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Insights into nitromethane combustion from detailed kinetic modeling – Pyrolysis experiments in jet-stirred and flow reactors

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Highlights

- New Pyrolysis experiments of nitromethane in jet-stirred and flow reactors.
- A detailed mechanism for the pyrolysis and oxidation of nitromethane is developed.
- $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$ is found to be the most important reaction.
- $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$ is the main reaction for nitromethane decomposition.

Abstract

The pyrolysis of nitromethane highly diluted in helium was studied in a plug flow reactor and in a jet-stirred reactor at 1.07 bar and over the temperature range from 500 to 1100 K. Mole fraction profiles of major products and of intermediates were identified with gas chromatography and Fourier transform infrared spectroscopy. Using these experimental data, as well as published ones, we have developed a newly compiled model for the prediction of the pyrolysis and of the

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oxidation of nitromethane in jet-stirred and flow reactors, freely propagating, and burner-stabilized premixed flames, as well as in shock-tubes. The experimental results from the present work and from the literature are interpreted with the help of the kinetic model derived here. This study mainly focuses on the analysis of speciation in different reactors. Among the nitrogenous species, NO is found to be a major product for pyrolysis and oxidation. The model suggests that for nitromethane pyrolysis and oxidation the thermal dissociation channel to CH₃ and NO₂ is the main reaction path for the nitromethane degradation followed by the H-atom abstraction channel. The most sensitive reactions for nitromethane pyrolysis in a flow reactor and during pyrolysis and oxidation in a jet-stirred reactor are found to be CH₃NO₂(+M) ⇌ CH₃ + NO₂(+M) and CH₃ + NO₂ ⇌ CH₃O + NO. The reaction CH₃ + NO₂ ⇌ CH₃O + NO is found to be the most important reaction for all conditions studied. In a burner-stabilized premixed flame, as the mixture gets richer, the thermal dissociation channel CH₃NO₂(+M) ⇌ CH₃ + NO₂(+M) becomes more important as the contribution of the H-atom abstraction channel is decreased. Furthermore, in the burner-stabilized premixed flames, it was found that NO is mainly formed via NO₂: NO₂ + H ⇌ NO + OH, NO₂ + CH₃ ⇌ CH₃O + NO. The model provided an overall reasonable agreement with the experimental data. However, for pyrolysis conditions, future work is desirable to improve predictions of intermediate species. This work extends the kinetic database and helps to improve the understanding of nitromethane chemistry. The kinetic model presented in this work can serve as a base model for hydrocarbons and oxygenated fuels higher than C₂ and nitrogen-containing compounds higher than C₁ as well as for pure nitrogen compounds.

Keywords: Nitromethane, Pyrolysis, Oxidation, Jet-stirred reactor, Flow reactor, Kinetic modeling,

1. Introduction

The diminishing availability of conventional fuels and stricter regulations on pollution control and CO₂ emission targets have led scientists and engineers to look for alternative fuels. Recently nitromethane has slowly gained interest as an alternative fuel or as a blend for internal combustion (IC) engines. Historically, a few detailed experimental studies have been conducted to determine the effect of engine and fuel operating parameters on the power output of an internal combustion engine using nitroparaffins as fuel. Nitromethane is one of the simplest nitroparaffins. In the past, it was mostly used as rocket propellant or as explosive and has been used as a reference fuel to understand the combustion mechanism of monopropellants [1-3]. Nitromethane is an energetic compound with a wide variety of applications, including its use as a monopropellant, a liquid explosive, a solvent for chemical processing and analysis, and a high-performance fuel additive for internal combustion engines and pulsed detonation engines [4,5]. The chemical formula of nitromethane is CH₃NO₂, it is liquid at room temperature (melting point -29°C) with a density of 1.138 g/cm³ at 20 °C and has a boiling point of 101.2°C [6]. In essence, it is the simplest of possible energetic C-H-O-N molecules that contain a nitro group, which is why it is often used in reaction studies as a prototype for more complex energetic materials. It is partly soluble in aliphatic hydrocarbons, mostly soluble in aromatics, and highly soluble in lighter (e.g methanol) alcohols. If subjected to shock, it has a tendency to decompose with considerable

violence and to explode in the manner of certain other organic nitrogen compounds [4]. Internal combustion engines are the commonly used propulsion system in passenger cars and racing cars. Glow plug two-stroke engines are often used to propel small-unmanned aerial vehicles and use a fuel mixture composed of methanol and a small quantity (5–20%) of nitromethane [7]. The effect of nitromethane blends on the engine performance was studied by Starkman [4] and an increase in engine power was reported with nitromethane addition. The above finding was supported with other studies as well [7,8]. In drag racing, nitromethane is blended with gasoline to enhance the engine power output [9]. The lower heating value of nitromethane is 11.3 MJ/kg while for gasoline it is 42–44 MJ/kg. The amount of air required to burn 1 kg of gasoline is approximately 14.7 kg while 1 kg of nitromethane only requires 1.7 kg of air. For a given engine cylinder volume 8.6 times more nitromethane can be burned than gasoline in one stroke. Hence, 2.3 times more power is generated compared to gasoline when combined with a given amount of oxygen [8]. The group of Yokoo et al. [10] investigated the effect of various fuel blends on the thermal efficiency of a spark-ignited engine. They stated that the addition of 5% of nitromethane extended the lean limit operation and improved the thermal efficiency by 1–2%. On the other hand, the research octane number (RON) is reduced by up to 5 points and the knock probability is increased. Similar characteristics are also reported by the group of Cracknell et al. [11] who studied the effect of 5% addition of nitromethane to gasoline in a homogeneously charged compression ignition engine. However, detailed combustion aspects of nitromethane are not addressed in these works. As a pollutant from the automotive exhaust, nitromethane is also observed both experimentally in the laboratory and on-site from automobiles [12–14]. To investigate nitromethane combustion characteristics in engines through modeling as a pure fuel or as a blend, a reliable and comprehensive kinetic mechanism is needed.

Additionally, several explosives such as Trinitrotoluene or High Melting Explosive (also known as octogen) are major environmental contaminants as a result of their manufacture and deployment in both industrial and military applications, beside from the obvious physical hazards associated with explosives [15,16]. While incineration is the most common technique to treat soils contaminated by explosives [17], it can be a source of dioxins or other harmful compounds which then require additional disposal or further treatments [18]. Consequently, it is of interest to study an alternative process such as thermal decomposition. That is why it is of interest to study the thermal decomposition of nitromethane as a surrogate of octogen and also to better understand the chemistry of nitro compounds.

Though studies on nitromethane are not new, there are a number of studies on nitromethane detonation [19–21] and nitromethane decomposition [22–25]. None of these studies took into account the detailed kinetic study of CH_3NO_2 combustion. Glarborg et al. [26] calculated the reaction rate for the thermal dissociation of CH_3NO_2 ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$) based on literature data and proposed a nitromethane mechanism validated only against the shock tube speciation from literature. Boyer and Kuo [2] proposed a detailed mechanism for nitromethane with 47 species and 250 reactions. However, their mechanism was dedicated to flame structure and burning rate only. When the work of Glarborg et al. [26] and Boyer and Kuo [2] was published, experimental data on the basic combustion features (ignition delay time, laminar flame speed)

were not available in literature. Naulé et al. [27] performed ignition delay time experiments of $\text{CH}_3\text{NO}_2/\text{O}_2/\text{N}_2$ in a shock tube and used two mechanisms from literature [28,29] to compare against their experimental data and mechanism predictions. Similarly, Brackmann et al. [30] measured the burning velocity of $\text{CH}_3\text{NO}_2/\text{O}_2$ diluted with CO_2 and the concentrations of NO , CO and CH_2O in two CH_3NO_2 flames (experimental condition shown in Table 1). They used the same two mechanisms [28,29] as Naulé et al. and compared model predictions and measurements. It was concluded in both studies that none of the mechanisms is able to accurately predict the experimental data [27,30]. They reached the similar conclusion that the nitromethane combustion chemistry is not fully understood, and further work is required to enable the construction of a kinetic mechanism with predictive capacity over a wide range of conditions.

Table 1: Experimental studies from literature and this work. py-: pyrolysis, ox-: oxidation Note: Temperature in burners are given for mixture inlet.

| Experimental devices | Measured properties | Experimental conditions | References |
|------------------------------|----------------------|---|-----------------------|
| Shock tube | Ignition delay times | 2.0 - 30.0 atm, 875 - 1595 K, $\phi = 0.5 - 2.0$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ | Mathieu et al. [29] |
| Shock tube | Ignition delay times | 8.0 - 32.0 atm, 947-1333 K, $\phi = 0.5 - 2.0$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{N}_2$ | Naulé et al. [27] |
| Shock tube | Ignition delay times | 2.0 - 10.0 atm, 1200 - 2000 K, $\phi = 0.5 - 2.0$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ | Gao et al. [34] |
| Flow reactor (py-) | Speciation | 30, 150 and 760 Torr, 750 - 1400 K, for 5 mol% CH_3NO_2 / 95 mol% Ar | Jia et al. [35] |
| Flow reactor (py-) | Speciation | 5 Torr, 735 - 1476 K, for 5 mol% CH_3NO_2 / 95 mol% Ar | Weng et al. [36] |
| Flow reactor (py-) | Speciation | 1.07 bar, 500 - 1100 K, for 1 mol % CH_3NO_2 / 99 mol% He | This work |
| Jet stirred reactor (ox-) | Speciation | 1 atm, 600 - 875 K, $\phi = 0.4$ and 2.0 for $\text{CH}_3\text{NO}_2 / \text{O}_2 / \text{Ar}$ | Weng et al. [36] |
| Jet stirred reactor (py-) | Speciation | 1.07 bar, 500 - 1100 K, for 1 mol% CH_3NO_2 / 99 mol% He | This Work |
| Spherical combustion chamber | Laminar flame speed | 0.5 - 3.0 bar, 423 K, $\phi = 0.5 - 1.3$ for $\text{CH}_3\text{NO}_2/\text{air}$ | Brequigny et al. [28] |
| Heat flux burner | Laminar flame speed | 1 atm, 338 - 358 K, $\phi = 0.8 - 1.6$ for $\text{CH}_3\text{NO}_2/\text{air}$ | Naulé et al. [37] |
| Heat Flux burner | Laminar flame speed | 1 atm, 348 and 358 K, $\phi = 0.8 - 1.6$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{CO}_2$ | Brackmann et al. [30] |
| McKenna Burner | Speciation | 4.67 kPa, 273 K, $\phi = 1.39$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ | Tian et al. [38] |
| McKenna Burner | Speciation | 4.655 kPa, 298 K, $\phi = 1.0 - 2.0$ for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ | Zhang et al. [39] |
| Heat Flux burner | Speciation | 1 atm, 363 K, $\phi = 0.8$ and 1.2 for $\text{CH}_3\text{NO}_2/\text{air}/\text{N}_2$ | Brackmann et al. [30] |

Historically, there are only experimental studies or combined (numerical and experimental) studies, often dedicated to a single reactor at specific conditions or experimental setups. Recently, there have been various experimental studies on CH_3NO_2 due to which the development and validation of CH_3NO_2 models can be improved. They are summarized in Table 1. However, to our best knowledge, there is no generally applicable kinetic model for CH_3NO_2 combustion covering

also the NO_x chemistry over a wide range of experimental conditions. This situation calls for an in-depth analysis of nitromethane – NO_x chemistry interactions.

The aim of the present work is twofold:

- 1) to perform an experimental study of nitromethane pyrolysis in a tubular flow reactor and in a jet-stirred reactor, which will serve to extend the existing database and can be used for nitromethane model development.
- 2) to extend our recently published mechanism [31], which describes the oxidation of CH₃OH and C₂H₅OH and its interaction with NO_x (including CH₃NO₂ as a fuel), to subsequently study their interaction with hydrocarbon and nitrogen chemistry by considering the available experiments in the validation process.

To our best knowledge, no experimental pyrolysis study of nitromethane in a jet-stirred reactor has been published before and this work is the first to report such data. This work is an ongoing effort to develop a reliable and comprehensive mechanism for fuel/NO_x interaction taking into account detailed nitrogen chemistry. In addition, the model derived in this work is critically tested for laminar flame speeds (LFS), for ignition delay times (IDT), for speciation in jet-stirred reactors (JSR), in plug flow reactors (PFR), and burner stabilized flames (BSF). The development and compilation strategy for our baseline mechanism is described in our previous work [31,32] (which focused on NH₃/H₂/CO/CH₄ oxidation and NO_x chemistry interaction). The present study is conducted in a similar manner. This makes the kinetic model more robust and reliable for combustion modeling under engine conditions (see section on *n*-heptane spray ignition in [33]).

2. Experimental

Experiments were carried out in two continuous flow reactors: a jet-stirred reactor (JSR) and a tubular reactor (TR) at a constant pressure of 1.07 bar, an inlet nitromethane mole fraction of 0.01 (high dilution in helium) and a residence time at the set-point temperature of 2.0 s. The experimental setup used in this study is shown in Figure 1. In both reactors, species were sampled at the outlet and analyzed using gas chromatography and Fourier-Transform Infrared (FTIR) spectroscopy. The features of the two reactors and analytical procedure are described in the following paragraphs.

Helium was provided by Messer (purity of 99.999%) and nitromethane was purchased from Sigma-Aldrich (purity of 98.5%). Gas flow rates were controlled by mass flow controllers and the liquid flow rate by a Coriolis flow controller. The uncertainty in the flow measurements was around 0.5% for each controller, so about 1% on the residence time.

2.1. Tubular reactor

The tubular reactor set-up was developed recently in LRGP for kinetic studies at elevated temperatures [40] (up to 1300 K) because this type of reactor is often used in industry for a variety of purposes. It consisted of an alumina tube (recrystallized alumina, volume of 294 cm³)

with an inner diameter of 20 mm and an outer diameter of 25 mm. The tubular reactor was heated by an electrical furnace from Vecstar. Temperature gradients along the tube were measured using a S-type thermocouple and are shown in Figure 2.

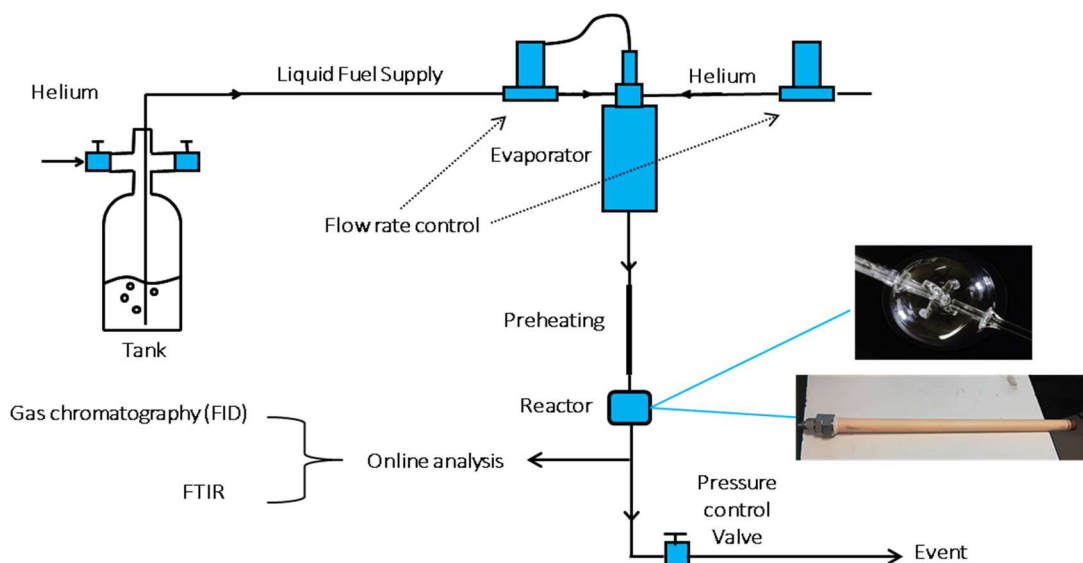


Figure 1: Experimental setup.

