig. 8 and Fig. 10(b) was to that of Group (II), respectively, although their sample situations were different in terms of their constituent and concentration.



Scheme 2. Chemical structure and synthesis of 6''-O-dodecylraffinose (5: C12Raf).

C12Raf was synthesized from raffinose in four steps of tritylation, benzylation, detritylation, dodecylation and subsequent debenylation, as shown in Scheme 2.

In two DSC charts in Fig. 10, an irregular deviation pointed out by an arrow was recognized on the base line. It appeared at -40°C in (a) and -50°C in (b). They were corresponding to a glass transition at this temperature, respectively. Fig. 10(a) suggested that the unfrozen phase was converted into the glass state after ice was built up during the cooling process. The exothermic peak appeared at around -40°C during the heating process, in Fig. 10(b). It indicated that the devitrification conclusively occurred immediately after the glass transition. The unfrozen phase in a ternary sample became a glass state by freeze-condensation during a cooling process at -70°C . Consequently, the formation of eutectic has been depressed under the kinetics.

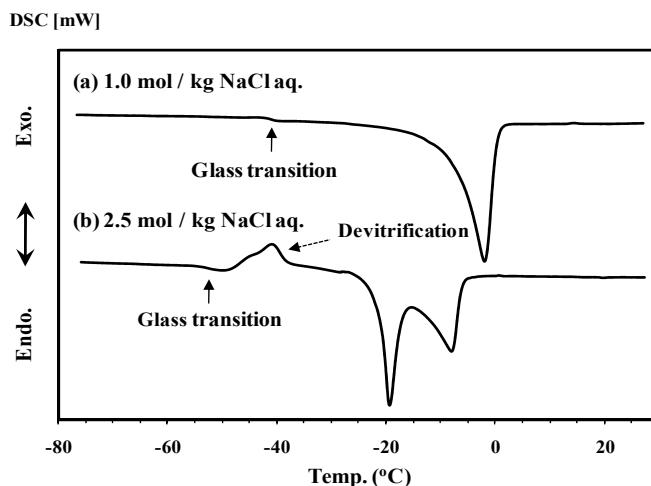


Fig. 10. DSC thermograms of C12Raf/NaCl/water systems in the thawing process.

The weight ratio between C12Raf and water was constant (C12Raf : water = 25 : 75 wt%). NaCl concentration was as follows; (a) 1.0 mol / kg of pure water and (b) 2.5 mol / kg of pure water.