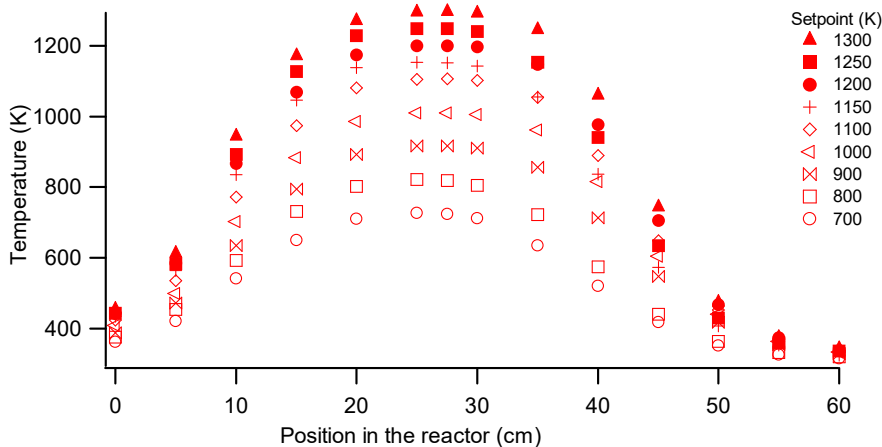


Figure 2: Measured temperature profiles in TR in the absence of reaction.

Simulations were performed assuming that the TR can be modeled as a PFR under our operating conditions (using the measured temperature profiles). This assumption has already been checked elsewhere [41].

2.2. Jet-stirred reactor

The spherical JSR (volume of 85 cm³), made of fused silica, is a type of ideal continuously stirred tank reactor which is well suited for gas phase kinetic studies. Reactants enter through an injection cross made of four nozzles which were located at its center. High turbulence was created

by gas jets from the nozzles and led to homogeneity in composition. The isothermal JSR was preceded by a fused silica annular (ring-shaped) preheating zone in which the temperature of the gas was increased up to the reactor temperature allowing a homogenous gas phase temperature in the reactor. The residence time of the gas mixture inside the annular preheater was very short compared to its residence time inside the reactor (a few percents). Both the reactor and the preheating zone were heated by means of Thermocoax resistances rolled up around the reactor. The reaction temperature was measured at the center of the reacting gas phase by a K-type thermocouple located in a glass finger.

2.3. Analytical procedure

A heated transfer line maintained at 433 K was used to avoid product condensation between the reactor outlet and the analytical devices. The gases leaving the reactors were then analyzed using:

Gas chromatography (GC): a chromatograph (fitted with a PlotQ capillary column) a thermal conductivity detector (TCD) and a flame ionization detector (FID), was used for the quantification of light products and the reactant. The identification of reaction products was performed using a gas chromatograph equipped with a plotQ capillary column and coupled to a mass spectrometer. Response factors were determined by injecting calibration mixtures or using the effective carbon number method. Relative uncertainties in mole fractions were estimated to be $\pm 5\%$ for species which were calibrated using standards and $\pm 10\%$ for other ones.

A Fourier Transform Infrared gas analyzer (FTIR) from Thermo Scientific Antaris equipped with a Mercure Cadmium Tellure photoelectric detector was used to identify and quantify CH_2O , CO , H_2O , HCN , NO , CO_2 and CH_3OH . FTIR calibrations were obtained by injecting standards. A typical FTIR spectrum obtained during nitromethane pyrolysis is shown in Figure S1 in the supplementary material. Relative uncertainties were slightly higher than in GC since interferences may occur between bands of absorbing species.

3. Kinetic mechanism

The mechanism proposed here extends our recent work [31], which was developed to predict CH_3OH and $\text{C}_2\text{H}_5\text{OH}$ oxidation and NO_x chemistry interaction. The present model also takes into account NH_3 as fuel. The mechanism of Shrestha et al. [31] was validated for other C_1 - C_2 fuels as well (methane (CH_4), acetylene (C_2H_2), ethylene (C_2H_4) and ethane (C_2H_6)) for a wide range of experimental conditions as available from the literature. We have demonstrated that the model captures C_1 - C_2 hydrocarbon/oxygenated NO_x chemistry interaction as well as NO formation in C_1 - C_2 premixed flames. The here developed mechanism is based on published literature and the NO_x sub-mechanism is augmented to include cross-reactions between nitrogen and carbon species. The main aim of this work is to include the CH_3NO_2 species as fuel molecule and to extend the validation range of our nitrogen chemistry. The nitromethane sub-mechanism includes CH_3NO_2 , CH_2NO_2 , CH_3ONO , CH_3NO , CH_2NO , CH_3NH_2 , CH_3NH , CH_2NH_2 , CH_2NH , and HCNH species and their related chemistry. Among them, CH_3NO_2 , CH_2NO_2 , CH_3ONO , and CH_3NO are already

included in our recently published mechanism [31]. However, for CH₃NO the complete set of decomposition reactions were not included in [31], different to this work. Further, six additional nitrogen containing species (and their related reactions) were considered in this work, namely CH₂NO, CH₃NH₂, CH₃NH, CH₂NH₂, CH₂NH, and HCNH. The inclusion of these species and their reactions have required updates. The complete reaction mechanism (CHEMKIN format) is provided in the supplementary material (SM).

A detailed description of the sub-mechanisms of the additional species, which are included in the present nitrogen chemistry, and the adaptation of the rate parameters are explained below highlighting the important reactions and their kinetics. The reactions and related rate parameters are mostly adopted from the work of Dean and Bozzelli [42].

3.1. Brief overview of recent work on the thermal decomposition of CH₃NO₂

As mentioned above, the mechanism proposed in this work is an extension of our recent work [31], and parts of the nitromethane sub-mechanisms are already included in that work. For the broad range of experimental conditions studied here, it was found that the unimolecular thermal dissociation reaction of CH₃NO₂ ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$) is the reaction that exhibits the highest sensitivity (see also Section 4). The adopted rate parameters of this and other reactions in the CH₃NO₂ sub-mechanism are discussed in detail in our recent work [31]. However, in [31], several recent studies on the thermal CH₃NO₂ decomposition were not taken into account. In these studies [23,25,43-45], it was suggested that CH₃NO₂ not only thermally dissociates to CH₃ + NO₂ but also isomerizes to methyl nitrate (CH₃ONO) via a roaming type transition state finally dissociating to CH₃O + NO. In 2009, Zhu and Lin [43] studied the CH₃NO₂ decomposition theoretically and suggested that the channel to CH₃O + NO could be detected in experiments. Later, Homayoon and Bowman [44] came to similar conclusions in their density functional theory (DFT) based theoretical analysis. However, both studies did not provide information about the conditions under which these products are formed. These were identified by Zhu et al. [45] suggesting that the roaming channel is active at pressures lower than 2 Torr over a wide temperature range (400–3000 K). This finding was confirmed by the succeeding study of Annesley et al. [23] who also performed shock tube experiments. Their theoretical analysis suggested that at pressures above 2 Torr the product distribution undergoes a sharp transition from the roaming dominated (CH₃ONO → CH₃O + NO) to the bond-fission dominated (CH₃ + NO₂) channel. Hence the bond fission channel is dominating for combustion environments featuring significantly higher pressures in agreement with all of the studies mentioned above and the studies that were reviewed in our recent work [31]. Furthermore, this finding is supported by the later study of Matsugi and Shiina [25] proposing branching ratios of 0.97 and 0.03 in favor of the C-N bond fission channel. Despite its low branching ratio, it is clearly recommended to take into account the roaming type channel in kinetic modeling studies [46]. Therefore, in the present work, we have included this channel using the rate parameters proposed by Matsugi and Shiina [25]. The available kinetic data on the dominating C-N fission channel are summarized in Figure S24 for low and high pressure conditions (see Supplementary material). The rate constant used in this work is adopted from [22], complemented with the results of [26] and well within the range of

available literature data. This choice is supported by the recent studies of Matsugi and Shiina [25] and Vlasov et al. [46] and leads to fractions of the roaming type channel from 0.02 to 0.04, increasing with temperature for the conditions studied in this work.

3.2. CH₃NO and CH₂NO sub-mechanisms

Only four reactions related to CH₃NO were included and described in the recent work of Shrestha et al. [31], which are CH₃NO₂ + H ⇌ CH₃NO + OH, CH₃ + HNO ⇌ CH₃NO + H, CH₃ + HNO ⇌ CH₄ + NO and CH₃ + NO(+M) ⇌ CH₃NO(+M). Furthermore, the CH₃NO H-atom abstraction reactions with H and O-atoms form OH, CH₃ and NH₂ radicals with CH₂NO as the main complementary product. However, the reactions CH₃NO + O (a) and CH₃NO + OH (b) are multi-channel reactions, with CH₃ + NO₂ as the second channel for (a) and CH₃ + HONO for (b). The formed CH₂NO from the above process can thermally decompose to HNCO + H. Furthermore, CH₂NO can react with H- and O-atoms as well as OH, CH₃, and NH₂ radicals, which are also multichannel reactions. These are the H-atom abstraction and the NO elimination. The common product of the H-atom abstraction channel is HCNO, the specific products for NO elimination can be found in Table 2. In addition, CH₂NO can also react with O₂ to produce CH₂O + NO₂. The reaction rate parameters of these reactions are adopted from the review of Dean and Bozzelli [42].

Table 2: NO elimination reactions from CH₂NO.

| No. | Reaction |
|-----|---|
| 1. | CH ₂ NO+H⇌CH ₃ +NO |
| 2. | CH ₂ NO+O⇌CH ₂ O+NO |
| 3. | CH ₂ NO+OH⇌CH ₂ OH+NO |
| 4. | CH ₂ NO+CH ₃ ⇌C ₂ H ₅ +NO |
| 5. | CH ₂ NO+NH ₂ ⇌CH ₂ NH ₂ +NO |

3.3. CH₃NH₂, CH₃NH, and CH₂NH₂ sub-mechanisms

CH₃NH₂ is mainly formed via the recombination reaction of CH₃ and NH₂ radicals, other products of CH₃ and NH₂ recombination are CH₂-3 (triplet methylene radical) + NH₃, CH₂NH₂ + H, CH₃NH + H, and CH₂NH + H₂. The rate parameters of these reactions are pressure dependent and are adopted from the work of Dean and Bozzelli [42], who calculated the rates using quantum Rice–Ramsperger–Kassel (QRRK) theory over the temperature range 600–2500 K. The formed CH₃NH₂ can thermally decompose to CH₂NH and H₂ (CH₃NH₂ + M ⇌ CH₂NH + H₂ + M). The reaction rate is adopted from Zhang et al. [47], who calculated the rate constant using ab initio Rice–Ramsperger–Kassel–Marcus (RRKM) theory over the temperature range from 500 to 2000 K. CH₃NH₂ can also undergo H-atom abstraction reacting with H and O-atoms, OH, CH₃ and NH₂ radicals. The reaction product is CH₂NH₂ when the H-atom is abstracted from a carbon-centered site. The common product is CH₃NH when the abstraction takes place from the nitrogen-centered position. The respective complementary products are found in Table 3. The rate parameters of these H-atom abstraction channels are adopted from the work of Dean and Bozzelli [42], who

estimated the values using the Direct hydrogen transfer (DHT) method based on the C-H and N-H bond dissociation energies (395 kJ/mol and 423 kJ/mol, respectively).

Table 3: H-atom abstraction reactions from CH_3NH_2 .

| No. | Reaction |
|-----|--|
| 1. | $\text{CH}_3\text{NH}_2 + \text{H} = \text{CH}_2\text{NH}_2 + \text{H}_2$ |
| 2. | $\text{CH}_3\text{NH}_2 + \text{H} = \text{CH}_3\text{NH} + \text{H}_2$ |
| 3. | $\text{CH}_3\text{NH}_2 + \text{O} = \text{CH}_2\text{NH}_2 + \text{OH}$ |
| 4. | $\text{CH}_3\text{NH}_2 + \text{O} = \text{CH}_3\text{NH} + \text{OH}$ |
| 5. | $\text{CH}_3\text{NH}_2 + \text{OH} = \text{CH}_2\text{NH}_2 + \text{H}_2\text{O}$ |
| 6. | $\text{CH}_3\text{NH}_2 + \text{OH} = \text{CH}_3\text{NH} + \text{H}_2\text{O}$ |
| 7. | $\text{CH}_3\text{NH}_2 + \text{CH}_3 = \text{CH}_2\text{NH}_2 + \text{CH}_4$ |
| 8. | $\text{CH}_3\text{NH}_2 + \text{CH}_3 = \text{CH}_3\text{NH} + \text{CH}_4$ |
| 9. | $\text{CH}_3\text{NH}_2 + \text{NH}_2 = \text{CH}_2\text{NH}_2 + \text{NH}_3$ |
| 10. | $\text{CH}_3\text{NH}_2 + \text{NH}_2 = \text{CH}_3\text{NH} + \text{NH}_3$ |

The formed CH_2NH_2 radical can thermally decompose to CH_2NH and H ($\text{CH}_2\text{NH}_2 \rightleftharpoons \text{CH}_2\text{NH} + \text{H}$) or undergo H-atom abstraction reacting with H-atom and O-atom, OH, CH_3 and O_2 . The H-atom abstraction channel mainly yields the CH_2NH radical. The rate parameters of these reactions are also adopted from the work of Dean and Bozzelli [42], where they performed QRRK calculations for temperatures ranging from 600 to 2500 K for the thermal decomposition of CH_2NH_2 , while for the H-atom abstraction the DHT approach was used. Reactions of CH_2NH_2 with O-atoms as well as OH and CH_3 radicals have other routes as forming $\text{CH}_2\text{O} + \text{NH}_2$, $\text{CH}_2\text{OH} + \text{NH}_2$, and $\text{C}_2\text{H}_5 + \text{NH}_2$, respectively. The rate parameters of these reactions are also adopted from Dean and Bozzelli [42], in analogy with the reactions $\text{C}_2\text{H}_5 + \text{O}$, $\text{CH}_3 + \text{OH}$, and $\text{CH}_3 + \text{C}_2\text{H}_5$ respectively relying on the rate parameters from Baulch et al. [48]. Similarly, CH_3NH formed in the above process can either thermally decompose to CH_2NH and H ($\text{CH}_3\text{NH} \rightleftharpoons \text{CH}_2\text{NH} + \text{H}$) or undergo H-atom abstraction reacting with H and O-atoms, OH, and CH_3 radicals mainly forming CH_2NH and related products. The rate parameters of these reactions are again adopted from the work of Dean and Bozzelli [42]. For the thermal decomposition of CH_2NH_2 , they performed QRRK calculations for temperatures between 600 and 2500 K. The direct H-atom abstraction parameters were derived with the DHT method. We note here that Dean and Bozzelli [42] concluded that thermal dissociation channels are more important than H-atom abstraction channels.

3.4. CH_2NH and HCNH sub-mechanisms

CH_2NH is mainly formed from the thermal dissociation of CH_3NH_2 , CH_2NH_2 , and CH_3NH . The other formation channel of CH_2NH is via the recombination of $\text{CH}_3 + \text{NH}_2$ ($\text{CH}_3 + \text{NH}_2 \rightleftharpoons \text{CH}_2\text{NH} + \text{H}_2$) and $\text{CH}_3 + \text{NH}$ ($\text{CH}_3 + \text{NH} \rightleftharpoons \text{CH}_2\text{NH} + \text{H}$). Hydrogen atom abstraction from CH_2NH can lead to two different radicals: H_2CN and HCNH . The H_2CN radical is formed when the H-atom is abstracted from the N-atom, while abstraction from the C-atoms leads to HCNH formation. The rate

parameters of these multichannel reactions are adopted from the work of Dean and Bozzelli [42] and derived with the DHT method. The reaction of CH_2NH with O-atoms has another route yielding $\text{CH}_2\text{O} + \text{NH}$ ($\text{CH}_2\text{NH} + \text{O} \rightleftharpoons \text{CH}_2\text{O} + \text{NH}$). This reaction is an exothermic reaction and the rate parameters are again adopted from Dean and Bozzelli [42] estimating the A-factor to be half of the value of the addition rate of O-atom to ethylene ($\text{O} + \text{C}_2\text{H}_4$) and took the rate expression from Baulch et al. [49]. HCNH produced in the above process can thermally dissociate to $\text{HCN} + \text{H}$ ($\text{HCNH} \rightleftharpoons \text{HCN} + \text{H}$). Dean and Bozzelli [42] performed a QRRK study for this reaction with N_2 as bath gas over the temperature range 600–2000 K and proposed pressure dependent rate parameters, which were adopted in this work. Furthermore, the reaction of HCNH with H-atoms is involved in two channels: one is the isomerization reaction ($\text{HCNH} + \text{H} \rightleftharpoons \text{H}_2\text{CN} + \text{H}$) and the second is the H-atom abstraction ($\text{HCNH} + \text{H} \rightleftharpoons \text{HCN} + \text{H}_2$). Dean and Bozzelli [42] determined the rate parameters of the reaction $\text{HCNH} + \text{H} \rightleftharpoons \text{H}_2\text{CN} + \text{H}$ in analogy with $\text{H} + \text{C}_2\text{H}_5 \rightleftharpoons \text{CH}_3 + \text{CH}_3$ and took half of the value proposed by Baulch et al. [48]. Dean and Bozzelli [42] estimated the H-atom abstraction channel ($\text{HCNH} + \text{H} \rightleftharpoons \text{HCN} + \text{H}_2$) using the DHT method. The reaction of HCNH with O-atoms is also a multichannel reaction producing $\text{HNCO} + \text{H}$ and $\text{HCN} + \text{OH}$ respectively. Dean and Bozzelli [42] estimated the reaction rate of the first reaction $\text{HCNH} + \text{O} \rightleftharpoons \text{HNCO} + \text{H}$ in analogy with the reaction $\text{O} + \text{C}_2\text{H}_5$ whose rate expression is adopted from Baulch et al. [48]. For the direct H-atom abstraction reaction ($\text{HCNH} + \text{O} \rightleftharpoons \text{HCN} + \text{OH}$), a DHT estimation was performed by Dean and Bozzelli [42]. HCNH reactions with OH and CH_3 radicals are direct H-atom abstractions giving $\text{HCN} + \text{H}_2\text{O}$ and $\text{HCN} + \text{CH}_4$, whose rate parameters were estimated by Dean and Bozzelli [42] using the DHT method.

3.5. Thermochemistry

The thermochemistry of the species in the mechanism was presented in our recent work [31]. The thermochemistry of the six additional species (CH_2NO , CH_3NH_2 , CH_3NH , CH_2NH_2 , CH_2NH , and HCNH) has been adopted from Goos et al. [50], being available in Burcat's thermochemical database [51].

4. Results and discussion

The $\text{CH}_3\text{NO}_2/\text{NO}_x$ model proposed in this work has been validated against a large set of experimental data, newly measured in the present work and taken from the literature. The validation targets include: speciation in flow and jet-stirred reactors, and premixed burner stabilized flames, laminar flame speeds and ignition delay times. This study investigates the $\text{CH}_3\text{NO}_2/\text{NO}_x$ kinetics in a multi set-up approach. In the figures shown below, the lines represent the simulation results using our proposed kinetic model and symbols represent the experimental data from the present work and the published literature unless stated differently. All simulations were performed using the LOGEresearch 1.10.0 [52] software package.

4.1. Pyrolysis of nitromethane

4.1.1. Speciation in a flow reactor

Figure 3, Figure 4 show the temperature dependence of species mole fractions in the reactive mixture during the pyrolysis experiment (inlet conditions: CH_3NO_2 1%/He 99% (on a mole basis)), residence time (τ) of 2.0 s). The lines in the graph represent the predicted species profiles computed with the model developed in this work. It can be seen in Figure 3(a) that the fuel (CH_3NO_2) consumption is very well captured by the model. However, at 1100 K (Figure 3(c)), CO and CO_2 mole fractions are underpredicted by 33%. It can be observed in Figure 3(b) that there is a large experimental scatter for H_2O over the temperature range 850–1000 K and that the model underpredicts its mole fractions by 33%. However, at 1100 K, the model predicts mole fractions close to the experimental data. Water detection could be affected by condensation and/or adsorption in the tube preceding the FTIR cell, partly explaining the scatter observed for this species (despite the high dilution, water condensation could be possible in the short non heated part between the heated transfer line and the heated FTIR gas cell). The experimental trend for NO formation is well captured by the model. However, at 850 K, the model predicts mole fractions 28% larger than the experimental ones, while at a higher temperature (1100 K) the overprediction is reduced to 15%. It is interesting to see for CH_2O (Figure 3(e)) that the model captures the experimental trend very well over the whole temperature range. The proposed model overpredicts CH_2O by around 15% over the temperature range 975–1100 K. Interestingly, the model is able to capture the small decrease in CH_2O concentration around 875 K. For the mole fraction of CH_4 (Figure 3(f)), over the temperature range 700–925 K, the model shows a decreasing trend of the CH_4 mole fraction starting at 925 K whereas the experimental mole fractions reach a plateau. In Figure 4, it can be seen that the model overpredicts C_2H_6 by a factor of 3, and underpredicts CH_3OH and HCN by a factor of 3 and 8 respectively over the temperature range 1000–1100 K.

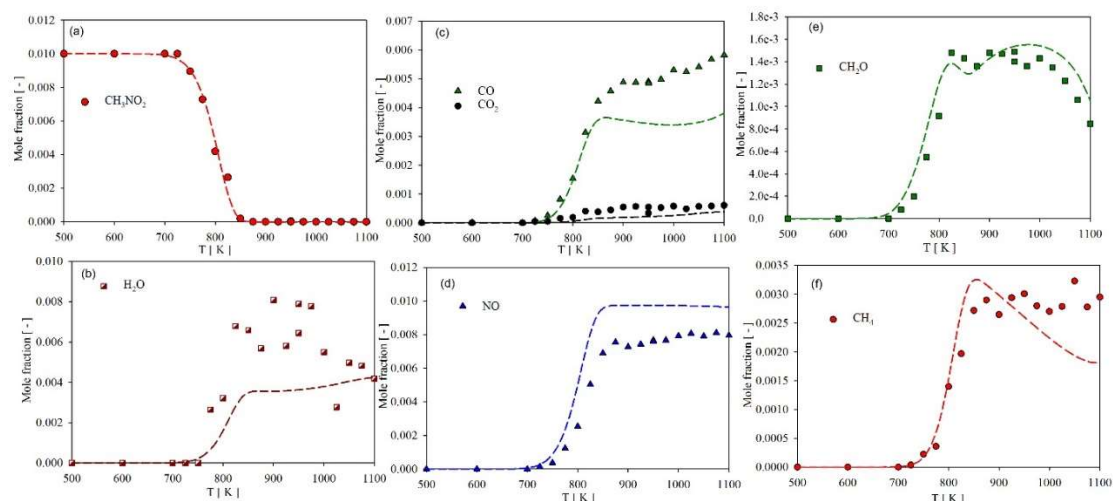


Figure 3: Pyrolysis of CH_3NO_2 1%/He 99% (on a mole basis) in a flow reactor at 1.07 bar, a residence time of 2.0 s in the isothermal zone of the reactor. Symbols: present work, dashed lines: data computed with the present model.

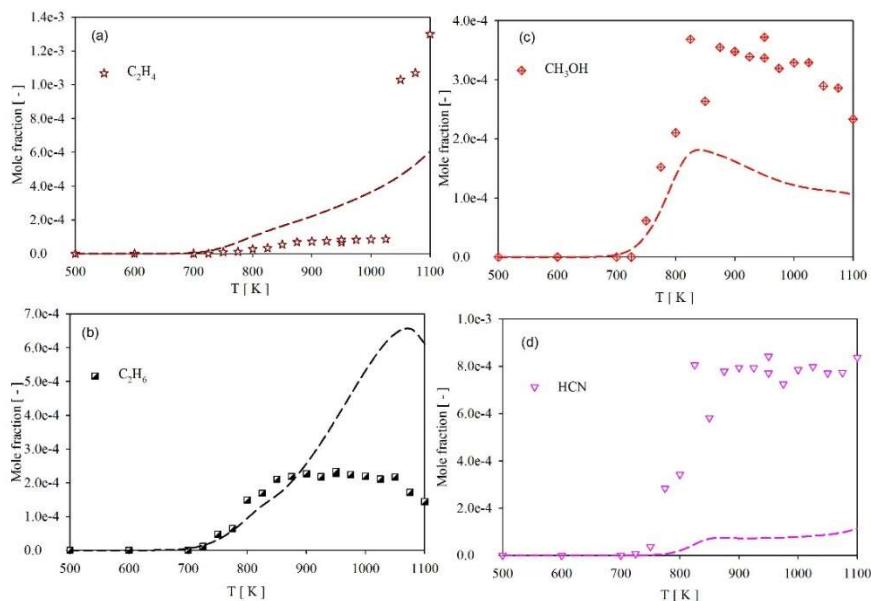


Figure 4: Pyrolysis of CH_3NO_2 (1%/He 99%, on a mole basis) in a flow reactor under the conditions of Figure 3. Symbols: present work, dashed lines: data computed with the present model.

Since the purpose of this work is to have a comprehensive model for both pyrolysis and oxidation, currently no attempt is made to optimize the model based on pyrolysis experimental data only. The model presented here performs much better under CH_3NO_2 oxidation conditions, which will be shown and discussed in Section 4.2 (Oxidation of nitromethane).

We now briefly discuss the reaction paths and the important reactions during CH_3NO_2 pyrolysis under the conditions of Figure 3 at 780 K (where 30% of CH_3NO_2 is consumed) with the help of the mass flux analysis shown in Figure 5 and the sensitivity analysis shown in Figure 6. In Figure 5, black solid lines represent the integrated mass flux based on C-atoms and blue solid lines represent the integrated mass flux based on N-atoms. Based on the nitrogen atom mass flux, CH_3NO_2 mainly decomposes ($\sim 62.5\%$) to CH_3 and NO_2 via a thermal dissociation channel ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$). Furthermore, $\sim 18.5\%$ of CH_3NO_2 undergoes H-atom abstraction reacting with H-atoms, OH, and CH_3 radicals forming CH_2NO_2 and $\sim 11\%$ react to HNO_2 which completely isomerizes to HONO . The $\sim 5\%$ CH_3NO_2 goes to CH_3NO , where the H-atom abstraction is the only channel. The remaining fraction of 3% goes to $\text{CH}_3\text{O} + \text{NO}$ via the another reaction channel $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$ (not shown in Figure 5). This observation is in line with the sensitivity analysis shown in Figure 6 where the reaction $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$ shows the highest sensitivity and consecutively the H-atom abstraction channel shows a much smaller sensitivity. Furthermore, NO_2 produced from thermal dissociation of CH_3NO_2 is mainly converted to NO , via reactions with H-atom, CH_3 , and HNO radical. The reaction of NO_2 with the CH_3 radical ($\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$) has the highest contribution for converting NO_2 to NO and this reaction is the second most sensitive reaction regarding the fuel decomposition (Figure 6). Then CH_3O , formed from the reaction $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$ and $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$, thermally dissociates to produce CH_2O and a H-atom ($\text{CH}_3\text{O}(+\text{M}) \rightleftharpoons \text{CH}_2\text{O} + \text{H}(+\text{M})$). The formed H-atom is mainly responsible for most of the H-atom abstraction from CH_3NO_2 . In line with this observation, the

reaction $\text{CH}_3\text{O}(+\text{M}) \rightleftharpoons \text{CH}_2\text{O} + \text{H}(+\text{M})$) is also one of the most sensitive reactions of the nitromethane decomposition, appearing at third position (Figure 6). It can be seen in Figure 5 that the formed CH_3NO , CH_2NO_2 , NO_2 , and HONO either thermally dissociate or react with radicals to give NO as the main product. This observation is in line with the NO concentration profile shown in Figure 3(d) where the NO concentration is highest at higher temperatures and CH_3NO_2 (Figure 3(a)) is completely consumed in the same temperature region. The sensitivity analysis (Figure 6) also reveals that most of the sensitive reactions under these conditions are related to nitrogen chemistry.