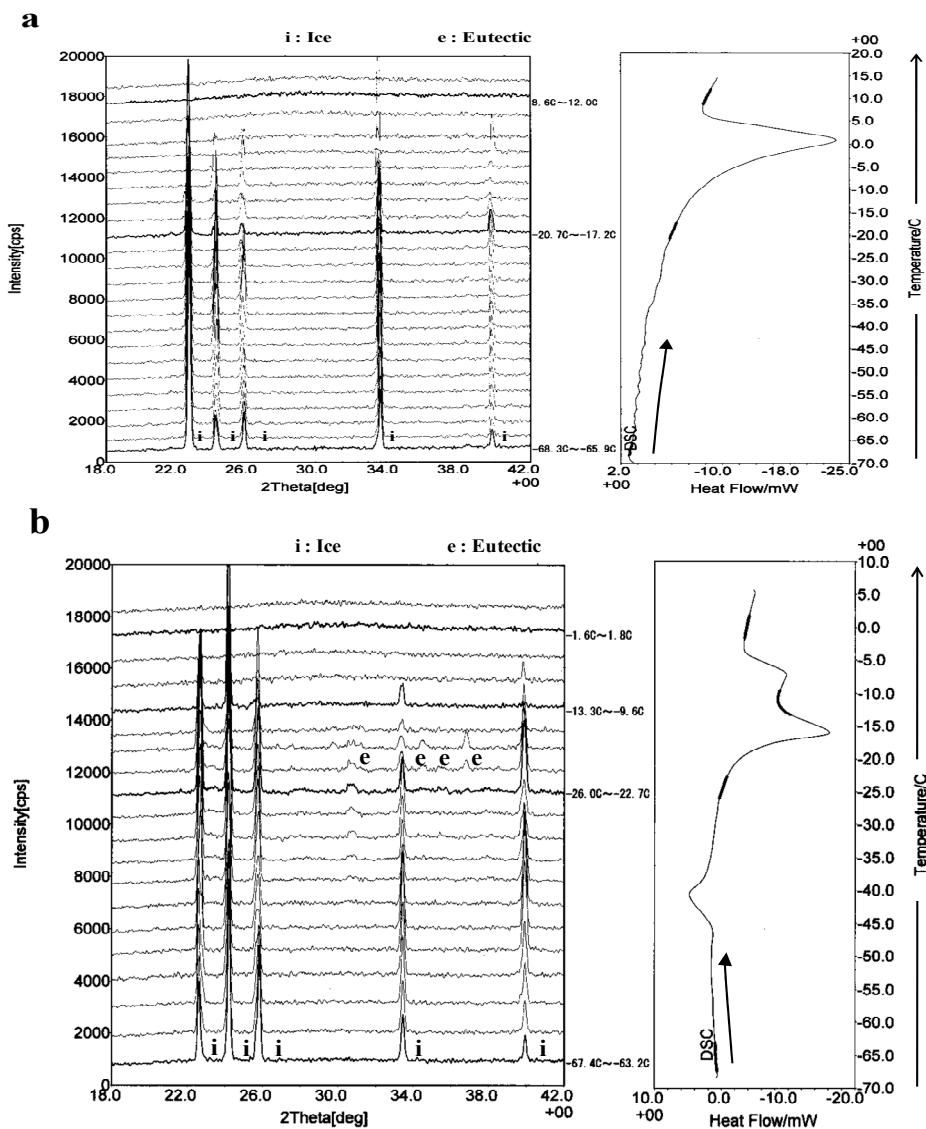


Fig. 11. Simultaneous XRD-DSC measurement of the thawing process in C12Raf/NaCl/water systems.

C12Raf and water was constant (C12Raf : water = 25 : 75 wt%). NaCl concentration was as follows; (a) 1.0 mol / kg of pure water and (b) 2.5 mol / kg of pure water. Apparatus; XRD-DSC II (RIGAKU) was used for measurement. Details of this apparatus are found elsewhere (Arii, T.; Kishi, A. & Kobayashi, Y. (1999). Kishi, A.; Otsuka, M. & Matsuda, Y. (2002).) Measurement conditions were as follows; 1. Cooled to -70°C at $-6^{\circ}\text{C}/\text{min}$. 2. Heated to 15°C at $2^{\circ}\text{C}/\text{min}$. e, eutectic diffraction peak; i, ice diffraction peak.

The two samples in Fig. 10, (a) and (b) were examined by simultaneous XRD-DSC measurement. The results were summarized in Fig. 11. DSC chart of Fig. 11(a) showed only one peak due to the fusion of ice. The XRD-DSC chart demonstrated that when the system was cooled until $-70\text{ }^{\circ}\text{C}$ the ice was definitely formed. Five peaks at $2\theta = 22.5, 24.1, 25.8, 33.4, 39.8$ [deg] were observed during the experiment. All diffraction peaks could be indexed to the standard hexagonal ice (Nishimoto, Y.; Kaneki, Y. & Kishi, A. (2004).). These peaks disappeared in the region above $0\text{ }^{\circ}\text{C}$. No peaks other than the ice were observed throughout the each and every temperature examined. It meant that formation of eutectic was completely depressed by C12Raf at this NaCl concentration.

Fig. 11(b) showed the XRD-DSC profiles for the sample prepared under a concentrated NaCl solution, its molality was 2.5 mol/kg . Highly meaningful results could be obtained by this method. In a cooler region of temperature between $-67\text{ }^{\circ}\text{C} \sim -30\text{ }^{\circ}\text{C}$, five peaks due to a hexagonal system of ice appeared at the same 2θ angles as in Fig.11(a) in a similar manner. At higher temperature after an exothermic peak appeared at about $-40.5\text{ }^{\circ}\text{C}$, four peaks newly emerged at $2\theta = 30.7, 34.5, 35.8, 36.8$ [deg]. This peak pattern was in fair consistent with the authentic diffraction data of the eutectic, $\text{NaCl} \cdot 2\text{H}_2\text{O} / \text{ice}$ (Kajiwara, K.; Motegi, A. & Murase, N. (2001).). That is, it was found that the devitrification induced the formation of eutectic after the occurrence of the glass transition at $-50\text{ }^{\circ}\text{C}$. These four peaks were extinguished accompanied by fusion of the eutectic above $-21\text{ }^{\circ}\text{C}$. Further increment of temperature also resulted in a complete disappearance of the diffractive peaks of the ice.

These experiments were able to be summarized as follows; in a circumstance of dilute NaCl solution such as Group (I), the formation of eutectic was depressed by amphiphilic sugar derivatives such as C8Glu and C12Raf during both cooling and heating processes. On the other hand, in a medium concentrated NaCl solution designated Group (II), the formation of eutectic was restricted during the cooling process, but during the heating process, the devitrification induced the formation of eutectic after the occurrence of the glass transition. As could be seen from the Fig. 10, both in Group (I) and (II), the glass transition was confirmed during the heating process.

The glass formation plays a main role for this phenomenon, such as depression of eutectic formation. Non-amphiphilic free sugars and certain polymers have properties to change an aqueous solution into the glass state and inhibit the eutectic formation (Nicolajsen, H. & Hvidt, A. (1994). Izutsu, K.; Yoshioka, S. & Kojima, S. (1995). Kajiwara, K.; Motegi, A. & Murase, N. (2001).). The amphiphilic sugars would exhibit more effective capabilities except for depressing the formation of eutectic because of the versatile characteristics based on their interface active properties.

3.4 Effects of hydrophobic length and sugar structure on inhibition of eutectic formation

Two different kinds of sugars with a hydrophobic group or without it were examined to make clear the influence of the hydrophobic groups on the inhibition effect for eutectic formation. In other word, the effect of formation of aggregate of the specimen was examined. The results were summarized in Fig. 12. 6''-O-Dodecylraffinose (C12Raf) and 6'-O-dodecanoylsucrose (C12Suc) were used as specimens. The former linked the hydrophobic dodecyl group through ether linkage and the latter combined it through dodecanoyl ester linkage.

We separately confirmed their aggregation behavior using an automatic digital Kyowa Surface Tensiometer, CBVP-3 (Kyowa Kaimen Kagaku Ltd) by Wilhelmy-plate method. It was found that these amphiphilic sugar derivatives, C12Raf and C12Suc, had a critical micelle concentration (cmc) in pure water at 0.49 mM and 0.16 mM, respectively under room temperature. It meant these sugars formed the aggregate in the measuring conditions.

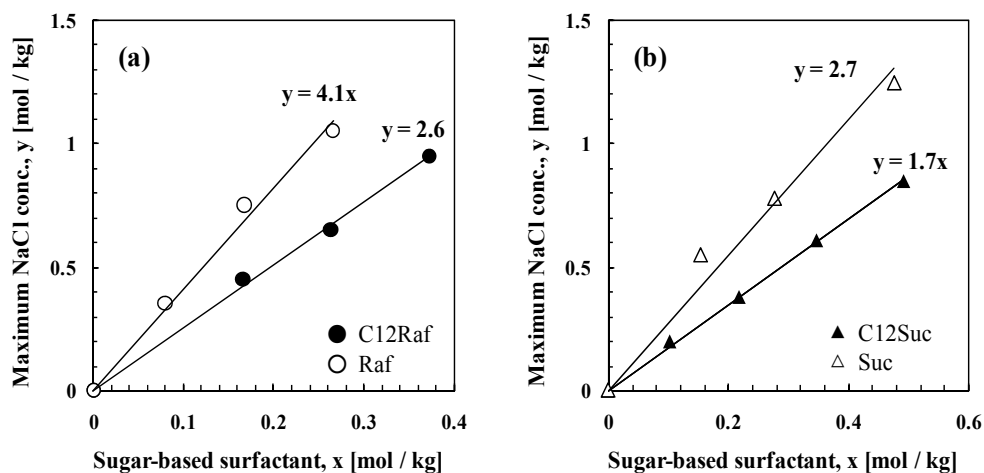


Fig. 12. NaCl concentration range of Group (I) with various concentrations of sugar-based amphiphiles. (a) C12Raf and Raf. (b) C12Suc and Suc.

Fig. 12 demonstrated relationships between the inhibited maximum NaCl concentration and the amphiphilic sugar concentration of the system in terms of two kinds of sugar. The ordinate suggested the maximum concentration of NaCl where the formation of eutectic was completely inhibited corresponding to the concentration of the sugar on the abscissa. This NaCl concentration refers to the boundary one between Group (I) and (II) shown in Fig. 9(a).

As can be seen from Fig. 12, depression ability for the formation of eutectic was clearly proportional to concentration of the sugar. The amphiphilic trisaccharide (C12Raf) and disaccharide (C12Suc) showed smaller depression ability than the corresponding non-amphiphilic free sugar. Its ratio was about 0.63 for all sugars examined.