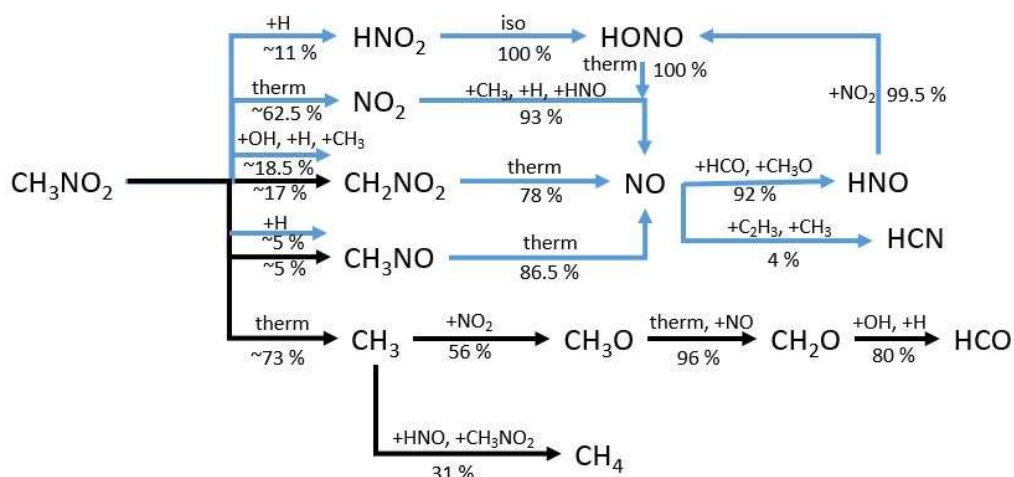


Figure 5: Reaction path analysis during $\text{CH}_3\text{NO}_2/\text{He}$ pyrolysis for the conditions shown in Figure 3 at 780 K. Blue lines (based on N-atom); black lines (based on C-atom). “therm” and “iso” in the figure mean thermal dissociation and isomerization, respectively.

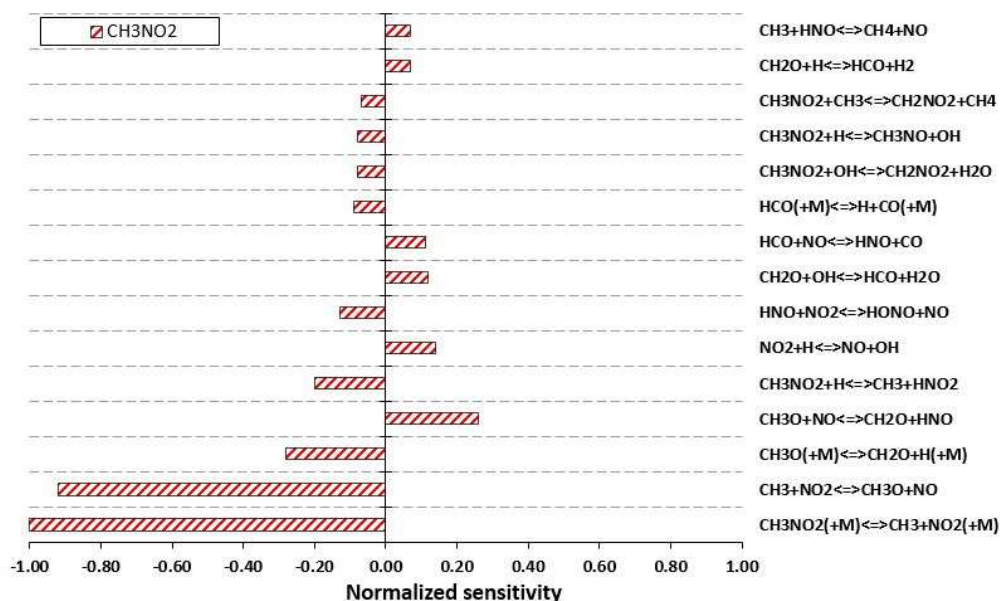


Figure 6: Sensitivity analysis during $\text{CH}_3\text{NO}_2/\text{He}$ pyrolysis for the conditions shown in Figure 3 at 780 K towards CH_3NO_2 .

To our best knowledge, there are no experimental data available in the literature for the oxidation of CH_3NO_2 in a flow reactor. Therefore, the model presented here is also validated against the published pyrolysis experimental data from the literature. Figure 7 shows the species profile for the pyrolysis of CH_3NO_2 (experimental conditions: 5% CH_3NO_2 /95% Ar at 5 Torr (0.666 kPa) studied by Weng et al. [36] (symbols) over the temperature range 735–1475 K using tunable synchrotron vacuum ultraviolet photoionization and molecular-beam mass spectrometry. It can be observed in Figure 7(a) that the onset temperature of fuel consumption is well predicted by the model, which starts around 1000 K. Further, the model is able to capture the experimental trend for the CH_3NO_2 decomposition over the whole temperature range. However, at higher temperatures ($T > 1200$ K), the model predicts the fuel decomposition onset at a temperature 50 K lower than observed in experiments. Moreover, for major species (H_2O , and NO), the model predictions against the experiment data are within reasonable uncertainty. At 1475 K, the model overpredicts the NO mole fraction by 20% and underpredicts the H_2O mole fraction by 50% respectively, while model significantly overpredicts H_2 mole fraction, which is almost factor 3 higher at 1375 K. For intermediate species like NO_2 , the model predictions agree well with the experiments. For CH_2O , the model overpredicts the experimental mole fractions by a factor of 4, which is not consistent when compared to the results from our present study (see Figure 3(e)), where the model slightly overpredicts CH_2O mole fractions at higher temperature, however still within the experimental uncertainty. It is worth noting that in both pyrolysis cases (Figure 3, Figure 7), CH_2O is formed via the same reaction paths which are the thermal dissociation reactions of CH_2NO_2 ($\text{CH}_2\text{NO}_2 \rightleftharpoons \text{CH}_2\text{O} + \text{NO}$) and CH_3O ($\text{CH}_3\text{O} + \text{M} \rightleftharpoons \text{CH}_2\text{O} + \text{H} + \text{M}$) (see Figures 5 and S3 in supplementary material). In fact, this over prediction of CH_2O can be explained by the significantly higher contribution of reaction channel $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$ forming CH_3O (see Figure S3 in supplementary material) at lower pressure which further leads to $\text{CH}_2\text{O} + \text{H}$ compared to case in Figure 3 (at 1.07 bar). Furthermore, in Figure 7(b), at 1475 K, the model overpredicts the CH_4 mole fraction by a factor of 2 but is in agreement for CH_3 at this temperature. The model predicts an earlier formation of CH_3 compared to experiments. The nitrogenated species HCNO and HNCO (Figure 7(d)) are significantly underpredicted by the model. The model suggests that HCNO is mainly formed via the reaction $\text{CH}_2 + \text{NO} \rightleftharpoons \text{HCNO} + \text{H}$ and that HNCO is formed via the two reactions $\text{CH}_2\text{NO} \rightleftharpoons \text{HNCO} + \text{H}$ and $\text{NCO} + \text{CH}_2\text{O} \rightleftharpoons \text{HNCO} + \text{HCO}$. At higher temperatures ($T > 1200$ K), although HCNO is within the experimental uncertainty, HNCO is still underpredicted by the model. Additional model validation for the pyrolysis of nitromethane is provided in the Supplementary material (see Figures S4–S6) where model predictions are good for major species but we see still have room for improvement for intermediate species predictions. Note that the use of tunable synchrotron vacuum ultraviolet photoionization and molecular-beam mass spectrometry allows Weng et al. [36] to detect species such as CH_3 , HCNO , and HNCO , which cannot be observed by gas chromatography in our experimental study.

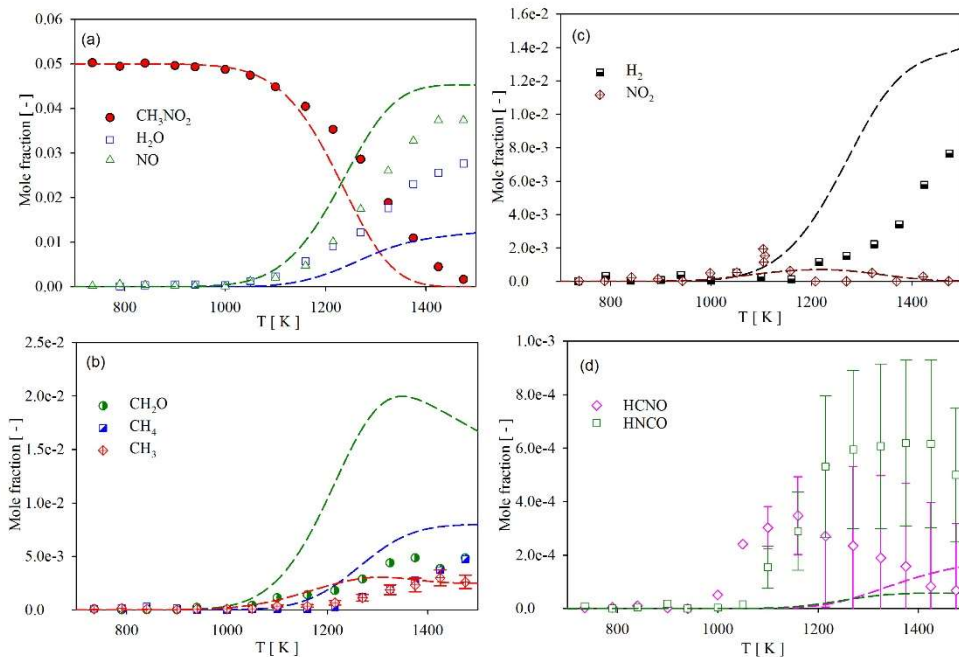


Figure 7: Pyrolysis of CH_3NO_2 5%/Ar 95% (on a mole basis) in a flow reactor at 5 Torr (0.666 kPa). Symbols: experimental data from Weng et al. [36]; dashed lines: this model. Error bars are only shown for CH_3 , HCNO, and HNCNO due to their much higher uncertainty compared to other species.

4.1.2. Speciation in a jet-stirred reactor

Figure 8 shows the species concentration vs temperature measured for the pyrolysis of CH_3NO_2 (experimental conditions: 1% CH_3NO_2 /99% He (on a mole basis) in a JSR at $\tau = 2.0$ s). Figure 8(a) shows that the fuel (CH_3NO_2) consumption is well captured by the model for the whole temperature range. The onset of fuel consumption is observed at a temperature of about 700 K, similarly to what was observed in the PFR (Figure 3(a)). However, in the JSR, CH_3NO_2 is almost completely consumed at around 925 K (Figure 8(a)), whereas CH_3NO_2 is completely consumed at 850 K in the PFR (Figure 3(a)) despite having similar initial conditions. This interesting observation is briefly discussed below. In both reactors (PFR and JSR), the proposed model is able to capture these trends very well. Similar observations as for the PFR case can be found for the comparison of model predictions of the intermediate species formation (CH_4 , C_2H_4 , C_2H_6 , and CH_3OH in Figure 8) and the experimental data: where CH_4 is overpredicted and C_2H_4 , C_2H_6 , and CH_3OH are underpredicted.

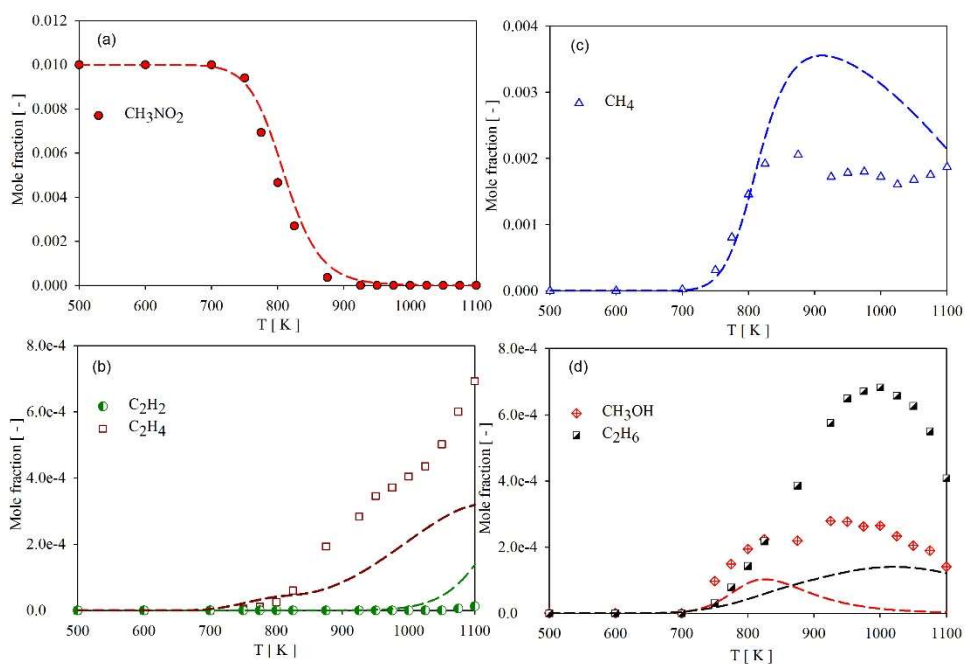


Figure 8: Pyrolysis of CH₃NO₂ 1%/He 99% (on a mole basis) in a JSR at 1.07 bar, $\tau = 2.0$ s. Symbols: present work; dashed lines: this model.

A mass flux analysis (based on N-atoms, see Supplementary material Figure S7) reveals that the major pathways for the formation and consumption of the fuel, as well as intermediate species, remain the same in both reactors. In the case of the JSR, the thermal dissociation channel ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$) is slightly stronger (by $\sim 9\%$) compared to the PFR case (Figure 5) and consecutively these 9% are compensated by the H-atom abstraction channel becoming less important (see Figures 5 and S7 in supplementary material). Furthermore, it can be observed in the sensitivity analysis (Figure 9) that in the case of the JSR the most sensitive reaction is also the thermal dissociation reaction ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$). It can also be observed in Figure 9 that the sensitivity of the reaction $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$ is slightly lower (compared to the PFR case, in Figure 5) while still being the second most sensitive reaction. However, the majority of sensitive reactions appearing for the JSR (Figure 9) and PFR (Figure 6) cases remains the same despite slight changes in their order. To explain the higher consumption of CH₃NO₂ in a flow reactor at lower temperature compared to JSR conditions we analyze the profiles of important species (see Figure S8 in supplementary material for a CH₃NO₂ concentration profile comparison). This early consumption is due to the fact that the H-atom abstraction channel forming HNO₂ ($\text{CH}_3\text{NO}_2 + \text{H} \rightleftharpoons \text{CH}_3 + \text{HNO}_2$) is more favored in a flow reactor compared to JSR conditions. HNO₂ readily isomerizes to HONO which is therefore formed earlier in case of the flow reactor (see Figure S8 in supplementary material). HONO decomposes to the two reactive radicals NO and OH. NO also reacts with the HCO radical to produce HNO + CO and this reaction ($\text{HCO} + \text{NO} \rightleftharpoons \text{HNO} + \text{CO}$) has a higher sensitivity under flow reactor conditions. The HNO formed in this process reacts with NO₂ mainly forming HONO (see Figure 5) leading to NO and OH formation via the same thermal decomposition path, adding more reactive radicals. Mole fraction profiles of NO, and NO₂ are also compared in Figure S8 in the supplementary material. In both the pyrolysis cases (JSR and PFR), the most sensitive reactions are part of the nitrogen chemistry. The

model suggests that H-atom abstraction reactions from CH_3NO_2 are contributing a lot to the pyrolysis chemistry of CH_3NO_2 under the conditions studied in this work.

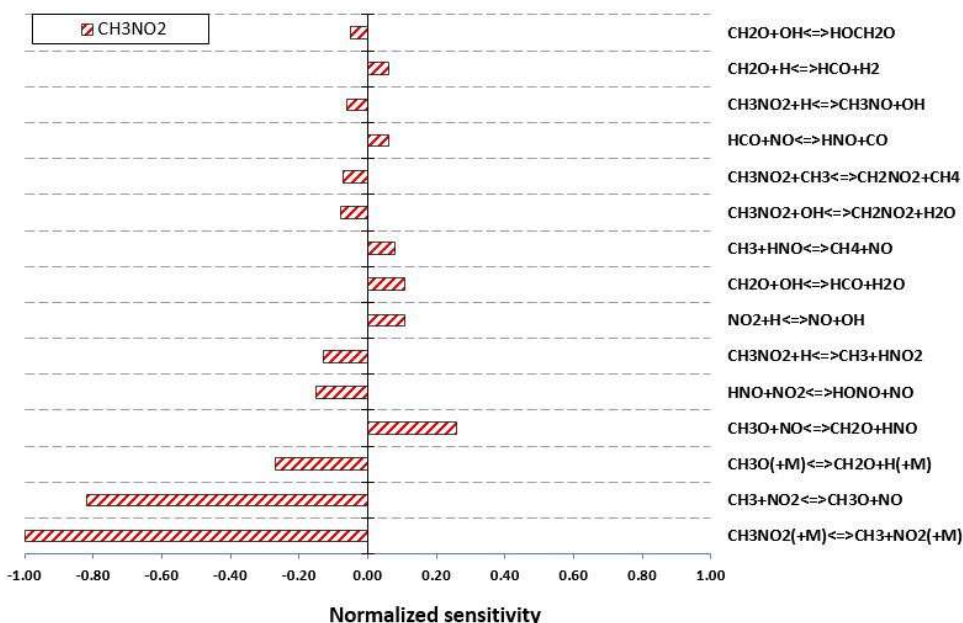


Figure 9: Sensitivity analysis during $\text{CH}_3\text{NO}_2/\text{He}$ pyrolysis for the condition shown in Figure 8 at 780 K towards CH_3NO_2 .

Figure 10 shows the fuel (CH_3NO_2) conversion during the pyrolysis of CH_3NO_2 (experimental conditions: 1% $\text{CH}_3\text{NO}_2/99\%$ He (on a mole basis) in the JSR at different residence times, at a pressure of 1.07 bar and a constant temperature of 775 K). Symbols represent experimental data and the dashed line is for data computed using the model built in this study. It can be observed (in Figure 10) at $\tau = 2.0$ s that the model predicts a 10% lower conversion of the fuel. As the residence time increases ($\tau > 2.0$ s), the difference between the measurements and model predictions gets smaller and at $\tau = 7.0$ s, the model accurately predicts the experimental CH_3NO_2 conversion.

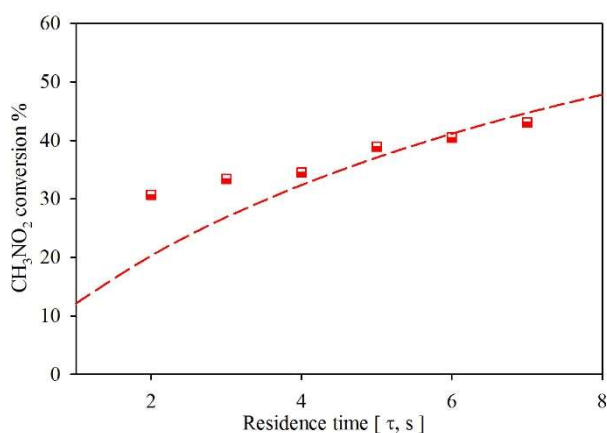


Figure 10: Conversion of CH_3NO_2 as a function of residence time during the pyrolysis of CH_3NO_2 1%/He 99% (on a mole basis), $P = 1.07$ bar, $T = 775$ K). Symbols: present work; dashed lines: present model.

4.2. Oxidation of nitromethane

4.2.1. Speciation in a jet-stirred reactor

Figure 11, Figure 12 show the evolution of species mole fractions with temperature during the oxidation of CH_3NO_2 (experimental conditions: see captions of Figure 11, Figure 12) in the JSR study by Weng et al. [36]. The experimental data (symbols) are compared against the model predictions (dashed lines). It can be observed in Figure 11(a) and (b) that the onset temperature of CH_3NO_2 and O_2 consumption is well predicted by the model. However, for temperatures ranging from 780 to 850 K, the model slightly underpredicts the CH_3NO_2 and O_2 concentrations but the values are still within the experimental uncertainty (10%). Further, it can be seen that for other major species (CO , CO_2 , NO , and H_2O , Figure 11) the model is in good agreement with the experimental measurements. For NO_2 , the model captures the experimental trend very well and agrees with the NO_2 peak concentration but it overpredicts the experimental values at most of the temperature points. Formation and consumption of nitrogenous intermediates species (HCN , HNO , and HONO) are well captured by the proposed model (Figure 12(a)). The model overpredicts CH_3OH mole fractions (see Figure 12(b)), which is significantly underpredicted by the model for the case of pyrolysis (see Figure 4, Figure 8). Furthermore, the formation of CH_2O (Figure 12(b)), which is the direct product of the thermal decomposition of CH_2NO_2 and CH_3O , is well captured by the model. Figure 12(c, d), shows that the model overpredicts the CH_4 mole fraction and underpredicts C_2H_4 and C_2H_6 mole fractions at a higher temperature. The fuel decomposition pathways and the important reactions will be discussed with the help of mass flow analysis and sensitivity analysis to see the difference between pyrolysis and oxidation.

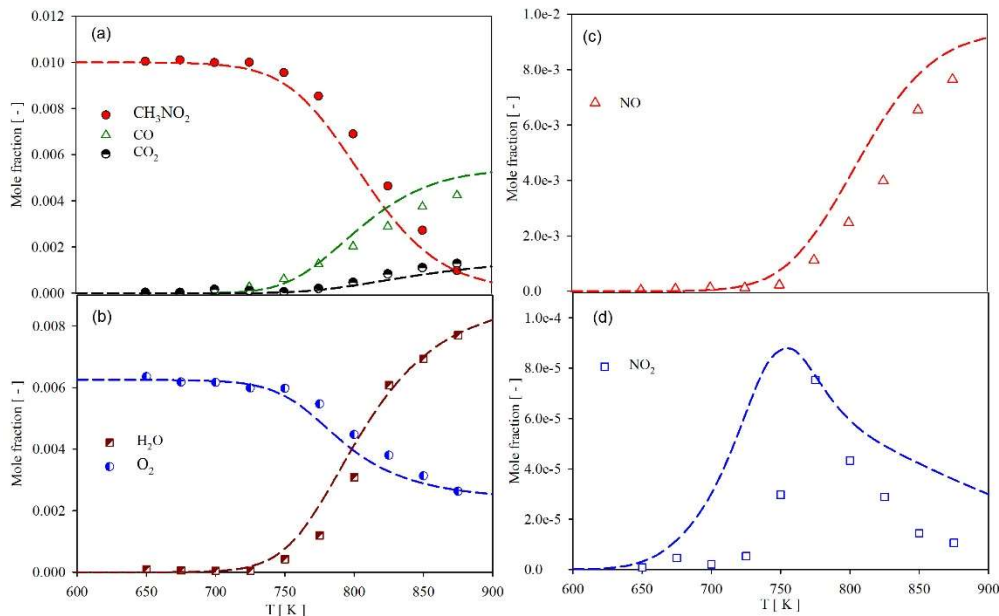


Figure 11: Oxidation of CH_3NO_2 (1% $\text{CH}_3\text{NO}_2/\text{O}_2$ 0.625%/Ar 98.375% at 1 atm, $\phi = 2.0$, $\tau = 2.0$ s) in a JSR. Symbols: experimental data from Weng et al. [36]; dashed lines: present model.

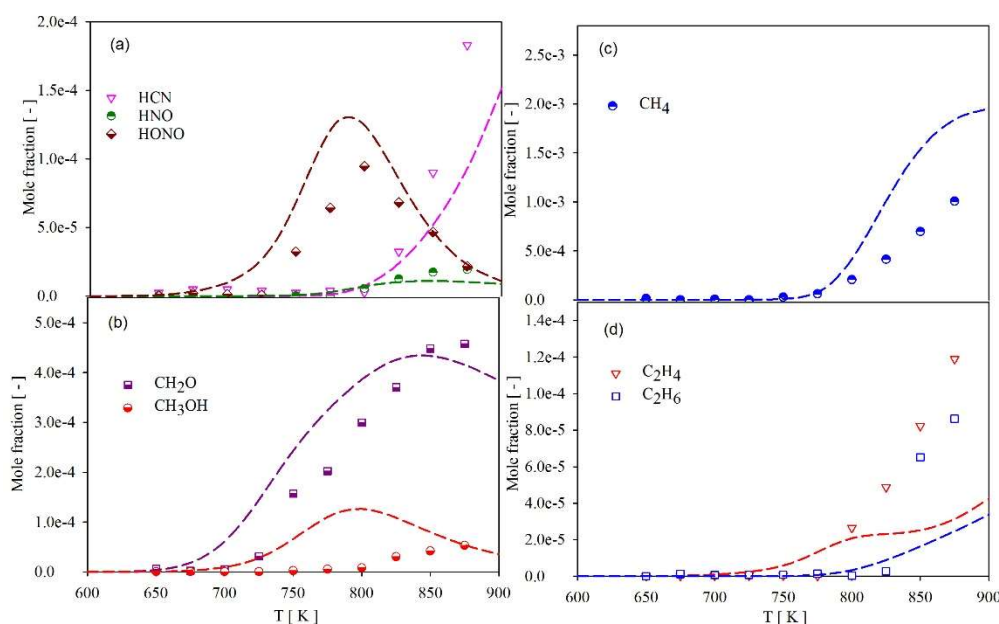


Figure 12: Oxidation of CH_3NO_2 (CH_3NO_2 1%/ O_2 0.625%/Ar 98.375%) for conditions of Figure 11 in a JSR. Symbols: experimental data from Weng et al. [36]; dashed lines: present model.

Figure 13 shows the integrated mass flux analysis based on nitrogen atoms and Figure 14 shows the normalized sensitivity for the 15 most sensitive reactions towards CH_3NO_2 during the $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ oxidation in the JSR for the conditions shown in Figure 11 at 780 K. It can be observed in Figure 13 that CH_3NO_2 is mainly decomposed via the thermal dissociation channel ($\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$). However, in the case of oxidation, thermal dissociation is weaker by 11% and the H-atom abstraction channel forming CH_2NO_2 is stronger by $\sim 15\%$ compared to pyrolysis conditions (see Figure S7 in supplementary material). The other minor channels forming HNO_2 and CH_3NO are less favored by 2% and 1% respectively compared to pyrolysis conditions (see Figure S7 in supplementary material). Furthermore, a small fraction $\sim 2.5\%$ of CH_3NO_2 decomposes to $\text{CH}_3\text{O} + \text{NO}$, similar to the pyrolysis case (Figure 8). However, the reaction path pattern for the formation and consumption of the major and intermediate species remains the same for both conditions. It can be observed in Figure 14 that the most sensitive reaction is the chain propagation channel $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$, followed by the thermal dissociation channel $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$ contrary to pyrolysis conditions (see Figure 9). During nitromethane oxidation, the third most sensitive reaction is $\text{CH}_3\text{NO}_2 + \text{OH} \rightleftharpoons \text{CH}_2\text{NO}_2 + \text{H}_2\text{O}$ (Figure 14), which is in the 10th rank under pyrolysis conditions (Figure 9). This observation is in line with the flow analysis (Figure 13) where this path ($\text{CH}_3\text{NO}_2 + \text{OH} \rightleftharpoons \text{CH}_2\text{NO}_2 + \text{H}_2\text{O}$) is stronger by a factor of 2 compared to pyrolysis conditions (see Figure S7 in supplementary material). The reaction $\text{CH}_3\text{O}(+\text{M}) \rightleftharpoons \text{CH}_2\text{O} + \text{H}(+\text{M})$ is also one of the most sensitive reactions since this reaction is mainly responsible for the formation of H-atoms, which react with NO_2 to give the OH radical ($\text{NO}_2 + \text{H} \rightleftharpoons \text{NO} + \text{OH}$). This reaction ($\text{NO}_2 + \text{H} \rightleftharpoons \text{NO} + \text{OH}$) contributes to 30% of OH formation, eventually influencing the radical pool in the system. The sensitivity analysis (Figure 14) shows that the OH radical takes part in most of the reactions and we, therefore, discuss other pathways forming OH: $\text{NO} + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{OH}$, which contributes by 30% to the OH formation and 27% contribution stems from the direct thermal

dissociation of the HONO radical ($\text{NO} + \text{OH} + \text{M} \rightleftharpoons \text{HONO} + \text{M}$). The thermal dissociation of HONO also gives NO, which mainly forms HNO in reactions with CH_3O , HCO , HONO , and HO_2 radicals. The formed HNO completely goes to HONO in the reaction with NO_2 ($\text{HNO} + \text{NO}_2 \rightleftharpoons \text{HONO} + \text{NO}$) and this reaction is the 8th most sensitive reaction (Figure 14). It can be observed in the reaction flow analysis (Figure 13) that there is a small cycle where NO is continuously recycled back ($\text{NO} \rightarrow \text{HNO} \rightarrow \text{HONO} \rightarrow \text{NO}$) in the process forming reactive OH radicals.

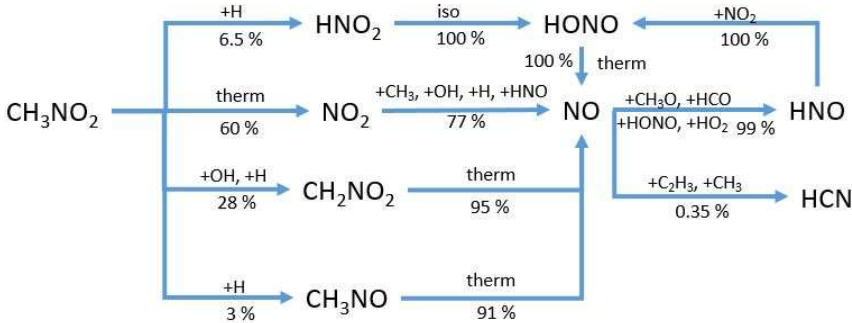


Figure 13: Major reaction paths during CH_3NO_2 1%/ O_2 0.625%/Ar 98.375% oxidation in a JSR for the conditions of Figure 11 at 780 K (based on N-atom). “therm” and “iso” in the figure mean thermal dissociation and isomerization, respectively.