The slope of the graph in Fig. 12 suggests the magnitude of the depression ability expressed in units per sugar molality. Fig. 13 showed the comparison of various kinds of sugars on the depression effect for eutectic formation. Amphiphilic glucose (C8Glu), mannose (C8Man) and gulose (C8Gul) are monosaccharide, sucrose (C12Suc) and maltose (C12Mal) are disaccharides, and raffinose (C12Raf) is trisaccharide. As can be seen from Fig. 13, the depression ability for the formation of eutectic of the sugar derivatives was proportional to the number of saccharide unit that constituted the hydrophilic part of the amphiphiles. The formation of eutectic made from about 0.8 ~ 0.9 molality of NaCl solution was inhibited by a unit molality of the sugar derivative per single unit of the

saccharide in a proportional manner. In contrast to this, the epimeric isomerism and the structural isomerism between aldose and ketose gave little influence on the capability of inhibition of eutectic formation.

In Group (I) region, the depression effect for the eutectic formation resulted from the vitrification of an unfrozen aqueous phase during the cooling process. T_g of anhydrous amphiphilic sugar derivatives of which the number of sugar unit are different were as follows: C8Glu = 11.2 °C (Ogawa, S.; Asakura, K & Osanai, S. (2010).); C8Mal = 50.4 °C (Kocherbitov, V. & Söderman, O. (2004).; C12Maltotrioside = 100 °C (Ericsson, C. A.; Ericsson, L. C. & Ulvenlund, S. (2005).)). As can be seen from this, the T_g of the sugar derivatives increased as the number of sugar unit increased. It was confirmed that the facility making vitrification was closely associated with the number of the sugar per a unit volume of the system or density of it.

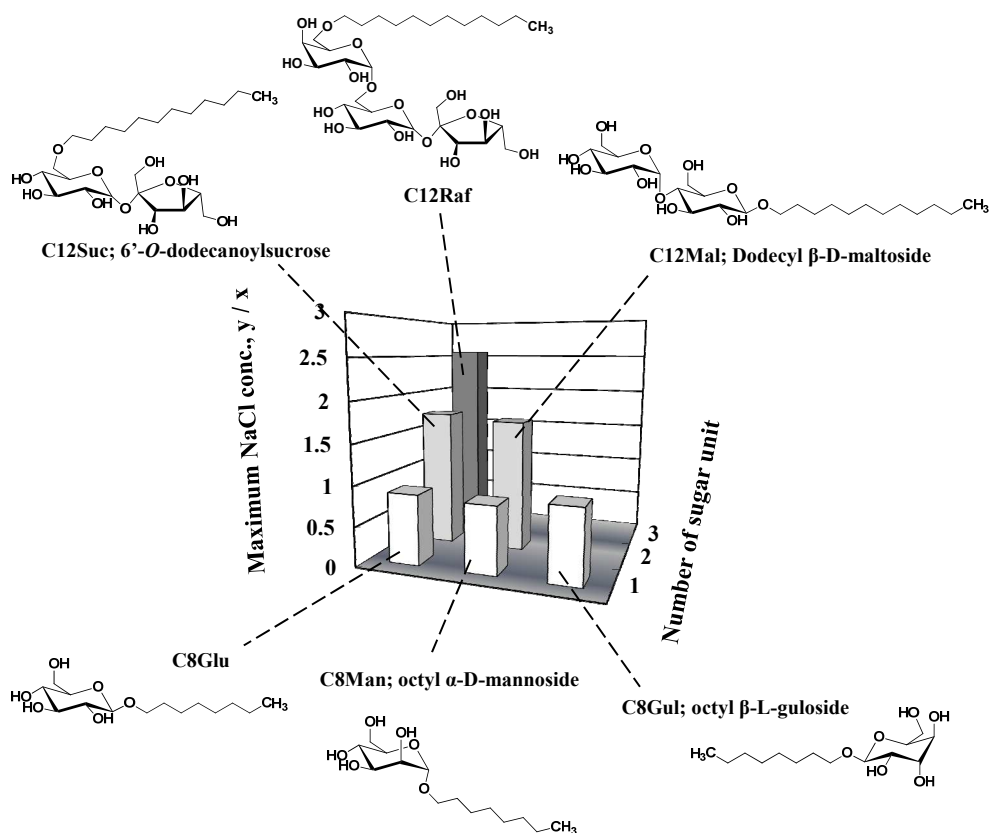


Fig. 13. Comparison of inhibition effect on eutectic formation with sugar structure.