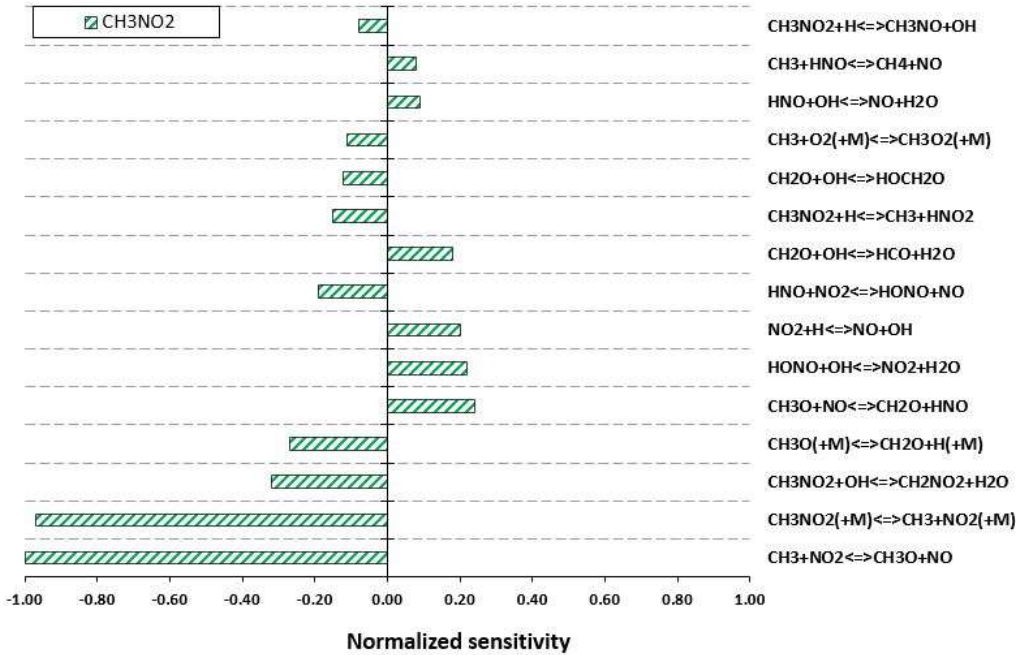


Figure 14: Sensitivity analysis during $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ oxidation in a JSR for the conditions of Figure 11 at 780 K towards CH_3NO_2 .

It is worth noting at this point that ignition delay times for $\text{CH}_4/\text{CH}_3\text{NO}_2$ blends were recently studied in a shock tube by Shang et al. [53] and they reported that CH_3NO_2 can dramatically reduce the ignition delay time of CH_4 , and the reduction effect increases with increasing blending ratio. This is due to the fact that with CH_3NO_2 in the system, the concentration of the radical pool

increases (see discussion in the section on reaction pathway analysis in Shang et al. [53]). This finding from Shang et al. [53] is in line with the work of Jia et al. [35] who also observed that the addition of nitromethane greatly increases the conversion rate of n-decane. NO has been long used as an ignition/reactivity promoter [54] for hydrocarbons but detailed studies about the enhancement effect of CH₃NO₂ are scarce. Additional model validation for nitromethane oxidation in a JSR can be found in the Supplementary material (Figures S9–S11).

4.2.2. Speciation in burner-stabilized flames

Burner stabilized flames provide important information regarding the flame structure, species formation and consumption in the reaction zone and subsequently about emission formation pathways. Therefore, we give special attention to the speciation predictions for burner-stabilized flames. The diffusion of small radicals is the determining process for the flame structure in premixed flames. The radical species O, H and OH are the most important species controlling the reactivity of the system during fuel oxidation.

For the burner-stabilized flame modeling, a mixture average transport model, thermal diffusion and radiative heat transfer were considered. The temperature profile is calculated solving the energy conservation equation, which considers the heat flux to the burner. We calculate the concentrations of all stable species on the burner surface from the gas composition in the feed, and the species diffusion close to the boundary. For all major radicals, we consider recombination at the burner surface. The undisturbed flame stabilizes above the burner through heat loss, which is calculated from the temperature gradient at the burner surface. The heat loss determines the maximum flame temperature, which is captured well.

Zhang et al. [39] experimentally studied a low-pressure (4.666 kPa) CH₃NO₂/O₂/Ar premixed burner stabilized flame at three different equivalence ratios (ϕ), namely 1.0, 1.5 and 2.0. Figure 15, Figure 16 show species concentrations vs height above the burner (HAB) in a CH₃NO₂/O₂/Ar flame at $\phi = 1.5$ [39]. Calculations were performed imposing the experimental temperature profile (dashed lines), and by solving the energy conservation equation (solid lines). Comparison between the measured and calculated temperature profiles is shown in the Supplementary material of this work (Figure S12). The difference in the calculated and experimental temperature profiles is observed mainly in the post-flame zone where it is governed by the thermochemistry and radiative heat loss. To reduce the over prediction of the calculated temperature in the post-flame zone, the radiation loss needs to be increased by a factor of 8 as shown in Figure S12. This was done because we cannot quantify the heat loss in post flame zone from other possible sources (e.g cooling due to thermocouple) and we use an increased radiative heat loss to account for such effects. It can be observed in Figure 15, Figure 16 that the model prediction follows the experimental trend. However, the prediction solving the energy conservation equations shifts the species profile towards the burner surface by 3–4 mm. This shift is due to the influence of the probe which causes disturbance of the flame [55]. In Figure 15(a), imposing the experimental temperature profile, experimental and modeling results for major species are in good agreement. However, for HAB around 7 mm, a peak can be seen in the CO profile for the experiment, which is

in contradiction with the model prediction, where we do not observe such steep peaks. The model underpredicts the experimental values by 15% in peak position. Such a peak is also observed for NO at same HAB (=7 mm) and simulation does not show such a steep peak, both for imposing the temperature profile as well as for solving the energy conservation equation. Further, the mole fractions of the intermediate species NO_2 is underpredicted by a factor 3 and C_2H_6 are overpredicted by a factor of 2 in the simulation, whereas other intermediate species are over- or underpredicted by a factor of 2. It should be noted that the authors of [39] mentioned that uncertainty of a factor of 2 is expected for intermediates species. The uncertainty of estimated photoionization cross-sections is even larger for radicals due to their quenching effect on the probe surface. Considering this, the model predictions show a reasonable agreement within the experimental uncertainty. In the following, based on the N- and C-atom mass flow analysis, the main degradation pathways of nitromethane for the flame shown in Figure 15 are discussed. Figure 17 shows the integrated mass flux imposing the experimental temperature profile. Special attention is paid to the species formation and consumption pathways, where nitrogen chemistry is involved because of their complex kinetics. The mass flow analysis in Figure 17(a) indicates that:

- 50% of CH_3NO_2 is dissociated to CH_3 and NO_2 via the reaction $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$.
- 27.5% of CH_3NO_2 reacts to CH_2NO_2 via H-atom abstraction by H-atoms, OH and CH_3 radicals yielding to CH_2NO_2 .
- 18% of CH_3NO_2 goes to CH_3O and NO via the reaction $\text{CH}_3\text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$
- 1.3% of CH_3NO_2 reacts with H atoms forming CH_3NO .
- 3.2% of CH_3NO_2 reacts with H-atoms to HNO_2 (not shown in Figure 3(a)) via the route $\text{CH}_3\text{NO}_2 + \text{H} \rightleftharpoons \text{HNO}_2 + \text{CH}_3$.

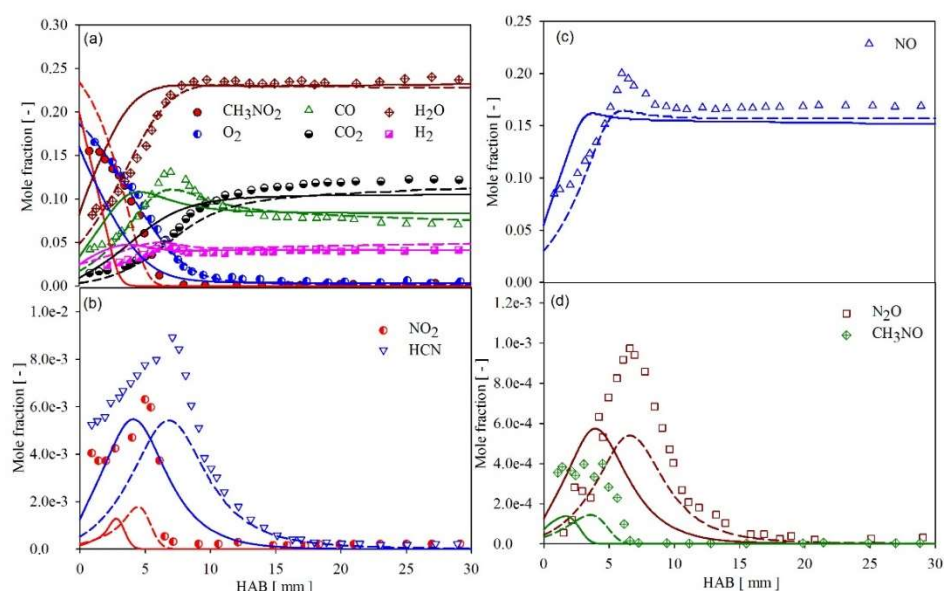


Figure 15: Speciation of premixed $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ burner stabilized flame at 4.666 kPa, $T = 298 \text{ K}$, $\phi = 1.5$. Symbols: experiments from Zhang et al. [39]; dashed lines: model predictions imposing the experimental temperature profile; solid lines: model predictions solving the energy conservation equation.

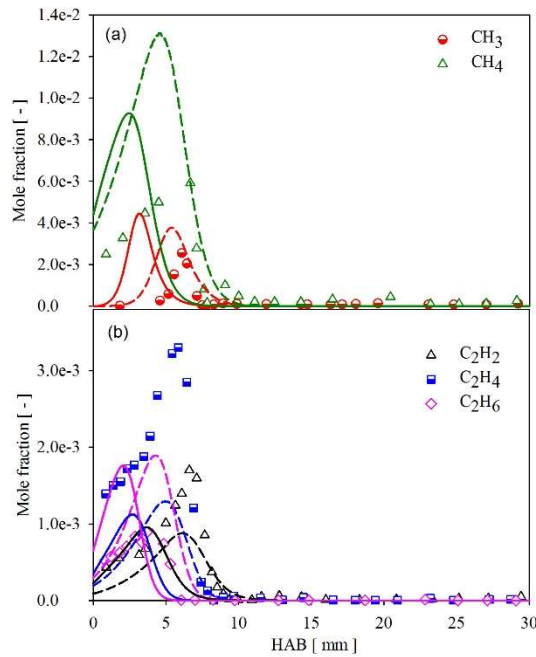


Figure 16: Speciation of premixed $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ burner stabilized flame shown in Figure 15. Symbols: experiments from Zhang et al. [39]; dashed lines: model predictions imposing the experimental temperature profile; solid lines: model predictions solving the energy conservation equation.

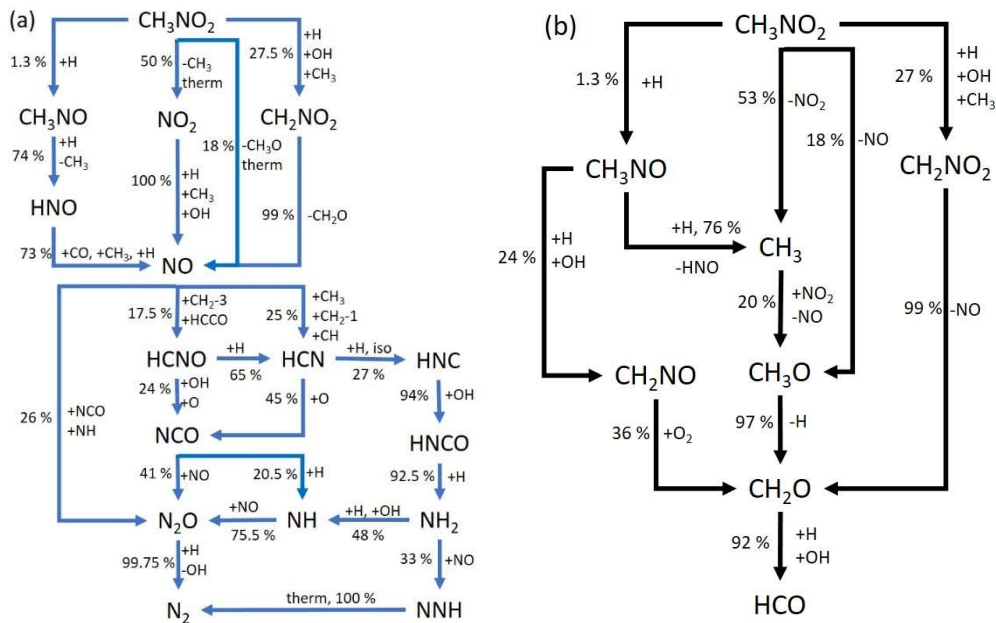


Figure 17: Integrated mass flux analysis for the flame shown in Figure 15. Based on the N-atoms (a), based on C-atoms (b). “therm” and “iso” in the figure mean thermal dissociation and isomerization, respectively.

The formed NO_2 completely reacts to NO via the following reactions: $\text{NO}_2 + \text{H} \rightleftharpoons \text{NO} + \text{OH}$, $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$, and $\text{NO} + \text{HO}_2 \rightleftharpoons \text{NO}_2 + \text{OH}$. $\text{NO}_2 + \text{H} \rightleftharpoons \text{NO} + \text{OH}$ is the major contributor to

NO₂ loss. All most all of the CH₂NO₂ (99%) is converted into NO and CH₂O via CH₂NO₂ ⇌ CH₂O + NO.

CH₃NO indirectly forms NO via HNO and methyl radicals in the reaction CH₃NO + H ⇌ CH₃ + HNO. Then, HNO mainly reacts with CO (HNO + CO ⇌ HCO + NO), with CH₃ (CH₃ + HNO ⇌ CH₄ + NO) and with H (HNO + H ⇌ NO + H₂) contributing to the formation of NO. Figure 15(c) shows that the NO mole fraction is as high as that of H₂O (Figure 15(a)) and is, therefore, one of the major species during CH₃NO₂ oxidation.

Next, NO decomposes through different pathways: 26% of NO is consumed by the reactions NCO + NO ⇌ N₂O + CO and NH + NO ⇌ N₂O + H. 25% of NO is consumed by the following reactions: CH₃ + NO ⇌ HCN + H₂O, CH₂-1 + NO ⇌ HCN + OH, and CH + NO ⇌ HCN + O, other minor reactions contributing to the formation of HCN are CH₂-3 + NO ⇌ HCN + OH, C₂H₃ + NO ⇌ HCN + CH₂O, C₂H + NO ⇌ HCN + CO and HCCO + NO ⇌ HCN + CO₂ (not shown here). Further, about 17.5% of NO reacts with CH₂-3 and HCCO leading to HCNO mainly via the reactions CH₂-3 + NO ⇌ HCNO + H and HCCO + NO ⇌ HCNO + CO. Around 65% of HCNO reacts with H to form HCN via the reaction HCNO + H ⇌ HCN + OH and 28% of HCNO forms NCO via these 3 channels: HCNO + OH ⇌ NCO + H + OH, HCNO + OH ⇌ NCO + H₂O, and HCNO + O ⇌ NCO + OH. HCN can then react by isomerization to HNC (27%) via HCN(+M) ⇌ HNC(+M) and HNC + H ⇌ HCN + H. HCN can also react with O-atoms (45%) producing NCO (HCN + O ⇌ NCO + H). HNC formed via the HCN route almost completely reacts with OH forming HNCO (HNC + OH ⇌ HNCO + H) and 92.5% of HNCO reacts with H-atoms yielding NH₂ via the route HNCO + H ⇌ NH₂ + CO. The last 7.5% react to NCO via minor routes (not shown here). NH₂ mainly goes to NH and NNH. Indeed, 48% of NH₂ is consumed to form NH via the reactions NH₂ + H ⇌ NH + H₂ and NH₂ + OH ⇌ NH + H₂O and 33% of NH₂ reacts to NNH via the reaction NH₂ + NO ⇌ NNH + OH. Then, NH reacts with NO (NH + NO ⇌ N₂O + H) producing N₂O. Almost all of the N₂O produced during the above process is converted to N₂ via the reaction N₂O + H ⇌ N₂ + OH and also all of the NNH reacts to N₂ via thermal dissociation path (NNH ⇌ N₂ + H).

Overall, the nitrogen atom mass flux analysis in Figure 17(a) can be seen as a two-step process:

- In the first step, almost all of the primary products lead to NO.
- The second step is directed to the decomposition of intermediate species mainly leading to N₂O which finally goes to N₂.

The main pathways for the formation and consumption of species for the three different equivalence ratios ($\phi = 1.0, 1.5$ and 2.0) remain the same. Mass flux analyses for $\phi = 2.0$ and 1.0 are shown in the Supplementary material (Figures S14-S15 and S17-S18). The main difference is observed for the NH₂ → NH/NNH route for the fuel-rich flame ($\phi = 2.0$), here 41% of NH₂ is consumed to form NH and 43% is decomposed to form NNH (see Figure S15 in supplement material) which is in contradiction with the flames at $\phi = 1.5$ and 1.0 (see Figures 17(a) and S18). This leads to the conclusion that at higher equivalence ratios, the NH₂ → NNH route gets stronger compared to the NH₂ → NH route. Additional validation of the burner stabilized premixed flame is provided in the Supplementary material (see Figures S13 and S16).

4.2.3. Ignition delay time and laminar flame speed

Finally, we demonstrate the capability of our model to predict essential features of nitromethane combustion over a wide range of conditions: the ignition timing, and the laminar flame speeds. Figure 18 compares the predicted ignition delay times against shock tube experimental data from the literature. Figure 18(a) shows results of the shock tube experiments by Mathieu et al. [29] at three different equivalence ratios ($\phi = 0.5, 1.0$ and 2.0) at 2 atm for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ mixtures. The model captures the equivalence ratio dependency of ignition delay times very well and agrees with the experimental data for the temperature range studied. Similarly, Figure 18(b) shows ignition delay times measured by Nauc ler et al. [27] in a shock tube compared to simulations at different equivalence ratios with changing pressure for $\text{CH}_3\text{NO}_2/\text{O}_2/\text{N}_2$ mixtures. We can observe that the model is in excellent agreement with the experimental data. It shows the model capability to predict ignition delay timing with changing mixture composition and pressure.

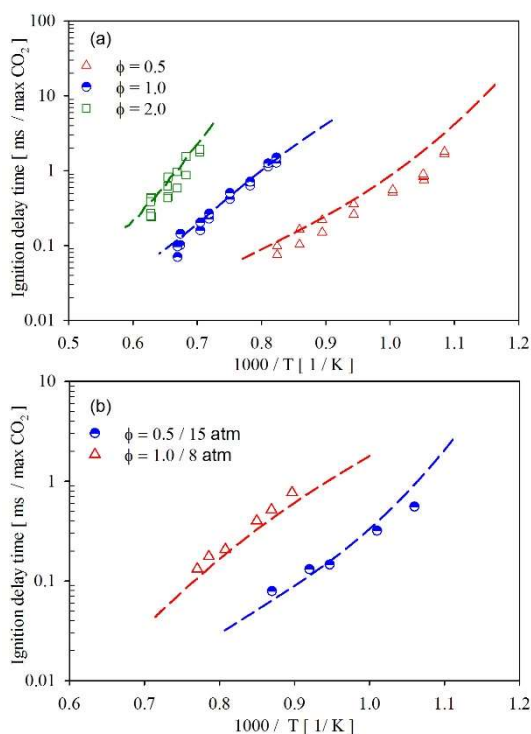


Figure 18: Ignition delay times of $\text{CH}_3\text{NO}_2/\text{O}_2/\text{Ar}$ (a) and $\text{CH}_3\text{NO}_2/\text{O}_2/\text{N}_2$ (b) at different ϕ and pressures. $\phi = 0.5$ (CH_3NO_2 4%/ O_2 6%), $\phi = 1.0$ (CH_3NO_2 2.857%/ O_2 2.143%), $\phi = 2.0$ (CH_3NO_2 1.455%/ O_2 0.545%) (a); $\phi = 0.5$ (CH_3NO_2 4%/ O_2 10%), $\phi = 1.0$ (CH_3NO_2 4%/ O_2 5%) (b). Balance Ar and N_2 for above mixtures. Symbols: experiments from [29] (a) and [27] (b); lines: simulations of this model.

Further ignition delay time sensitivity analysis (towards temperature) is performed and the 20 most sensitive reactions at $\phi = 0.5$ and $T = 1100$ K for two pressures, 2 and 15 atm, are shown in Figure 19. A positive sensitivity means that the reaction promotes the reactivity and negative sensitivity means that the reaction inhibits the reactivity. It can be seen for both pressures that

the most sensitive reactions are $\text{HCO} + \text{NO} \rightleftharpoons \text{HNO} + \text{CO}$ and $\text{O}_2 + \text{H} \rightleftharpoons \text{OH} + \text{O}$ which is in line with the results from the work of Mathieu et al. [29]. Unlike for the PFR and JSR experiments, the thermal dissociation reaction of CH_3NO_2 is not among the most sensitive reactions. However, the reaction still appears in sensitivity analyses and its sensitivity decreases at 15 atm compared to 2 atm. The recombination reactions of NO with H and O-atoms forming HNO and NO_2 respectively have much higher sensitivity at 15 atm than at 2 atm. It can be seen in Figure 19 that the sensitivity of most of the reactions increases at higher pressure, while only a few reactions ($\text{HCO} + \text{O}_2 \rightleftharpoons \text{CO} + \text{HO}_2$, $\text{HCO} + \text{NO} \rightleftharpoons \text{HNO} + \text{CO}$, and $\text{CH}_2\text{O} + \text{OH} \rightleftharpoons \text{HCO} + \text{H}_2\text{O}$) show a decreased sensitivity. It is also interesting to see (in Figure 19) that one of the reaction ($\text{CH}_4 + \text{OH} \rightleftharpoons \text{CH}_3 + \text{H}_2\text{O}$) alter the sensitivity direction at higher pressure. Overall, it can be seen that most of the sensitive reactions appear within the nitrogen chemistry. Additional model validation for the ignition delay times of nitromethane as well as speciation in shock tube is shown in the Supplementary material (Figures S19–S21).

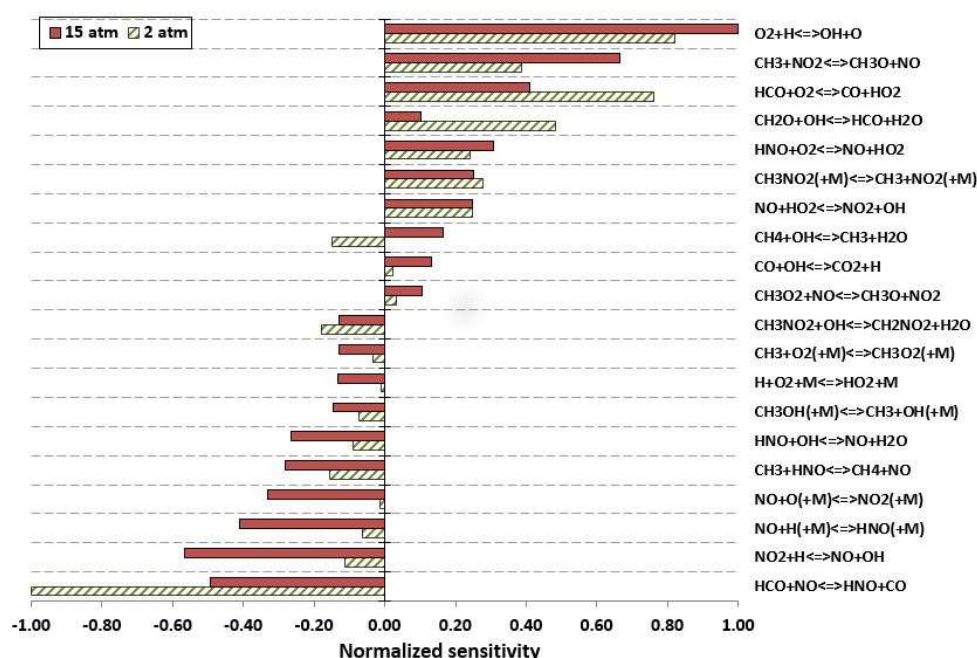


Figure 19: Ignition delay time sensitivity analysis at $\phi = 0.5$ and $T = 1100$ K. For 2 atm (mixture condition from [29]), for 15 atm (mixture condition from [27]) see Figure 18.

In Figure 20, laminar flame speeds of a CH_3NO_2 /air flame at 423 K are compared against the model predictions for initial pressures of 1, 2, and 3 bar. The measurements (symbols) were performed by Brequigny et al. [28] for propagating spherical flames. To determine the equivalence ratio, they used the equation $\text{CH}_3\text{NO}_2 + 0.75(\text{O}_2 + 3.76 \text{N}_2) \rightleftharpoons \text{CO}_2 + 1.5 \text{H}_2\text{O} + 3.32 \text{N}_2$ (Equation (1)). The maximum flame speed was observed (experimentally) for lean mixtures around $\phi = 0.7$ – 0.8 (see Figure S22 in supplement material). From the studies listed in Table 1, [28] is the only one to use Equation (1) for determining equivalence ratios. All other studies listed in Table 1 used (and/or suggested) $\text{CH}_3\text{NO}_2 + 1.25 \text{O}_2 \rightleftharpoons \text{CO}_2 + 1.5 \text{H}_2\text{O} + \text{NO}$ (Eq. (2)) as the correct equation for determining equivalence ratios where NO is a major product instead of N_2 . It is seen in the present work as well as in other studies discussed above, at different conditions and for different reactors,

that NO is a more stable product and found in higher concentrations. Nevertheless, the equivalence ratios according to Equations (1) and (2) have a linear relationship of ϕ Equation (1) = $0.6 \times \phi$ Equation (2) and thereby they can be directly converted into each other which is also suggested by Faghhi and Chen [56]. For the laminar flame speed shown in Figure 20, we recalculated the equivalence ratio of Brequigny et al. [28] using Equation (2). We can see (Figure 20) that the maximum laminar flame speed is shifted towards the rich side (to around $\phi = 1.25$) which is generally observed in hydrocarbon laminar flame speeds. The model derived in this work is able to capture the experimental trends with regard to the variation of equivalence ratio and pressure. However, the model overpredicts the laminar flame speed for very rich conditions ($\phi > 1.8$). Further, the laminar flame speeds at 1 atm and at elevated temperature (338–358 K) were measured by Nauc ler et al. [37] (see Figure S23 in supplementary material). Here, in contrast to the observation in Figure 20, the model underpredicts the laminar flame speed by 4 cm/s at $\phi = 1.2$ but this discrepancy is smaller for the lean and rich sides. Nevertheless, we can state that the model is able to predict the temperature dependency of the laminar flame speed for the range of equivalence ratios studied here (see Figure S23 in supplementary material). At present, we cannot say whether the remaining discrepancy is from the experiments or from the kinetic model. In general, the model is in good agreement with the measurements from Brequigny et al. [28] but underpredicts the measurements of Nauc ler et al. [37].

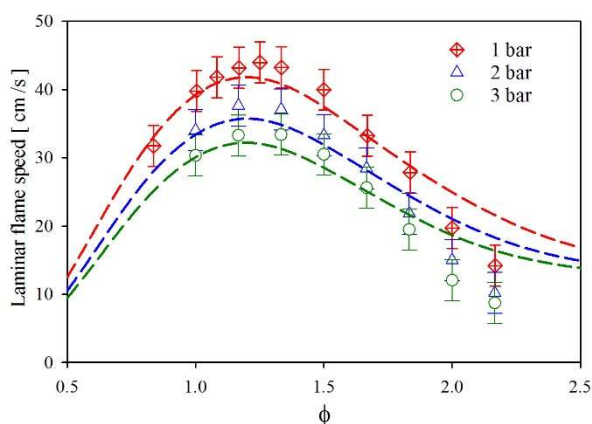


Figure 20: Laminar flame speed of $\text{CH}_3\text{NO}_2/\text{air}$ at 423 K and three different pressures 1, 2 and 3 bar. Symbols: measurements from [28]; lines: this model.

5. Effect of reaction rate uncertainty on model predictions

In the present work, we observed that the model is unable to predict several intermediate species profiles under pyrolysis conditions for both flow reactor and JSR experiments despite predicting the fuel decomposition very well (Figure 3, Figure 4, Figure 7, Figure 8). We have performed an uncertainty analysis to investigate if a better choice of the kinetic parameters of the elementary reactions in the C/H/O/N chemistry can improve the model performance. Before discussing the results of this uncertainty analysis, we present a brief overview of the model and discuss the present choice of rate parameters. In our previous work [32], we developed the H/O/N chemistry and the kinetic model was extensively validated against a large set of published experimental data

set and the choice of rate parameters for each reaction involved in this (H/O/N) system has been explained there. Furthermore, the H/O/N chemistry was extended in our recent study [31] to include C₁-C₂ kinetics and cross-reactions between C and N chemistry. It was demonstrated in [31] that the model performed very well for a broad range of experimental conditions (speciation in a flow reactor, JSR and BSF as well as for laminar flame speed and ignition delay times) and was able of describing the C₁-C₂/NO_x interaction reasonably well, even for very rich mixtures. For detailed information, the reader can refer to two recent studies [31,32]. The model development approach in all of our previous work is hierarchically in nature as it is in this work. It is well known that the nitrogen chemistry is more complex than that of simple hydrocarbons, in which reaction rate determination often based on analogy works quite well. Since experimental or high-level theoretical studies are unfortunately scarce, most of the reaction rates used for the elementary reactions in the CH₃NO₂ sub-set are derived from analogy or from the QRRK estimation, and are definitely prone to high uncertainty. The available literature is focused on the unimolecular dissociation of CH₃NO₂ as discussed above in Section 3.1. To investigate the discrepancy between model predictions and experimental data, we choose the pyrolysis case shown in Figure 3, Figure 4. A sensitivity analysis was performed towards HCN and C₂H₆ and the sensitive reactions from the C/H/O/N chemistry with regard to these targets are listed in Table 4 and the order of sensitivity is shown in Figure 21. HCN and C₂H₆ have been chosen because these species were significantly under- and overpredicted. Table 4 also provides the rate parameters for the chosen sensitive reactions and their rate constant uncertainty where available in literature. Where not available, we use 50% which is reasonable compared to other values listed in Table 4. First, the reaction rate parameters of sensitive reactions (listed in Table 4) are modified (within the uncertainty) to improve HCN, shown by the solid lines in Figure 22, Figure 23. It can be observed that the HCN concentration increases as expected (moving closer to the experimental data), but the maximum concentration is still significantly underpredicted. However, these changes have a negative impact on the C₂H₆ concentration profile (moving away from the experimental data) but has positive impact on CH₃OH (moving closer to the experimental data). Similar observations can be seen for other species. In the second case, the reaction rate parameters of sensitive reactions (listed in Table 4) are modified (within the uncertainty) in order to improve C₂H₆ shown by the dotted lines in Figure 22, Figure 23. As expected, the C₂H₆ concentration profile decreases (moving closer to the experimental data) but similar to the previous case, there are other species, for which the agreement worsens. From this analysis, we conclude that with the present understanding on nitromethane degradation kinetics we are not able to reasonably predict all the intermediate species for pyrolysis conditions and this problem should be addressed by the combustion community in future work. Further studies on elementary reactions along the nitromethane degradation pathways are highly desirable.