Material; C12Suc, C8Man, C8Gul, and C12Mal were prepared according to published procedures (Ferrer, M.; Cruces, M. A.; Bernabé, M., Ballesteros, A. & Plou, F. J. (1999). Bryan, M. C.; Plettenburg, O.; Sears, P.; Rabuka, D.; Wacowich-Sgarbi, S. & Wong, C.-H. (2002).).

4. Conclusion

Although sugar-based surfactants possess extraordinarily high T_g in an anhydrous state, little is known about the actual application for its excellent glass forming ability. So far as we, authors know, this is the first attempt to apply it in aqueous system and description on it.

In this chapter, we presented the fundamental behavior of glass formation of sugar-based surfactant/water binary system and the inhibition effects of the sugar-based amphiphiles on the formation of eutectic that caused a lot of damage to a variety of bio-organisms from cells to proteins.

In Section 2, the outline of the glass transition behavior of C8Glu, which is one of the representative sugar-based surfactant, and water mixture system was described and summarized. It was clarified the formation of "lyotropic liquid crystal glass" generated from the liquid crystal such as cubic (Q) and lamella (L) in this system. The experimental data for T_g of the lyotropic liquid crystal glass were in fair agreement with the theoretical values proposed as "original" equation by Couchman-Karasz. The peculiar behavior of the system observed through the change of specific heat (ΔC_p) during the glass transition and enthalpy (ΔH) of the phase transition from lamella to isotropic solution or fused liquid was discussed from the standpoint of the permeability of water molecule in the bimolecular membrane structure.

In Section 3, we mentioned the key aspects of the relationships between the inhibiting effect of the sugar-based surfactants and the generating of eutectic in the system. It was also confirmed that increasing saccharide unit of sugar-based surfactant induced an excellent inhibiting effect to the formation of eutectic. Although we focused the increment of the inhibiting power for the formation of eutectic on the introduction of hydrophobic group into the free sugar, the resulted sugar-based surfactant showed only 0.63 times ability for it comparing with the original free sugar.

Because the sugar-based surfactants possess not only the glass forming ability but also the interface active property in the same time, we could expect the possibility that these surfactants show some useful characteristic which could not be obtained by the ordinary free sugars. For example, various kinds of surfactant exhibit abilities that they can depress the deactivation of the protein during the freezing and thawing (Chang, B. S.; Kendrick, B. S. & Carpenter, J. F. (1996). Hillgren, A.; Lindgren, J. & Aldén, M. (2002).). But some surface active agents do not always show their contribution to maintain activities of the water soluble proteins such as LDH (Lactate Dehydrogenase) and β -Galactosidase in the freeze-drying treatment. In contrast to this, when a little amount of a certain sugar derivative was added to a system, it exhibited excellent effects for appreciable retention of the protein activities not only during freeze-thawing but also during freeze-drying processes (Izutsu, K.; Yoshioka, S. & Terao, T. (1993, 1994). Izutsu, K.; Yoshioka, S. & Kojima, S. (1995).).

It has been well known that carbohydrates or sugars are materials that can easily form glass state. (Dave, H.; Gao, F.; Lee, J.-H.; Liberatore, M.; Ho, C.-C. & Co, C. C. (2007).). The sugar-based surfactants could be considered as excellent multiple function surfactants, because

they have two representative properties of the glass forming and the interface activity. Their application in an aqua-system expands its availability in the fields of foods, medicine and functional materials. Although they have the potential to play an advisable role, their application in a multicomponent system remains underdevelopment state. Under the current situation, the sugar-based surfactant has been applied in the bio-science fields, such as a preservation agent of proteins by freeze-drying method, a solubilizing agent for the preparation of reconstituted protein etc.

We expect that the research mentioned here would be further studied and contribute to their practical application of the sugar-based surfactants including the analytical development on the physico-chemical properties.

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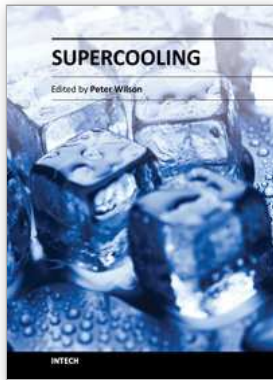
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Supercooled liquids are found in the atmosphere, in cold hardy organisms, in metallurgy, and in many industrial systems today. Stabilizing the metastable, supercooled state, or encouraging the associated process of nucleation have both been the subject of scientific interest for several hundred years. This book is an invaluable starting point for researchers interested in the supercooling of water and aqueous solutions in biology and industry. The book also deals with modeling and the formation subsequent dendritic growth of supercooled solutions, as well as glass transitions and interface stability.

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