Table 4: Selected sensitive reactions for the pyrolysis experiment shown in Figure 22, Figure 23. Est.: estimated value, Exp.: experimentally determined, Theory: theoretically calculated, N/A: not available. Units: cm³-mol-s-cal.

| Reactions | A | n | Ea | Method | Uncertainty | | Ref. |
|--|----------|------|----------|--------|-------------|--|---------|
| | | | | | % (±) | | |
| CH ₂ NO ₂ +CH ₃ ⇌C ₂ H ₅ +NO ₂ | 5.00E+13 | 0 | 0.00E+00 | Est. | N/A | | [39] |
| CH ₂ NO ₂ ⇌CH ₂ O+NO | 1.00E+13 | 0 | 3.60E+04 | Exp. | N/A | | [57] |
| CH ₃ NO ₂ (+M) ⇌CH ₃ +NO ₂ (+M) | 1.80E+16 | 0 | 5.85E+04 | Exp. | 33 | | [22,26] |
| Low pressure rate | 1.26E+17 | 0 | 4.20E+04 | Exp. | 25 | | |
| CH ₃ O+NO⇌CH ₂ O+HNO | 1.30E+14 | -0.7 | 0.00E+00 | Eva. | N/A | | [58] |
| CH ₃ +HNO⇌CH ₄ +NO | 1.50E+11 | 0.76 | 3.48E+02 | Theory | N/A | | [59] |
| CH ₃ +NO⇌H ₂ CN+OH | 1.50E-01 | 3.52 | 3.95E+03 | Theory | N/A | | [60] |
| CH ₃ +NO⇌HCN+H ₂ O | 1.50E-01 | 3.52 | 3.95E+03 | Theory | N/A | | [60] |
| CH ₃ NO ₂ +H⇌CH ₂ NO ₂ +H ₂ | 5.40E+02 | 3.5 | 5.20E+03 | Est. | N/A | | [26] |
| CH ₃ NO ₂ +H⇌CH ₃ +HNO ₂ | 3.30E+12 | 0 | 3.73E+03 | Exp. | 25 | | [61] |
| C ₂ H ₃ +NO⇌HCN+CH ₂ O | 7.00E+21 | -3.4 | 1.03E+03 | Theory | N/A | | [62] |
| C ₂ H ₃ +NO ₂ ⇌CH ₂ CHO+NO | 7.70E+14 | -0.6 | 0.00E+00 | Exp. | 25 | | [63] |
| CH ₃ NO ₂ +H⇌CH ₃ NO+OH | 1.40E+12 | 0 | 3.73E+03 | Exp. | 25 | | [61] |
| CH ₃ NO ₂ +OH⇌CH ₂ NO ₂ +H ₂ O | 5.00E+05 | 2 | 1.00E+03 | Est. | N/A | | [26] |
| CH ₃ +NO ₂ ⇌CH ₃ O+NO | 4.00E+13 | -0.2 | 0.00E+00 | Theory | 30 | | [26] |
| CH ₃ NO ₂ +OH⇌CH ₃ OH+NO ₂ | 2.00E+10 | 0 | 1.00E+03 | Est. | N/A | | [26] |
| CH ₃ NO ₂ +C ₂ H ₅ ⇌CH ₂ NO ₂ +C ₂ H ₆ | 3.00E+11 | 0 | 1.17E+04 | Est. | N/A | | [29] |
| CH ₃ NO ₂ +CH ₃ ⇌CH ₂ NO ₂ +CH ₄ | 5.50E-01 | 4 | 8.30E+03 | Est. | N/A | | [26] |

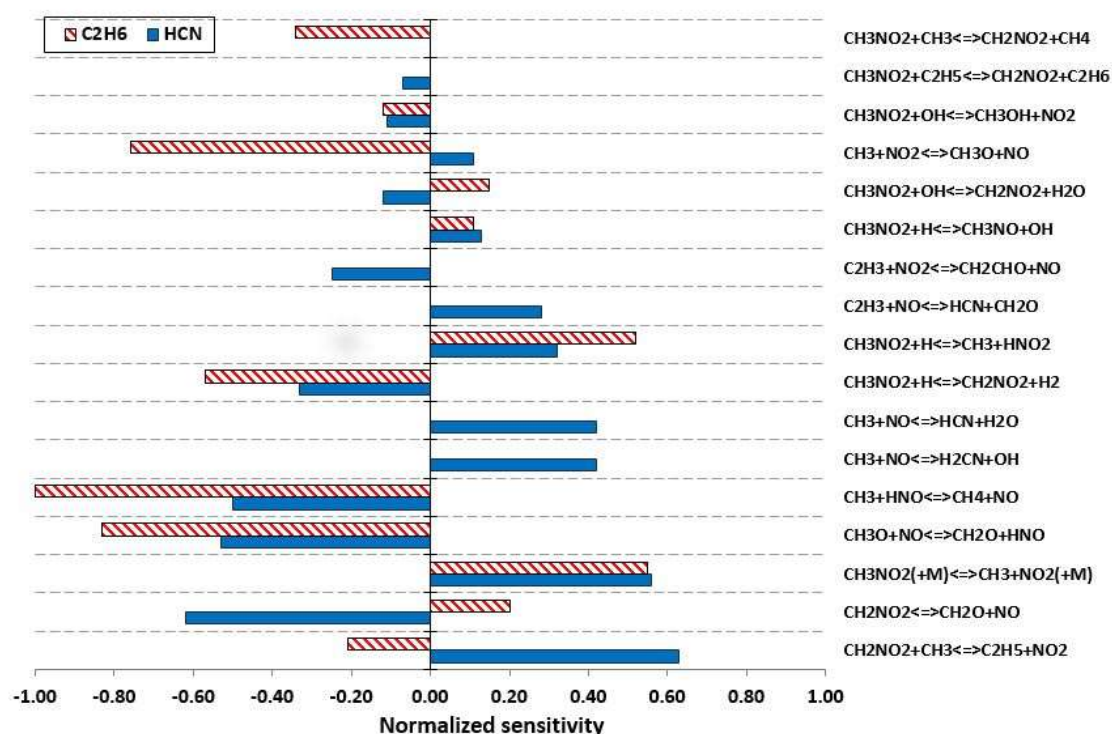


Figure 21: Sensitivity analysis for pyrolysis condition shown in Figure 22 at 780 K, towards HCN (blue) and towards C₂H₆ (red). (For interpretation of the references to colour in this figure legend, the reader is referred to the web version of this article.)

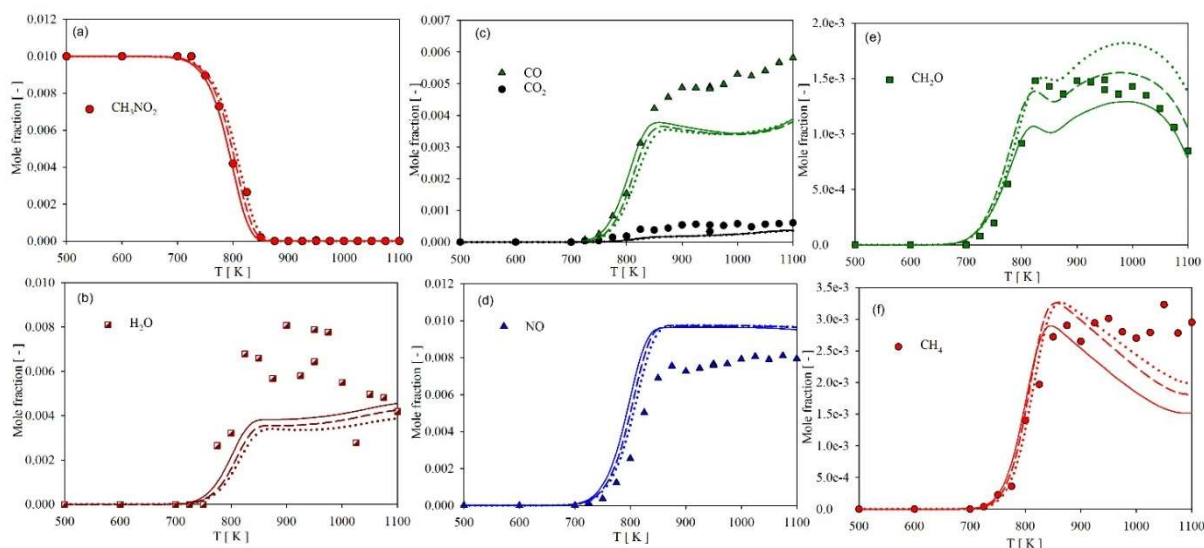


Figure 22: Pyrolysis of CH_3NO_2 1%/He 99% (on a mole basis) in a flow reactor at 1.07 bar, a residence time of 2.0 s in the isothermal zone of the reactor. Symbols: present work; dashed lines: prediction with the proposed model, solid lines: model predictions with modified rate constant for reactions listed in Table 4 (targeting HCN), dotted lines: model predictions with modified rate constant for reactions listed in Table 4 (targeting C_2H_6).

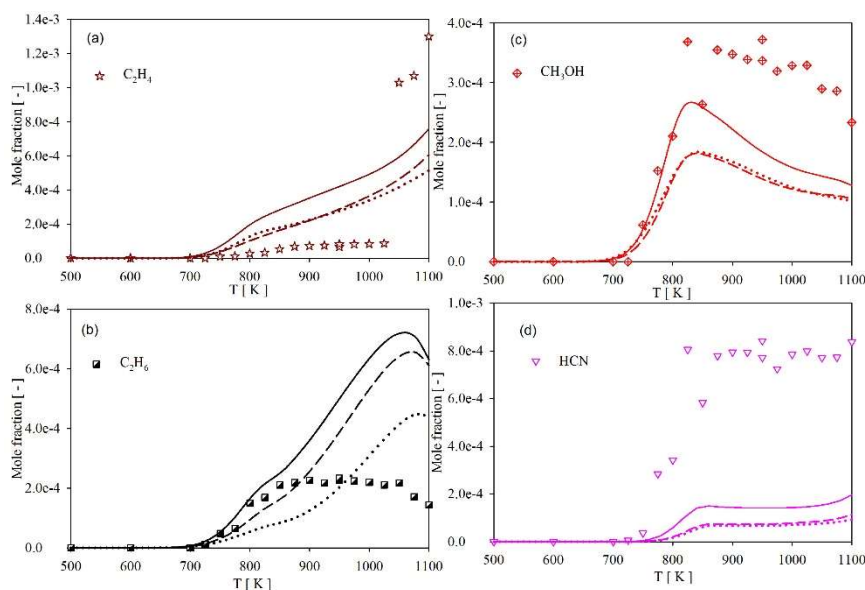


Figure 23: Pyrolysis of CH_3NO_2 1%/He 99% (on a mole basis) in a flow reactor at 1.07 bar, a residence time of 2.0 s in the isothermal zone of the reactor. Symbols: present work; dashed lines: prediction with the proposed model, solid lines: model predictions with modified rate constant for reactions listed in Table 4 (targeting HCN), dotted lines: model predictions with modified rate constant for reactions listed in Table 4 (targeting C_2H_6).

6. Conclusions

The pyrolysis of nitromethane was studied at 1.07 bar in a flow reactor and in a jet-stirred reactor. Mole fraction profiles of products were determined using gas chromatography and FTIR

spectroscopy. Based on the present measurements and published experimental data, a detailed reaction mechanism for the pyrolysis and oxidation of nitromethane has been developed. A number of published experiments have been selected to demonstrate that important features of the CH_3NO_2 chemistry are well captured by the broadly validated reaction mechanism presented in this work. Overall, there is a good agreement between model predictions and measurements for the wide range of experimental conditions investigated. Among the nitrogenous species, NO is found to be a major product for both pyrolysis and oxidation. The kinetic analysis suggests that the thermal dissociation channel is the main path for the initial nitromethane decomposition followed by the H-atom abstraction channel. The most sensitive reactions for nitromethane pyrolysis in a flow reactor and during pyrolysis and oxidation in a jet-stirred reactor are found to be $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$ and $\text{CH}_3 + \text{NO}_2 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$. The simple decomposition pathway (with regard to the C-atom flow) of nitromethane during pyrolysis can be represented as $\text{CH}_3\text{NO}_2 \rightarrow \text{CH}_3 \rightarrow \text{CH}_3\text{O} \rightarrow \text{CH}_2\text{O} \rightarrow \text{HCO}$. For the burner stabilized premixed flame the thermal dissociation channel $\text{CH}_3\text{NO}_2(+\text{M}) \rightleftharpoons \text{CH}_3 + \text{NO}_2(+\text{M})$ gains importance as mixtures become richer while the H-atom abstraction channel is weakened. Under flame conditions, NO is mainly formed via the NO_2 route: $\text{NO}_2 + \text{H} \rightleftharpoons \text{NO} + \text{OH}$, $\text{NO}_2 + \text{CH}_3 \rightleftharpoons \text{CH}_3\text{O} + \text{NO}$. Based on the mass flux analysis, several important reactions were identified involving hydrocarbon radicals and nitric oxides. Published rate parameters of reactions related to nitromethane chemistry are limited and the ones available in the literature are decades old. Further work is required in this area to enhance the model development and improve the understanding of nitromethane combustion and pyrolysis. The here derived model is capable to serve as the base mechanism for hydrocarbon fuels and for oxygenated fuels higher than C_2 and nitrogen-containing compounds higher than C_1 as well as for compounds with nitrogen only.

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Supplementary material

Additional comparisons and kinetic analyses

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